

CBSE TEST PAPER-02
CLASS - X (Chemical Reactions and equations)

1. Some crystals of copper sulphate were dissolved in water. The colour of the solution obtained would be (1)
(a) green
(b) red
(c) blue
(d) brown.
2. When dilute HCl is added to zinc pieces taken in a test tube (1)
(a) No change takes place
(b) the colour of the solution becomes yellow.
(c) A pungent smelling gas gets liberated.
(d) small bubbles of H_2 gas appear on the surface of zinc pieces
3. PbS reacts with ozone (O_3) and forms $PbSO_4$. As per the balanced equation, (1)
molecules of ozone required for every one molecule of PbS is / are
(a) 4 (b) 3
(c) 2 (d) 1
4. Chemically rust is (1)
(a) Hydrated ferrous oxide
(b) hydrated ferric oxide
(c) only ferric oxide
(d) none of these
5. Which of the following reactions is not correct (1)
(a) $Zn + CuSO_4 \rightarrow ZnSO_4 + Cu$
(b) $2Ag + Cu(NO_3)_2 \rightarrow 2AgNO_3 + Cu$
(c) $Fe + CuSO_4 \rightarrow FeSO_4 + Cu$
(d) $Mg + 2HCl \rightarrow MgCl_2 + H_2$

6. Identify the type of chemical reaction (2)
 - (i) $A \rightarrow B + C$
 - (ii) $AD + CD \rightarrow AD + CB$
7. Why does not silver evolve hydrogen on reacting with dil H_2SO_4 ? (2)
8. Way do diamond and graphite, the two allotropic forms of carbon evolve different amounts of heat on combustion? (2)
9. What is the sole of oxidizing agent is a reaction? (2)
10. (a) Define Rusting (3)
 - (b) Why do you apply paint an iron articles?
11. White the balanced reactions for the following (3)
 - (i) Potassium Bromide (aq) + Barium iodide (aq) \rightarrow Potassium iodide (aq) + Barium Bromide(aq)
 - (ii) Zinc carbonate (s) \rightarrow Zinc oxide (s) + carbon dioxide (g)
 - (iii) Hydrogen (g) + chlorine (g) \rightarrow Hydrogen chloride
12. The reaction is given by (3)

$$Zn + H_2SO_4 \rightarrow ZnSO_4 + H_2$$
 - (i) White the ionic equation for the reaction
 - (ii) The ionic equations can be represented by two half equations.
Write these equations.
 - (iii) Explain why this is a redox reaction
13. What are neutralization reactions? Why are they named so? Give one example? (3)
14. You are given with (5)

(a) Iron Nails	(b) $CuSO_4$ solution
(c) $BaCl_2$	(d) Cu powder
(e) Ferrous sulphate crystal	(f) Quick lime.

Make five reactions that can take place from these materials.