

CBSE TEST PAPER-01
CLASS - X (Acid Base and Salt)

1. An acid can react with (1)
 - (a) $AgCl$
 - (b) Na_2CO_3
 - (c) $PbSO_4$
 - (d) Na_2SO_4
2. Which of the following gives CO_2 on heating? (1)
 - (a) Slaked
 - (b) Quick lime
 - (c) Lime stone
 - (d) Soda ash.
3. Plaster of Paris is made from (1)
 - (a) Lime stone
 - (b) Slaked Lime
 - (c) Quick lime
 - (d) Gypsum
4. Which is a base and not alkali? (1)
 - (a) $NaOH$
 - (b) KOH
 - (c) $Fe(OH)_3$
 - (d) None
5. Chemical formula of baking soda is (1)
 - (a) $MgSO_4$
 - (b) Na_2CO_3
 - (c) $NaHCO_3$
 - (d) $MgCO_3$

6. What happens to the crystals of washing soda when exposed to air? (2)
7. What is the chemical name of washing soda? Name three raw materials used in making washing soda by Solvay process? (2)
8. What is efflorescence? Give an example? (2)
9. Why is sodium hydrogen carbonate an essential ingredient in antacids? (2)
10. (a) Name the raw materials used in the manufacture of sodium carbonate by Solvay process? (3)
(b) How is sodium hydrogen carbonate formed from a mixture of NH_4Cl and $NaHCO_3$?
11. Write equations for the following reactions (3)
 - (i) Dilute sulphuric acid reacts with zinc granules
 - (ii) Dilute hydrochloric acid reacts with magnesium ribbon.
 - (iii) Dilute sulphuric acid reacts with aluminum powder.
12. (a) An aqueous solution has a pH value of 7.0. Is this solution acidic, basic or neutral? (3)
(b) If H^+ concentration of a solution is $1 \times 10^{-2} \text{ mol L}^{-1}$ what will be its pH value?
(c) Which has higher pH value: 1-M HCl or 1-M $NaOH$
13. What will you observe when: (3)
 - (i) Red litmus is introduced into a solution of sodium sulphate.
 - (ii) Methyl orange is added to dil HCl .
 - (iii). Blue litmus is introduced into a solution of ferric chloride
14. (a) The pH of rain water collected from two cities A and B was found to be 6 and 5 respectively. Water of which city is more acidic? Find out the ratio of hydrogen ion concentration in the two samples of rain water? (5)
(b) Arrange the following in order (ascending) of their pH values.
 $NaOH$ solution, Blood, lemon Juice.