CY1018: Environmental Chemistry Theory

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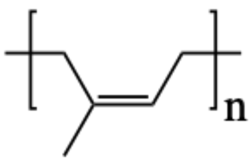
Microbial degradation of organics

Environmental degradation of polymers

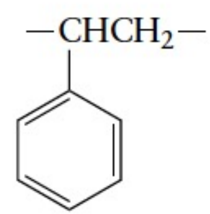
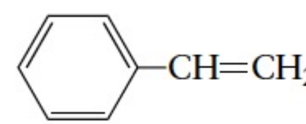
Proteins: polymers of amino acids

Starch and Cellulose: polymers of sugars

Natural rubber is not easily biodegradable



***cis*-1,4- polyisoprene**

Polymer	Repeat unit	Monomer(s)
Polyethylene	$-CH_2CH_2-$	$H_2C=CH_2$
Polypropylene	$-CH(CH_3)CH_2-$	$CH_3HC=CH_2$
Polystyrene		

- Globally, only 18% of plastics waste are recycled, and 24% are incinerated
- 58% is placed in landfills (dumping yards)
- Relatively little decomposition takes place in landfills, so very little of this plastic material disappears with time (stable for more than 500 years)

- Biodegradable Polymers

  water-soluble polymers : Polymers that tend to be used only once

  Microorganisms use degradable polymers for their growth

- Many hydrocarbon polymers have chemical reactivity similar to that of high boiling petroleum fractions

Environmental Degradation of Polymers

Nonbiodegradable Polymers

❖ **Goal of the polymer chemist was to design a polymer that degrades very slowly**

  ❖ use in vinyl siding for homes : should be resistant to light air, water etc

  May appear to degrade but actually due to the loss of their phthalate ester plasticizer

  vinyl products become brittle and crack, but the actual degradation of the the polymer proceeds more slowly.

Photooxidation

Polymers will be degraded if they absorb sunlight at wavelengths greater than that of the radiation that is not absorbed by stratospheric ozone

Direct photolysis is not usually a major degradative pathway

Most polymers do not have functionality that results in light absorption at wavelengths greater than 290 nm,

- Indirect photochemical processes, such as oxidation by photochemically generated singlet oxygen can also result in polymer degradation

$$R\text{—}R \xrightarrow{\text{mechanical shearing}} 2R^{\cdot}$$

$$R^{\cdot} + O_2 \rightarrow ROO^{\cdot}$$

$$ROO^{\cdot} + RH \rightarrow ROOH + R^{\cdot}$$

Radical formed by polymer processing

Hydroperoxides absorb light at wavelengths greater than 290nm and are cleaved to hydroxyl radicals and alkoxy radicals or alternatively cleave to carbonyls

$$ROOH \xrightarrow{h\nu} RO^{\cdot} + {}^{\cdot}OH$$

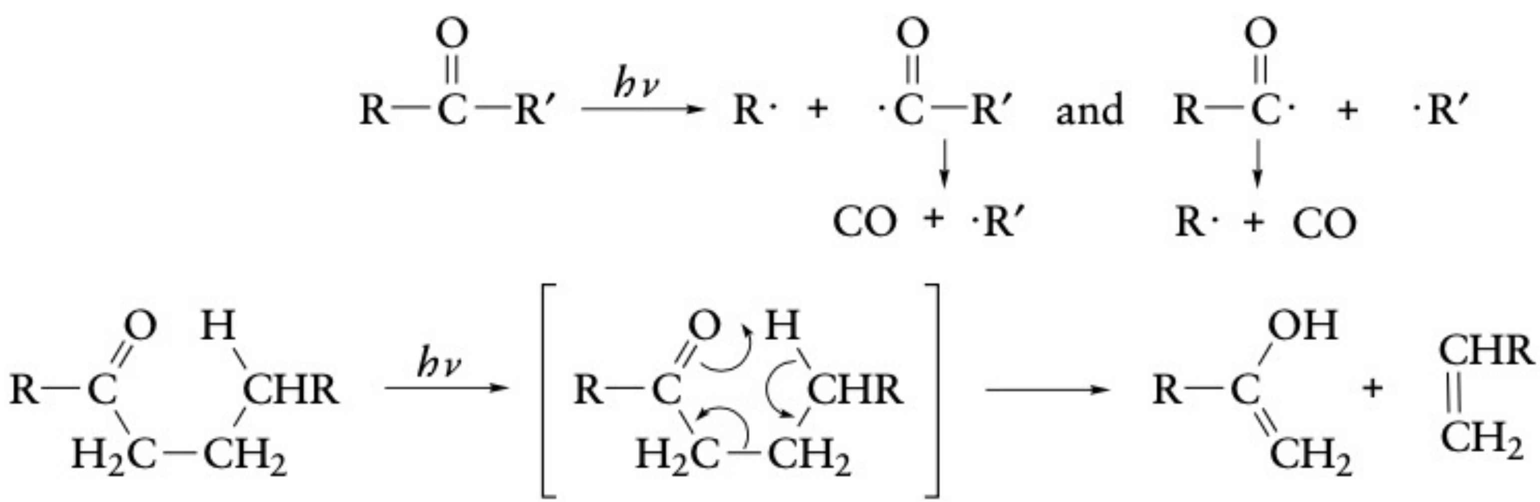
$$ROOH \xrightarrow{h\nu} R'(CO)R'', R'''(CO)CH_3$$

