

Introduction to Nuclear Magnetic Resonance Spectroscopy

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Introduction. Magnetic moment in magnetic field.

- Introduction to Magnetic Resonance
- Applications of NMR
- Magnetic moments in magnetic field.
- Orbital angular momentum operator
- Spin angular momentum operator
- Spin in a magnetic field
- Equilibrium magnetization
- Resonant energy absorption.
- Populations dynamics in two-level system.

1.1. Introduction to Magnetic Resonance

- Magnetic resonance (MR) is a phenomenon of resonant energy absorption by a system of nuclei (and electrons).
- Nuclear magnetic resonance (NMR) results from the intrinsic magnetic moment of the nuclei of some atoms. Magnetic moments of electrons are exploited in electron spin resonance (ESR).
- Magnetic resonance (MR) generally involves placing a sample in a strong magnetic field (to generate polarisation at a fixed resonant frequency) and detecting signals produced following application of pulsed radio-frequency electromagnetic fields (RF pulses).
- MR is a very powerful method for studying the structure of materials: used in physics, chemistry, biology, medicine etc.

1.2. Applications of NMR

- NMR spectroscopy is used for chemical analysis and for molecular structure determination

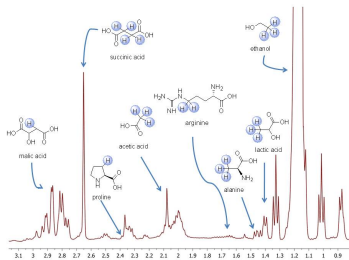


Fig.1: ^1H NMR spectrum of a sample of Spanish wine (<http://www.unirioja.es/gsoe/NMR.htm>)

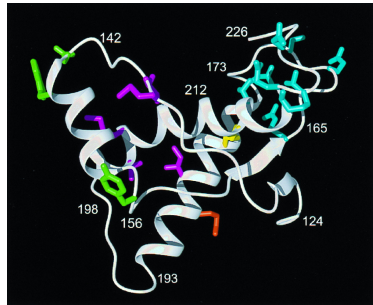


Fig.2: NMR-derived structure of a prion <http://www.pnas.org/content/94/14/7281.full>

1.2. Applications of NMR

- NMR relaxometry can be used to monitor molecular environment

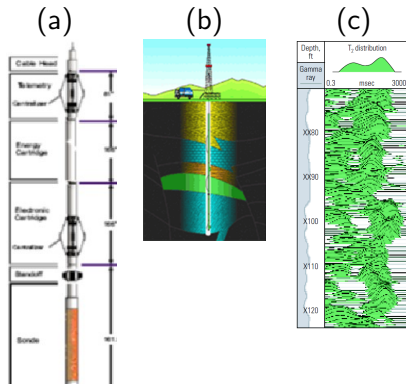


Fig.3: (a) NMR-logging probe, (b) Schematic positioning of the probe in a well, (c) T_2 -relaxation profile along the bore. Sources: 1) Allen et al. Oilfield review, Autumn 2000; 2) Coates, Xiao NMR Logging Principles and Applications, Halliburton

1.2. Applications of NMR

- NMR forms the basis for magnetic resonance imaging (MRI)

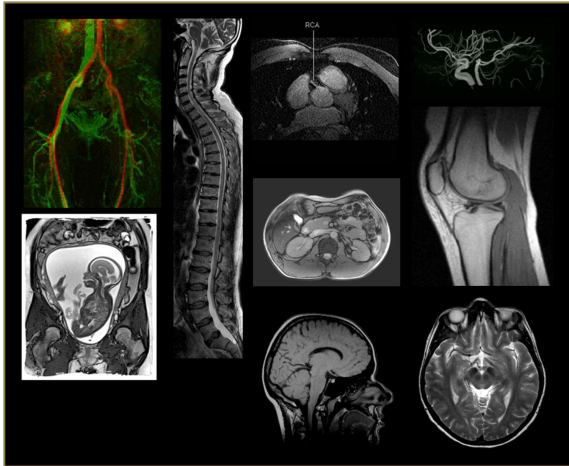


Fig.4: Example magnetic resonance images of blood vessel (in legs), fetus in utero, spine, heart, abdomen, head, blood vessels (in brain), knee, brain (courtesy of Prof. Richard Bowtell)

1.3. Magnetic moments in magnetic field.

- Consider charges moving in a limited volume. The position of a charge \mathbf{e}_n will be given by a vector \mathbf{r}_n and its velocity by \mathbf{v}_n . The overall magnetic moment of such a system is defined as:

$$\mathbf{M} = \frac{1}{2} \sum_n \mathbf{e}_n \mathbf{r}_n \times \mathbf{v}_n \quad (1)$$

- If all the charges and masses are the same, then \mathbf{M} can be rewritten as:

$$\mathbf{M} = \frac{e}{2m} \sum_n m \mathbf{r}_n \times \mathbf{v}_n = \gamma \mathbf{L}, \quad (2)$$

where

$$\mathbf{L} = \sum_n \mathbf{p}_n \times \mathbf{r}_n \quad (3)$$

is the mechanical angular momentum.

- Gyromagnetic ratio** (or magnetogyric):

$$\gamma = \frac{e}{2m} \quad (4)$$

1.3. Magnetic moments in magnetic field.

- When a magnetic moment \mathbf{M} is placed into an external uniform permanent magnetic field \mathbf{B} , its energy is given by:

$$E = -\mathbf{M} \cdot \mathbf{B} \quad (5)$$

- The torque acting on the system:

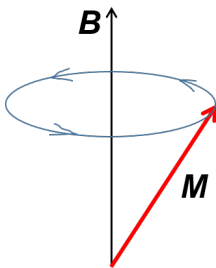
$$\frac{d\mathbf{L}}{dt} = \mathbf{M} \times \mathbf{B} \quad (6)$$

- Now using equation 2 we can obtain the equation describing the motion of vector \mathbf{M} :

$$\frac{d\mathbf{M}}{dt} = \gamma \mathbf{M} \times \mathbf{B} \quad (7)$$

1.3. Magnetic moments in magnetic field.

- In a uniform magnetic field directed along z-axis $\mathbf{B} = (0, 0, B_0)$, the equation for individual components of \mathbf{M} follow the equations:



$$\begin{aligned}\frac{dM_x}{dt} &= \omega_L M_y \\ \frac{dM_y}{dt} &= -\omega_L M_x \\ \frac{dM_z}{dt} &= 0,\end{aligned}\quad (8)$$

where $\omega_L = \gamma B_0$ - Larmor frequency.

