Exporment 1 Experiment - to delenment the amount of Noor and Apparatus - Pipette, burette, beakers, vonirde glask, funnel, burette blond and clamp Chemicals - Sodium conferate (Na; 103), Sodium rydhoside (Na OH), hydraethloris aid (Ha) metry orange and prendphallie Chemical Strutules M3 C yellow hed Methyl orange The solution of the solution o

Pronolphalein

	Date
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	and the same of th
experiment - To determ	ine the amount of Na OH and
Na, Koz phese	ine the amount of Na 031 and in the same solution.
Apparatus - Pipetto, b	wrette, bearers, conical flat
furrel, b	wrette stand and clamp.
The same of the sa	The same of the sa
Chemisals - Sodium la	(Na 0+1), hydractoric and (to
Nydroside	(Na OH), hydrocloric and
Methyl o	hange and phenolphtallin.
10.00 .00	11. 11. 11. 11.
theory - when a mixt	we so MaOH and Na 103 15
titaated o	gainst a standard Hill golution
word of the	Solution hanges show
yeuw. do	pink (using methyle change a) due to complete neutralisation
an indicate	airables at pH24 ruchen to
of our in	titerated with phenophibale
111 16 12	icator, the whous of solution
the see how	em pink to colowless due to
ranges to	interalisation of know and half
2019 + Sulis	wion of Na, 103 (i.e upto th
1. on lers o	in of Na, 103 to Nasseo3) at
DH & 8. 1	to difference of two titants
julyes of	ing the amount of old soquet
Los help	newteralisation of Naz (03
ushile t	he difference of first Julie &

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amount of Mel enquiled from 140 of newtralisation

Manied Equation - Naon + MU - Nall + H20

Na, 103 + Mel - Non 203 + Nall

My H203 + MU - Nall + 20, + H20

Indicator - Methyl shange, Chandlothalein End point - yellow to shange

Observation - i) Sturdordization of tel sal

volume of 01 N Na, 103 Lowier taken for lack

S. No	Burette Re	ading (ml)	I solum of the HU used (ml)
	Doiting	Final	all almi la
	0	9.7	9.7
2	0	9.7	4.7
3	0	9.7	97

Mean value of HU(Vo) = 9.7 ml.

(ii) Determination of Na OH And Na, 103 in the mixture galum of mixture of Na OH and Na, 103 salwide tuber get each pix hadion = 10 ml

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4	The second of th
-	Equations: - Na OH + HICH -> Nacl + H, O
_	Nu, 103 + HU - Na H 103 + Nacl
	Na Hroz + Mul + roz + Mg
	Procedure: - 1) Standardisation of HCl
	Transfer 10 mls of Standard 0.1 N Ma, 60, solution in a clear conical plass using a sipette.
COT NO.	Add 2 drops of so melly dronge indicator. Fithall the solution against Hel purses from the burette
in	the whole of the solution changes from yellow to pink (and point)
(U)	Note the golume of the bollion used and repeat the experiment fithation at local 4 times and take the mean of the closely related readings (Va)
(2)	Determination of Ma OH and Ma, 10, philent.
i	a Javonsfer 10 ml of nintures of alkali solution into a warried fronk.
ji.	Add 2-3 derops of pronolphthalin indication. The
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SNO	Burette Reading (ml)			Volume of Her used (m)	
	9 Aitial (A)	holostos with phenolophhalein (B)	with methy		M= 1-A
- 200	000	14-03=13.7	16.5-03=162	13.7	16.2

mean value of Hel used for P = 13.7 ml mean geller of Hel used for M = 16.2 ml

As P workesponds to & newthalisation Na, 10, and NaOH Half of Na, 10, = M-P = 2.5 ml.

So, volume of HCl required for rentralization of My (03 = 2 (M-P) = V, = 2 × 2.5 = 5 ml

volume dreguere gele rendentisation of Maosi = 16.2-5=11.2 ml

Calculation - 1) Standardization of MUI

You of alpali salwier (Mu, 103) taken = 10 ml

Normality (Ni) of Ma, 103 = 0.1

Solume is MU we d= 9.7 m(1)

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	solution becomes pink in color.
(iii)	Note the initial reading of the from the burette (A) titerate the solution with standard old while the solution becomes colourles (B)
	titeate the solution with standard old while
	the solution becomes colourles (B)
(V)	Note the the sixter weller and this is the
-	phenolphtrales and point (p) , to the Same Solution
-	and 2-3 derops of methyl orange indicator
-	and continue the tileration with yel, institut
	grange at the end point(x).
	grange so ma series.
WI	This titere value i.e, the total value to Ha
	hun been from from the pegipping of the
	experiment to the methyl brange and point is
	hun seems from grow the beginning of the experiment to the methyl brange or end point is noted and this is the methyl brange and point (M)
	Results: - The your about nixture contains
	Results: - The yiven about nixture contains No.0H = 4.6 y/L
	The giver albali mix, contains May 103 = 2.741
	Exacted to 1 Daily til Andication of the fall
	Expected dos / Daily life Applications - The testal
	Experted dos / Daily life Applications - the testal basic content of an antacid tables con be determined in a similar manner
	som so was minus of a sympas munici.
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N = N × 10 = 01 × 10 = 1 = 0.103 N. (2) Detelorization of No OH and No, 105 11 Determination of MADY IL of INtel = 40 y of Mach resemblity of the used = 0.103 N 11.4 ml. of 0 = 103 NHCl = 40x 11.4 x 0.103 = of Mans of Ma OH = 0.046 if Strongto of No ON = y, x100 = 0.046×100 = 4.69/L UII Determination of May 103 1 L of I N MU = 53 g of MOON 5 ml of 0 103 MHU = 53 x 5 x 0 · 103 = y, grang of 1000 Muy 103 = 0.027-y Stronger of Mu, 103= y, X100= 2.79/L Result - The year where mintener contains know= 4.6916 the given shall mistrule weatheries Man 103= 27 911

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Precaution.	- (i) Riase the	pipette and burette
	(ii) Always put derap veise	to the correct place.
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