

1. What is the definition of equilibrium?

- A. A state of balance between opposing forces
- B. A state of change in which no net change occurs
- C. A state of maximum disorder
- D. A state of minimum energy

2. What is the definition of a chemical reaction?

- A. A process that results in the formation of new products
- B. A process that results in the rearrangement of atoms
- C. A process that results in the release of energy
- D. A process that results in the consumption of energy

3. What is the definition of a reactant?

- A. A substance that is consumed in a chemical reaction
- B. A substance that is produced in a chemical reaction
- C. A substance that is necessary for a chemical reaction to occur
- D. A substance that is not involved in a chemical reaction

4. What is the definition of a product?

- A. A substance that is consumed in a chemical reaction
- B. A substance that is produced in a chemical reaction
- C. A substance that is necessary for a chemical reaction to occur
- D. A substance that is not involved in a chemical reaction

5. What is the definition of a catalyst?

- A. A substance that increases the rate of a chemical reaction
- B. A substance that decreases the rate of a chemical reaction
- C. A substance that is necessary for a chemical reaction to occur
- D. A substance that is not involved in a chemical reaction

6. What is the definition of activation energy?

- A. The minimum amount of energy required for a chemical reaction to occur
- B. The amount of energy released in a chemical reaction
- C. The amount of energy required to break the bonds of a reactant

D. The amount of energy required to form the bonds of a product

7. What is the definition of a reaction rate?

A. The speed at which a chemical reaction occurs

B. The amount of product formed in a chemical reaction

C. The amount of reactant consumed in a chemical reaction

D. The amount of energy released in a chemical reaction

8. What is the definition of a rate law?

A. An equation that describes the relationship between the rate of a chemical reaction and the concentrations of the reactants

B. An equation that describes the relationship between the rate of a chemical reaction and the temperature

C. An equation that describes the relationship between the rate of a chemical reaction and the pressure

D. An equation that describes the relationship between the rate of a chemical reaction and the catalyst

9. What is the definition of an equilibrium constant?

A. A value that describes the position of equilibrium for a chemical reaction

B. A value that describes the rate of a chemical reaction

C. A value that describes the amount of product formed in a chemical reaction

D. A value that describes the amount of reactant consumed in a chemical reaction

10. What is the definition of Le Chatelier's principle?

A. The principle that states that a change in temperature will shift the position of equilibrium to the side of the reaction that absorbs heat

B. The principle that states that a change in concentration will shift the position of equilibrium to the side of the reaction that consumes reactant

C. The principle that states that a change in pressure will shift the position of equilibrium to the side of the reaction that has the smaller number of moles of gas

D. The principle that states that a change in catalyst will shift the position of equilibrium to the side of the reaction that has the higher activation energy