$$T(\kappa) = \theta(^\circ C) + 273$$

$$E = mc\Delta\theta \qquad E = mL$$

internal energy

mass

specific heat capacity

$\to$ different for each material and its state

$$P\underset{\text{volume}}{V} \simeq \frac{1}{3} N \underset{\text{mass of particle}}{m} \overline{c^2} \to \text{mean of the square of the velocity}$$

pressure

number particles

$$\frac{P}{T} = \text{const} \qquad \frac{V}{T} = \text{const} \qquad PV = \text{const}$$

$$PV = nRT = N\kappa T$$

$\to$ Molar gass constant

$\to$ Boltzman constant

$$\frac{1}{2}m\overline{c^2} = \left(\frac{3}{2}\kappa\right)T$$

(const

$$\to Ke \propto T$$

Test : Monday 8th September