

**Cations**

<b>1+</b>		<b>2+</b>		<b>3+</b>	
ammonium	$\text{NH}_4^+$	barium	$\text{Ba}^{2+}$	aluminum	$\text{Al}^{3+}$
cesium	$\text{Cs}^+$	beryllium	$\text{Be}^{2+}$	chromium(III)	$\text{Cr}^{3+}$
gold(I)	$\text{Au}^+$	cadmium	$\text{Cd}^{2+}$	cobalt(III)	$\text{Co}^{3+}$
hydrogen	$\text{H}^+$	calcium	$\text{Ca}^{2+}$	gold(III)	$\text{Au}^{3+}$
lead(I)	$\text{Pb}^+$	cobalt(II)	$\text{Co}^{2+}$	iron(III)	$\text{Fe}^{3+}$
lithium	$\text{Li}^+$	copper(II)	$\text{Cu}^{2+}$	manganese(III)	$\text{Mn}^{3+}$
potassium	$\text{K}^+$	iron(II)	$\text{Fe}^{2+}$	<b>4+</b>	
silver	$\text{Ag}^+$	lead(II)	$\text{Pb}^{2+}$		
sodium	$\text{Na}^+$	magnesium	$\text{Mg}^{2+}$		
copper(I)	$\text{Cu}^+$	manganese(II)	$\text{Mn}^{2+}$		
		mercury(I)	$\text{Hg}_2^{2+}$		
		mercury(II)	$\text{Hg}^{2+}$		
		nickel(II)	$\text{Ni}^{2+}$		
		strontium	$\text{Sr}^{2+}$	tin(IV)	$\text{Sn}^{4+}$
		zinc	$\text{Zn}^{2+}$	nickel(IV)	$\text{Ni}^{4+}$
		tin(II)	$\text{Sn}^{2+}$	lead(IV)	$\text{Pb}^{4+}$

*Roman numeral notation indicates charge of ion when element commonly forms more than one ion. For example, iron(II) has a 2+ charge; iron(III) a 3+ charge.*

**Anions**

1-				2-		3-	
acetate	$\text{C}_2\text{H}_3\text{O}_2^-$	cyanide	$\text{CN}^-$	carbonate	$\text{CO}_3^{2-}$	arsenate	$\text{AsO}_4^{3-}$
amide	$\text{NH}_2^-$	cyanate	$\text{OCN}^-$	chromate	$\text{CrO}_4^{2-}$	arsenite	$\text{AsO}_3^{3-}$
hydrogen carbonate		fluoride	$\text{F}^-$	dichromate	$\text{Cr}_2\text{O}_7^{2-}$	citrate	$\text{C}_6\text{H}_5\text{O}_7^{3-}$
(bicarbonate)	$\text{HCO}_3^-$	hydride	$\text{H}^-$	oxide	$\text{O}^{2-}$	ferricyanide	$\text{Fe}(\text{CN})_6^{3-}$
hydrogen sulfate		hydroxide	$\text{OH}^-$	oxalate	$\text{C}_2\text{O}_4^{2-}$	nitride	$\text{N}^{3-}$
(bisulfate)	$\text{HSO}_4^-$	hypochlorite	$\text{ClO}^-$	silicate	$\text{SiO}_3^{2-}$	phosphate	$\text{PO}_4^{3-}$
bisulfide	$\text{HS}^-$	iodate	$\text{IO}_3^-$	sulfate	$\text{SO}_4^{2-}$	phosphite	$\text{PO}_3^{3-}$
bisulfite	$\text{HSO}_3^-$	iodide	$\text{I}^-$	sulfide	$\text{S}^{2-}$	phosphide	$\text{P}^{3-}$
bromate	$\text{BrO}_3^-$	nitrate	$\text{NO}_3^-$	sulfite	$\text{SO}_3^{2-}$	CHEMISTRY EXPERT RAKESH KUMAR "CULTIVATING EXCELLENCE IN EVERY STUDENT" +919814516618	
bromide	$\text{Br}^-$	nitrite	$\text{NO}_2^-$	tartrate	$\text{C}_4\text{H}_4\text{O}_6^{2-}$		
chlorate	$\text{ClO}_3^-$	perchlorate	$\text{ClO}_4^-$	tetraborate	$\text{B}_4\text{O}_7^{2-}$		
chlorite	$\text{ClO}_2^-$	permanganate	$\text{MnO}_4^-$	thiosulfate	$\text{S}_2\text{O}_3^{2-}$		
chloride	$\text{Cl}^-$	thiocyanate	$\text{SCN}^-$				