Presence of iron (III) or Titanium (IV) impurities can also trigger the same radical processes

- Antioxidants, which react with free radicals, are added to many polymers to slow their rate of degradation

Photochemical decomposition Triggered by Carbonyl Groups

absorb UV light at 300-325nm and initiate reactions leading to the cleavage of the polymer backbone: Norrish Type I and Norrish Type II



➢ Polymers have been prepared in which carbonyl groups were purposely incorporated into the chain to facilitate their extensive photochemical degradation.

❖ Poly- styrene cups in which 1% of the styrene units also contained a ketone grouping broke down to a wettable powder after standing outside in the sun for 3 weeks

NOT COMPLETELY DEGRADED BUT FRAGMENTED

❖ copolymer of ethylene and carbon monoxide is being used to make the plastic strap that holds beverage cans

Biodegradation

most of the biodegradable polymers contain functional groups that are subject to attack by microbial enzymes

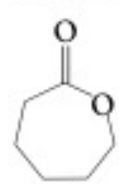
contain ester or amide groups that can be hydrolyzed or linear chains that can be oxidatively cleaved

Problem: inability of microorganisms to ingest high polymers through their cell walls

500 Da

- Some microorganisms that degrade polyesters secrete esterases, enzymes that catalyze the hydrolysis of ester groups, which break the polymer into smaller fragments that can then be ingested.

Some typical polymers that are biodegradable under aerobic conditions

Polymer	Repeat unit	Monomer(s)
Poly(vinyl acetate)	$-CH_2CH(OC(=O)CH_3)-$	$H_2C=CH-O-C(=O)CH_3$
Poly(vinyl alcohol)	$-CH_2CH(OH)-$	
Poly(ethylene oxide)	$-CH_2CH_2O-$	H_2C-CH_2 (with O bridging, epoxide ring)
Polyglycolate	$-OCH_2C(=O)-$	$HOCH_2CO_2H$
Polylactate	$-OCH(CH_3)C(=O)-$	$HOCH(CH_3)CO_2H$
Poly(β-hydroxybutyrate)	$-OCH(CH_3)CH_2C(=O)-$	$HOCH(CH_3)CH_2CO_2H$
Poly(β-hydroxyvalerate)	$-OCH(CH_2CH_3)CH_2C(=O)-$	$HOCH(CH_2CH_3)CH_2CO_2H$
Polycaprolactone	$-O(CH_2)_5C(=O)-$	

Poly β-hydroxybutyrate (PHB) and the copolymer of β-hydroxybutyrate and β-hydroxyvalerate (PHBV) are unique in that they are synthesized by microorganisms

It has been possible to induce the microorganisms to prepare the copolymer PHVB by adding some b-hydroxyvalerate to their growth medium.

In 1999 it was reported that a genetically engineered plant (oilseed rape) produced PHVB directly. This process is not economically feasible at present

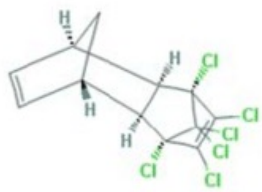
A polyethylene containing 5-20% starch granules is used in the manufacture of shopping bags that are claimed to be biodegradable.