- A solution to this system of differential equations with initial values of $M_x(0)$, $M_y(0)$, $M_z(0)$ has the following form:

$$\begin{aligned}M_x(t) &= M_x(0) \cos(\omega_L t) + M_y(0) \sin(\omega_L t) \\ M_y(t) &= -M_x(0) \sin(\omega_L t) + M_y(0) \cos(\omega_L t) \\ M_z(t) &= M_z(0)\end{aligned}\quad (9)$$

1.4. Orbital angular momentum operator

- In quantum mechanics physical quantity A is represented by an operator \hat{A} . The mechanical angular momentum is replaced by its corresponding operator:

$$\mathbf{L} = \sum_n \mathbf{p}_n \times \mathbf{r}_n \longleftrightarrow \hat{\mathbf{L}} = \frac{1}{\hbar} \sum_n \hat{\mathbf{r}}_n \times \hat{\mathbf{p}}_n = -i \sum_n \hat{\mathbf{r}}_n \times \nabla_n \quad (10)$$

- Angular momentum operator properties. Commutation:

$$[\hat{L}_y, \hat{L}_z] = i\hat{L}_x, [\hat{L}_z, \hat{L}_x] = i\hat{L}_y, [\hat{L}_x, \hat{L}_y] = i\hat{L}_z \quad (11)$$

- Angular momentum squared, and its commutation properties:

$$\hat{L}^2 = \hat{L}_x^2 + \hat{L}_y^2 + \hat{L}_z^2 \quad (12)$$

$$[\hat{L}^2, \hat{L}_x] = [\hat{L}^2, \hat{L}_y] = [\hat{L}^2, \hat{L}_z] = 0 \quad (13)$$

1.4. Orbital angular momentum operator

- Eigenfunctions of both \hat{L}^2 and \hat{L}_z operators can be characterized by integer quantum numbers l and m respectively. These eigen functions will be denoted as $|lm\rangle$. Their eigenvalues are:

$$\hat{L}_z|lm\rangle = m|lm\rangle \quad (14)$$

$$\hat{L}^2|lm\rangle = l(l+1)|lm\rangle \quad (15)$$

- Another useful operator are raising and lowering operators:

$$\hat{L}_+ = \hat{L}_x + i\hat{L}_y, \hat{L}_- = \hat{L}_x - i\hat{L}_y \quad (16)$$

$$\langle lm|\hat{L}_+|l(m-1)\rangle = \langle l(m-1)|\hat{L}_-|lm\rangle = \sqrt{(l+m)(l-m+1)} \quad (17)$$

Problem

Calculate $[\hat{L}_+, \hat{L}_x] = ?$, $[\hat{L}_+, \hat{L}_-] = ?$

$$[\hat{L}_y, \hat{L}_z] = i\hat{L}_x, [\hat{L}_z, \hat{L}_x] = i\hat{L}_y, [\hat{L}_x, \hat{L}_y] = i\hat{L}_z$$

$$\hat{L}^2 = \hat{L}_x^2 + \hat{L}_y^2 + \hat{L}_z^2$$

$$[\hat{L}^2, \hat{L}_x] = [\hat{L}^2, \hat{L}_y] = [\hat{L}^2, \hat{L}_z] = 0$$

$$\hat{L}_+ = \hat{L}_x + i\hat{L}_y, \hat{L}_- = \hat{L}_x - i\hat{L}_y$$

1.4. Orbital angular momentum operator

- Classical magnetic moment will have its own quantum analogue, the operator of angular momentum:

$$\mathbf{M} = \gamma \mathbf{L} \longleftrightarrow \hat{\mu} = \gamma \hbar \hat{\mathbf{L}} \quad (18)$$

- Given the electron charge $e = 1.6 \cdot 10^{-19} \text{C}$, and mass $m = 9.1 \cdot 10^{-31} \text{kg}$ **Bohr magneton**:

$$\beta_e = \gamma \hbar = \frac{e \hbar}{2m} \approx 9.27 \cdot 10^{-24} \text{J} \cdot \text{T}^{-1} \quad (19)$$

- Similarly a **nuclear magneton** could be calculated for a proton (^1H nucleus):

$$\beta = \gamma_N \hbar = \frac{e \hbar}{2m_p} \approx 5.05 \cdot 10^{-27} \text{J} \cdot \text{T}^{-1} \quad (20)$$

1.5. Spin angular momentum operator

However, real nuclei and electrons have spins (intrinsic magnetic moment). Their z-axis projection m takes integer and half-integer values: $m = \frac{1}{2}, 1, \frac{3}{2}, 2$ etc. Similar to the equation for the orbital angular momentum Eq.18. For nuclei spins we get its magnetic moment as:

$$\hat{\mu}_{\mathbf{N}} = \gamma_N \hbar \hat{\mathbf{I}}, \quad (21)$$

where $\hat{\mathbf{I}}$ stands for the nuclear spin operator. All the properties of angular momentum operators listed in Eqs.11-17 will be true for $\hat{\mathbf{I}}$.

1.5. Spin angular momentum operator

- Many nuclei in the periodic table are magnetic, i.e. have spin $I \neq 0$. Their magnetic moments could be measured in units of β_N :

$$\hat{\mu}_N = \gamma_N \hbar \hat{\mathbf{I}} = g_N \beta_N \hat{\mathbf{I}}, \quad (22)$$

where g_N - dimensionless g-factor.

Nucleus	Natural abundance %	Nuclear spin (I)	g_N , g-factor	γ_N , Gyromagnetic ratio (10^7 rad/T*s)
^1H	99.98	$\frac{1}{2}$	5.585	26.7519
^2H	$1.5 \cdot 10^{-2}$	1	0.857	4.1066
^{13}C	1.108	$\frac{1}{2}$	1.405	6.7283
^{14}N	99.635	1	0.403	1.9338
^{15}N	0.365	$\frac{1}{2}$	-0.567	-2.712

- Electron magnetic moments can be measured in units of Bohr magnetons: $\hat{\mu}_S = -\gamma_e \hbar \hat{\mathbf{S}} = -g_e \beta_e \hat{\mathbf{S}}$, and for a free electron spin $g_e \approx 2.0023$.

Summary of Lecture 1

- Applications of NMR: chemistry, biology, medicine, industry ...
- Magnetic moment in magnetic field: Classical description
- Recap of angular momentum operator properties: commutation properties.
- Nuclei have their own nuclear magnetic moment. Described using spin angular momentum operator.

