



THE KENYA NATIONAL EXAMINATIONS COUNCIL
Kenya Certificate of Secondary Education

233/2

CHEMISTRY (Theory)

Nov. 2023 - 2 hours



Paper 2

Serial No.

14848389

Name: **Index Number:**

Candidate's signature: **Date:**

Instructions to candidates

- (a) Write your name and index number in the spaces provided above.
- (b) Sign and write the date of examination in the spaces provided above.
- (c) Answer **all** the questions in the spaces provided in the question paper.
- (d) **Non-programmable** silent electronic calculators and KNEC mathematical tables may be used.
- (e) All working **must** be clearly shown where necessary.
- (f) **This paper consists of 16 printed pages.**
- (g) **Candidates should check the question paper to ascertain that all the pages are printed as indicated and that no questions are missing.**
- (h) **Candidates should answer the questions in English.**



For Examiner's use only

Question	Maximum Score	Candidate's Score
1	13	
2	11	
3	12	
4	10	
5	12	
6	12	
7	10	
Total Score	80	



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- 1 Table 1 gives some properties of the elements in period 3 of the periodic table.

Table 1

Element	Na	Mg	Al	Si	P	S	Cl	Ar
Atomic number	11	12	13	14	15	16	17	18
Atomic radius (nm)	0.186	0.160	0.143	0.117	0.110	0.104	0.099	0.097



- (a) Give the formula and name of the compound formed by the reaction between Al and S.

Formula (1 mark)

Name (1 mark)

- (b) Explain the variations in the atomic radius of the elements across the period. (2 marks)

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- (c) Select the element with the highest ionisation energy. Give a reason. (2 marks)

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- (d) Write the electron arrangement of phosphorus in PCl_3 . (1 mark)

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- (e) Select an element that forms an ion with the smallest ionic radius. Give a reason. (2 marks)

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- (f) Table 2 gives the melting points ($^{\circ}\text{C}$) of some of the elements.

Table 2

Element	Na	Mg	Cl	Ar
Melting point ($^{\circ}\text{C}$)	98	650	-101	-189

Explain, in terms of structure and bonding, the differences in the melting points of:

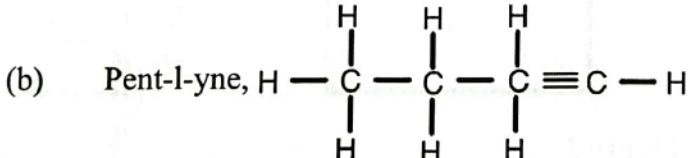
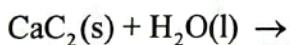
- (i) Na and Mg; (2 marks)

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- (ii) Cl and Ar. (2 marks)

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- 2 (a) Complete the following equation: (1 mark)



reacts with bromine to form compounds B and C as shown in Figure 1.

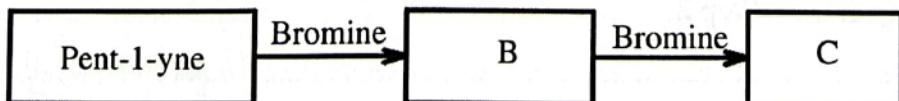


Figure 1

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Draw the structures of compounds **B** and **C**.

Compound B

(1 mark)

Compound C

(1 mark)

- (c) Study the flow chart in **Figure 2** and answer the questions that follow.

