1 Answers

1.1 Unit 1

Problem 1

The molar mass of $\mathrm{C_9H_8O_4}$ is $1.008 * 8 + 12.01 * 9 + 16.00 * 4 = 180.2\frac{g}{mol}$

$$7.89kg \times \frac{1g}{10^{-3}kg} \times \frac{1mol}{180.2g} = 43.8mol \tag{1}$$

Problem 2

(b), the tallest peak of the graph is the one at $64u$.

Problem 3

In one mole of $\mathrm{C_{13}H_{18}O_2}$ is $206.31g$.

$$1mol\,\mathrm{C_{13}H_{18}O_2} \times \frac{13mol\,\mathrm{C}}{1mol\mathrm{C_{13}H_{18}O_2}} \times \frac{12.01g}{1mol\,\mathrm{C}} = 156.31g \tag{2}$$

Thus, the percent composition by weight is $\frac{156.31}{206.31} = 75.764\%$

Problem 4

Take $100g$ of the substance such that there are $32.38g$ sodium, $22.65g$ sulfur, and $44.99g$ oxygen.

$$\begin{aligned} 32.38g\,\mathrm{Na} \times \frac{1mol\,\mathrm{Na}}{22.99g} &= 1.408mol\,\mathrm{Na} \\ 22.65\,g\mathrm{S} * \frac{1\,mol\mathrm{S}}{32.07g} &= 0.7063\,mol\mathrm{S} \\ 44.99\,g\mathrm{O} * \frac{1\,mol\mathrm{O}}{16g} &= 2.812\,mol\mathrm{O} \end{aligned} \tag{3}$$

Take the ratio of each compound with the smallest quantity.

$$\begin{aligned} S : \frac{0.7063}{0.7063} &= 1 \\ Na : \frac{1.408}{0.7063} &= 2 \\ O : \frac{2.812}{0.7063} &= 4 \end{aligned} \tag{4}$$

Therefore, the empirical formula is Na_2SO_4

Problem 5

$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10}$

Problem 6

Be. The peak location of the peak on the x-axis means that there is less binding energy for the electrons in element X. Be has fewer protons and both electrons are in the same shell, so it peak must belong to Be.

Problem 7

- The electronegativity increases from left to right across a period. This is because if a valence shell of electrons is less than half full than it requires less energy to lose an electron than gain one. If if the valence shell of electrons is more than half full, it is easier to pull and electron into the valence shell. The electronegativity decreases from the top to the bottom of a group. This is beause there is a greater atomic radius lower on the group.
- The ionization energy increases from left to right in a period. This is because of greater valence shell stability also because of smaller atomic radius. The ionization energy also decreases from top to bottom of a group. This is because of greater electron shielding and greater atomic radius.
- Atomic radius decreases from left to right within a period. This is because there are more protons to the right of the period. Atomic radius increases from top to bottom within a group. This is because of electron shielding and there are more electron shells in the atom.

1.2 Unit 2

Problem 8

The ionic character increase the greater the electronegativity difference. In this case, Na and O had the greatest electronegativity difference.

Problem 9

(c) $\mathrm{Cl} - \mathrm{F} > \mathrm{H} - \mathrm{I} > \mathrm{Se} - \mathrm{N}$

Problem 10

Cs^+ has a larger atomic radius than K^+. So the distance between the cation and anion is greater than in CsF than in KF

Problem 11

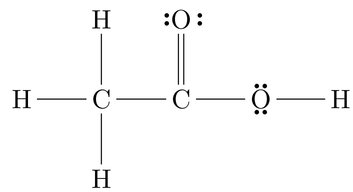
Most metallic elements form crystalline solids at room temperature. Their bonds are metallic bonds due to electrostatic attraction between metal cations and delocalized electrons.

Problem 12

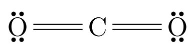
- Substitutional alloys. These alloys form when one atom of a similar size to the host metal replaces an atom of the host metal. The substitute atom must be of similar size. These alloys have good thermal and electrical conductivity.

- Interstitial alloys. These alloys are formed when smaller atoms fill in the gaps between the larger host atoms. This makes the metal harder and less malleable.

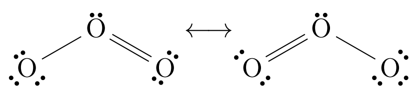
Problem 13



Problem 14

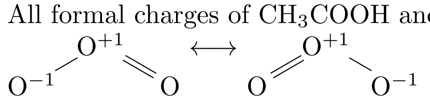


Problem 15



Problem 16

All formal charges of CH_3COOH and CO_2 are zero.



Problem 17

The electron geometry is tetrahedral. The molecular geometry is trigonal pyramidal. Hybridization of N atom is sp^3 since it has tetrahedral electron geometry.

1.3 Unit 3

Problem 18

Dipole-diple and london dispersion forces. The C — O bond is polar and the molecule is asymmetrical so it is polar. There are no H — F, H — O, or H — N bonds, so there is no hydrogen bonding.