Suggested reading: Harris: 1.1, 1.2, 1.3, 1.4, 1.6, 2.4

1.6. Spin in a magnetic field

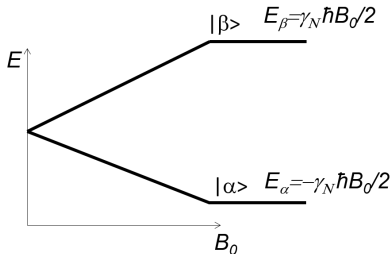
- Let's quantum mechanically describe the system of spins in the magnetic field. Eq. 5 can be rewritten in a form of Hamiltonian:

$$E = -\mathbf{M} \cdot \mathbf{B} \longleftrightarrow \mathcal{H} = -\hat{\boldsymbol{\mu}}_{\mathbf{N}} \cdot \mathbf{B} \quad (23)$$

- when the magnetic field is directed along z-axis $\mathbf{B} = (0, 0, B_0)$:

$$\mathcal{H} = -\hat{\boldsymbol{\mu}}_{\mathbf{N}} \cdot \mathbf{B} = -\gamma_N \hbar B_0 \hat{I}_z \quad (24)$$

- For spin $I = \frac{1}{2}$ such Hamiltonian produces a two-level system. Its energy levels corresponding to eigenfunctions $|\alpha\rangle$ and $|\beta\rangle$:



1.6. Spin in a magnetic field

- The transition between the two states requires an energy quantum¹:

$$h\nu_L = \gamma\hbar B_0, \omega_L = \gamma B_0, \nu_L = \frac{\gamma B_0}{2\pi} \quad (25)$$

ω_L and ν_L is the Larmor frequency (angular and cyclic respectively)

Nucleus	Natural abundance %	Nuclear spin (I)	Larmor frequency at 11.744T, MHz	γ_N , Gyro-magnetic ratio (10^7 rad/T*s)
^1H	99.98	$\frac{1}{2}$	500	26.7519
^2H	$1.5 \cdot 10^{-2}$	1	76.753	4.1066
^{13}C	1.108	$\frac{1}{2}$	125.721	6.7283
^{14}N	99.635	1	36.118	1.9338
^{15}N	0.365	$\frac{1}{2}$	50.664	-2.712

1.7. Equilibrium magnetization

NMR measurements are generally made on bulk samples which contain very large numbers of nuclear spins (e.g. 1 cm³ contains $N \approx 6.7 \cdot 10^{22}$ ¹H atoms) The measured signals therefore result from the collective effect of a large number of magnetic moments that can be described using a bulk magnetization. At thermal equilibrium, the numbers of nuclei in the $|\alpha\rangle$ state N_α and $|\beta\rangle$ state N_β follow Boltzmann distribution:

$$\frac{N_\alpha}{N_\beta} = e^{-\frac{\gamma B_0}{kT}} \approx (1 - \frac{\gamma B_0}{kT}), \quad (26)$$

when $\gamma B_0 \ll kT$. Overall magnetization then can be calculated as:

$$M_z = N_\alpha(-\frac{1}{2}\gamma\hbar) + N_\beta(\frac{1}{2}\gamma\hbar) = N\frac{\gamma^2\hbar^2 B_0}{4kT} \quad (27)$$

Problem

- What is the value of $\frac{\gamma_N \hbar B_0}{kT}$ for proton nuclei (^1H) at 9.4 T magnetic field at 300 K?
- What is the value of $\frac{\gamma_e \hbar B_0}{kT}$ electron (^1H) at 9.4 T magnetic field at 4 K?

$$\text{Electron charge } e = 1.602 \cdot 10^{-19} \text{ C}$$

$$\text{Electron mass } m_e = 9.109 \cdot 10^{-31} \text{ kg}$$

$$\text{Proton mass } m_p = 1.673 \cdot 10^{-27} \text{ kg}$$

$$\text{Plank constant } \hbar = 1.054 \cdot 10^{-34} \frac{\text{J} \cdot \text{s}}{\text{rad}}$$

$$\text{Proton g-factor } g_p = 5.585$$

$$\text{Electron g-factor } g_e = 2.0023$$

$$\text{Nuclear magneton } \beta_N = 5.05 \cdot 10^{-27} \text{ J} \cdot \text{T}^{-1}$$

$$\text{Bohr magneton } \beta_e = 9.27 \cdot 10^{-24} \text{ J} \cdot \text{T}^{-1}$$

$$\text{Proton gyromagnetic ratio } \gamma_N = 26.7519 \cdot 10^7 \frac{\text{rad}}{\text{T} \cdot \text{s}}$$

$$\text{Electron gyromagnetic ratio } \gamma_e = 1.76 \cdot 10^{11} \frac{\text{rad}}{\text{T} \cdot \text{s}}$$

$$\text{Boltzmann constant } k = 1.38 \cdot 10^{-23} \text{ J} \cdot \text{K}^{-1}$$

$$\text{Avogadro's constant } N_A = 6.023 \cdot 10^{23} \text{ mol}^{-1}$$

1.8. Resonant energy absorption.

Let's apply oscillating magnetic field to our system. A spin system Hamiltonian becomes time-dependent and for an oscillation along the x-axis we obtain:

$$\begin{aligned}\mathcal{H}(t) &= -\hat{\mu}(\mathbf{B}_0 + \mathbf{B}(t)) = \\ &= -\gamma\hbar\hat{I}_z(B_0 + B_1(t)) = \\ &= -\gamma\hbar\hat{I}_z B_0 - \gamma\hbar\hat{I}_x B_1 \cos(\omega t),\end{aligned}\tag{28}$$

where H_1 and ω are the amplitude and the frequency of the oscillating magnetic field.

According to perturbation theory the transition probability between the initial state $|a\rangle$ and the final state $|b\rangle$ with a time dependent Hamiltonian $\hat{V}(t) = 2\hat{F} \cos(\omega t)$ is (Fermi's golden rule):

$$P_{ab} = \frac{2\pi}{\hbar} |\langle a|\hat{F}|b\rangle|^2 \delta(E_{ab} - \hbar\omega),\tag{29}$$

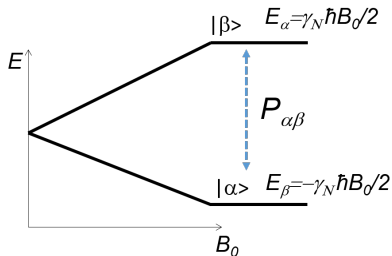
where $E_{ab} = E_a - E_b$ is an energy difference between the energies of levels a and b .

Note! The matrix element $|\langle a|\hat{F}|b\rangle|$ represents the selection rules for various levels, if it equals to zero, the transition is forbidden, otherwise - allowed.