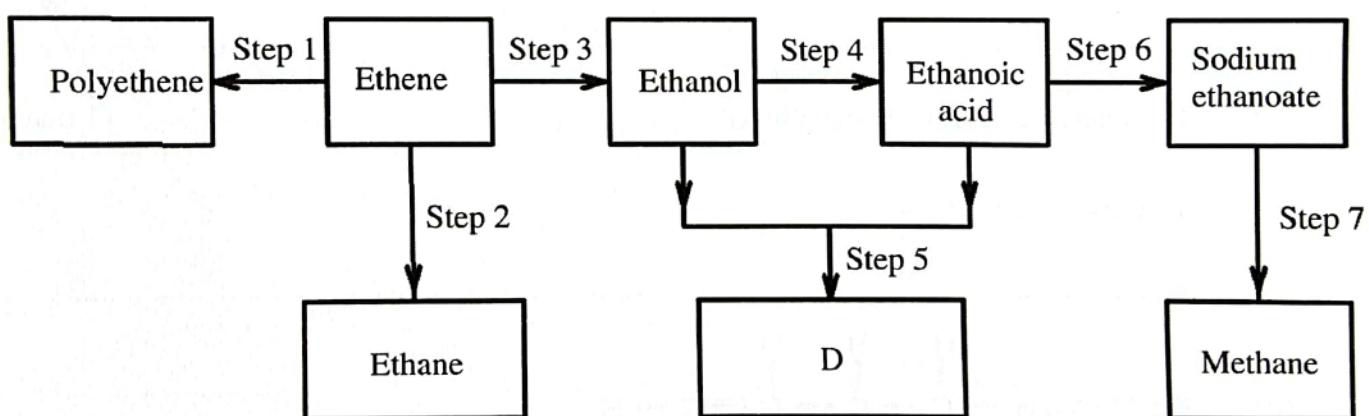


Figure 2

- (i) Give the reagents and conditions used in:

I. Step 2;

(1 mark)

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II. Step 7. (1 mark)



(ii) Write an equation for the reaction that takes place in:

I. Step 1; (1 mark)

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II. Step 3. (1 mark)

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(iii) Name the type of reaction that takes place in:



I. Step 4; (1 mark)

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II. Step 2. (1 mark)

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(iv) I. Draw the structure of organic compound D. (1 mark)

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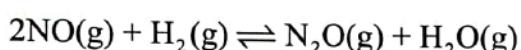
II. Give the name of compound D. (1 mark)

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- 3 (a) Explain how an increase in temperature affects the rate of a chemical reaction. (2 marks)

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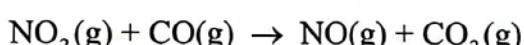
- (b) Consider the following gaseous reaction:



Explain how an increase in pressure affects the rate of this reaction. (2 marks)

.....

- (c) At high temperatures, NO₂ and CO gases react as shown in the following equation:



The reaction was monitored by measuring the changes in the concentration of NO(g) with time. **Table 3** shows the data obtained.



Table 3

Time/seconds	Concentration of NO × 10 ³ / moles per litre
0	0
50	16
100	22
150	26
200	29
250	31
300	32

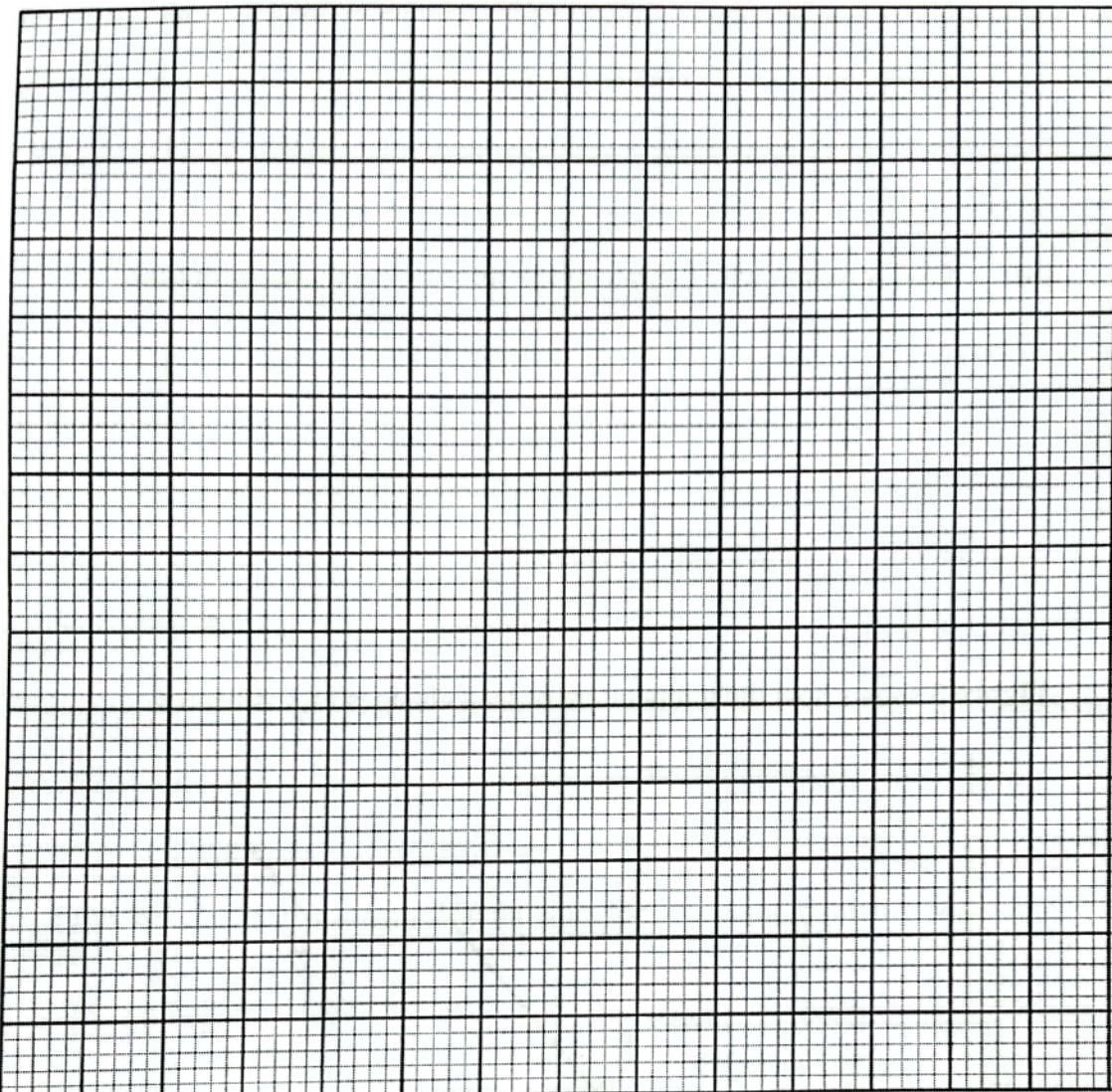
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- (i) Plot on the grid provided, a graph of concentration of NO (vertical axis) against time.
(3 marks)