Other additives such as transition metals may also be added to the polyethylene to enhance the photodegradation of the polyethylene via hydroperoxides

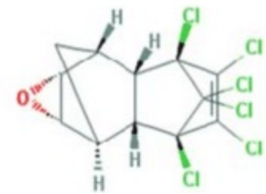
Organic Chemicals in the Environment

- pharmaceutical agents
- Fibers
- building materials
- agricultural chemicals
- solvents
- cleaning agents

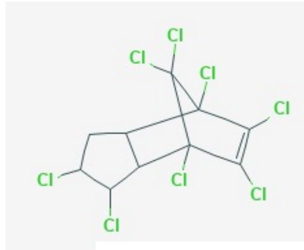
- ➢ inexpensive to manufacture
- ➢ used in large quantities



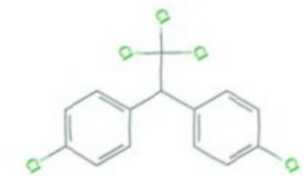
Aldrin



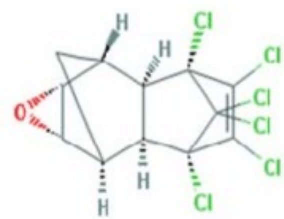
Dieldrin



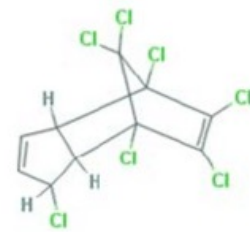
Chlordane



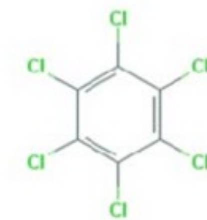
DDT



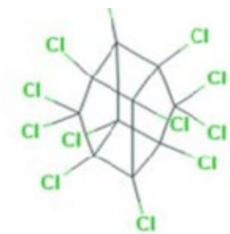
Endrin



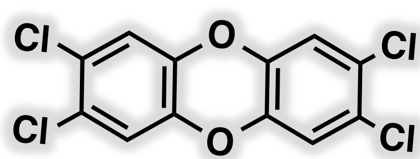
Heptachlor



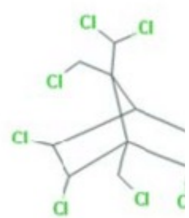
HCB



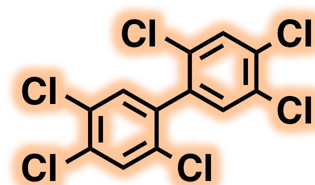
Mirex



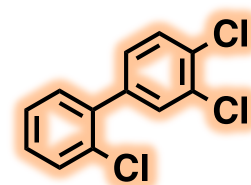
Dioxins



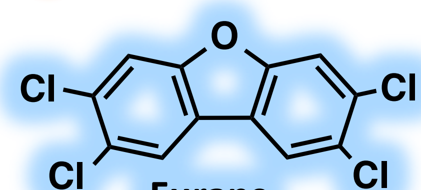
Toxaphene



PCB_1



PCB_2



Furans

Polychlorinated compounds

2,3,7,8-tetrachlorodibenzo[*b*,*d*]furan

Stockholm Convention In 2001

The 12 Key POPs—the Dirty Dozen

POP	Use
Aldrin	crop insecticide (corn, cotton)
Chlordane	crop insecticide (vegetables, citrus, cotton, potatoes)
DDT	crop insecticide (cotton)
Dieldrin	crop insecticide (cotton, corn)
Endrin	crop insecticide (cotton, grains)
Heptachlor	insecticide (termites and soil insects)
Hexa-chlorobenzene	fungicide for seed treatment
Mirex	insecticide (termites, fire ants)
Toxaphene	insecticide (livestock and crops)
PCBs	industrial chemical (heat exchange fluid for electrical transformers, paint and plastic additive)
Dioxins	unintentionally produced during combustion
Furans	unintentionally produced during combustion

ENVIRONMENTAL DEGRADATION

Hydrolysis with water:

Water reacts with the functional groups;

makes the organic part: more soluble in water,

hence more amenable to degradation by microorganisms

Reaction on mineral surfaces: Many organics bind to the surfaces of minerals in sediments

May catalyse the decomposition of substituted organics

For example, sedimentary clay minerals catalyze the hydrolysis of some chloroorganics.

