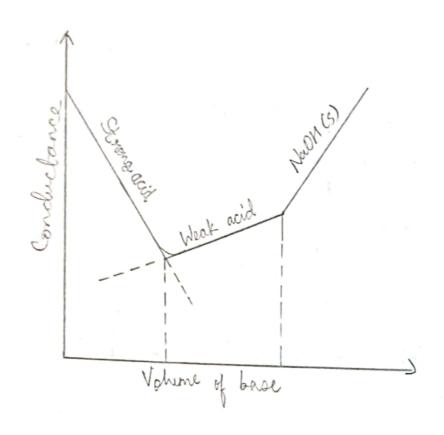
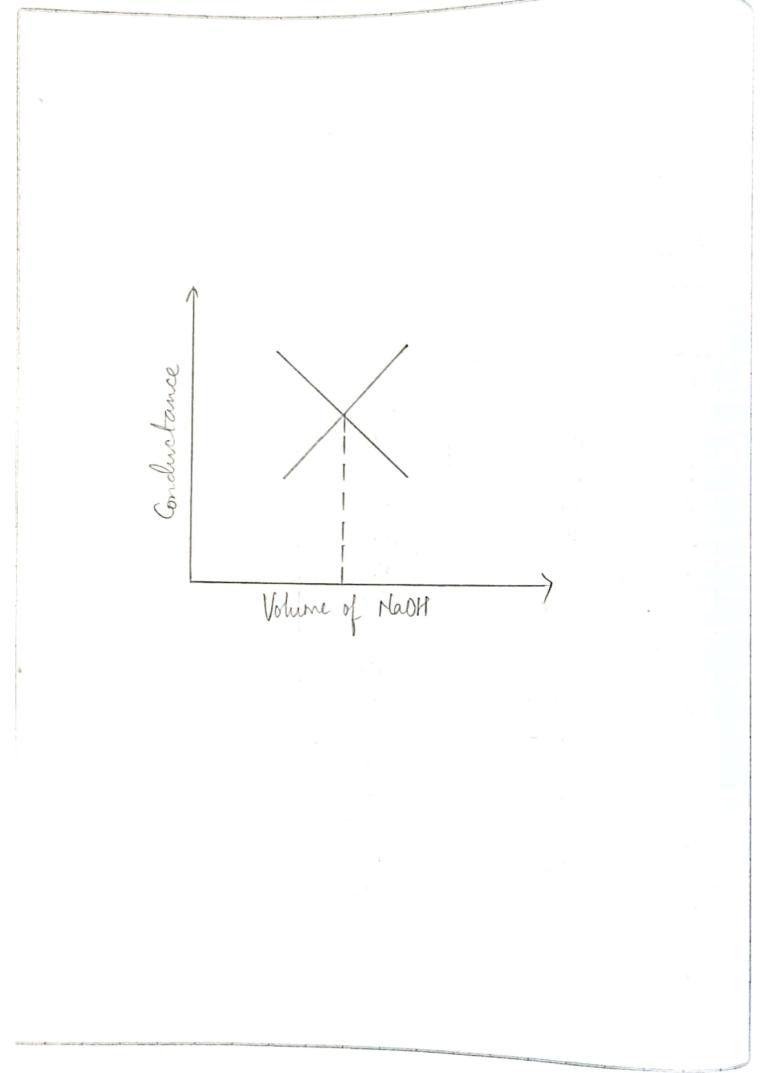
_	Date
Expt	No. 7 Page No
	DETERMINATION OF THE STRENGTH OF A MIXTURE OF ACETIC ACID
	COLD HADBOCHADAS ASSESSED THE STREETS IN DE A PILL TOPE DE METTE
	AND HYDROCHLORIC ACID BY CONDUCTOMETRY.
74	AIM:
7	To estimate the strength of the nixture of acetic acid and
	To estimate the strength of the nixture of acetic acid and hydrochloric acid present in a given mixture by conductometrically.
	conductometrically.
4	PRINCIPLE:
7	The conductivity of the solution is related to the mobility
-	I imme which it to an related with the size of the
_	The conductivity of the solution is related to the mobility of ions which in turn related with the size of the
	ABTN.
7	NaOH + HCl -> Nacl + HD
	CH3COOH + Na+ + DH> CH3COO-+Na+ +420
X	PROCEDURE:
->	The given mixture of acids is made up to 100 ml using
	distilled water.
\rightarrow	10 ml of this made up solution is pipetted out intralean
	beaker and 100 ml of distilled water is added.
-	The conductively cell is dipped into the test solution and
	tiliated against NaOH (0.5 ml) interval with proper stiring.
7	The conductance is measured after each 0.5 ml addition of NaOH at various stages of neutralization.
	NaOH at various stages of neutralization.
7	After complete neutralization, the amount of acid present in
	After complete neutralization, the amount of acid present in the given mixture is determined based on the volume of
	NAOW amagined
7	Volume of base consumed for strong acid and weak acid are determined by platting a graph between conductance
	are determined by plotting a graph between conductance
	Teacher's Signature

