

CBSE Test Paper-01

Chapter 01 Chemical Reactions and Equations

1. Match the following with correct response. **(1)**

(1) Combination reaction	(A) $ZnO + C \rightarrow Zn + CO$
(2) Decomposition reaction	(B) $2H_2O \rightarrow 2H_2 + O_2$
(3) Displacement reaction	(C) $2Mg + O_2 \rightarrow 2MgO$
(4) Redox reaction	(D) $Fe + CuSO_4 \rightarrow FeSO_4 + Cu$

- a. 1-D, 2-A, 3-C, 4-B
- b. 1-C, 2-B, 3-D, 4-A
- c. 1-B, 2-D, 3-A, 4-C
- d. 1-A, 2-C, 3-B, 4-D

2. The chemical formula of lead sulphate is **(1)**

- a. Pb_2SO_4
- b. $PbSO_4$
- c. $Pb_2(SO_4)_3$
- d. $Pb(SO_4)_2$

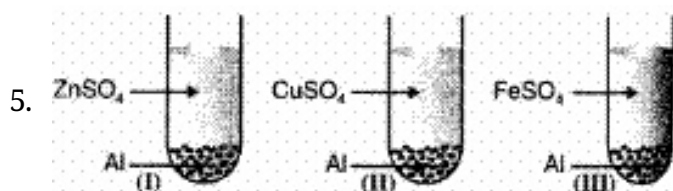
3. The colour of zinc metal is **(1)**

- a. Red dish brown
- b. silvery
- c. white
- d. Grey

4. When you place an iron nail in copper sulphate solution, the reddish brown coating formed on the nail is **(1)**

- a. smooth and shiny
- b. rough and granule
- c. soft and dull

d. hard and flaky



Observation	I	II	III
Solution after reaction	Colourless	Colourless	Colourless
Metal Deposited	Zn	Cu	Fe

Which of the following is correct conclusion? **(1)**

- Al is more reactive than Cu and Fe but less reactive than Zn
 - Al is more reactive than Cu but less reactive than Zn and Fe
 - Al is more reactive than Zn and Cu but less reactive than Fe
 - Al is more reactive than Zn, Cu, Fe
- Oil and fat containing food items are flushed with nitrogen. Why? **(1)**
 - What happens when copper powder is heated ? **(1)**
 - What is the principle of balancing of chemical equation? **(1)**
 - Why are food items packed in aluminium foils? **(1)**
 - Predict whether gold can displace copper from copper sulphate solution. **(3)**
 - What does one mean by exothermic and endothermic reactions ? Give examples. **(3)**
 - What are double displacement reaction? **(3)**
 - In the reaction $\text{PbS (s)} + 4\text{H}_2\text{O}_2(\text{aq}) \rightarrow \text{PbSO}_4(\text{s}) + 4\text{H}_2\text{O (l)}$, which is oxidizing agent and which is reducing agent? **(3)**
 - Write the formula and then balance the following equations. **(5)**