

(III) Other physical properties such as solubility in a particular solvent decreases with increasing molecular mass.

(IV) The chemical properties, which are determined solely by the functional group, remain similar in a homologous series.

(a) (II) and (III)

(b) (II) and (IV)

(c) (I), (III) and (IV)

(d) (I), (II), (III) and (IV)

(v) The table shows the formulae of three organic compounds that belong to the same homologous series.

First member of the homologous series	$\text{CH}_3\text{—O—CH}_3$
Second member of the homologous series	$\text{CH}_3\text{CH}_2\text{—O—CH}_3$
Third member of the homologous series	$\text{CH}_3\text{CH}_2\text{CH}_2\text{—O—CH}_3$

What is the general formula of this series?

(a) $\text{C}_n\text{H}_{2n}\text{O}$

(b) $\text{C}_n\text{H}_{2n+2}\text{O}$

(c) $\text{C}_n\text{H}_{2n}\text{OH}$

(d) $\text{C}_n\text{H}_{2n+2}\text{OH}$

2

Read the following and answer any four questions from 2(i) to 2(v).

When an element exists in two or more different forms in the same physical state, these different forms are called allotropes and the phenomenon is known as allotropy. Allotropes have similar chemical properties but they differ in their physical properties. Carbon exists in crystalline and amorphous forms. In crystalline form, it occurs as diamond, graphite and fullerenes. Diamond is a colourless, transparent substance having extraordinary brilliance. It is the hardest natural substance known. It is used for cutting marble, granite and glass. Graphite is a greyish-black, opaque substance. It is lighter than diamond *i.e.*, it has lower density. It has sheet like structure having hexagonal layers. One layer slides over the other layer which makes it soft to touch. It is the reason that graphite is used as a lubricant.

(i) Substance X is a moderate conductor of electricity. Substance X has the structure shown below :



Which statements about substance X are correct?

(I) It is a covalent compound.

(II) It has a giant molecular structure.

(III) It has the same structure as graphite.

(IV) It has the same structure as diamond.

(a) (I) and (III)

(b) (II) and (III)

(c) (II) and (IV)

(d) (I), (II) and (IV)

(ii) Which of the following is correct about the structure of diamond?

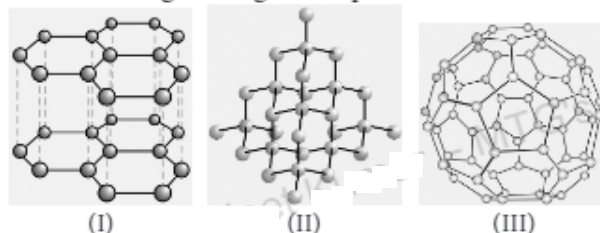
(a) Carbon atoms are held together by single covalent bonds.

(b) Electrons move freely through the structure.

(c) Layers of atoms slide easily over each other.

(d) Carbon atoms conduct electricity in the molten state.

(iii) Which three allotropes of carbon, do the given figures represent?



(I)

(II)

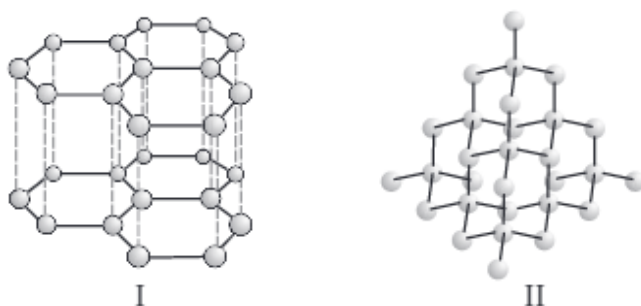
(III)

- | | | |
|--------------|-----------------------|-----------------------|
| (a) Diamond | Graphite | Buckminster fullerene |
| (b) Graphite | Buckminster fullerene | Diamond |
| (c) Diamond | Buckminster fullerene | Graphite |
| (d) Graphite | Diamond | Buckminster fullerene |

(iv) Identify the incorrect statement(s).