1.9. Populations dynamics in two-level system.

- For a two level system described before, the matrix element $\langle\alpha|\hat{I}_x|\beta\rangle = \frac{1}{2}$. The transition probability then becomes:

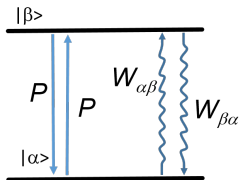
$$P_{\alpha\beta} = \frac{\pi}{2\hbar}(\gamma H_1)^2\delta(E_{\alpha\beta} - \hbar\omega), \quad (30)$$



- The effect of resonant absorption (and emission) of electromagnetic irradiation at the frequency matching the energy difference in a nuclear system is called Nuclear Magnetic Resonance (NMR).**

1.9. Populations dynamics in two-level system.

- In a two-level system the transition will take place due to the action of external irradiation, but also due to interaction with the environment.



P - the rate of transitions driven by external field, $W_{\alpha\beta}$, $W_{\beta\alpha}$ - rates of spontaneous spin flips due to interaction with environment.

- In thermal equilibrium:

$$N_{\alpha}^0 W_{\alpha\beta} = N_{\beta}^0 W_{\beta\alpha}, \text{ i.e.} \quad (31)$$

$$\frac{W_{\beta\alpha}}{W_{\alpha\beta}} = \exp\left(-\frac{\gamma\hbar B_0}{kT}\right) \approx 1 - \frac{\gamma\beta B_0}{kT} \quad (32)$$

1.9. Populations dynamics in two-level system.

- Equation for populations of levels:

$$\begin{aligned}\frac{dN_\alpha}{dt} &= -N_\alpha(P + W_{\alpha\beta}) + N_\beta(P + W_{\beta\alpha}) \\ \frac{dN_\beta}{dt} &= N_\alpha(P + W_{\alpha\beta}) - N_\beta(P + W_{\beta\alpha})\end{aligned}\tag{33}$$

- If we introduce the average rate of spontaneous transitions

$$W = \frac{1}{2}(W_{\alpha\beta} + W_{\beta\alpha}), \text{ then } W_{\alpha\beta} = W(1 + \frac{\gamma\hbar B_0}{2kT}) \text{ and}$$

$$W_{\beta\alpha} = W(1 - \frac{\gamma\hbar B_0}{2kT}), \text{ the equations can be rewritten:}$$

$$\begin{aligned}\frac{dN_\alpha}{dt} &= (N_\beta - N_\alpha)P + (N_\beta - N_\alpha)W - W\frac{\gamma\hbar B_0}{2kT}N \\ \frac{dN_\beta}{dt} &= -(N_\beta - N_\alpha)P - (N_\beta - N_\alpha)W + W\frac{\gamma\hbar B_0}{2kT}N\end{aligned}\tag{34}$$

1.9. Populations dynamics in two-level system.

- Denote the population difference as $n = N_\beta - N_\alpha$ and thermal equilibrium population difference $n_0 = N_\beta^0 - N_\alpha^0 \approx N \frac{\gamma \hbar B_0}{2kT}$ the equations can be rewritten as:

$$\frac{dn}{dt} = -2nP - 2nW + 2Wn_0, \quad (35)$$

or

$$\frac{dn}{dt} = -2nP - \frac{(n - n_0)}{T_1}, \quad (36)$$

where $T_1 = \frac{1}{2W}$ is called **spin-lattice relaxation time** determines how quickly a spin system reaches a thermal equilibrium with environment.

- In equilibrium, when $\frac{dn}{dt} = 0$:

$$n = \frac{n_0}{1 + 2PT_1} \quad (37)$$

when the power is very large $PT_1 \gg 1$, $n \rightarrow 0$, i.e. the system is **saturated** and no signal can be observed.

Summary of Lecture 2

- Spin $I = \frac{1}{2}$ in a magnetic field. Two-level system.
- System of spins in a magnetic field is capable of absorbing radiation at a resonant frequency.
- Population dynamics in a two-level system. Signal as function of radiation power and saturation.

Suggested reading: Harris 1.5, 1.7, Slichter 1.3

Harris 1.20 - CW NMR spectrometer

Chemical shifts and J-couplings. NMR spectra in liquids.

- Chemical shifts.
- J-couplings.
- J-couplings in AX system.
- Interpreting simple NMR spectra
- Range of J-couplings. Are they useful for structure determination?
- Heteronuclear J-couplings.
- J-coupling in equivalent system.
- J-coupling in AB system.

2.1. Chemical shifts.

- Electrons in atoms in molecules interact with external magnetic field and in turn produce their own magnetic field B_0 . The Larmor frequency get shifted in a chemical specific manner - this is known as the **chemical shift**. The spin Hamiltonian for a nucleus is:

$$\hat{H} = -\gamma_N \hbar (1 - \sigma) B_0 \hat{I}_z, \quad (38)$$

- Classical illustration of diamagnetic chemical shift:

$$\omega = \frac{e}{2m_e} B_0$$

$$\mathbf{j} = -e[\boldsymbol{\omega} \times \mathbf{r}] \rho_e = -\frac{e^2}{2m_e} [B_0 \times \mathbf{r}] \rho_e$$

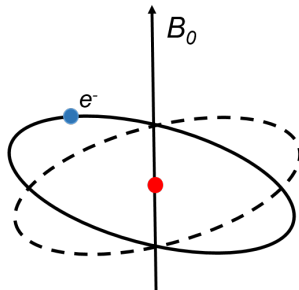
$$d\mathbf{B}_i = \frac{\mu_0}{4\pi r^3} [\mathbf{j} \times \mathbf{r}] dV$$

$$d\mathbf{B}_i = -\frac{\mu_0 e^2}{8\pi m_e r^3} [[B_0 \times \mathbf{r}] \times \mathbf{r}] \rho_e dV$$

$$B_{iz} = -B_0 \frac{\mu_0 e^2}{8\pi m_e} \int \rho_e \frac{x^2 + y^2}{r^3} dV$$

Quantum mechanical result:

$$\sigma = -\frac{\mu_0 e^2}{8\pi m_e} \langle \Psi | \frac{x^2 + y^2}{r^3} | \Psi \rangle$$



2.1. Chemical shifts.

Let's calculate the effect of ring current in cyclic aromatic molecules. Consider benzene molecule:

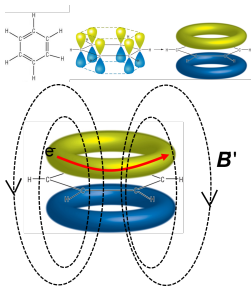


Fig.5: (top) Schematic representation of electron orbitals in a benzene molecule, (bottom) local fields in a benzene molecule produced by electron currents induced by a magnetic field.