Photolysis: mainly with volatile organic compounds,

Direct photolysis,

By the radicals produced by the photolysis of other atmospheric gases.

Oxidation by molecular oxygen: Unsaturated compounds are usually the most susceptible to oxidation by molecular oxygen (singlet excited state)

Microbial degradation

MICROBIAL DEGRADATION OF ORGANICS

Xenobiotics

Microorganisms mutate rapidly,

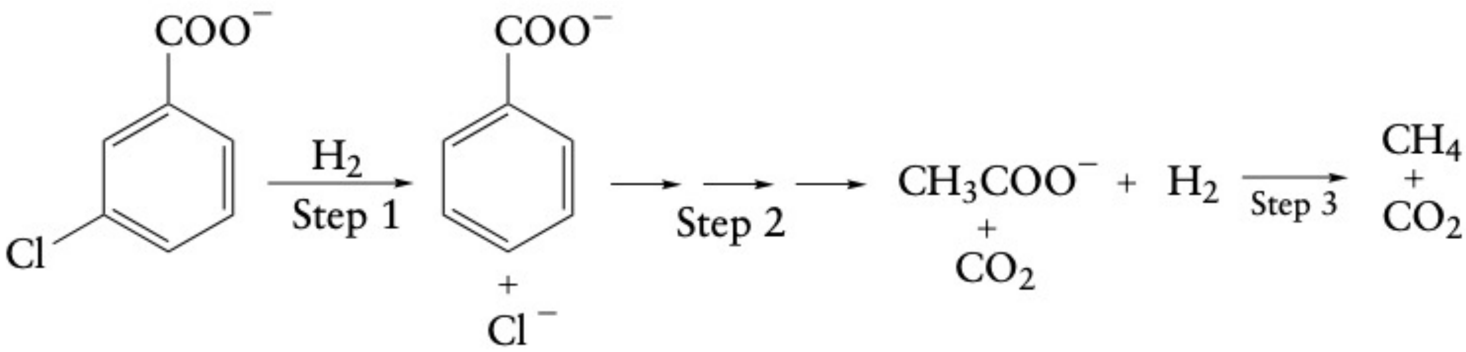
so in some cases there is a mutant that can live by metabolizing the xenobiotic

usually carried out by a cooperating group of microorganisms: consortium

Different steps in the degradation are carried out by different microorganisms

Degradation of 3-chlorobenzoate

performed by three different groups of bacteria



The hydrogen generated in step 2 is utilized by the group of microorganisms doing the reductive dehalogenation in step 1

Biodegradation may result in the intermediate conversion of the carbon to new microorganisms, not achieving mineralization until these intermediate products die

The realistic goal of biodegradation is the conversion of a xenobiotic compound to a biotic one

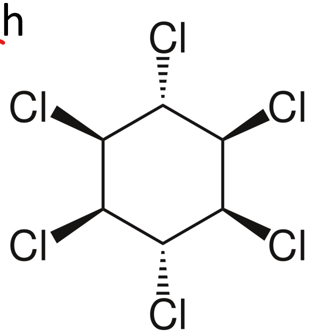
Both anaerobes and aerobes degrade xenobiotics.