- (I) Diamond is the hardest substance known while graphite is smooth and slippery.
- (II) Diamond is made up of billions of carbon atoms. Each carbon atom is bonded to four other carbon atoms in a tetrahedral manner to form a giant lattice. All carbon atoms are bonded by strong covalent bonds.
- (III) Graphite is a poor conductor of electricity unlike other non-metals.
- (IV) Graphite has a giant covalent structure that is made up of layers of carbon atoms. In each layer, each carbon atom is bonded to three other carbon atoms to form hexagonal rings of carbon atoms.
- (a) (I) and (III) (b) Only (III) (c) (II) and (IV) (d) (I), (II) and (IV)

(v) Structures of two different forms of carbon are given below :



Identify the two forms (I and II respectively) and how are they related to each other?

- | | |
|--|--|
| (a) Diamond, Graphite, Isotopes of carbon | (b) Graphite, Diamond, Allotropes of carbon |
| (c) C^{12} , C^{14} , Allotropes of carbon | (d) C^{14} , C^{12} , Isotopes of carbon |

3

Read the following and answer any four questions from 3(i) to 3(v).

As neutral atom carbon has electronic configuration $K \ L \ 2, 4$. To gain inert gas configuration carbon can either

donate 4 valence electrons (helium gas configuration) or gain 4 electrons (neon gas configuration), but it cannot do so. To acquire inert gas configuration carbon can only share its 4 valence electrons with other atoms forming covalent bonds. A covalent bond can be defined as a chemical bond formed between two atoms by mutual sharing of valence electrons so that each atom acquires the stable electronic configuration of the nearest noble gas. The concept of covalent bonds was given by Langmuir and Lewis to explain bonding in non-ionic

compounds. The covalent bonds are of three types. If each atom contributes one electron, the covalent bond formed is called a single covalent bond and is represented by a single line (–) and if each atom contributes two electrons, the covalent bond formed is called a double bond and is represented by a double line (=) and if each atom contributes three electrons, the covalent bond formed is called a triple bond and is represented by a triple line (≡).

(i) Which of the following do not contain a double bond?

- | | | | |
|-------------------|---------------------|---------------------|-------------------|
| I. SO_2 | II. NH_3 | III. HCl | IV. O_2 |
| (a) I and II only | (b) II and III only | (c) III and IV only | (d) I and IV only |

(ii) Which of the following contains a triple bond?

- | | | | |
|------------------|------------------|-------------------|------------------|
| (a) N_2 | (b) O_2 | (c) CO_2 | (d) H_2 |
|------------------|------------------|-------------------|------------------|

(iii) The shared pair of electrons is said to constitute a _____ bond between two hydrogen atoms.

- | | | | |
|------------|------------|------------|-----------|
| (a) single | (b) double | (c) triple | (d) ionic |
|------------|------------|------------|-----------|

(iv) Which of the following molecules has all its atoms joined together by double covalent bonds?

- | | | | |
|-------------|-----------|--------------------|--------------------------|
| (a) Methane | (b) Water | (c) Carbon dioxide | (d) Nitrogen trichloride |
|-------------|-----------|--------------------|--------------------------|

(v) Chlorine forms a diatomic molecule, Cl_2 . The electron dot structure for this molecule is

- | | | | |
|---|---|--|---|
| (a)  | (b)  | (c)  | (d)  |
|---|---|--|---|

4

Read the following and answer any four questions from 4(i) to 4(v).

Two allotropic forms of carbon which are crystalline in nature, are diamond and graphite. They differ physically but chemically they are similar. Diamond is the hardest crystalline form of carbon. In diamond, each carbon atom is linked to four other carbon atoms by covalent bonds. In graphite, each carbon atom is linked to three other carbon atoms by covalent bond. Graphite is relatively soft and greasy. It is also a good conductor of electricity. The C—C bond length in graphite is 141.5 pm while in diamond it is 154 pm.

(i) Which of the following is a good conductor of heat and electricity?

