

## CBSE Test Paper-02

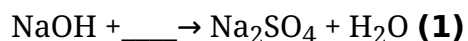
### Chapter 01 Chemical Reactions and Equations

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1. When zinc metal is dipped in copper sulphate solution **(1)**
  - a. No reaction takes place
  - b. The solution remains blue and copper metal gets deposited
  - c. The solution becomes colourless and reddish brown copper metal gets deposited
  - d. The solution becomes green and copper metal gets deposited
2. Which gases are given out when Lead nitrate is heated? **(1)**
  - a.  $\text{NO}_2$ ,  $\text{O}_2$
  - b.  $\text{N}_2\text{O}_4$ ,  $\text{O}_2$
  - c.  $\text{PbO}$ ,  $\text{O}_2$
  - d.  $\text{NO}$ ,  $\text{O}_3$
3. Find the incorrect statement : **(1)**
  - (I) Oxygen is highly combustible and hydrogen is supporter of combustion,
  - (II) Oxygen and hydrogen both are highly combustible,
  - (III) Oxygen and hydrogen both are supporters of combustion,
  - (IV) Hydrogen is highly combustible and oxygen is supporter of combustion
  - a. I, II and III
  - b. I, III and IV
  - c. IV, I and II
  - d. I, II and IV
4. Which of the following decolourizes a blue solution of copper sulphate? **(1)**
  - A. Al
  - B. Zn
  - C. Fe
  1. (A), (B) and (C)
  2. (B) only
  3. (A) only
  4. (C) only
5. The reaction  $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$  represents: **(1)**

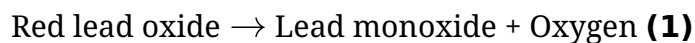
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- a. Decomposition
  - b. Oxidation
  - c. Combination
  - d. Reduction

6. Complete and balance the following chemical reaction.



7. Why does the colour of heated copper powder becomes black when air is passed over it? **(1)**

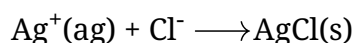
8. Write the formula and then balance the following equation.



9. Define a displacement reaction. **(1)**

10. Two beakers A and B contain Iron (II) sulphate solution. In the beaker A is placed a small piece of copper and in the beaker B is placed a small piece of zinc. It is found that a grey deposit forms on the zinc but not on the copper. What can be concluded from these observations? **(3)**

11. When solutions of silver nitrate and sodium chloride are mixed, white precipitate forms. The ionic equation for the reaction is **(3)**



i. a. What is the name of the white precipitate?

b. Is it a soluble or insoluble compound?

ii. Is the precipitation of silver chloride a redox reaction?

12. Name the method used to balance a chemical equation. **(3)**

13. What is a balanced chemical equation ? Why should chemical equations be balanced ? **(3)**

14. What are the types of combination reactions ? Give example of each type. **(5)**

15. How will you write a chemical equation ? **(5)**