General Chemistry 2 | 3rd Quarter

MPT2: Concentration

January 19 2021

Virtual Lab Link

Virtual Lab Guide

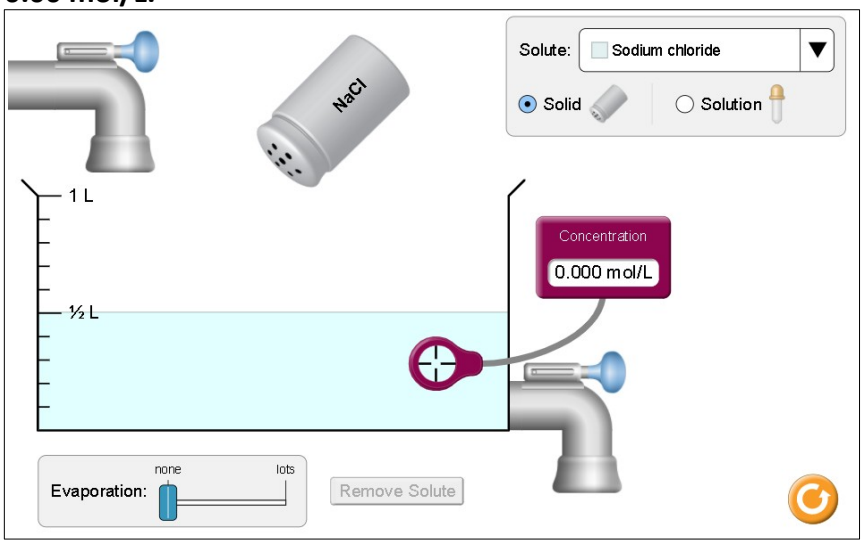
Guide Questions:

1. What solution concentration is used in the Virtual Lab in computing for the concentration of a solution, MOLARITY, MOLALITY, or PERCENTAGE CONCENTRATION BY MASS? What made you say so?

   Molarity, because of the unit it displays which is mol/L.

2. What is the concentration of pure water in the Virtual Lab?

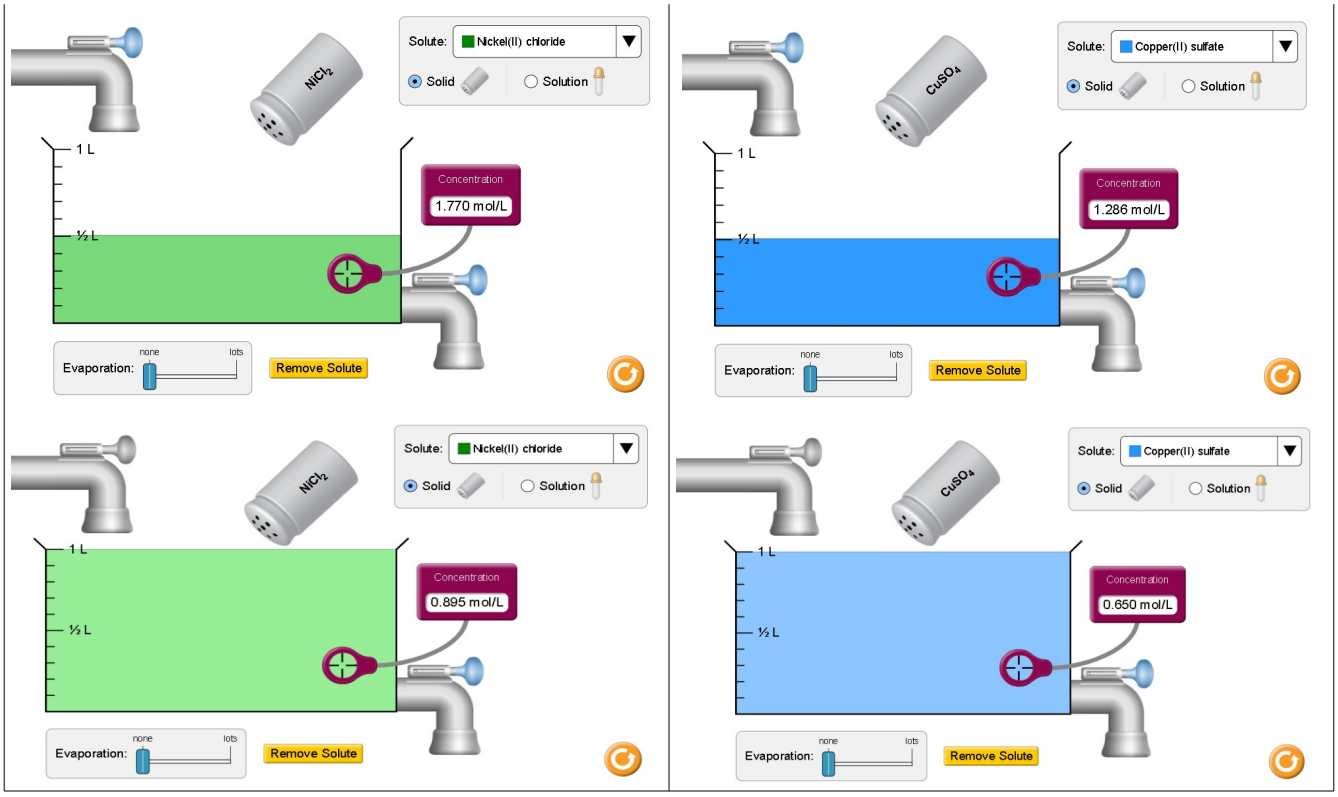
   0.00 mol/L.



