


WALCHAND COLLEGE OF ENGINEERING, SANGLI.
(An Autonomous Institute)

First Year B.Tech. (Civil, Electrical & Electronics)
END SEMESTER EXAMINATION SEM. I DECEMBER-2017
ENGINEERING CHEMISTRY (3CH101)

ESE

Print Question Paper

SELECT COURSE CO
3CH101

Exam Seat Number: _____

FACULTY

Dr. Pendse M.H. - Powar

 Max Marks: **50**
IMP: Verify that you have received question paper with correct course, code, branch etc.

- i) All questions are compulsory. Writing question number is compulsory. The answers may not be assessed if question number is not written.
- ii) Figures to the right of question text indicate full marks.
- iii) Assume suitable data wherever necessary, Write the answers with neat handwriting.
- iv) Only FX82 series non programmable Calculator is allowed.

Students 208

Number of Prints
(Regular + Re-reg
+ Extra)

			Marks	Options
Obj	Q1 A)	What are Primary Standard? Discuss criteria/Requirements for choosing primary standard substances towards preparation of standard solutions used in titrimetry.	6 CO3	Module 1
An-S	Q1 B)	List different sources of natural water and compare them with respect to Purity, Impurities present and their significance. --OR-- Define hardness of water? Why it is caused. List different units used to measure hardness of water with interrelation among them.	6 CO2	Module 2
Desc	Q1 C)	With neat labeled diagram, discuss Pb-Ag system.	6 CO1	Module 3
Obj	Q2 A)	Define alloy and discuss with example purposes of making alloys. --OR-- What are carbon steels? How are they classified? Give composition, properties and uses of High Carbon Steel.	4 CO3	Module 4
Obj	Q2 B)	State compound composition of Portland cement and mention functions of each constituent in setting and hardening of Portland cement with reactions.	6 CO3	Module 4

Desc	Q3 A)	With neat labeled diagram discuss construction and working of TGA equipment. With proper example describe TGA thermogram. List at least four applications.	10	CO1	Module 6
An-M	Q3 B)	<p>Following data was recorded while determining calorific value of gaseous fuel by Boy's gas calorimeter.</p> <p>i) Volume of gas burnt at STP = 0.09m³</p> <p>ii) Mass of water used in time 't' = 24 Kg</p> <p>iii) Temperature of inlet water = 22 °C</p> <p>iv) Temperature of outlet water = 35 °C</p> <p>v) Mass of steam condensed = 0.03 Kg</p> <p>Calculate Higher and Lower calorific value.</p> <p>(Given: Latent heat of condensation of water vapour 587 Kcal/Kg)</p>	3	CO2	Module 5
An-M	Q3 C)	<p>A sample of coal containing 5% hydrogen was tested in Bomb Calorimeter for its calorific value, following data were recorded. Weight of coal burnt = 0.98gm, Acid correction = 55 Cal. Rise in temperature = 2.52 °C, Water equivalent of bomb & calorimeter = 550 gm, Weight of water taken in copper calorimeter = 2200 gm</p> <p>Latent heat of condensation of steam = 587 cal/gm</p> <p>Calculate Gross & Net calorific value.</p>	4	CO2	Module 5
An-M	Q3 D)	Compare solid and gaseous fuels with respect to Calorific value, Ignition temperature, Velocity of combustion, Control of combustion, Risk of fire hazard, use for IC engines, Transport and storage, Products of combustion.	5	CO2	Module 2

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