- (ii) Use the graph, to determine the rate of the reaction:
I. In the time interval 25 seconds and 75 seconds; (2 marks)



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II. At the 175th second.

(2 marks)



(iii) Give a reason why the rate of the reaction decreases with time. (1 mark)

- 4 (a) Figure 3 shows how the temperature of lead changes as it is heated.

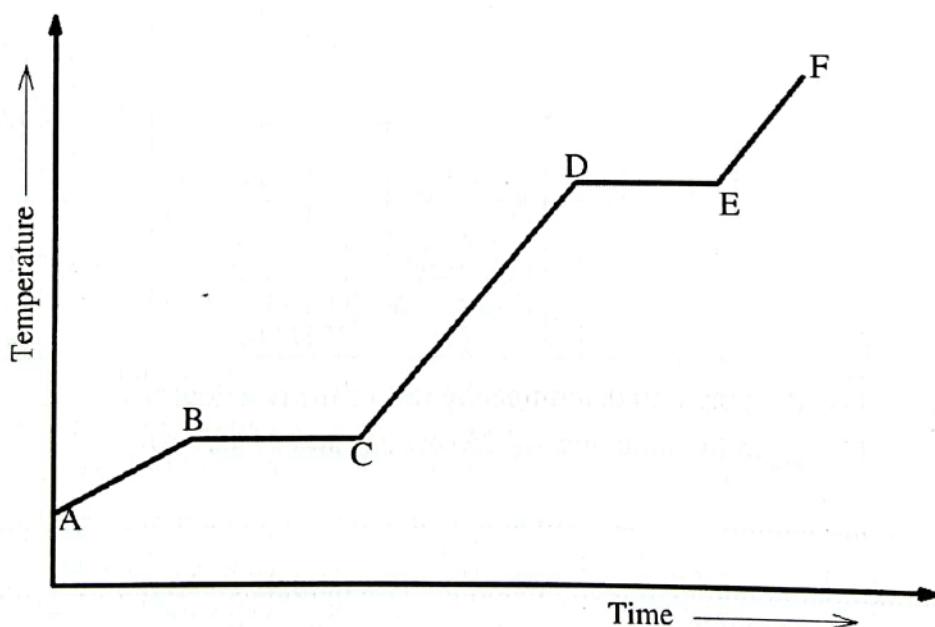


Figure 3

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(i) Label on the diagram the states present on the regions:

I. CD;

(1 mark)

.....
II. EF.