Anaerobes often use reductive processes to degrade organics

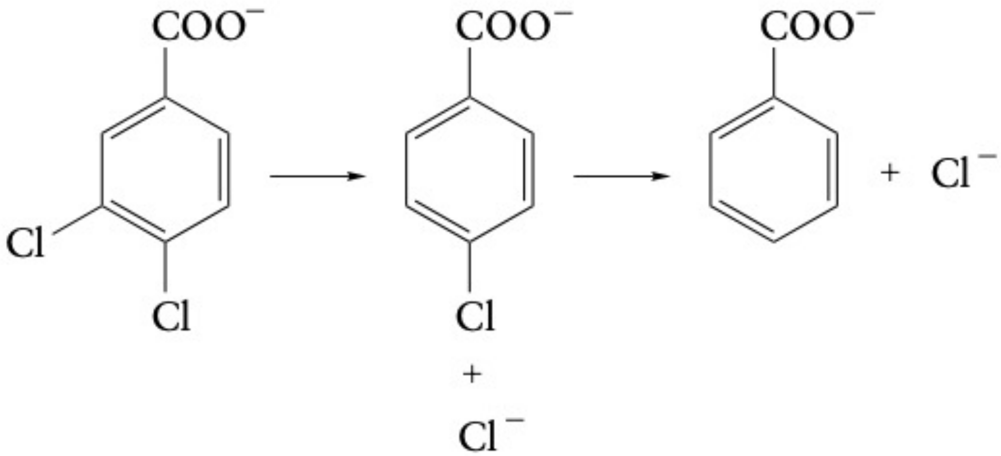
Aerobes use oxidative processes

Reductive Degradation

❖ Anaerobic dehalogenation of a pesticide was discovered during an investigation of the loss of potency of γ-hexachlorocyclohexane (Lindane) in a cattle tick bath



Principal pathway for the reductive dechlorination of 2,3-dichlorobenzoate

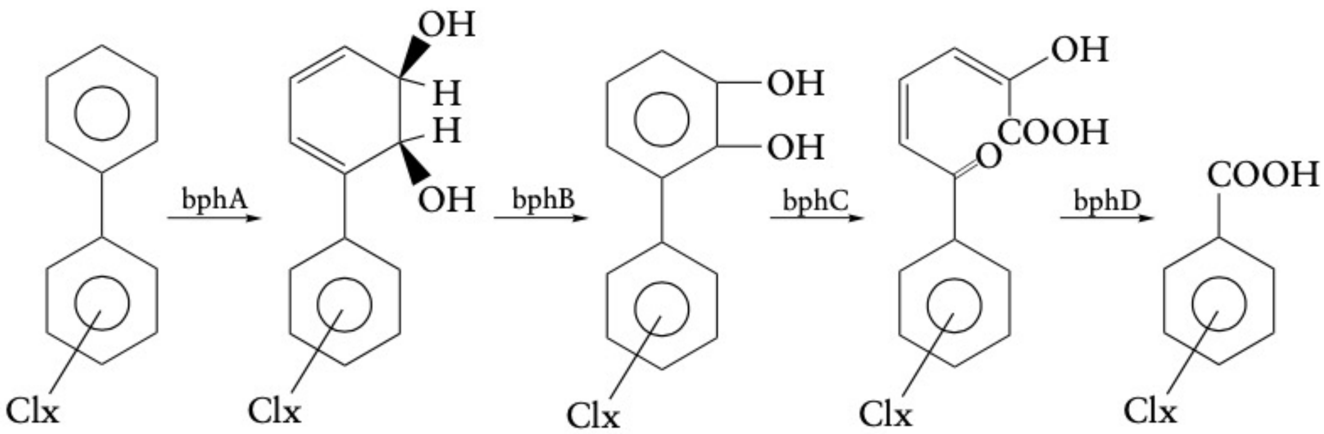


Pentachlorophenol (PCP) is rapidly converted to tetrachlorophenols,

but the subsequent loss of chloride proceeds much more slowly

Polychlorinated biphenyls (PCBs)

members of the genus Pseudomonas, degrade PCBs in the presence of oxygen



The metabolic pathways used by these diverse microorganisms for PCB destruction are similar

Aliphatic compounds

The reductive dehalogenation of aliphatic compounds has also been observed

For example, 1,1,1-trichloroethane (methylchloroform) can be reduced to the corresponding dichloro derivative

$$(Cl)_3CCH_3 + H^+ + 2e^- \longrightarrow (Cl)_2CHCH_3 + Cl^-$$

Anaerobic microorganisms reduce the chloro groups from the highly substituted biphenyls to give lightly chlorinated biphenyls, and the aerobic microorganisms destroy the lightly substituted biphenyls