- | | | | |
|----------|-------------|--------------|--------------|
| (a) Coal | (b) Diamond | (c) Charcoal | (d) Graphite |
|----------|-------------|--------------|--------------|

(ii) Graphite is a good conductor of electricity because

- | | | | |
|---------------------------|-----------------------|-----------------------|----------------------------|
| (a) it has free electrons | (b) it has free atoms | (c) it is crystalline | (d) it is soft and greasy. |
|---------------------------|-----------------------|-----------------------|----------------------------|

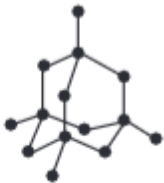

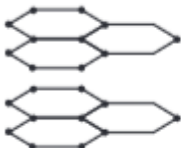
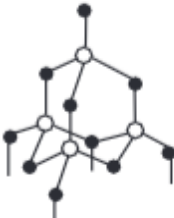
(iii) Which of the following types of binding forces is present in the structure of diamond?

- | | | | |
|-----------|--------------------|--------------|-------------------|
| (a) Ionic | (b) van der Waals' | (c) Covalent | (d) None of these |
|-----------|--------------------|--------------|-------------------|

(iv) Diamond is not a good conductor of electricity because

- | | |
|-----------------------------|-----------------------------------|
| (a) it is very hard | (b) its structure is very compact |
| (c) it is not water soluble | (d) it has no free electron. |

(v) Which of the following is the structure of diamond?

- | | | | |
|---|---|--|---|
| (a)  | (b)  | (c)  | (d)  |
|---|---|--|---|

Read the following and answer any four questions from 5(i) to 5(v).

The compounds which have the same molecular formula but differ from each other in physical or chemical properties are called isomers and the phenomenon is called isomerism. When the isomerism is due to difference in the arrangement of atoms within the molecule, without any reference to space, the phenomenon is called structural isomerism. In other words, structural isomers are compounds that have the same molecular formula but different structural formulas, i.e., they are different in the order in which different atoms are linked. In these compounds, carbon atoms can be linked together in the form of straight chains, branched chains or even rings.

- (i) Which of the following sets of compounds have same molecular formula?
- (a) Butane and *iso*-butane (b) Cyclohexane and hexene
(c) Propanal and propanone (d) All of these
- (ii) In order to form branching, an organic compound must have a minimum of
- (a) four carbon atoms (b) three carbon atoms
(c) five carbon atoms (d) any number of carbon atoms.
- (iii) Which of the following is an isomeric pair?
- (a) Ethane and propane (b) Ethane and ethene
(c) Propane and butane (d) Butane and 2-methylpropane
- (iv) Among the following the one having longest chain is
- (a) *neo*-pentane (b) *iso*-pentane
(c) 2-methylpentane (d) 2, 2-dimethylbutane.
- (v) The number of isomers of pentane is
- (a) 2 (b) 3 (c) 4 (d) 5

Read the following and answer any four questions from 6(i) to 6(v).

Study the table related to three hydrocarbons A, B, C and answer the questions that follow.

Organic compound	Molecular formula
A	C_3H_8
B	C_5H_{10}
C	C_4H_6

- (i) A, B and C are classified as hydrocarbons because
- (a) they contain hydrogen (b) they contain carbon
(c) they contain both carbon and hydrogen (d) none of these.
- (ii) Which of these organic compounds is an alkyne?
- (a) A (b) B (c) C (d) All of these
- (iii) C_5H_{10} belongs to
- (a) C_nH_{2n+2} series (b) C_nH_{2n} series
(c) C_nH_{2n-2} series (d) none of these.

(iv) Identify the incorrect statement about these three hydrocarbons.

