title

author

2018-09-22

目录

第一章	Introduction	5
第二章	Demos	7
2.1	Chemical formulae	7
2.2	Chemical equations	8
2.3	Structural formulae	9

4 目录

第一章 Introduction

第二章 Demos

2.1 Chemical formulae

```
H_2O, Sb_2O_3
H^+,\,CrO_4^{2-},\,[AgCl_2]^-,\,Y^99^+,\,Y^{99+}
\mathrm{Fe^{II}Fe^{III}2O_4},\, 2\,\mathrm{H_2O},\, 2\,\mathrm{H_2O},\, 0.5\,\mathrm{H_2O},\, \tfrac{1}{2}\,\mathrm{H_2O},\, (_{1/2})\mathrm{H_2O}\,\,,\, n\,\mathrm{H_2O}
^{227}_{90}\text{Th}^{+}, ^{2}27_{90}\text{Th}^{+}, ^{0}_{-1}\text{n}^{-}, ^{0}_{-1}\text{1}
H^3HO, H^3HO
(NH_4)_2S, [\{(X_2)_3\}_2]^{3+}
H_2(aq), CO_3^2 - {}_{(aq)}, NaOH(aq, \infty)
OCO · -, NO<sup>(2 · )-</sup>
\mathrm{NO_{x}}, \texttt{\sc family bfseries ce \{NO_x\}}, \mathrm{Fe^{n+}}, \texttt{\sc family bfseries ce \{Fe^n+\}}
\ce{\mu-Cl}, \ce{[Pt(\eta^2-C2H4)Cl3]-}
KCr(SO_4)_2 \cdot 12 H_2O, KCr(SO_4)_{2,12}H_2O, KCr(SO_4)_2 * 12 H_2O
C_6H_5-CHO, A-B=C\equiv D, \sffamily\bfseries\ce{A-B=C\#D}
A-B=C\equiv D, A-B=C, A\equiv B\equiv C\equiv D, A\cdots B\cdots C, A\rightarrow B\leftarrow C
```

2.2 Chemical equations

$$\mathbf{A} \longrightarrow \mathbf{B}$$

$$A \longleftarrow B$$

$$\mathbf{A} \longleftrightarrow \mathbf{B}$$

$$\mathbf{A} \longleftarrow -> \mathbf{B}$$

$$A \Longrightarrow B$$

$$A \rightleftharpoons B$$

$$A \rightleftharpoons B$$

$$\mathbf{A} \xrightarrow{\mathbf{H_2O}} \mathbf{B}$$

$$\mathbf{A} \xrightarrow[text below]{} \mathbf{B}$$

$$\mathbf{A} \xrightarrow[x_i]{x} \mathbf{B}$$

$$\mathbf{A} \xrightarrow{x} \mathbf{B}$$

$$A + B$$

$$A-B$$

$$A = B$$

$$A \pm B$$

$$SO_4^{2-} + Ba^{2+} \longrightarrow BaSO_4 \downarrow$$

$$A \downarrow B \downarrow \longrightarrow B \uparrow B \uparrow$$

$$CH_4 + 2 \left(O_2 + \frac{79}{2} N_2\right)$$

$$xNa(NH_4)HPO_4 \xrightarrow{\Delta} (NaPO_3)_x + xNH_3 \uparrow + xH_2O$$

$$\mathrm{CO}_2 + \mathrm{C} \longrightarrow 2\,\mathrm{CO}$$

$$\mathrm{Hg}^{2+} \xrightarrow{I-} \mathrm{HgI}_2 \xrightarrow{I-} [\mathrm{Hg}^{\mathrm{II}}\mathrm{I}_4]^{2-}$$

$$\operatorname{Zn}^{2+} \xrightarrow[+2\,\operatorname{H}^+]{}^{+2\,\operatorname{OH}^-} \operatorname{Zn}(\operatorname{OH})_2 \downarrow \xrightarrow[+2\,\operatorname{H}^+]{}^{+2\,\operatorname{OH}^-} \left[\operatorname{Zn}(\operatorname{OH})_4\right]^{2-}$$

$$K = \frac{[{\rm Hg}^{2+}][{\rm Hg}]}{[{\rm Hg}_2^{2+}]}$$

$$K = \frac{[Hg^2 +][Hg]}{[Hg2^2 +]}$$

2.3 Structural formulae

Br

$$\begin{array}{c} H \\ | \\ | \\ | \\ + \\ | \\ | \\ H \end{array}$$

$$\begin{array}{c|c} O & \operatorname{CH}_3 \\ \hline & F & \\ \operatorname{Cl} & \\ \end{array}$$

$$\operatorname{CH}_3$$
 Cl
 Cl
 C
 C

$$\begin{array}{c} \text{HOCH}_2 \\ \\ \text{H}_3\text{C} \\ \end{array} \begin{array}{c} \text{Cl} \\ \\ \text{S} \\ \end{array}$$

$$O$$
 CH_3
 CH_3
 CH_3

$$\begin{array}{c|cccc} O & H & H \\ \parallel & \parallel & \parallel + \\ Cl - Cl - Cl & CH_3 - N - H \\ & \parallel & H \end{array}$$

$$O = C \xrightarrow{F} Cl \xrightarrow{F} Cl \xrightarrow{F} Cl \xrightarrow{G} Cl \xrightarrow{G} Cl$$

$$C \xrightarrow{G} Cl \xrightarrow{G} Cl \xrightarrow{G} Cl$$

$$C \xrightarrow{G} Cl \xrightarrow{G} Cl$$

$$C \xrightarrow{G} Cl \xrightarrow{G} Cl$$

$$C \xrightarrow{G} C$$