- Larmor precession frequency of electrons $\omega_L = \frac{eB_0}{2m_e}$. Current can be calculated as charge $6e$, divided by precession period $\frac{2\pi}{\omega_L}$:

$$i = \frac{3e^2 B_0}{2\pi m_e} \quad (39)$$

- Circular conductor creates a magnetic moment $\mu = i \cdot \pi r^2$, totalling in:

$$\mu = - \frac{3e^2 B_0 r^2}{2m_e} \quad (40)$$

- The magnetic field created by a magnetic moment

$$\mathbf{B} = \frac{\mu_0}{4\pi} \left(\frac{\mathbf{r}(\mathbf{m} \cdot \mathbf{r})}{r^5} - \frac{\mathbf{m}}{r^3} \right) \text{ reduces to:}$$

$$B_i = - \frac{\mu_0}{4\pi} \frac{m}{r^3} = \frac{3\mu_0 e^2}{8\pi} \frac{r^2}{(r+d)^3} B_0 \quad (41)$$

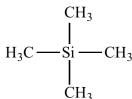
$$\sigma = - - \frac{\mu_0}{4\pi m_e} \frac{m}{r^3} = \frac{3\mu_0 e^2}{8\pi} \frac{r^2}{(r+d)^3} \quad (42)$$

- Given benzene molecule radius $r = 1.4 \text{ \AA}$, CH-bond length $d = 1.1 \text{ \AA}$ obtain:

$$\sigma \approx -5.3 \cdot 10^{-6}, \sigma_{iso} \approx -1.8 \cdot 10^{-6} \quad (43)$$

2.1. Chemical shifts.

- Chemical shifts are usually measured in ppm's and are referenced with respect to the signals of tetramethylsilane (TMS).



$$\delta = \frac{\nu_{\text{sample}} - \nu_{\text{TMS}}}{\nu_{\text{TMS}}} \times 10^6 \text{ ppm} \quad (44)$$

$$\delta = (\sigma_{\text{TMS}} - \sigma_{\text{sample}}) \times 10^6 \text{ ppm} \quad (45)$$

- In general the ^1H chemical shift is greater for nuclei to atoms/bonds that reduce the electron density at the atom.

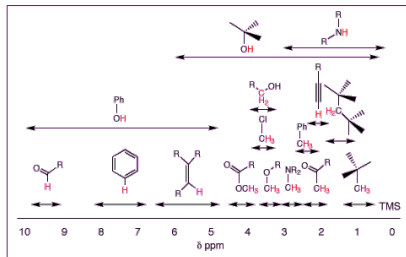


Fig.6: Chemical shifts in various organic molecules Source: <http://orgchem.colorado.edu/Spectroscopy/nmrtheory/protonchemshift.html>

2.1. Chemical shifts.

Consider the ^1H -spectrum of methyl acetate.

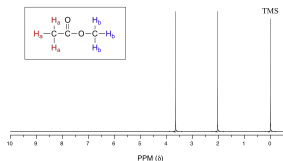


Fig.7: ^1H spectrum of methyl acetate. Source: http://chemwiki.ucdavis.edu/Organic_Chemistry

- the TMS appears at 0 ppm
- chemical shift increases from right to left
- two resonances correspond to protons of the two methyl groups
- the presence of electronegative oxygen atom in methoxy group produces weaker shielding (σ) thus makes bigger chemical shift δ
- three nuclei of methyl groups resonate at the same frequency. Peak heights are the same.

2.1. Chemical shifts.

For other nuclei, paramagnetic shifts which arise from mixing of excited state with the ground state due to the effect of the applied field, B_0 on the Hamiltonian can be important, and the range of chemical shifts is usually larger than for ^1H nuclei.

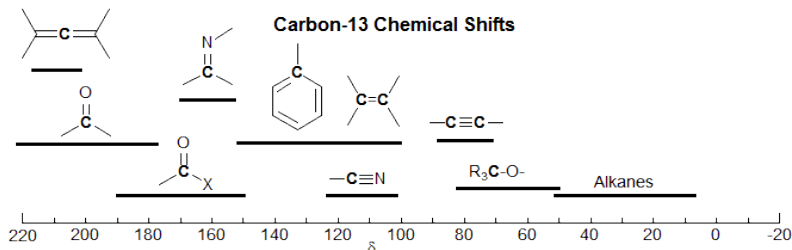


Fig.8: ^{13}C chemical shifts, <https://www.chem.wisc.edu/areas/reich/nmr/>

2.2. J-couplings.

Nuclear spins interact with one another. In solution, one prominent spin-spin interaction is called **J-coupling**.

This intra-molecular scalar coupling is caused by the combination of two effects: the Pauli principle means that the electrons in the bond have opposite spin-state (spin-up and spin-down), while hyperfine couplings (specifically Fermi contact interaction) mean that it is energetically favourable for each nuclear spin to be anti-parallel to the electron spin.

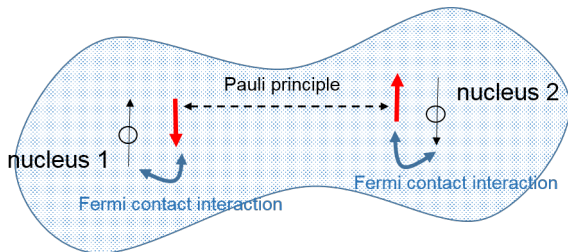


Fig.9: Origin of J-couplings. Low energy configuration in which nuclear spins are antiparallel

2.2. J-couplings.

- Hamiltonian of J-coupling in solution (all anisotropy is averaged):

$$\begin{aligned}\hat{H} &= J\hat{\mathbf{I}}_1 \cdot \hat{\mathbf{I}}_2 = J(\hat{I}_{1x}\hat{I}_{2x} + \hat{I}_{1y}\hat{I}_{2y} + \hat{I}_{1z}\hat{I}_{2z}) = \\ &= J\hat{I}_{1z}\hat{I}_{2z} + \frac{J}{2}(\hat{I}_{1+}\hat{I}_{2-} + \hat{I}_{1-}\hat{I}_{2+})\end{aligned}\tag{46}$$

J is usually measured in units of frequency, i.e. Hz.

- In frequency units the full spin Hamiltonian for a system of two nuclei A and B of the same kind would be:

$$\hat{H} = -\nu_0(1-\sigma_A)\hat{I}_{Az} - \nu_0(1-\sigma_B)\hat{I}_{Bz} + J\hat{I}_{Az}\hat{I}_{Bz} + \frac{J}{2}(\hat{I}_{A+}\hat{I}_{B-} + \hat{I}_{A-}\hat{I}_{B+})\tag{47}$$

2.2. J-couplings.

- For two coupled spins $\frac{1}{2}$ the Hamiltonian matrix in the basis of functions $|1\rangle = |\alpha_1\alpha_2\rangle, |2\rangle = |\alpha_1\beta_2\rangle, |3\rangle = |\beta_1\alpha_2\rangle, |4\rangle = |\beta_1\beta_2\rangle$ The Hamiltonian matrix of such a system are:

$$H_{ik} = \langle i|\hat{H}|k\rangle = \begin{bmatrix} -\frac{\nu_A + \nu_B}{2} + \frac{J}{4} & 0 & 0 & 0 \\ 0 & -\frac{\nu_A - \nu_B}{2} - \frac{J}{4} & \frac{J}{2} & 0 \\ 0 & \frac{J}{2} & \frac{\nu_A - \nu_B}{2} - \frac{J}{4} & 0 \\ 0 & 0 & 0 & -\frac{\nu_A + \nu_B}{2} + \frac{J}{4} \end{bmatrix} \quad (48)$$

when $|\nu_A - \nu_B| \gg J$ the off-diagonal terms due to $\hat{I}_{A\pm}\hat{I}_{B\mp}$ can be neglected. That is a called **AX system**. When off-diagonal terms cannot be neglected, we deal with a so called **AB system**.