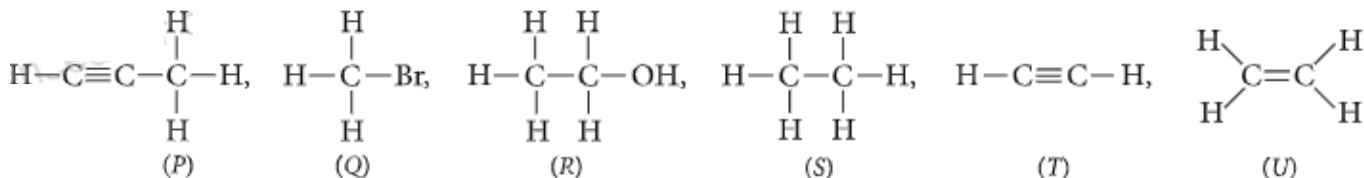
- (a) All have different general formula. (b) *A* and *B* differ by $-\text{CH}_2$ unit.
(c) *C* is an alkyne. (d) *B* is an alkene.

(v) General formula for alkane is

- (a) C_nH_{2n} (b) $\text{C}_n\text{H}_{2n+2}$ (c) $\text{C}_n\text{H}_{2n-2}$ (d) C_nH_n

7

Read the following and answer any four questions from 7(i) to 7(v).



(i) Which of the following compounds belong to same homologous series?

- (a) *S* and *T* (b) *T* and *U* (c) *P* and *U* (d) *P* and *T*.

(ii) The functional group of compound (*R*) is

- (a) alcohol (b) aldehyde (c) ketone (d) carboxylic acid.

(iii) Compound (*T*) belongs to homologous series of

- (a) alkynes (b) alkenes (c) alkanes (d) none of these.

(iv) Which of the following compounds is unsaturated hydrocarbon?

- (a) *S* (b) *Q* (c) *U* (d) *R*

(v) Which of the following compounds belongs to alkane series?

- (a) *P* (b) *S* (c) *T* (d) *U*

8

Read the following and answer any four questions from 8(i) to 8(v).

The table given below shows six organic compounds *A*, *B*, *C*, *D*, *E* and *F* having different molecular formula :

Organic compound	Molecular formula
<i>A</i>	C_7H_{16}
<i>B</i>	C_8H_{16}
<i>C</i>	C_4H_6
<i>D</i>	C_6H_{10}
<i>E</i>	C_5H_{10}
<i>F</i>	C_9H_{20}

(i) Which of the following compounds belong to same homologous series?

- (a) *E* and *F* (b) *B* and *C* (c) *A* and *B* (d) *C* and *D*

(ii) Which of the following is the member of the same homologous series as *E*?

- (a) *D* (b) *A* (c) *F* (d) *B*

(iii) Identify the correct statements.

- (a) *A* and *F* are saturated hydrocarbons while all others are unsaturated hydrocarbons.
- (b) *C* and *D* belong to a homologous series having general formula C_nH_{2n} .
- (c) *B* and *E* are alkynes.
- (d) All the compounds have same physical and chemical properties.

(iv) Compound *B* is

- (a) an alkane
- (b) an alkene
- (c) an alkyne
- (d) none of these.

(v) Compound (*F*) has a general formula

- (a) C_nH_{2n+2}
- (b) C_nH_{2n}
- (c) C_nH_{2n+4}
- (d) C_nH_{2n+2}

9

Read the following and answer any four questions from 9(i) to 9(v).

A hydrocarbon (*P*) has the molecular formula $C_{10}H_{22}$. A hydrocarbon (*Q*) has two carbon atoms less than (*P*) and belong to the same homologous series. A hydrocarbon (*R*) has two carbon atoms more than (*P*) and belong to the same homologous series.

(i) What is the molecular formula of (*Q*)?

- (a) $C_{12}H_{26}$
- (b) C_8H_{16}
- (c) C_8H_{18}
- (d) C_8H_{14}

(ii) To which homologous series do the compound (*P*), (*Q*) and (*R*) belong?

- (a) C_nH_{2n}
- (b) C_2H_{2n-2}
- (c) C_nH_{2n+2}
- (d) C_nH_{2n+1}

(iii) What is the molecular formula of (*R*)?

