

**BIRLA INSTITUTE OF TECHNOLOGY AND SCIENCE, PILANI-HYDERABAD CAMPUS**  
**FIRST SEMESTER 2019-2020**  
**Course Handout (Part II)**

01/08/2019

In addition to part I (General Handout for all courses appended to the time table) this portion gives further specific details regarding the course.

**Course No :** CHE F211

**Course Title:** Chemical Process Calculations

**Instructor-in-charge:** DR I SREEDHAR

**1. Objective and Learning Outcomes:**

Chemical Process Calculations is one of the most fundamental courses of Chemical Engineering and serves as an introduction to the field. The course aims at inculcating systematic problem solving skills in students.

At the end of the course, the student should be able to:

- Solve material balance problems with or without chemical reaction for single and multiple units
- Using the concepts of recycle, bypass and purge in material balances
- Solve energy balance problems with and without chemical reactions using thermodynamic tables, psychrometric charts and basic thermodynamic principles
- Solve material and energy balance problems simultaneously

**2. Text Book:**

Himmelblau, D. M. "Basic principles & calculations in chemical Engg", PHI, 7<sup>th</sup> ed., 1997.

**3. Reference Book:**

Felder, R. M. & R. W. Rousseau, "Elementary Principles of Chemical Processes", John Wiley & Sons, Inc., 3<sup>rd</sup> ed., 2000.

**4. Course Plan:**

Lecture No.	Learning Objectives	Topics to be covered	Chapter in the Text Book
1	Introduction to Chemical Engineering	-	-
2	Units, Dimensions and their conversion, Dimensional consistency in equations	Units and Dimensions	1.1-1.4
3-4	Mole, density and concentration	Density, specific gravity, mole fraction and mass fraction, concentration	Ch 2
5	Choosing a basis	How to choose a basis	Ch 3
6	Different units of temperature, pressure	Temperature and Pressure	Ch 4, 5
7-10	Introduction to material balances	Concept of material balance, open and closed systems, steady state/unsteady state	Ch 6, 7
11-12	Solving material balance problems without chemical	Material balances without chemical reactions	Ch 8

	reactions		
13-14	Stoichiometry	Chemical reaction equations and stoichiometry	Ch 9
15-17	Solving material balances with chemical reactions	Material balances with chemical reactions	Ch 10
18-19	Material balances involving multiple subsystems	Solving material balance problems involving multiple subsystems	Ch 11
20-23	Material balances with recycle and bypass	Recycle bypass and purge calculations	Ch 12
24-25	Calculation procedures for ideal gas systems	Ideal gas law calculations, ideal gas mixtures and partial pressure	Ch 13
26-27	Two phase systems	Vapor pressure, saturation, phase diagrams	16.1, 17.1, 17.3, 19.1
28-30	General energy balances for open and closed systems without chemical reactions	The General energy balance	Ch 21, 22
31-32	Calculation of enthalpy changes using heat capacity equations and enthalpy tables	Calculation of enthalpy changes	23.1-23.3, Ch 24
33-35	Enthalpy balances with chemical reaction	Energy balances that account for chemical reactions	Ch 25
36-37	Calculation of heat of mixing and heat of solution	Heats of solution and mixing	Ch 28
38-39	Humidity chart and its use in determining the properties of moist air	Humidity charts and their use	Ch 29
40-42	Solving simultaneous material and energy balances	Miscellaneous Examples	-

#### 5. Evaluation Scheme:

Component	Duration	Weightage (%)	Date & Time	Nature of component
Midsem Test	1 hr	35	30/9, 11.00 -- 12.30 PM	CB
Surprise Quiz (min 2)		10		CB
Tutorial tests (min 2)		10		OB
Comprehensive Exam.	3 hrs	45	4/12, AN	CB (15)+ OB (30)

- 6. Chamber Consultation Hour:** Will be announced in the class (Chamber: D-206)
- 7. Notices:** Notices concerning the course will be displayed on the CMS website
- 8. Make-up Policy:** Make-up will be granted only for genuine cases with valid justification and only with prior permission of Instructor-in-charge.
- 9. Academic Honesty and Integrity Policy:** Academic honesty and integrity are to be maintained by all the students throughout the semester and no type of academic dishonesty is acceptable.

**Instructor-in-charge**  
**CHE F211**