

## Ionic Charges Chart (Cations and Anions)

Name:

Date:

## **Cations**

1+		2+		3+	
ammonium	NH <sub>4</sub> +	barium	Ba <sup>2+</sup>	aluminum	<b>AI</b> <sup>3+</sup>
cesium	Cs <sup>+</sup>	beryllium	Be <sup>2+</sup>	chromium(III)	Cr <sup>3+</sup>
		cadmium	Cd <sup>2+</sup>	cobalt(III)	Co <sup>3+</sup>
gold(I)	Au⁺	calcium	Ca <sup>2+</sup>	gold(III)	Au <sup>3+</sup>
hydrogen	H+	cobalt(II)	Co <sup>2+</sup>	iron(III)	Fe <sup>3+</sup>
lead(I)	Pb⁺	copper(II)	Cu <sup>2+</sup>	manganese(III)	Mn³+
lithium	Li+	iron(II)	Fe <sup>2+</sup>		
		lead(II)	Pb <sup>2+</sup>	_	
potassium	K+	magnesium	Mg <sup>2+</sup>	4+	
silver	Ag⁺	manganese(II)	Mn <sup>2+</sup>	tin(IV)	Sn <sup>4+</sup>
sodium	Na⁺	mercury(I)	Hg <sub>2</sub> <sup>2+</sup>	nickel(IV)	Ni <sup>4+</sup>
		mercury(II)	Hg <sup>2+</sup>	lead(IV)	Pb <sup>4+</sup>
copper(I)	Cu⁺	nickel(II)	Ni <sup>2+</sup>		
		strontium	Sr <sup>2+</sup>		
		zinc	Zn²+		
		tin(II)	Sn <sup>2+</sup>		

Roman numeral notation indicates charge of ion when element commonly forms more than one ion. For example, iron(II) has a 2+ charge; iron(III) a 3+ charge.

## **Anions**

amide $NH_2^-$ cyanate $OCN^-$ hydrogen carbonate fluoride $F^-$ chromate $Cr0_4^{2-}$ citrate $C_6H_5O_7^{3-}$ (bicarbonate) $HCO_3^-$ hydride $H^-$ dichromate $Cr_2O_7^{2-}$ ferricyanide $Fe(CN)_6^{3-}$ hydrogen sulfate hydroxide $OH^-$ oxide $O^{2-}$ nitride $O^{2-}$ phosphate $O^{2-}$ bisulfide $O^{2-}$ bisulfide $O^{2-}$ silicate $O^{2-}$ phosphite $O^{2-}$ bisulfite $O^{2-}$ bisulfit	1-			2-		3-		
bromide Br $^{3}$ nitrite N0 $_{2}^{-}$ chlorate Cl0 $_{3}^{-}$ perchlorate Cl0 $_{4}^{-}$ chlorite Cl0 $_{2}^{-}$ permanganate Mn0 $_{4}^{-}$ chloride Cl $^{-}$ thiocyanate SCN $^{-}$ thiosulfate S0 $_{3}^{2-}$ CHEMISTRY EXPER RAKESH KUMAR "CULTIVATING EXCELLENCE EVERY STUDENT" +9198145166	amide hydrogen of (bicarbona hydrogen s (bisulfate) bisulfide bisulfite bromate bromide chlorate chlorite	NH <sub>2</sub> carbonate te) HCO <sub>3</sub> culfate HSO <sub>4</sub> HS- HSO <sub>3</sub> BrO <sub>3</sub> CIO <sub>2</sub>	cyanide cyanate e fluoride hydride hydroxide hypochlorite iodate iodide nitrate nitrite perchlorate	FHOH CIO IO3 I NO3 NO2 CIO4 MnO4	carbonate chromate dichromate oxide oxalate silicate sulfate sulfide sulfite tartrate tetraborate	$CO_3^{2-}$ $CrO_4^{2-}$ $Cr_2O_7^{2-}$ $O^{2-}$ $C_2O_4^{2-}$ $SiO_3^{2-}$ $SO_4^{2-}$ $S^{2-}$ $SO_4^{2-}$ $SO_3^{2-}$ $SO_4^{2-}$ $SO_4^{2-}$ $SO_4^{2-}$	arsenate arsenite citrate ferricyanide nitride phosphate phosphite phosphide  CHEMISTR RAKESH KUM "CULTIVATING EVERY STUDE	AsO <sub>4</sub> 3- AsO <sub>3</sub> 3- C <sub>6</sub> H <sub>5</sub> O <sub>7</sub> 3- Fe(CN) <sub>6</sub> 3- N3- PO <sub>4</sub> 3- PO <sub>3</sub> 3- P3-  EY EXPERT  IAR  EXCELLENCE IN NT"