- (a) $C_{12}H_{26}$
- (b) $C_{12}H_{24}$
- (c) $C_{12}H_{22}$
- (d) $C_{12}H_{28}$

(iv) Identify the correct statement about compounds (*P*), (*Q*) and (*R*).

- (a) They have same melting and boiling points.
- (b) They have same chemical properties.
- (c) They have different general formula.
- (d) They differ by $-CH_2$ unit.

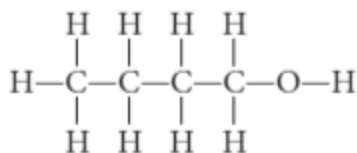
(v) Compounds (*P*), (*Q*) and (*R*) are

- (a) alkanes
- (b) alkenes
- (c) alkynes
- (d) none of these.

10

Read the following and answer any four questions from 10(i) to 10(v).

An organic molecule has the following structure :



(i) To which homologous series does this molecule belong?

- (a) Aldehydes
- (b) Ketones
- (c) Alcohols
- (d) Alkanes

(ii) What is the general formula of this homologous series?

- (a) $C_nH_{2n+1}OH$
- (b) C_nH_{2n+2}
- (c) $C_nH_{2n}O$
- (d) $C_nH_{2n+1}CHO$

(iii) Which is the next member of this series?

- (a) C_4H_9OH
- (b) C_3H_7OH
- (c) $C_5H_{11}OH$
- (d) $C_6H_{13}OH$

- (iv) Which is the third member of this series?
(a) C_3H_7OH (b) C_4H_9OH (c) C_2H_5OH (d) CH_3OH
- (v) Which is the second member of this series?
(a) Ethanol (b) Methanol (c) Propanol (d) Butanol

ASSERTION & REASON

For question numbers 11-25, two statements are given- one labelled Assertion (A) and the other labelled Reason (R). Select the correct answer to these questions from the codes (a), (b), (c) and (d) as given below :

- (a) Both A and R are true, and R is correct explanation of the assertion.
(b) Both A and R are true, but R is not the correct explanation of the assertion.
(c) A is true, but R is false.
(d) A is false, but R is true.
11. **Assertion :** Saturated hydrocarbons are chemically less reactive.
Reason : All the valencies of carbon atom are satisfied by single covalent bonds.
12. **Assertion :** Both aldehydes and ketones contain carbonyl group.
Reason : In aldehydes, the functional group is attached to atleast one hydrogen atom.
13. **Assertion :** In alkanes, alkenes and alkynes the valency of carbon is always four.
Reason : All hydrocarbons except alkanes contain double bonds.
14. **Assertion :** Graphite is a good conductor of electricity.
Reason : It has one free valence electron.
15. **Assertion :** The functional group present in alcohols is $-OH$.
Reason : It is the same group as present in water, hence water and alcohol have similar properties.
16. **Assertion :** Ethanol is first member of the alcohol homologous series.
Reason : A homologous series can be represented by a general formula.
17. **Assertion :** Carbon and its compounds can be used as fuels.
Reason : They are highly inflammable and have high calorific value.
18. **Assertion :** Diamond is not good conductor of electricity.
Reason : It has no free electrons.
19. **Assertion :** Covalent compounds are generally poor conductor of electricity.
Reason : They consist of molecules and not ions which can transfer charge.
20. **Assertion :** Carbon possesses property of catenation.
Reason : Carbon atoms form double as well as triple bonds during catenation.
21. **Assertion :** Two members of a homologous series have similar chemical properties.
Reason : Propane and butane are members of same homologous series.
22. **Assertion :** Olefins have the general formula C_nH_{2n+1} .
Reason : There is atleast one double bond between two carbon atoms in their molecules.
23. **Assertion :** Diamond is the hardest natural known substance.
Reason : Diamond is used for cutting marble, granite and glass.
24. **Assertion :** Diamond and graphite do not have the same crystal structure.
Reason : Diamond is crystalline while graphite is amorphous.
25. **Assertion :** Graphite is soft and slippery to touch.
Reason : Graphite has sheet like layered structure.

