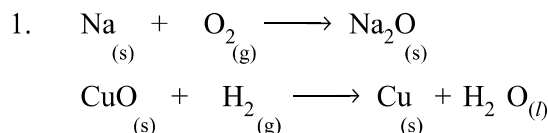


Answer the following questions briefly

(2 mark)

- Write down the observations which indicate the occurrence of a chemical reaction.
- Why is respiration considered as an exothermic reaction? Explain.
- Transfer the following statements into Chemical equations and then balance them.
 - Hydrogen gas combines with nitrogen to form ammonia.
 - Potassium metal reacts with water to give potassium hydroxide and hydrogen gas.
- Identify the substances that are oxidised and the substances that are reduced to the following reactions.



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- What happens when silver chloride is exposed to sunlight? Give one practical application of this reaction. Write the equation also.
- Why is the bag used for potato chips flushed with nitrogen gas?

Answer the following questions in detail

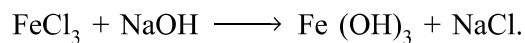
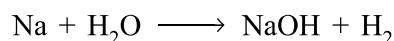
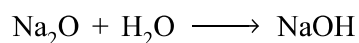
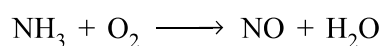
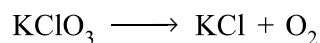
(3 marks)

- Write down the balanced chemical equations for the following reactions.
 - Zinc Carbonate_(s) \longrightarrow Zinc Oxide + Carbon Dioxide_(g)
 - Aluminium_(s) + Chlorine_(g) \longrightarrow Aluminium Chloride_(s)
 - Magnesium_(s) + Water_(l) $\xrightarrow{\text{Heat}}$ Magnesium Hydroxide_(l) + Hydrogen_(g)
- Choose combination, displacement and double displacement reactions out of the given reactions.
 - $\text{MnO}_{2(s)} + 4\text{HCl}_{(l)} \longrightarrow \text{MnCl}_{2(s)} + \text{Cl}_{2(g)} + 2\text{H}_2\text{O}_{(l)}$
 - $\text{CaO}_{(s)} + \text{CO}_{2(g)} \longrightarrow \text{CaCO}_{3(s)}$
 - $2\text{AgCl}_{(s)} \longrightarrow 2\text{Ag}_{(s)} + \text{Cl}_{2(g)}$

3. What happens when $\text{CO}_{2(g)}$ is passed through slaked lime? Write the balanced chemical equation. Write the type of reaction that has occurred.

Explain the following questions detail (5 marks)

1. Balance the following chemical equation and identify the type of reaction they represent



2. Define various types of chemical reactions. Write one chemical equation for each type.

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