1. What is the name for a solution that is neither acidic nor basic?

A. Acidic
B. Basic
C. Neutral

2. What is the name for a solution with a pH greater than 7?

A. Acidic
B. Basic
C. Neutral

3. What is the name for a solution with a pH less than 7?

A. Acidic
B. Basic
C. Neutral

4. What is the definition of a buffer?

A. A solution that resists changes in pH when small amounts of acid or base are added.
B. A solution that is neither acidic nor basic.
C. A solution that is used to calibrate pH meters.

5. What is the definition of a titration?

A. A method of measuring the pH of a solution.
B. A method of measuring the concentration of a solution.
C. A method of measuring the amount of acid or base in a solution.

6. What is the definition of an indicator?

A. A substance that changes color in the presence of an acid or base.
B. A substance that is used to calibrate pH meters.
C. A substance that is used to measure the pH of a solution.

7. What is the definition of a standard solution?

A. A solution with a known concentration of acid or base.
B. A solution with a known pH.
C. A solution that is used to calibrate pH meters.

8. What is the definition of a molarity?

A. The number of moles of solute per liter of solution.
B. The number of moles of solvent per liter of solution.
C. The number of moles of acid or base per liter of solution.

9. What is the definition of a normality?

A. The number of moles of solute per liter of solution.
B. The number of moles of solvent per liter of solution.
C. The number of moles of acid or base per liter of solution.

10. What is the definition of a molality?

A. The number of moles of solute per kilogram of solvent.

B. The number of moles of solvent per kilogram of solution.
C. The number of moles of acid or base per kilogram of solution.

11. What is the definition of a percent by mass?

A. The number of grams of solute per 100 grams of solution.
B. The number of grams of solvent per 100 grams of solution.
C. The number of grams of acid or base per 100 grams of solution.

12. What is the definition of a percent by volume?

A. The number of milliliters of solute per 100 milliliters of solution.
B. The number of milliliters of solvent per 100 milliliters of solution.
C. The number of milliliters of acid or base per 100 milliliters of solution.

13. What is the definition of a percent by weight?

A. The number of grams of solute per 100 grams of solution.
B. The number of grams of solvent per 100 grams of solution.
C. The number of grams of acid or base per 100 grams of solution.

14. What is the definition of a mole fraction?

A. The number of moles of solute per mole of solution.
B. The number of moles of solvent per mole of solution.
C. The number of moles of acid or base per mole of solution.

15. What is the definition of an equilibrium constant?

A. A value that represents the relative concentrations of reactants and products at equilibrium.
B. A value that represents the amount of acid or base in a solution.
C. A value that represents the pH of a solution.

16. What is the definition of an acid?

A. A substance that increases the concentration of H+ ions in a solution.
B. A substance that decreases the concentration of H+ ions in a solution.
C. A substance that reacts with a base to form a salt.

17. What is the definition of a base?

A. A substance that increases the concentration of H+ ions in a solution.
B. A substance that decreases the concentration of H+ ions in a solution.
C. A substance that reacts with an acid to form a salt.

18. What is the definition of a salt?

A. A substance that is formed when an acid and a base react.
B. A substance that is formed when a metal and an acid react.
C. A substance that is formed when a metal and a base react.

19. What is the definition of an electrolyte?

A. A substance that conducts electricity in a solution.
B. A substance that does not conduct electricity in a solution.
C. A substance that is used to make a solution conduct electricity.

20. What is the definition of a non-electrolyte?

A. A substance that conducts electricity in a solution.
B. A substance that does not conduct electricity in a solution.
C. A substance that is used to make a solution conduct electricity.

21. What is the definition of an acid-base reaction?

A. A reaction in which an acid and a base react to form a salt and water.
B. A reaction in which an acid and a base react to form an acid or a base.
C. A reaction in which an acid or a base is formed.

22. What is the definition of an oxidation-reduction reaction?

A. A reaction in which one substance gains oxygen and another substance loses oxygen.
B. A reaction in which one substance gains electrons and another substance loses electrons.
C. A reaction in which one substance loses electrons and another substance gains electrons.

23. What is the definition of a precipitation reaction?

A. A reaction in which two solutions are mixed and a solid is formed.
B. A reaction in which an acid and a base react to form a salt and water.
C. A reaction in which an acid and a base react to form an acid or a base.

24. What is the definition of a double replacement reaction?

A. A reaction in which two solutions are mixed and a solid is formed.
B. A reaction in which one substance gains electrons and another substance loses electrons.
C. A reaction in which one substance loses electrons and another substance gains electrons.

25. What is the definition of a decomposition reaction?

A. A reaction in which a substance is broken down into two or more simpler substances.
B. A reaction in which a substance is formed from two or more simpler substances.
C. A reaction in which an acid and a base react to form a salt and water.

1. C. Neutral
2. B. Basic
3. A. Acidic
4. A. A solution that resists changes in pH when small amounts of acid or base are added.
5. C. A method of measuring the amount of acid or base in a solution.
6. A. A substance that changes color in the presence of an acid or base.
7. A. A solution with a known concentration of acid or base.
8. A. The number of moles of solute per liter of solution.
9. C. The number of moles of acid or base per liter of solution.
10. A. The number of moles of solute per kilogram of solvent.
11. A. The number of grams of solute per 100 grams of solution.
12. A. The number of milliliters of solute per 100 milliliters of solution.
13. A. The number of grams of solute per 100 grams of solution.
14. A. The number of moles of solute per mole of solution.
15. A. A value that represents the relative concentrations of reactants and products at equilibrium.

16. A. A substance that increases the concentration of H+ ions in a solution.
17. B. A substance that decreases the concentration of H+ ions in a solution.
18. A. A substance that is formed when an acid and a base react.
19. A. A substance that conducts electricity in a solution.
20. B. A substance that does not conduct electricity in a solution.
21. A. A reaction in which an acid and a base react to form a salt and water.
22. B. A reaction in which one substance gains electrons and another substance loses electrons.
23. A. A reaction in which two solutions are mixed and a solid is formed.
24. B. A reaction in which one substance gains electrons and another substance loses electrons.
25. A. A reaction in which a substance is broken down into two or more simpler substances.