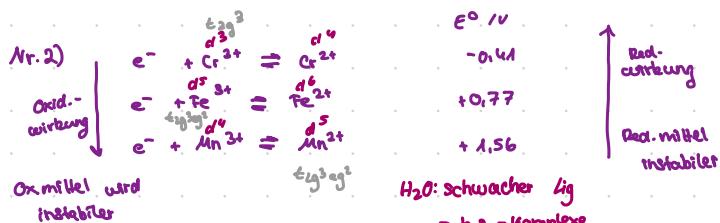
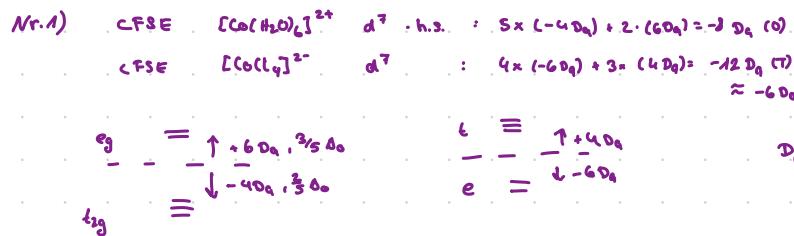
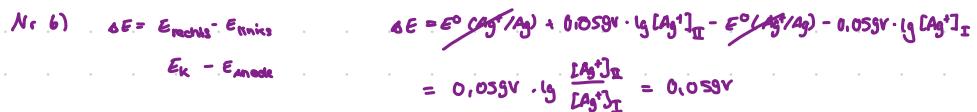
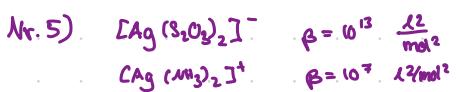
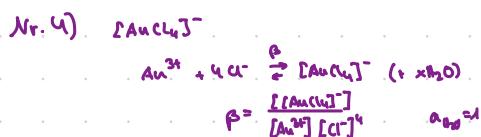


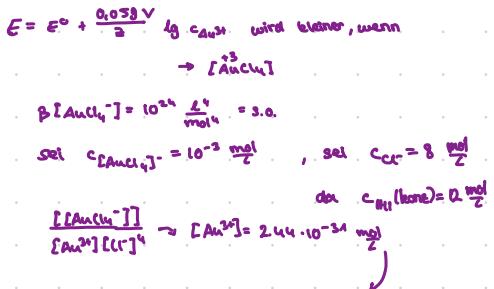
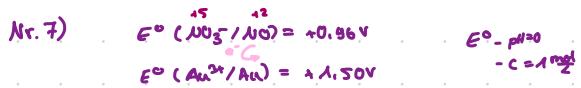
Übung 5



Nr. 3) $\Delta G^\circ = -R \cdot T \cdot \ln k$
 $= -2 \cdot 298 \text{ K} \cdot 25 = -195,1 \text{ kJ/mol}$



- a) $+ \text{NH}_3$ in I: $\rightarrow [\text{Ag}(\text{NH}_3)_2]^+$
 $[\text{Ag}^+]$ sinkt, α C wird größer $\rightarrow \Delta E$ wird größer
- b) $+ \text{S}_2\text{O}_3^{2-}$ in I: $\rightarrow [\text{Ag}(\text{S}_2\text{O}_3)_2]^{3-}$
 $[\text{Ag}^+]$ sinkt weiter $\rightarrow \Delta E$ wird noch größer



in Nernst: $E(\text{Au}^{3+}/\text{Au}) = 0,90V$
 \Rightarrow oberhalb von $E^\circ(\text{NO}_3^-/\text{NO})$