(1 mark)

(ii) Explain why the temperature remains constant in regions:

I. BC;

(1 mark)

.....
II. DE.

(1 mark)

(b) Figure 4 shows an energy cycle diagram for processes involving potassium bromide.

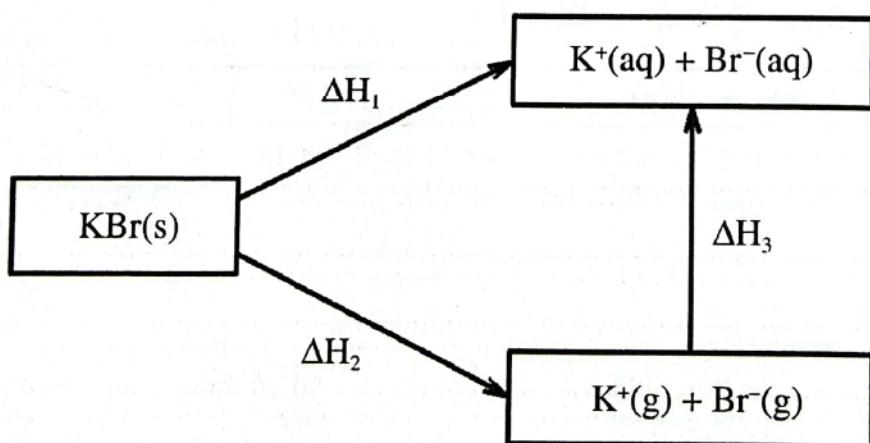


Figure 4

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(i) Name the following enthalpy changes:

I. ΔH_1 ;

(1 mark)

II. ΔH_2 ;

(1 mark)

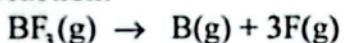
III. ΔH_3 .

(1 mark)

(ii) Write an expression showing how ΔH_3 can be obtained from ΔH_1 and ΔH_2 .

(1 mark)

(c) Using the thermochemical data given in Table 4, calculate the enthalpy change for the reaction:



(2 marks)

Table 4

Process	$\Delta H / \text{kJmol}^{-1}$
$B(s) \rightarrow B(g)$	590
$B(s) + \frac{3}{2} F_2(g) \rightarrow BF_3(g)$	-1111
$F_2(g) \rightarrow 2F(g)$	158



- 5 (a) Use the standard electrode potentials in **Table 5** to answer this question.

Table 5

Number	Electrode reaction	E^\ominus, V
I	$2\text{H}^+(\text{aq}) + 2\text{e} \rightarrow \text{H}_2(\text{g})$	0.00
II	$\text{Zn}^{2+}(\text{aq}) + 2\text{e} \rightarrow \text{Zn}(\text{s})$	- 0.76
III	$\text{Sn}^{2+}(\text{aq}) + 2\text{e} \rightarrow \text{Sn}(\text{s})$	- 0.14
IV	$\text{Cu}^{2+}(\text{aq}) + 2\text{e} \rightarrow \text{Cu}(\text{s})$	+ 0.34
V	$\text{Fe}^{2+}(\text{aq}) + 2\text{e} \rightarrow \text{Fe}(\text{s})$	- 0.44
VI	$\text{Pb}^{2+}(\text{aq}) + 2\text{e} \rightarrow \text{Pb}(\text{s})$	- 0.13
VII	$\text{Cu}^+(\text{aq}) + \text{e} \rightarrow \text{Cu}(\text{s})$	+ 0.52
VIII	$\text{Ag}^+(\text{aq}) + \text{e} \rightarrow \text{Ag}(\text{s})$	+ 0.80

- (i) Select **two** electrodes which when connected gives the cell with the lowest e.m.f. (1 mark)
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- (ii) Arrange the metals Ag, Fe, and Sn and in order of their reactivity with dilute hydrochloric acid, starting with the most reactive. Give a reason. (2 marks)
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(iii) An electrochemical cell is made up of electrode numbers IV and VII.

I. Calculate the e.m.f of the cell.

(1 mark)

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II. Write an equation for the cell reaction.

