

## B.Sc. 5th Semester (Honours) Examination, 2023 (CBCS)

Subject : Chemistry

Course : CC-XI  
(Inorganic Chemistry)

Time: 2 Hours

Full Marks: 40

*The figures in the margin indicate full marks.**Candidates are required to give their answers in their own words  
as far as practicable.*

1. Answer any five questions: 2×5=10
- State the conditions for orbital contribution to the magnetic moment.
  - What will be the increasing order of  $\Delta_{3p}$  for complexes  $[\text{Ni}(\text{CN})_4]^{2-}$ ,  $[\text{Pd}(\text{CN})_4]^{2-}$  and  $[\text{Pt}(\text{CN})_4]^{2-}$ ? Give reason(s).
  - For which complex spin state equilibrium may occur —  $[\text{Co}(\text{CN})_6]^{4-}$ ;  $[\text{CuCl}_6]^{4-}$ ;  $[\text{Ni}(\text{CO})_4]$  and  $[\text{PtCl}_4]^{2-}$ ? Give reason
  - Why do 4d and 5d metals form square planar complexes rather than tetrahedral complexes?
  - Write down general electronic configuration of transition elements and actinides. Hence indicate the effect of lanthanide type contraction in both cases.
  - Draw Orgel diagram of  $d^7 \text{O}_h$  and  $d^9 \text{T}_d$  complexes.
  - Why are compounds of lanthanides and actinides are coloured?
  - ' $[\text{Co}(\text{H}_2\text{O})_6]^{2+}$  forms  $[\text{CoCl}_4]^{2-}$  but  $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$  does not forms  $[\text{NiCl}_4]^{2-}$ .' — Why?
2. Answer any two questions: 5×2=10
- 'Tetrahedral complexes are always high spin.' — Explain.
    - 'Carbonyl' and 'Phosphine' — the bread and butter ligand for first transition series, are not good ligands for lanthanides.' — Explain. 2½+2½
  - Which will show greater  $\Delta_0$  value among the pairs and why? (A)  $[\text{Co}(\text{CN})_6]^{3-}$  and  $[\text{Co}(\text{CN})_6]^{4-}$  (B)  $[\text{Co}(\text{CN})_6]^{3-}$  and  $[\text{CoCl}_6]^{3-}$ .
    - Give example of one high spin square planar complex. Draw its structure and give its  $d$ -orbital splitting pattern. (½+½+½+½)+3
  - Electronic configuration of  $\text{Cu}^+$  and  $\text{Cu}^{2+}$  is  $[\text{Ar}] 3d^{10}$  and  $[\text{Ar}] 3d^9$ , respectively. Through  $\text{Cu}^+$  has filled valence shell but it is less stable than  $\text{Cu}^{2+}$ . — Why?
    - Why do transition elements show variable valency? 2½+2½
  - Why are common oxidation state of 3d series metal ions (except Sc and Cu) +2?
    - Why are absorption band of lanthanides very sharp where as that of transition metal compounds broad? 2½+2½



3. Answer any two questions:

10×2=20

- (a) (i) Why is the cause of Jahn-Teller distortion?  
 (ii) In octahedral geometry among  $\text{Co}^{2+}$  high spin or  $\text{Co}^{2+}$  low spin complex, which one will face more Jahn-Teller distortion?  
 (iii) Why  $\text{Pt(II)}$  and  $\text{Pd(II)}$ , though symmetrically field, undergoes Jahn-Teller distortion?  
 (iv) Why does  $\text{Cu(II)}$  experience distortion even for weak field ligand but  $\text{Ni(II)}$  system experience it only with strong field ligand?  
 (v) Which type of distortion is preferred by  $\text{Cu(II)}$  complexes? Give reason.  $2+2+2+2+2$
- (b) (i) Discuss the structure and bonding of  $[\text{Cr}(\text{CH}_3\text{COO})_2]_2 \cdot 2\text{H}_2\text{O}$  and  $[\text{Cu}(\text{CH}_3\text{COO})_2]_2 \cdot 2\text{H}_2\text{O}$ . Also discuss their magnetic moment.  
 (ii) Why is separation of lanthanide ions a difficult task?  
 (iii) '2nd and 3rd row transition metals are very much close in their chemical properties'—why?  $(1+1+1+1+1+1)+2+2$
- (c) (i) Draw the crystal field splitting pattern for  $d^8$  configuration in octahedral, tetrahedral and square planar electronic field.  
 (ii) Calculate CFSE in weak and strong crystal field for  $d^7$  configuration in octahedral, tetrahedral and square planar electronic field.  $(1\frac{1}{2}+1\frac{1}{2}+2)+(1\frac{1}{2}+1\frac{1}{2}+1+1)$
- (d) (i) What is magnetically concentrated compound and magnetically dilute compound? Explain with example.  
 (ii) Explain quenching of magnetic moment with example.  
 (iii) What is super exchange phenomenon—state with example.  
 (iv) Give elementary idea on antiferromagnetism with example.  $(2+2)+2+2+2$