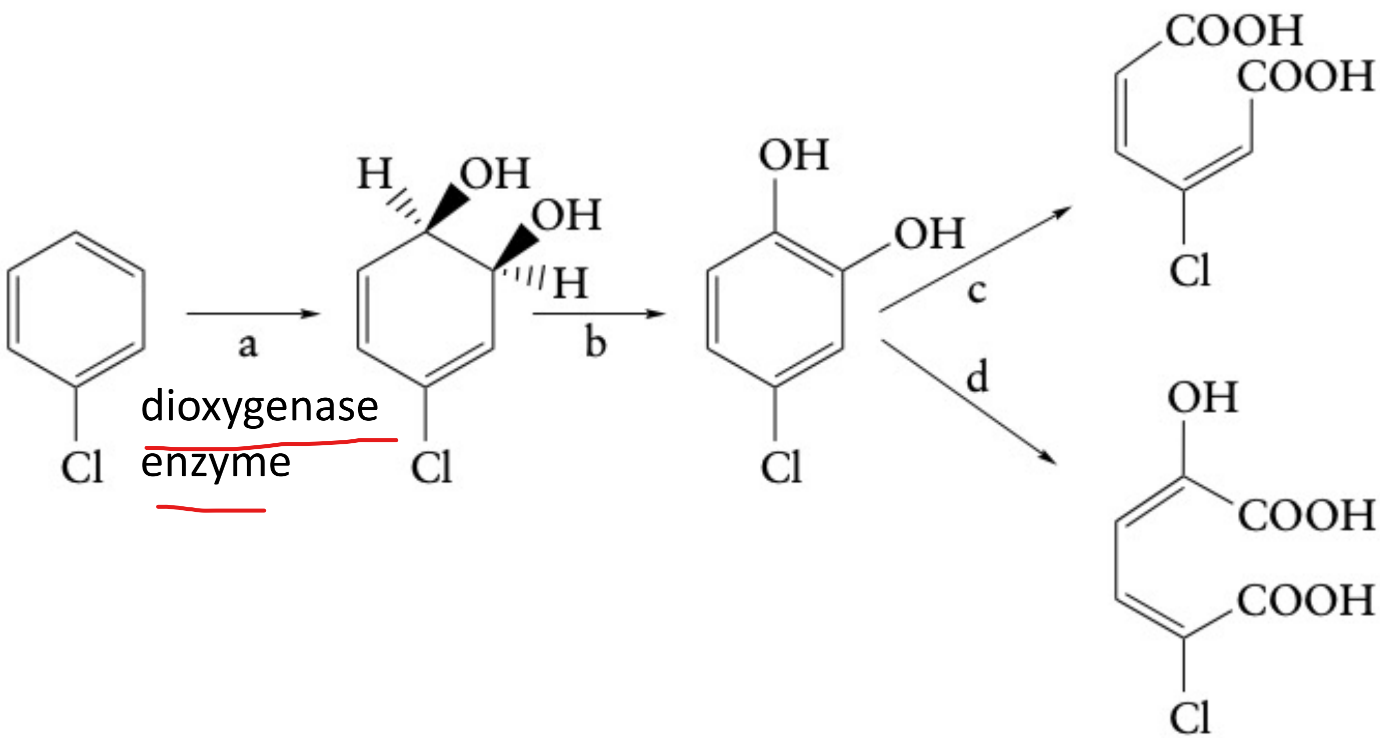
Oxidative Degradation

Occur mainly with lightly chlorinated chloroorganics

because they require the attack of an electrophilic oxygen on an electron-rich center Such as an aromatic ring or a double bond

The delocalization of these electrons by electronegative chlorine atoms decreases the electron density on the aromatic ring;

hence these compounds are less readily oxidized by molecular oxygen



readily degraded to simpler compounds and eventually to CO_2, H_2O, and Cl^-

Hydrolytic Degradation

Hydrolytic cleavage of halogens, esters, amides, and other groupings are catalyzed by both aerobes and anaerobes

$$\mathrm{G\text{-}SH} + \mathrm{CH_2Cl_2} \longrightarrow \underset{\substack{+\\ \mathrm{HCl}}}{\mathrm{G\text{-}SCH_2Cl}} \longrightarrow \underset{\substack{+\\ \mathrm{Cl^-}}}{{}^{+}\mathrm{G\text{-}S{=}CH_2}} \xrightarrow{\mathrm{H_2O}} \mathrm{G\text{-}SH} + \mathrm{CH_2O} + \mathrm{H^+}$$

Glutathione (G-SH)

microorganisms that catalyse this reaction utilize formaldehyde as an energy source.

