

[例题7-1]  $\text{H}_2$ 、 $\text{He}$ 共2.5g, 混合 $v$ 不变,  $\Delta T=1\text{K}$ , 共吸热 $2.25R$ ;

求(1)  $m_{\text{H}_2}$ ,  $m_{\text{He}}$ ;

(2) 若 $P$ 不变,  $\Delta T=1\text{K}$ , 混合气吸热?

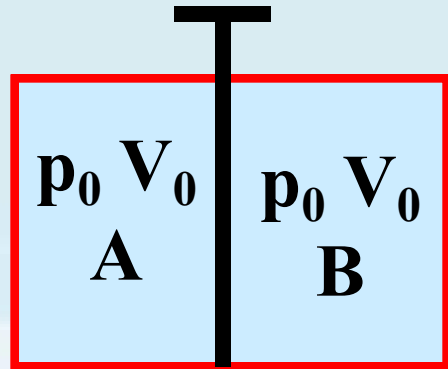
解(1)  $Q_V = \nu_{\text{H}_2} C_{V_{\text{H}_2}} \Delta T + \nu_{\text{He}} C_{V_{\text{He}}} \Delta T$

$$\Rightarrow \left. \begin{aligned} \frac{m_{\text{H}_2}}{2} \frac{5R}{2} + \frac{m_{\text{He}}}{4} \frac{3R}{2} &= 2.25R \\ m_{\text{H}_2} + m_{\text{He}} &= 2.5 \end{aligned} \right\} \begin{aligned} m_{\text{H}_2} &= 1.5\text{g} \\ m_{\text{He}} &= 1\text{g} \end{aligned}$$

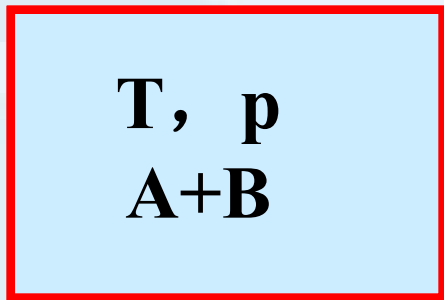
(2)  $Q_P = \nu_{\text{H}_2} C_{P_{\text{H}_2}} \Delta T + \nu_{\text{He}} C_{P_{\text{He}}} \Delta T$

$$= \frac{1.5}{2} \times \frac{7}{2} R + \frac{1}{4} \times \frac{5}{2} R = 3.25R$$

[讨论1] 已知  $p_0$ 、 $V_0$ , 抽去绝热板, 求平衡时  $T$ 、 $p$



1mol 2mol  
3原子 双原子



解  $E_A + E_B = E'_A + E'_B$

$$v_A \frac{6}{2} RT_A + v_B \frac{5}{2} RT_B = v_A \times \frac{6}{2} RT + v_B \times \frac{5}{2} RT$$

1mol      2mol

$$\frac{6}{2} p_0 V_0 + \frac{5}{2} p_0 V_0 = \frac{16}{2} RT \Rightarrow T = \frac{11 p_0 V_0}{16 R}$$

$$p = nkT = \frac{N_A + 2N_A}{2V_0} kT = \frac{3RT}{2V_0} = \frac{33}{32} p_0$$

or.  $p 2V_0 = (1+2)RT \rightarrow p = \frac{33}{32} p_0$

[例题7-2] 一定量  $N_2$ , 绝热升压  $p \rightarrow 2p$ ,  $\bar{v}$  变为原几倍?

$$\text{解: } \bar{v} = \sqrt{\frac{8RT}{\pi M_m}} \Rightarrow \frac{\bar{v}_2}{\bar{v}_1} = \sqrt{\frac{T_2}{T_1}} = 2^{1/7}$$

$$\left. \begin{array}{l} p^{\gamma-1} T^{-\gamma} = c_2 \\ \gamma = C_p / C_v = 7/5 \end{array} \right\} \Rightarrow \frac{T_2}{T_1} = \left( \frac{p_2}{p_1} \right)^{2/7} = 2^{2/7}$$

[讨论2] 若上题为绝热压缩:  $V \rightarrow V/2$ , 压缩前后  $\bar{z}$  如何变化?

$$\begin{aligned} \text{解: } \bar{z} &= \sqrt{2} n \pi d^2 \bar{v} = \sqrt{2} \frac{N}{V} \pi d^2 \sqrt{\frac{8RT}{\pi M_m}} = c_3 \frac{\sqrt{T}}{V} \xrightarrow{TV^{\gamma-1} = c_4} \bar{z} = c_5 V^{-\frac{1}{2} - \frac{\gamma}{2}} \\ \Rightarrow \frac{\bar{z}_2}{\bar{z}_1} &= \left( \frac{V_2}{V_1} \right)^{-\frac{1}{2} - \frac{\gamma}{2}} = 2^{\frac{6}{5}} \end{aligned}$$

若为  $H_2O$  蒸汽,  $\bar{z}$  如何变化?  $\gamma = \frac{4}{3}$