

# Experiment 4 -CH110 - 2020 Batch

Quiz- 10 marks - Time - 10 mins

...

1. Please write your roll number?

2001ME48

2. Please write your name?

Prateek Kumar

3. Which department you belong to?

ME



4. What are the oxidation states of O and Na in  
(1 Point)

$Na_2O_2$ ?

☐  $O = -2$  and  $Na = +2$

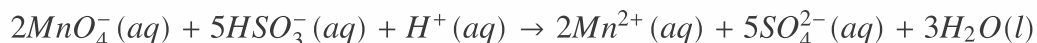
☐  $O = +2$  and  $Na = -2$

☒  $O = -1$  and  $Na = +1$

☐  $O = +1$  and  $Na = -1$

5. Which element is oxidized in the following reaction, and how does its oxidation number change?

(1 Point)



- ☐ Mn its oxidation number changes from  $-1$  to  $+2$
- ☐ Mn its oxidation number changes from  $+7$  to  $+2$
- ☐ S its oxidation number changes from  $-1$  to  $-2$
- ☒ S its oxidation number changes from  $+4$  to  $+6$

6. You are asked to prepare 250 mL of 0.25 N solution of acetic acid (molecular weight 60 g/mol). How much acetic acid will you need?

(1 Point)

- ☐ 60 g
- ☐ 15 g
- ☒ 3.75 g
- ☐ 1.875 g

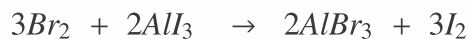
7. The standardization of sodium thiosulfide with potassium dichromate can be represented by the following reaction. How much volume of 0.25 N thiosulfide is required to for the end point of titration with 10 mL of 0.1 M dichromate solution.

(1 Point)



- ☐ 4 mL
- ☐ 12 mL
- ☐ 16 mL
- ☒ 24 mL

8. Identify the oxidizing agent of the following reaction  
(1 Point)



- ☒  $\text{Br}_2$
- ☐  $\text{AlI}_3$
- ☐  $\text{I}_2$
- ☐ None

9. What are the oxidation states of Ca and C in  
(1 Point)



- ☐  $\text{Ca} = +8$  and  $\text{C} = -4$
- ☐  $\text{Ca} = +4$  and  $\text{C} = -2$
- ☒  $\text{Ca} = +2$  and  $\text{C} = -1$
- ☐  $\text{Ca} = 0$  and  $\text{C} = 0$

10. Which oxide does not react as a reducing agent?  
(1 Point)

- ☐  $\text{NO}$
- ☐  $\text{NO}_2$
- ☐  $\text{N}_2\text{O}$
- ☒  $\text{N}_2\text{O}_5$

11. You are asked to prepare 100 mL of 0.25 N solution of  $\text{KMnO}_4$  (molecular weight 158 g/mol) at pH 8. How much  $\text{KMnO}_4$  will you need?  
(1 Point)

☐ 1.975 g

☒ 1.32 g

☐ 0.99 g

☐ 0.79 g

12. What will be end point reading of a titration of 25 ml 0.1 N of oxalic acid with 0.1 M potassium permanganate in acidic medium.

(1 Point)

☐ 25 mL

☐ 12.5 mL

☒ 5 mL

☐ 2.5 mL

13. You are asked to prepare 50 mL of 0.1 N solution of  $\text{H}_3\text{PO}_3$  (molecular weight 82 g/mol). How much  $\text{H}_3\text{PO}_3$  will you need?

(1 Point)

☐ 0.41 g

☒ 0.205 g

☐ 0.137 g

☐ 1.23 g

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