TOXICITY

- Acute or immediate toxicity to humans and animals is not usually a problem for most commercial haloorganic compounds.

  - The toxicity associated with some haloorganics is of the chronic type, where the deleterious effect appears 2-30 years after the initial exposure

- For example, vinyl chloride evidently caused cancer (Liver tumour) in factory worker twenty years after they started working with it

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Most commercial halocarbons are nonpolar : removed from the bloodstream by the liver -enzyme cytochrome P-450 (responsible for oxidative degradation of non-polar organic)

Heavily chlorinated compound : high oxidation potentials :oxidized slowly, or not at all

- high levels of this enzyme : catalyze the oxidation of other nonpolar organics such as steroids can change the normal hormonal balance
  can cause endocrine, immune, and neurological effects

❖ Decline in the number of eagles and hawks in the 1960s and 1970s

that correlated with the buildup of chloroorganics in the environment

Resurgence of the birds in the 1980s following the ban of the use of DDT in 1972

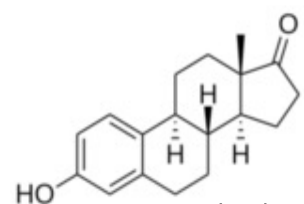
❖ Abnormal sexual development: higher levels of chloroorganics in the same animals

some industrial chemicals have hormonal activity that mimics or inhibits the activities of the sex hormones

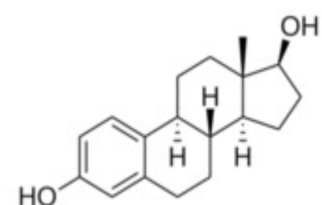
- feminization of seagulls was attributed to the higher levels of chloroorganics in these birds
- Polychlorinated bi- phenyls (PCBs) demonstrate estrogen-like activity in turtles .

Estrogen, a mixture of three steroid hormones,

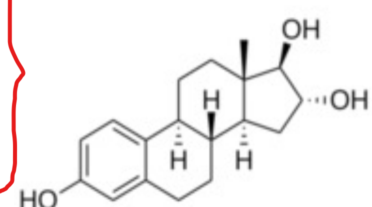
A specific ratio of estrogen to androgen (male hormones) is required during prenatal and postnatal development for sex differentiation and for the development of reproductive organs.



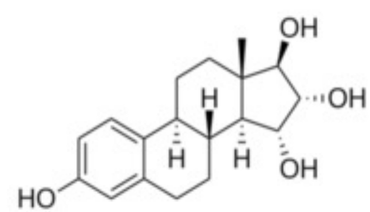
Estrone (E1)



Estradiol (E2)



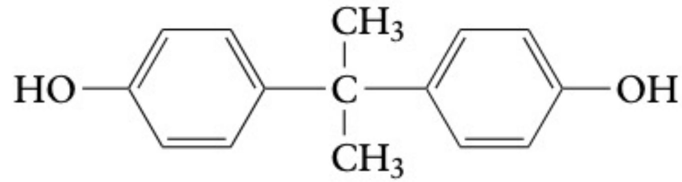
Estriol (E3)



Estetrol (E4)

produced only during pregnancy

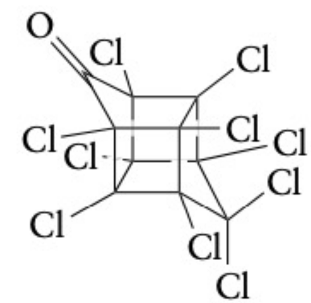
- the increases in breast cancer
- 40% decline in sperm counts in men
- increase in testicular cancer



Bisphenol A

Proposed Hormonal Activity of Environmental Compounds

Estrogen-like	Antiestrogens	Antiandrogen
PCBs	2,3,7,8-Tetrachlorodioxin	DDE
o,p-DDT	Polychlorinated benzofurans	
Methoxychlor	Benzo(*a*)pyrene	
Kepone		
Toxaphene		
Dieldrin		
Endosulfan		
Atrazine		
Bisphenol A		
Nonylphenol		



Kepone

❖ more data are required before conclusive statements can be made

❖ The toxicity associated with some haloorganics is of the chronic type, where the deleterious effect appears 2-30 years after the initial exposure.

