

1. What is thermochemistry?

- A. The study of the relationship between heat and chemical reactions
- B. The study of the relationship between heat and physical changes
- C. The study of the relationship between heat and energy
- D. The study of the relationship between heat and temperature

2. What is the most important factor that determines the amount of heat released or absorbed in a chemical reaction?

- A. The nature of the reactants
- B. The amount of reactants
- C. The rate of the reaction
- D. The temperature of the reaction

3. What is the unit for heat?

- A. Joules
- B. Calories
- C. Kilojoules
- D. Kilocalories

4. What is the first law of thermodynamics?

- A. Energy can be neither created nor destroyed, but it can be converted from one form to another
- B. The universe is constantly expanding
- C. The entropy of the universe is always increasing
- D. The speed of light is constant

5. What is the second law of thermodynamics?

- A. Energy can be neither created nor destroyed, but it can be converted from one form to another
- B. The universe is constantly expanding
- C. The entropy of the universe is always increasing
- D. The speed of light is constant

6. What is the standard enthalpy of formation of a substance?

- A. The heat released when one mole of a substance is formed from its elements in their standard states
- B. The heat released when one mole of a substance is formed from its elements in their standard states at 25 degrees Celsius
- C. The heat released when one mole of a substance is formed from its elements in their standard states at 1 atmosphere pressure
- D. The heat released when one mole of a substance is formed from its elements in their standard states at 1 mole per liter concentration

7. What is the standard enthalpy of combustion of a substance?

- A. The heat released when one mole of a substance is burned in oxygen at 25 degrees Celsius
- B. The heat released when one mole of a substance is burned in oxygen at 1 atmosphere pressure
- C. The heat released when one mole of a substance is burned in oxygen at 1 mole per liter concentration
- D. The heat released when one mole of a substance is burned in oxygen

8. What is the standard enthalpy of reaction?

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- C. The heat released or absorbed when one mole of a substance reacts at 1 atmosphere pressure
- D. The heat released or absorbed when one mole of a substance reacts at 1 mole per liter concentration

9. What is the heat of fusion of a substance?

- A. The heat released when one mole of a substance melts
- B. The heat released when one mole of a substance melts at 25 degrees Celsius
- C. The heat released when one mole of a substance melts at 1 atmosphere pressure
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10. What is the heat of vaporization of a substance?

- A. The heat released when one mole of a substance vaporizes
- B. The heat released when one mole of a substance vaporizes at 25 degrees Celsius
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11. What is the heat of condensation of a substance?

- A. The heat released when one mole of a substance condenses
- B. The heat released when one mole of a substance condenses at 25 degrees Celsius
- C. The heat released when one mole of a substance condenses at 1 atmosphere pressure
- D. The heat released when one mole of a substance condenses at 1 mole per liter concentration

12. What is the heat of solution of a substance?

- A. The heat released when one mole of a substance dissolves
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13. What is the heat of neutralization of a substance?

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14. What is the heat of precipitation of a substance?

- A. The heat released when one mole of a substance precipitates
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15. What is the heat of sublimation of a substance?

- A. The heat released when one mole of a substance sublime
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