HINTS & EXPLANATIONS

1. (i) (a): All the members of homologous series show similar chemical properties.

(ii) (c): Alkynes have the general formula C_nH_{2n-2} , e.g., Ethyne (C_2H_2), Propyne (C_3H_4), Butyne (C_4H_6).

(iii) (b): Two consecutive members of a homologous series differ by a $-CH_2-$ group

(iv) (c): The melting and boiling points increase with increasing molecular mass.

(v) (b): Molecular formula of first member : C_2H_6O

Molecular formula of second member : C_3H_8O

Molecular formula of third member : $C_4H_{10}O$

Thus, the general formula of the homologous series is $C_nH_{2n+2}O$.

2. (i) (c): Each atom is covalently bonded to four other atoms, which in turn, are bonded to four more atoms. Thus, X is a giant molecule and has a structure similar to that of diamond. Substance X is not a compound as it consists of only one type of atoms. Thus, X is an element. Graphite has layers of carbon atoms.

(ii) (a) (iii) (d)

(iv) (b): In graphite only three valence electrons are used for bond formation and hence fourth electron is free to move which makes it a good conductor of electricity.

(v) (b): Given structures are of graphite and diamond and these are allotropes of carbon.

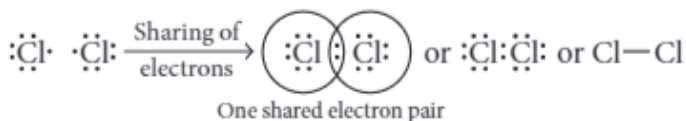
3. (i) (b): Both NH_3 and HCl have single bonds.

(ii) (a): $N \equiv N$

(iii) (a)

(iv) (c): $O = C = O$

(v) (c): In chlorine molecule, both chlorine atoms contribute one electron and thus share single electron pair to form single covalent bond. As shared pair is shared by both atoms, they acquire inert gas configuration of argon atom in valence shell.



4. (i) (d) (ii) (a)

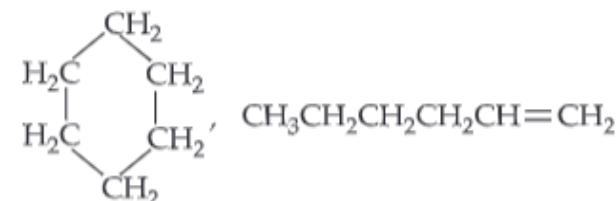
(iii) (c)

(iv) (d): In diamond, one carbon is attached to four other carbon atoms hence it has no free electron.

(v) (a)

5. (i) (d): $CH_3CH_2CH_2CH_3$, $\begin{array}{c} CH_3CHCH_3 \\ | \\ CH_3 \end{array}$

(Butane and *iso*-Butane- C_4H_{10})



(Cyclohexane and hexene- C_6H_{12})

CH_3CH_2CHO , CH_3COCH_3

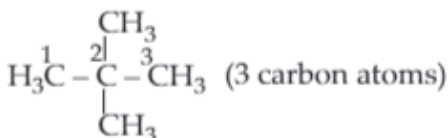
(Propanal and propanone- C_3H_6O)

(ii) (a)

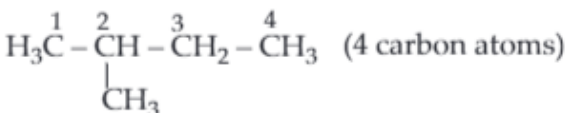
(iii) (d): $CH_3CH_2CH_2CH_3$ and $\begin{array}{c} CH_3CHCH_3 \\ | \\ CH_3 \end{array}$ have

different structural formulas and same molecular formula.

