Rate Law Expressions

The rate law expresses the rate of a reaction as a function of the concentration of its reactants.

First Order Reactions

Graphical methods involve plotting data to determine the order of reaction from the slope.

Second Order Reactions

First-order reactions have a constant half-life and follow exponential decay.

Arrhenius Equation

The Arrhenius equation shows the temperature dependence of the rate constant:

$k = A * e^{(-Ea/RT)}$