Problem 19

Condensation. Both have no regular arrangement, the one on the left is separated by the one on the right is close together, so the molecules are transitioning from gas to liquid.

Problem 20

(a) As you increase the temperature, the average kinetic energy or speed increases as well.
(b) $n_1 = n_2$ and $V_1 = V_2$, so $P_1V_1 = n_1RT_1$ and $P_2V_2 = n_2RT_2$. $\frac{P_2V_2}{RT_2} = \frac{P_1V_1}{RT_1} \implies P_2 = \frac{P_1T_2}{T_1} = \frac{0.7*425}{299} = 0.99atm$
(c) As the temperature increases, the average kinetic energy increases, so the molecules undergo more collisions with the walls of the container.

Problem 21

$$60.3g \times \frac{1mol\,\mathrm{Ba(OH)_2}}{171.35g} \times \frac{2mol\,\mathrm{OH^-}}{1mol\,\mathrm{Ba(OH)_2}} * \frac{1}{1.75L} = 0.402M$$

Problem 22

The photoelectric effect occurs when light of a certain minimum frequency/energy hits the surface of a metal and electrons are ejected.

1.4 Unit 4

Problem 23

$2\,C_5H_{10} + 15\,O_2 \longrightarrow 10\,CO_2 + 10\,H_2O$

Problem 24

The oxidation reaction:

$$\begin{aligned}
&I^- \longrightarrow I_2 \\
&2\,I^- \longrightarrow I_2 \\
&2\,I^- \longrightarrow I_2 + 2\,e^- \\
&10\,I^- \longrightarrow 5\,I_2 + 10\,e^-
\end{aligned}$$

The reduction reaction:

$$\mathrm{MnO_4^-} \longrightarrow \mathrm{Mn^{2+}}$$
$$\mathrm{MnO_4^-} \longrightarrow \mathrm{Mn^{2+}} + 4\,\mathrm{H_2O}$$
$$\mathrm{MnO_4^-} + 8\,\mathrm{H^+} \longrightarrow \mathrm{Mn^{2+}} + 4\,\mathrm{H_2O}$$
$$5\,\mathrm{e^-} + \mathrm{MnO_4^-} + 8\,\mathrm{H^+} \longrightarrow \mathrm{Mn^{2+}} + 4\,\mathrm{H_2O}$$
$$10\,\mathrm{e^-} + 2\,\mathrm{MnO_4^-} + 16\,\mathrm{H^+} \longrightarrow 2\,\mathrm{Mn^{2+}} + 8\,\mathrm{H_2O}$$

Add the reactions together:
$10\,\mathrm{I^-} + 16\,\mathrm{H^+} + 2\,\mathrm{MnO_4^-} \longrightarrow 5\,\mathrm{I_2} + 2\,\mathrm{Mn^{2+}} + 8\,\mathrm{H_2O}$

Problem 25

$\mathrm{Fe^{3+}(aq)} + 3\,\mathrm{OH^-(aq)} \longrightarrow \mathrm{Fe(OH)_3(s)}$

Problem 26

Chemical processes are characterized by changes in intramolecular forces, while physical processes are characterized by changes only in intermolecular forces.

Problem 27

Find the limiting reactant:

$$\begin{aligned} 36.0\,gH_2O * \frac{1\,molH_2O}{18.02\,gH_2O} * \frac{1\,molFe_3O_4}{4\,molH_2O} &= 0.49945\,molFe_3O_4 \\ 67.0\,gFe * \frac{1\,molFe}{55.85\,gFe} * \frac{1\,molFe_3O_4}{3\,molFe} &= 0.39988\,molFe_3O_4 \end{aligned} \tag{5}$$

Therefore, the limiting reactant is $\mathrm{Fe_3}$
Find how much iron oxide is produced: Use the limiting reactant

$$0.39988\,mol\mathrm{Fe_3} * \frac{231.55\,gFe_3O_4}{1\,mol\mathrm{Fe_3}} = 92.6\,gFe_3O_4 \tag{6}$$

Find how much excess reactant is left over:

$$67.0\,gFe * \frac{1\,molFe}{55.85\,gFe} * \frac{4\,molH_2O}{3\,molFe} * \frac{18.02\,gH_2O}{1\,molH_2O} = 28.8\,gH_2O \tag{7}$$

28.8 grams of water is used out of 32.0 grams. So there is 7.2 grams left over of the excess reagent.