- Let's consider nuclei of the same type (i.e. both ^1H or both ^{13}C). When $\gamma_N\hbar |\omega_1 - \omega_2| \ll J$,

2.3. J-couplings in AX system.

Eq.29 is non-zero when corresponding matrix element is not zero.
The selection rules for two nuclei then : $\langle i | \hat{I}_{1x} + \hat{I}_{2x} | k \rangle \neq 0$.

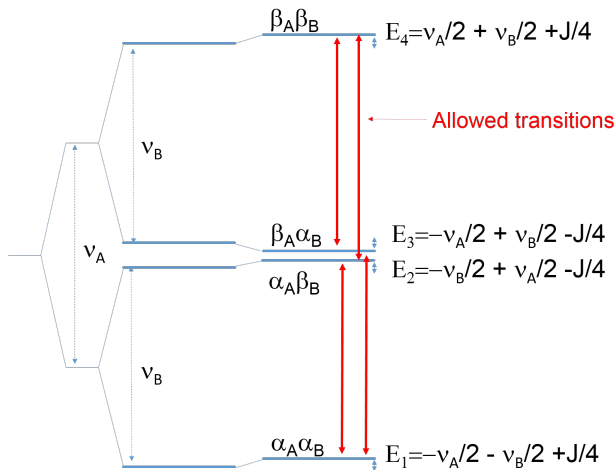


Fig.10: Level diagram for two J-coupled spins in AX system

2.3. J-couplings in AX system.

In a more general case, when $I_A, I_B \neq \frac{1}{2}$, the energy levels follow the following equation:

$$E = -\nu_A m_1 - \nu_0(1 - \sigma_B)m_2 + Jm_1m_2 \quad (49)$$

The allowed transitions for spin A have the following frequencies:

$$\nu = \nu_A + m_B J, \quad (50)$$

where $m_B = -I_B, -(I_B - 1) \dots (I_B - 1), I_B$ is the projection of nuclear spin B . The resonance line is therefore being split into several components. Similarly, for spin B :

$$\nu = \nu_B + m_A J, \quad (51)$$

2.4. Interpreting simple NMR spectra

For one coupled nucleus:

$$\nu = \nu_A + m_B J, \quad (52)$$

For two coupled nuclei:

$$\nu = \nu_A + m_B J + m_C J, \quad (53)$$

For three:

$$\nu = \nu_A + m_B J + m_C J + m_D J, \quad (54)$$

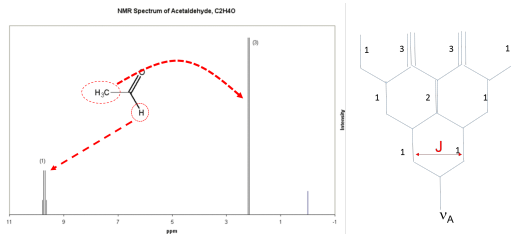
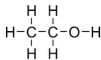


Fig.11: Spectrum of acetaldehyde. Source: <https://chem242.wikispaces.com>

2.4. Interpreting simple NMR spectra

Problem

Draw schematically ^1H spectrum of ethyl alcohol? Proton Larmor frequency is 90 MHz (i.e. treat as AX system).



$\delta_{\text{CH}_3} = 1.22 \text{ ppm}$, $\delta_{\text{CH}_2} = 3.68 \text{ ppm}$, $\delta_{\text{OH}} = 2.61 \text{ ppm}$,
consider only $J_{\text{CH}_2-\text{CH}_3} = 7.29 \text{ Hz}$.

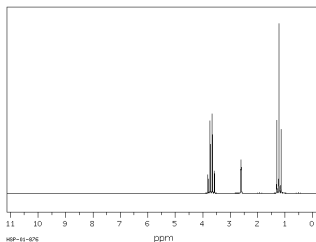


Fig.12: Ethanol spectrum at 90 MHz in CDCl_3

2.5. Range of J-couplings. Are they useful for structure determination?

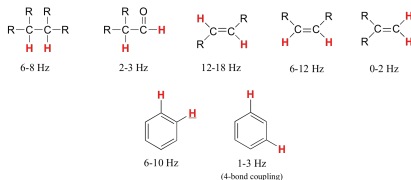


Fig.13: Typical J-couplings https://chem.libretexts.org/Textbook_Maps/Organic_Chemistry_Textbook_Maps/

Is there any simple meaning to J-couplings? **Karplus equation** is an empiric formula for J coupling as a function of a dihedral angle ϕ .

$$J = A \cos \phi + B \cos 2\phi + C \quad (55)$$

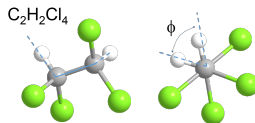


Fig.14: Dihedral angle in 1,1,2,2-tetrachlorethane

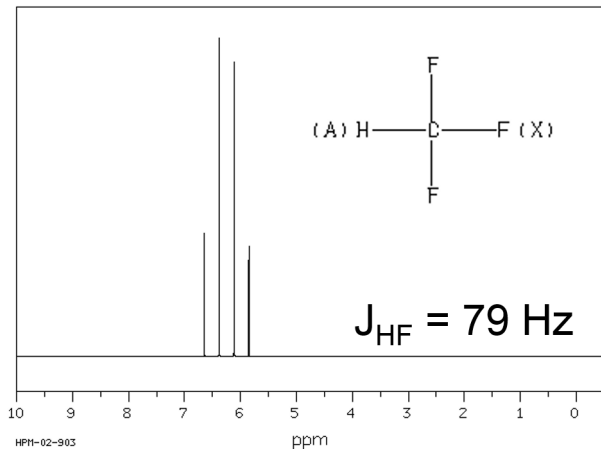
2.6. Heteronuclear J-couplings.