The United Nations sponsored and adopted a treaty in 2000 intended to end the production of these and other persistent chlorinated compounds throughout the world and to destroy their stocks

The ``dirty dozen'' compounds include

- the insecticides aldrin, chlordane, DDT, dieldrin, endrin, heptachlor, hexachlorobenzene, mirex, and toxaphene plus
- the industrially used polychlorinated biphenyls and
- the industrial by-products, dioxins and benzo-furans

❖ The use of DDT for the control of malaria is the one exception to the ban, since the insecticide is still used for this purpose in Africa, Asia, and South America

CHLOROFLUOROCARBONS, HYDROFLUOROCARBONS, AND PERHALOGENATED ORGANICS

- as aerosol propellants, refrigerants, solvents, and foaming agents, but this use has been curtailed

CFC (eg. CCl_3F, $C_2F_2Cl_4$, and $C_3F_2Cl_6$)banned from production in industrialized nations in 1996

- nontoxic, stable, colorless, and non-flammable : ideal substances for industrial applications.
  Two major environmental problems,
  global warming and
  the destruction of the ozone layer

- BromoFluorocarbons (halons), such as $CBrF_3$, used to extinguish fires in military and commercial aircraft
  Banned in 1994 in industrialized countries

- Thomas Midgely Jr., invented both CFC and tetraethyllead made possible the widespread use of refrigerators and automobiles

tetraethyllead prevents preignition in internal combustion engines

Atmospheric lifetime

refers to the duration of time a greenhouse gas remains in the **atmosphere** before being decomposed by chemical processes.

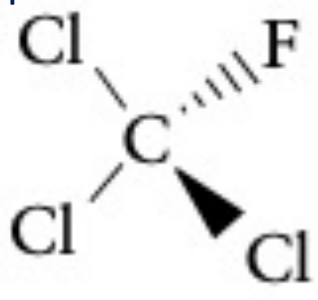
Chloro fuorocarbons (CFCs) and perhalogenated organics

were used for a long time as aerosol propellants, refrigerants, solvents, and foaming agents, **but this use has been curtailed**

So-called halons, which contain bromine as well as fluorines, are still used to extinguish fires in critical situations such as on aircraft.

nontoxic,
stable,
colorless, and
non flammable:

ideal substances for use in a variety of industrial applications



Two major environmental problems,

- Global warming and
- Destruction of the ozone layer

Atmospheric Lifetimes and Total Global Warming Potentials of Some Greenhouse Gases (Based on the Atmospheric Composition in 1990)

Gas	Atmospheric lifetime (years)	Time horizon (years) 20	100	500
CO_2		1	1	1
CH_4	10.5	63	21	9
N_2O	132	270	290	190
CFC-11[a]	55	4500	3500	1500
CFC-12[a]	116	7100	7300	4500
HCFC-22[a]	16	4100	1500	510

* $CFCl_3$ = CFC-11

Range of Lifetimes and Ozone Depletion Potentials (ODPs) of Selected Halogenated Gases

Compound	Lifetime (years)	ODP
CFC-11	50	1.0
CFC-12	102	0.82
CFC-113	85	0.90
HCFC-141b	9.4	0.10
HCFC-142b	19.5	0.05
HCFC-22	13.3	0.04
HFC-134a	14	$< 1.5 \times 10^{-5}$
HFC-152a	1.5	0
H-1301	65	12
H-1211	20	5.1
CF_4	>50,000	0
CH_3Br	~ 1.3	~ 0.6
CF_3I	< 2 days	0
HFE-7100	4.1	0

Perfluorocarbons (PFCs)
Since in general they absorb radiation in the near-infrared ``window,'' they are very powerful greenhouse gases

impact of the PFCs on global warming may be greater than 5000 times than that of carbon dioxide 100 years after their release into the atmosphere

a = number of carbon -1
b = number of hydrogen + 1
c = number of fluorine

$CFCl_3$ = CFC-11 or CFC-011

Add 90 to the numerical value

In this case
90+11= 101
def

No. of chlorine= 2d+2-e-f

d = number of carbon
e = number of hydrogen
f = number of fluorine

Quiz: Thursday, Jan 20, 2022

Instructions

- Number of questions: 10
- Total marks: 10
- All are multiple choice
- Some questions have more than one answer