

Climate change mitigation: Turning carbon dioxide into rock

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An international team of scientists have found a potentially viable way to remove anthropogenic (caused or influenced by humans) carbon dioxide emissions from the atmosphere - turn it into rock.

The study, published in *Science*, has shown for the first time that the greenhouse gas carbon dioxide (CO₂) can be permanently and rapidly locked away from the atmosphere, by <u>injecting</u> it into volcanic bedrock. The CO₂ reacts with the <u>surrounding</u> rock, forming environmentally benign minerals.

Measures to tackle the problem of <u>increasing</u> greenhouse gas emissions and the resultant global <u>warming</u> are numerous. One approach is Carbon Capture and Storage (CCS), where CO₂ is physically removed from the atmosphere and trapped underground. Geoengineers have long explored the possibility of sealing CO₂ gas in voids underground, such as in abandoned oil and gas reservoirs, but these are susceptible to leakage. So, attention has now turned to the mineralisation of carbon to permanently dispose of CO₂.



Section of rock core from the CO2 storage reservoir showing vesicular basalt with a well-defined fracture with calcium carbonate mineralisation (Annette K Mortensen)

Until now it was thought that this process would take several hundreds to thousands of years and is therefore not a practical option. But the current study -- led by Columbia University, University of Iceland, University of Toulouse and Reykjavik Energy -- has demonstrated that it can take as little as two years.

Lead author Dr Juerg Matter, Associate Professor in Geoengineering at the University of Southampton, says: "Our results show that between 95 and 98 per cent of the injected CO₂ was mineralised over the period of less than two years, which is amazingly fast."

The gas was injected into a deep well at the study site in Iceland. As a volcanic island, Iceland is made up of 90 per cent basalt, a rock rich in elements such as calcium, magnesium and iron that are required for carbon mineralisation. **The CO₂ is dissolved in water and carried down the well.** On contact with the target storage rocks, at 400-800 metres under the ground, the solution quickly reacts with the surrounding basaltic rock, forming carbonate minerals.

"Carbonate minerals do not leak out of the ground; thus, our newly developed method results in permanent and environmentally friendly storage of CO₂ emissions," says Dr Matter, who is also a member of the University's Southampton Marine and Maritime Institute and Adjunct Senior Scientist at Lamont-

Doherty Earth Observatory Columbia University. "On the other hand, basalt is one of the most common rock types on Earth, potentially providing one of the largest CO₂ storage capacity."

To monitor what was <u>happening</u> underground, the team also injected 'tracers', chemical compounds that literally trace the transport path and reactivity of the CO_2 . There were eight <u>monitoring</u> wells at the study site, where they could test how the chemical composition of the water had changed. The researchers discovered that by the time the groundwater had migrated to the monitoring wells, the concentration of the tracers -- and therefore the CO_2 -- had diminished, indicating that mineralisation had occurred.

"Storing CO₂ as carbonate minerals significantly enhances storage security which should improve public acceptance of Carbon Capture and Storage as a climate change mitigation technology," says Dr Matter.

"The overall scale of our study was relatively small. So, the obvious next step for CarbFix is to upscale CO₂ storage in basalt. This is currently <u>happening</u> at Reykjavik Energy's Hellisheidi geothermal power plant, where up to 5,000 tonnes of CO₂ per year are captured and stored in a basaltic reservoir."

The investigation is part of the CarbFix project, a European Commission and U.S. Department of Energy funded programme to develop ways to store anthropogenic CO₂ in basaltic rocks through field, laboratory and <u>modelling</u> studies (<u>http://carbfix.com</u>).

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<u>Materials</u> provided by <u>University of Southampton</u>. *Note: Content may be edited for style and length.*

Journal Reference:

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