Problem 28

$1.5 \times 10^3 mol$

Problem 29

$CH_4 \quad$ rate $= \frac{-\Delta[\mathrm{CH_4}]}{\Delta t}$
$O_2 \quad$ rate $= \frac{-1}{2}\frac{\Delta[\mathrm{O_2}]}{\Delta t}$
$CO_2 \quad$ rate $= \frac{\Delta[\mathrm{CO_2}]}{\Delta t}$
$H_2O \quad$ rate $= \frac{1}{2}\frac{\Delta[\mathrm{H_2O}]}{\Delta t}$

Problem 30

$10\frac{M}{s}$

Problem 31

$\frac{(0.04-0.1)M}{125ms} \times \frac{1ms}{10^{-3}s} = -0.48\frac{M}{s}$
rate $= -\frac{1}{2} \times -0.48 = 0.24\frac{M}{s}$

Problem 32

When concentration of B is held constant and the concentration of A is tripled, the initial rate is multiplied by 9 so the order with respect to A is 2. The order with respect to B is 0 because nothing changes when the concetration is increased. Thus the rate law is

$$rate = k[A]^2[B]^0 = k[A]$$

$1.0 \times 10^{-2} = k \times 0.1^2 \implies k = 1.0\frac{1}{ms}$, thus $rate = [A]^2$

Problem 33

$$ln[A]_t = -kt + ln[A]_0$$
$$ln[A]_t = -(4.8 \times 10^{-4})(825) + ln(0.0165)$$
$$[A]_t = e^{-4.50} = 0.0111M$$

Half-life:

$$ln(\frac{1}{2}[A]_0) - ln[A]_0 = -kt$$
$$ln(\frac{1}{2}) = -kt$$
$$t = \frac{ln(2)}{k}$$

Problem 34

$k = 4.0 \times 10^{-4}\frac{1}{Ms} = 35\frac{1}{M \times days}$
$\frac{1}{[A]_t} = 35 * 6 + \frac{1}{0.1} \implies [A]_t = 4.5 \times 10^{-3}$

Problem 35

$\frac{1}{0.085M} = 35t + \frac{1}{0.1M} \implies t = 0.05days$

Problem 36

$2\,H_2O_2 \longrightarrow 2\,H_2O + O_2$ The rate determining step is the slow reaction, so $rate = k[H_2O_2][I^-]$
I^- is the catalyst. IO^- is the intermediate.

Problem 37

Overall: $2\,NO + 2\,H_2 \longrightarrow 2\,H_2O + N_2$
Rate determining step: $rate = k[N_2O_2][H_2]$
Fast equilbium: rate forward = rate back $\implies [N_2O_2] = \frac{k_f}{k_r}[NO]^2$
Thus, $rate = \frac{kk_f}{k_r}[NO]^2[H_2]$

1.5 Unit 6

Problem 38

$50gH_2O \times \frac{1molH_2O}{18.01gH_2O} \times \frac{6.022\times10^{23}H_2O}{1molH_2O} \times \frac{2O-H}{1H_2O} \times \frac{1.8\times10^{-9}cal}{1O-H} \times \frac{4.184J}{1cal} \times \frac{1kJ}{10^3J} = 2500kJ$

Problem 39

$q_{water} = 7.3 \times 100 \times 4.184$
$q_{water} = -q_{metal}$
$q_{metal} = C_{metal}(27.3 - 100) * 120 \implies C_{metal} = 0.350\frac{J}{g°C}$

Problem 40

$5.0LH_2O \times \frac{1mLH_2O}{10^{-3}LH_2O} \times \frac{1gH_2O}{1mL} \times \frac{1mol}{18.01g} \times \frac{40.72kJ}{1mol} = 11302.32kJ$

Problem 41

$$\begin{aligned} &C_2H_2 + 2\,H_2 \longrightarrow C_2H_6 \\ &C_2H_6 \longrightarrow C_2H_4 + H_2 \qquad \implies C_2H_2 + H_2 \longrightarrow C_2H_4 \end{aligned}$$

Thus, $\Delta H = -175$

Problem 42

$1.5molO_2 \times \frac{1mol_{rxn}}{3molO_2} \times \frac{-1371kJ}{mol_{rxn}} = -685.5kJ$

Problem 43

Increase

Problem 44

$\Delta S = 205.0 + 2 \times 130.58 - 2 \times 188.83 = 88.5\frac{J}{K}$

Problem 45

$\Delta S_{surr} = \frac{-\Delta H}{T} = -\frac{-802.2*1000}{298.15} = 2691\frac{J}{K}$
$\Delta S_{sys} = 2 \times 188.7 + 213.7 - (186.1 + 2 \times 205.0) = -5.0\frac{J}{K}$
$\Delta S_{universe} = -5 + 2691 = 2686\frac{J}{K}$.

Problem 46

$\Delta G = -19800J - (1000) * (-197.3\frac{J}{K}) = 106000J$

Problem 47

More product at high temperature

1.6 Unit 7

Problem 48

$K_c = \frac{[\mathrm{CH_4}][\mathrm{H_2O}]}{[\mathrm{H_2}]^2[\mathrm{CO}]}$

Problem 49

$Q_c = 6.25$. $Q_c > K_c$ so the reaction will shift towards reactants.

Problem 50

$K_p = \frac{(3x)^3 x}{(14-2x)^2} \implies 2.0 \times 10^{-6} \approx \frac{(3x)^3 x}{14^2} \implies x = 0.062atm$. The partial pressure is $0.062atm$.

Problem 51

$x^2 = 0.16 \implies x = 0.4M$. Note that $250g = 4.89mol$ which is clearly enough to produce the $0.4M$ predicted by the equilibrium constant.