Heteronuclear systems are automatically AX systems, i.e.

$$|\nu_A - \nu_B| \gg J.$$

Example: trifluoromethane (fluoroform)

^1H NMR spectrum



Summary of lecture 3.

- J-couplings and chemical shift to an extent are unique in different molecules, and NMR spectra help in identifying chemical compounds by their spectra.
- Chemical shifts originate from magnetic field shielding by electrons.
- CS report on the chemical environment of a nucleus
- J-couplings cause spectral splittings.

Suggested reading: Harris 1.8, 1.9, 1.11, 1.12, 1.13, 8.2, 8.9., 8.23

For more detailed quantum mechanical theory of chemical shifts: Slichter 4.1-4.5, for detailed theory of J-couplings: Slichter 4.9

2.7. J-coupling in equivalent system.

Are protons in CH₃ or CH₂ groups coupled to one another?

Yes. But they are equivalent and therefore not observed.



$$\hat{H} = -\nu_0(\hat{I}_{1z} + \hat{I}_{2z} + \hat{I}_{3z}) + J(\mathbf{I}_1\mathbf{I}_2 + \mathbf{I}_2\mathbf{I}_3 + \mathbf{I}_1\mathbf{I}_3) \quad (56)$$

- Let's rewrite the last term as:

$$\mathbf{I}_1\mathbf{I}_2 + \mathbf{I}_2\mathbf{I}_3 + \mathbf{I}_1\mathbf{I}_3 = \frac{1}{2}(\mathbf{I}_1 + \mathbf{I}_2 + \mathbf{I}_3)^2 - \frac{1}{2}(\mathbf{I}_1^2 + \mathbf{I}_2^2 + \mathbf{I}_3^2) \quad (57)$$

- We can introduce new operator: $\mathbf{F} = \mathbf{I}_1 + \mathbf{I}_2 + \mathbf{I}_3$. Given that $I_1 = I_2 = I_3 = \frac{1}{2}$ the Hamiltonian can be rewritten using this new operator:

$$\hat{H} = -\nu_0\hat{F}_z + J(\mathbf{F}^2 - \frac{9}{4}) \quad (58)$$

- Eigenfunctions of this Hamiltonian are functions $|FM_F\rangle$, where $F = \frac{1}{2}$ or $\frac{3}{2}$.

2.7. J-coupling in equivalent system.



$$\hat{H} = -\nu_0 \hat{F}_z + J(\mathbf{F}^2 - \frac{9}{4}) \quad (59)$$

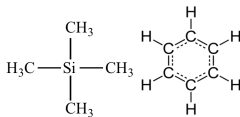
Selection rules for transitions: $\langle i | \hat{F}_x | k \rangle \neq 0$, since

$\hat{F}_x = \frac{\hat{F}_+ + \hat{F}_-}{2}$, the selection rules then are: $\langle i | \hat{F}_\pm | k \rangle \neq 0$

Since $\hat{F}_\pm = \text{const} |FM_F \pm 1\rangle$, the allowed transition does not change the second term in Eq.59.

- Strictly speaking if a molecule contains only nuclei of one type, and no other J -splitting will not be observed. Examples:

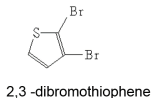
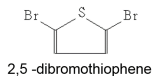
tetramethylsilane (TMS), benzene



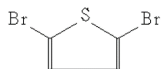
2.7. J-coupling in equivalent system.

Problem

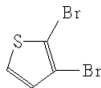
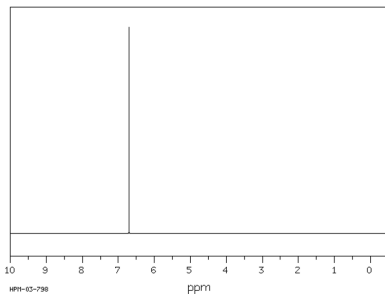
Sketch the ^1H NMR spectra of 2,3-dibromothiophene and 2,5-dibromothiophene? ^{79}Br and ^{81}Br are quadrupolar nuclei, and their J-couplings can be ignored.



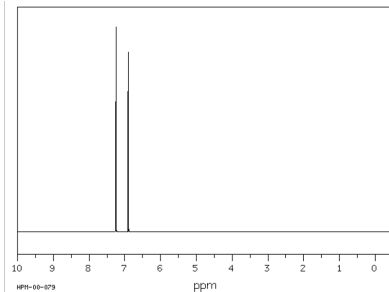
2.7. J-coupling in equivalent system.



2,5 -dibromothiophene



2,3 -dibromothiophene



2.8. J-coupling in AB system.

What happens if $\nu_A - \nu_B \gg J$ is no longer true?

- The matrix of spin Hamiltonian is:

$$\begin{bmatrix} -\frac{\nu_A + \nu_B}{2} + \frac{J}{4} & 0 & 0 & 0 \\ 0 & -\frac{\nu_A - \nu_B}{2} - \frac{J}{4} & \frac{J}{2} & 0 \\ 0 & \frac{J}{2} & \frac{\nu_A - \nu_B}{2} - \frac{J}{4} & 0 \\ 0 & 0 & 0 & -\frac{\nu_A + \nu_B}{2} + \frac{J}{4} \end{bmatrix} \quad (60)$$

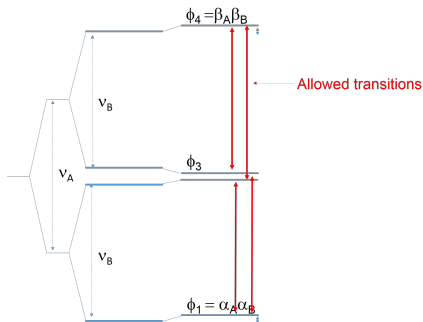
- The general way of solving is finding the eigenvalues and eigenvalues and eigenfunctions of this Hamiltonian.
- the solution can be represented by following replacements:

$$\begin{aligned} \frac{\nu_A - \nu_B}{2} &= \frac{\delta}{2} = C \cos 2\theta \\ \frac{J}{2} &= C \sin 2\theta \\ C &= \frac{1}{2} \sqrt{\delta^2 + J^2}, \tan 2\theta = \frac{J}{\delta}, \bar{\nu} = \frac{\nu_A + \nu_B}{2} \end{aligned} \quad (61)$$

2.8. J-coupling in AB system.

In this notation the eigenvalues and eigenfunctions are:

$$\begin{aligned}
 E_1 &= -\bar{\nu} + \frac{J}{4} & \phi_1 &= |\alpha_A \alpha_B\rangle \\
 E_2 &= -\frac{J}{4} - C & \phi_2 &= \sin \theta |\alpha_A \beta_B\rangle - \cos \theta |\beta_A \alpha_2\rangle \\
 E_3 &= -\frac{J}{4} + C & \phi_3 &= \cos \theta |\alpha_A \beta_B\rangle + \sin \theta |\beta_A \alpha_2\rangle \\
 E_4 &= \bar{\nu} + \frac{J}{4} & \phi_4 &= |\beta_A \beta_B\rangle
 \end{aligned} \tag{62}$$



Problem

Calculate the transition probability between levels 1 and 2? (Hint: use Eq.29)

Fig.15: Level diagram for AB system of J-coupled nuclei.