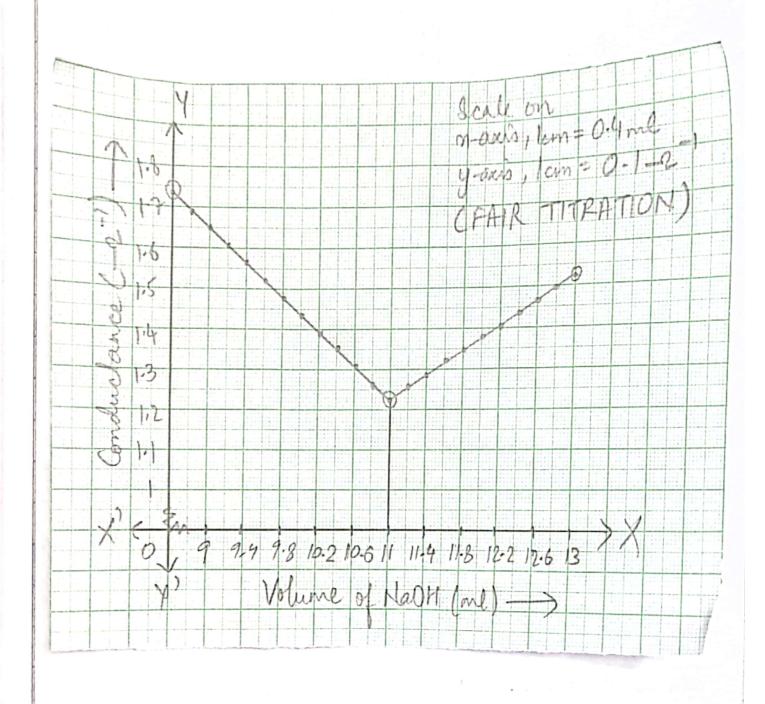


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EXI	t. No			Page No.			
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	1	, , ,	111	. I boint corresponds			
	and volume of base added, where first end point corresponds to strong acid and second end point corresponds to weak acid.						
	to strong acid and second end foint correspondes						
	weak acid.						
*	+ OBSERVATIONS:						
	0.1	Volume of NaOH	Conductance	(TABLE 1)			
	S.No.	added (ml)	(-2-1)	(FAIR TITRATION)			
		a	1.736				
	2	9.2	1.670				
_	3	9.4	1.605				
_	4	9.6	1.523				
_	5	9.8	1.503				
_	6	10	1.424				
_	7	10.2	1-363				
-	8	10.4	1.320				
	9	10.6	1.282				
_	10	10.8	1.270				
	11	11	1.236				
\dashv	12	11.2	1.260				
	13	11.4	1.302				
+	14	11.6	1.318				
+	15	11.8	1.370				
+	16	12	1.383				
1	17	12.2	1.387				
+	18	12.4	1.449				
+	19	12.5	1.457				
	-						
	Teacher's Signature						



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		A second of the		
Is.No.	Volume of NaOH	Conductance	(TABLE 2)	
	Volume of NaOH added (ml)	Conductance	(PILOT TITRATION)	
1	0	4.4		
2		4.2		
3	2_	3.8		
4	3	3-5	J'-	
6	4	3.2		
6	5	2-9	,	
7	6	2.5		
8	7	2.3		
9	8	2		
10	9	1-853		
11	lo	1.498		
12	11	1.215		
13	12 /	1.283		
14	13	1.40	3 , 1 1	
15	14	1.552		
16	15	1.704		
17	16	1.843		
18	17	1.973		
in [
001 0011	ATIONS:			
1/1	$\frac{1}{\sqrt{ V }} \int \frac{ V }{\sqrt{ V }} dV = \frac{1}{\sqrt{ V }} dV = \frac{1}{\sqrt{ V }} \int \frac{ V }{\sqrt{ V }} dV = \frac{1}{\sqrt{ V }} dV = \frac{1}{$	Oml	10 (10 × 0.11 ×	
Jeun	1:1 1 110 (N.	1= V2N2 =>	11 XO:1 = 0.11 N	
Horma	lily of All City	V	10	
	0 /	= Il ml (he	om graph)	
Volum	e of NON (V2)	-11 FOC C/S	1	
Volume of NaOH (V2) = 11 ml (from graph) Normality of NaOH (N2) = 0.1 N Strength of HCl acid = 0.11N Teacher's Signature				
strength of HCl acid = 0.111/				
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* CIRAPH:



* CORAPH: 5.0 4.0 8 10 12 14 16 18 Volume of MaOH (ml)-

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	A = 0 · 0 · T ·	·
*	PESULI:	
->	The strength of	giness HCl solution = 0.11 N/
		U
		•
		X
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