(1 mark)

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(iv) Draw a labelled diagram of an electrochemical cell that is used to measure the standard electrode potential for tin (Sn), electrode number III. (3 marks)

(b) The products of electrolysis of sodium chloride, depend on the conditions used. Give the products obtained under each set of conditions in Table 6.

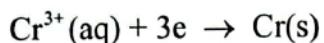
Table 6

Conditions	Product at:	
	Anode	Cathode
Dilute aqueous sodium chloride		
Concentrated aqueous sodium chloride		

(2 marks)



- (c) Aqueous chromium(III) sulphate was electrolysed using inert electrodes. The equation for the reaction is:



Calculate the time in seconds required to deposit 2.6 g chromium using a current of 5.5 amperes. (1 Faraday = 96,500 Coulombs; Cr = 52.0) (2 marks)

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- 6 (a) (i) State Charles' law of gases. (1 mark)

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- (ii) Table 7 shows the data obtained in an experiment using 0.012 moles of neon gas.

Table 7

Temperature/ K	Volume/ dm ³	Pressure/ atm
250	0.005	50
300	0.006	50



- Show that the data is consistent with Charles' law. (2 marks)

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- (b) (i) State Graham's law of diffusion of gases. (1 mark)

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- (ii) Given that 1 mole of a gas occupies a volume of 24.0 dm^3 at 298 K , calculate the density in grams per litre of:

I. oxygen gas ($\text{O} = 16$)

(1 mark)

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II. hydrogen gas ($\text{H} = 1.0$)

(1 mark)

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- (iii) Determine the rate of diffusion of hydrogen gas compared to that of oxygen gas at 298 K .

(2 marks)

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- (c) Ammonia gas was prepared in the laboratory by warming a mixture of solid ammonium chloride and solid calcium hydroxide. The equation for the reaction is:



The gas was dried and then collected. If the volume of ammonia collected was 1340 cm^3 measured at 312 K and 1 atmosphere pressure:

($\text{N} = 14.0$; $\text{Cl} = 35.5$; $\text{H} = 1.0$; Volume of one mole of gas at $298\text{K} = 24 \text{ dm}^3$)



- (i) Calculate the volume that ammonia gas will occupy at 298 K and 1 atmospheric pressure.

(2 marks)

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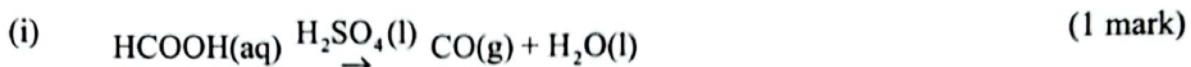
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- (ii) Determine the mass of ammonium chloride that reacted.

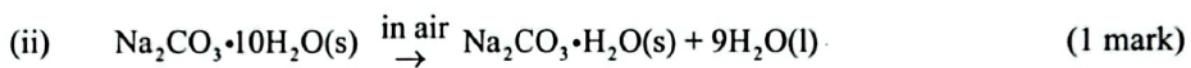
(2 marks)

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- 7 (a) Give the names of the processes represented by the following equations:



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- (b) Sodium carbonate is manufactured through a series of reactions involving sodium chloride, ammonia and carbon(IV) oxide.



- (i) Ammonia is obtained by reacting hydrogen and nitrogen in the Haber process. State how the other two materials are obtained:

I. Sodium chloride; (1 mark)

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II. Carbon(IV) oxide. (1 mark)

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- (ii) Concentrated sodium chloride solution, saturated with ammonia is passed into a carbonation tower in which carbon(IV) oxide is bubbled through. Reactions in the tower involve formation of ammonium hydrogen carbonate which then reacts with sodium chloride to form sodium hydrogen carbonate.

Write the equations for the formation of:

- I. Ammonium hydrogen carbonate; (1 mark)

-
II. Sodium hydrogen carbonate. (1 mark)

- (iii) Describe how the:

- I. sodium hydrogen carbonate is separated; (1 mark)

-
II. Sodium hydrogen carbonate is converted to sodium carbonate. (1 mark)

- (iv) One of the uses of sodium carbonate is in the removal of water hardness.

- I. Explain how sodium carbonate removes water hardness. (1 mark)

-
II. State one other industrial use of sodium carbonate. (1 mark)

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