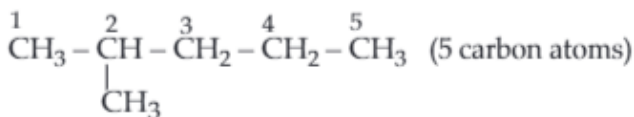
(iv) (c): *neo*-Pentane :



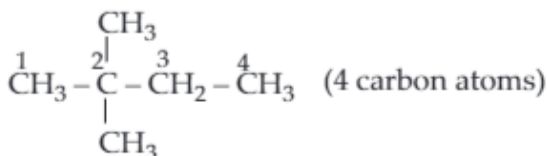
iso-Pentane :



2-Methylpentane :



2,2-Dimethylbutane :

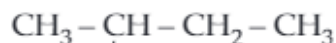


Hence, 2-methylpentane has the longest carbon chain.

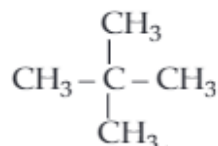
(v) (b): Pentane (C_5H_{12}) has three structural isomers:



n-Pentane



iso-Pentane



neo-Pentane

6. (i) (c): *A*, *B* and *C* are classified as hydrocarbons because these compounds are made up of carbon and hydrogen only.

(ii) (c): *C* is an alkyne.

(iii) (b): C_5H_{10} is an alkene having a general formula C_nH_{2n} .

(iv) (b): *A* and *B* do not belong to same homologous series. *A* is an alkane while *B* is an alkene.

(v) (b)

7. (i) (d): (*P*) and (*T*) are alkynes.

(ii) (a): Alcohol ($-OH$).

(iii) (a): (*T*) is an alkyne having general formula of C_nH_{2n-2} .

(iv) (c): (*U*) is an alkene.

(v) (b)

8. (i) (d): *A* and *F* are alkanes; *B* and *E* are alkenes; *C* and *D* are alkynes.

(ii) (d): *B* is an alkene having general formula C_nH_{2n} , the homologous series to which *E* belongs.

(iii) (a): *C* and *D* belong to a homologous series having general formula C_nH_{2n-2} . *B* and *E* are alkenes. All the compounds have different physical and chemical properties.

(iv) (b): (*B*) is alkene.

(v) (d): (*F*) is an alkane.

9. (i) (c): Molecular formula of (*Q*) is C_8H_{18} as it has two carbon atoms less than (*P*).

(ii) (c): Compounds (*P*), (*Q*) and (*R*) are alkanes having general formula C_nH_{2n+2} .

(iii) (a): Molecular formula of (*R*) is $C_{12}H_{26}$ as it has two carbon atoms more than (*P*).

(iv) (b): Compound (*P*), (*Q*) and (*R*) belong to same homologous series so they have different physical properties but similar chemical properties. They have same general formula C_nH_{2n+2} . They differ by 2 carbon atoms and 4 hydrogen atoms.

(v) (a)

10. (i) (c): Alcohol ($-OH$).

(ii) (a): $C_nH_{2n+1}OH$ is the general formula of the homologous series of alcohol.

(iii) (c)

(iv) (a)

(v) (a): Ethanol; C_2H_5OH is the second member of this series.

11. (a)

12. (b)

13. (c): Only alkenes contain double bond, alkynes contain triple bond, but in all hydrocarbons carbon remains tetravalent.

14. (a)

15. (c): In alcohols $-OH$ group is attached to an alkyl group hence water and alcohol have different properties.

16. (d): Methanol is first member of the alcohol homologous series.

17. (a)

18. (a): In diamond, one carbon atom is attached to four other carbon atoms, hence it has no free electron.

19. (a)

20. (b): Catenation property of carbon is primarily due to its small size, electronic configuration and unique strength of carbon-carbon bonds.

21. (b): All the members of homologous series contain the same functional group, show gradation in physical properties and similarity in chemical properties.

22. (d): Olefins are unsaturated hydrocarbons. There is at least one double bond between two carbon atoms in their molecules and they have the general formula C_nH_{2n} .

23. (b)

24. (c): In diamond, C-atoms are sp^3 hybridized while in graphite, they are sp^2 hybridized. Diamond and graphite both are crystalline forms of carbon.

25. (a)