2.8. J-coupling in AB system.

Problem

Calculate the transition probability between levels 1 and 2? (Hint: use Eq.29)

For transition $1 \leftrightarrow 2$ we have:

$$\begin{aligned} I_{1 \leftrightarrow 2} &\sim |\langle 1 | \hat{I}_{Ax} + \hat{I}_{Bx} | 2 \rangle|^2 = \\ &= \langle 1 | \frac{\hat{I}_{A+} + \hat{I}_{A-}}{2} + \frac{\hat{I}_{B+} + \hat{I}_{B-}}{2} | 2 \rangle^2 = \\ &= \left(\langle \alpha_A \alpha_B | \frac{\hat{I}_{A+} + \hat{I}_{A-}}{2} + \frac{\hat{I}_{B+} + \hat{I}_{B-}}{2} | \sin \theta | \alpha_A \beta_B \rangle + \cos \theta | \beta_A \alpha_B \rangle \right)^2 = \\ &= \left(\frac{\cos \theta + \sin \theta}{2} \right)^2 = \frac{1 + \sin 2\theta}{4} \end{aligned} \tag{63}$$

2.8. J-coupling in AB system.

For all the transitions we obtain:

Transition	Frequency	Intensity:
$4 \leftrightarrow 2$	$\bar{\nu} + C + \frac{J}{2}$	$1 - \sin 2\theta$
$3 \leftrightarrow 1$	$\bar{\nu} + C - \frac{J}{2}$	$1 + \sin 2\theta$
$4 \leftrightarrow 3$	$\bar{\nu} - C + \frac{J}{2}$	$1 + \sin 2\theta$
$2 \leftrightarrow 1$	$\bar{\nu} - C - \frac{J}{2}$	$1 - \sin 2\theta$

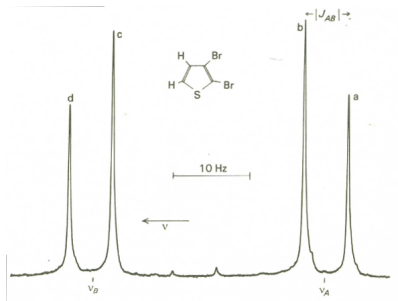


Fig.16: 100 MHz ^1H NMR spectrum of 2,3-dibromothiophene, $|\nu_A - \nu_B| = 30.5$ Hz, $J = 5.7$ Hz

2.8. J-coupling in AB system.

For all the transitions we obtain:

Transition	Frequency	Intensity:
$4 \leftrightarrow 2$	$\bar{\nu} + C + \frac{J}{2}$	$\frac{1 - \sin 2\theta}{4}$
$3 \leftrightarrow 1$	$\bar{\nu} + C - \frac{J}{2}$	$\frac{1 + \sin 2\theta}{4}$
$4 \leftrightarrow 3$	$\bar{\nu} - C + \frac{J}{2}$	$\frac{1 + \sin 2\theta}{4}$
$2 \leftrightarrow 1$	$\bar{\nu} - C - \frac{J}{2}$	$\frac{1 - \sin 2\theta}{4}$

Consider following scenarios:

1. $J \ll \delta$ means that $\sin 2\theta \ll 1$ and $C \approx \frac{\delta}{2}$. The leads to AX system shown Fig.51B
2. $\delta \ll J$ means that $\sin 2\theta \approx 1$. A strong double and two weak satellite lines in Fig.51C.

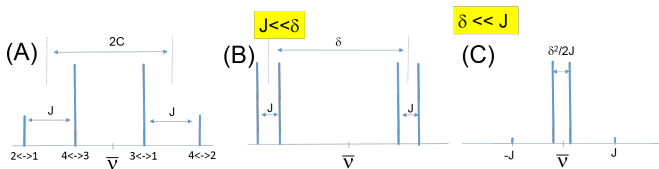


Fig.17: Schematic spectra in several cases (A) $|\nu_A - \nu_B| \sim J$, (B) weakly coupled nuclei $|\nu_A - \nu_B| \gg J$, (C) very weak J-coupling $|\nu_A - \nu_B| \ll J$

2.8. J-coupling in AB system.

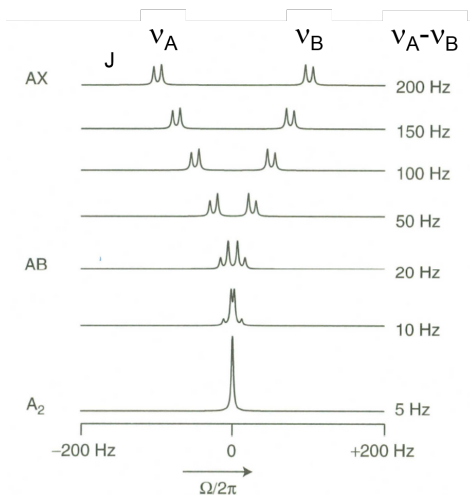


Fig.18: Spectra of spin $\frac{1}{2}$ pairs for $J = 10$ Hz as a function of Larmor frequency difference.

Summary of Lecture 4.

- Splittings in equivalent systems are not observed.
- The case of strong J-couplings can be treated exactly.

Suggested reading: Harris 1.14, 2.17, 2.10

NMR spectra of solids.

- Nuclear magnetic moments of nuclei interact via dipolar interaction. (It is averaged to zero in liquids). Classical expression for such interaction is:

$$E = \frac{\mu_0}{4\pi} \left(\frac{3(\mathbf{m}_1 \mathbf{r})(\mathbf{m}_2 \mathbf{r})}{r^5} - \frac{\mathbf{m}_1 \mathbf{m}_2}{r^3} \right) \quad (64)$$

- In quantum mechanics the magnetic moments $\mathbf{m}_1, \mathbf{m}_2$ are replaced with operators $g_{1,2}\beta_{1,2}\hat{l}_{1,2} = \gamma_{1,2}\hbar\hat{l}_{1,2}$, giving rise to a spin Hamiltonian:

$$\hat{H} = \frac{\mu_0}{4\pi} \frac{\gamma_1 \gamma_2 \hbar^2}{r^3} \sum_{i,j=x,y,z} (3\hat{l}_{1i}\hat{l}_{2j} - \hat{l}_1\hat{l}_2) \quad (65)$$

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