3. If you double the amount of solvent (pure water), what will happen to the concentration of solid nickel (II) chloride? Solid copper (II) sulfate?

   For $NiCl_2$, the solution became less concentrated significantly. While for the $CuSO_4$, the concentration didn't decrease significantly.

$NiCl_2$	$CuSO_4$

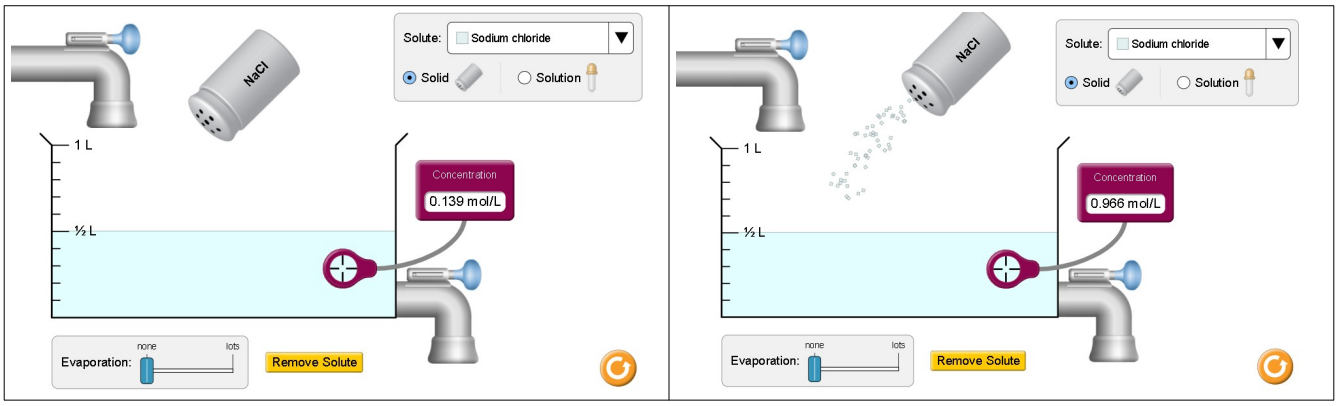


4. For the same amount of solvent, what will happen to the concentration of the solution if you continue adding a solute for the same temperature?

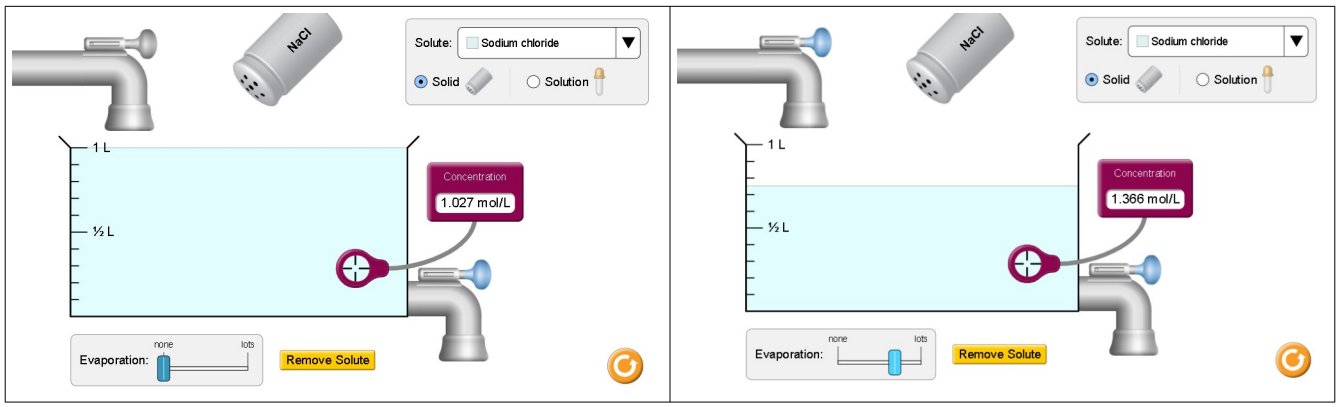
   If you continue adding a solute, the concentration will drastically go up.

5. Cite 2-3 steps you did to increase the concentration of the solution.

   1. **Add more solute to the solution.**



   2. **Evaporate the solution.**



3. **Drop liquid solute in it.**