

**CBSE TEST PAPER-01**

**CLASS - X Science (Chemical Reactions and Equations)**

1. Copper displaces which of the following metals from its salt solution: (1)
  - (a)  $\text{ZnSO}_4$
  - (b)  $\text{FeSO}_4$
  - (c)  $\text{AgNO}_3$
  - (d)  $\text{NiSO}_4$
2. In an electrolytic cell where electrolysis is carried, anode has: (1)
  - (a) Positive charge
  - (b) Negative charge
  - (c) Connected to negative terminal of the battery
  - (d) None of these is correct.
3. The reaction  $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$  represents: (1)
  - (a) Oxidation
  - (b) Reduction
  - (c) Decomposition
  - (d) Combination
4. In the reaction  $\text{PbO} + \text{C} \rightarrow \text{Pb} + \text{CO}$  (1)
  - (a) PbO is oxidised
  - (b) C act as an oxidising agent
  - (c) C act as a reduction agent
  - (d) Reaction does not represent redox reaction.
5. A substance which oxidizes itself and reduces other is known as (1)
  - (a) Oxidising agent
  - (b) reducing agent
  - (c) Both (a) and (b)
  - (d) None of these.

6. What happens chemically when quick lime is added to water? (2)
7. Why is a combustion reaction an oxidation reaction? (2)
8. Why are food particles preferably packed in aluminum foil? (2)
9. What happens to lime water when  $\text{CO}_2$  gas is bubbled through it in excess? (2)
10. Identify the type of reaction in the following (3)
  - (a)  $\text{ZnCO}_3 + 2\text{HCl (aq)} \rightarrow \text{ZnCl}_2 \text{ (aq)} + \text{H}_2\text{CO}_3 \text{ (aq)}$
  - (b)  $2\text{NaBr (aq)} + \text{Cl (g)} \rightarrow 2\text{NaCl (aq)} + \text{Br}_2 \text{ (aq)}$
  - (c)  $2\text{CuO (s)} \xrightarrow{\text{Heat}} 2\text{Cu (s)} + \text{O}_2 \text{ (g)}$
11. A student dropped few pieces of marble in dilute hydrochloric acid contained in a test tube. The evolved gas was then passed through lime water. What change would be observed in lime water? Write balanced chemical equation for both the change observed? (3)
12. In the reaction  $\text{MnO}_2 + 4\text{HCl} \rightarrow \text{MnCl}_2 + 2\text{H}_2\text{O} + \text{Cl}_2$  (3)
  - (a) Name the substance oxidised.
  - (b) Name the oxidising agent.
  - (c) Name the reducing agent and the substance reduced.
13. Give one example each of (3)
  - (1) Thermal decomposition
  - (2) Electrolytic decomposition
  - (3) Photo decomposition
14. A metal is heated with dil  $\text{H}_2\text{SO}_4$ . The gas evolved is collected by the method shown in the figure : Answer the following (5)
  - (a) Name the gas.
  - (b) Name the method of collection of gas.
  - (c) Is the gas soluble or insoluble in water?
  - (d) Is the gas lighter or heavier than air?

