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# **Electronic Supplementary Information**

Toward a mechanistic understanding of oxidative homocoupling: the Glaser–Hay reaction

Jesús Jover, Philipp Spuhler, Ligang Zhao, Ciarán McArdle and Feliu Maseras

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**Table S1.** Computed potential energies, enthalpies, free energies and dispersion correction terms for all the species involved in the catalytic cycles (in Hartrees).

Species	Potential	Enthalpy	Free energy	Dispersion
Phenylacetylene	-308.076609	-307.963100	-308.000964	-0.004231
Diphenyldiacetylene	-614.994097	-614.784693	-614.840135	-0.009103
$\mathrm{O}_2$	-150.236349	-150.229615	-150.252925	-0.000001
H <sub>2</sub> O	-76.380900	-76.356823	-76.378306	-0.000004
TMEDA	-347.371068	-347.145162	-347.190952	-0.011080
TMEDA-H <sup>+</sup>	-347.843450	-347.601824	-347.647779	-0.012538
$[Fe(CN)_6]^{3-}$	-1820.623761	-1820.560194	-1820.615095	-0.008841
$[Fe(CN)_6]^{4-}$	-1820.765989	-1820.703054	-1820.756364	-0.008848
G1	-544.632739	-544.401947	-544.451771	-0.016987
G2	-852.764499	-852.417613	-852.483578	-0.029129
DP_TS	-1200.116501	-1199.545258	-1199.638877	-0.047954
G3	-852.289850	-851.954905	-852.023633	-0.023680
G4	-1704.583454	-1703.912235	-1704.028188	-0.064477
G5 (O16)	-1704.303944	-1703.626756	-1703.733259	-0.070199
RedEl_TS	-1704.292304	-1703.617346	-1703.724353	-0.068613
NH <sub>3</sub>	-56.511871	-56.474881	-56.497793	-0.000017
DP_TS-NH <sub>3</sub>	-909.253106	-908.871636	-908.946675	-0.027978
OH.	-75.858236	-75.846758	-75.866347	0.000000
DP_TS-OH	-928.627826	-928.271280	-928.345091	-0.028344
G2-pF-Phenylacetylene	-951.949721	-951.610075	-951.678028	-0.029534
DP_TS-pF-Phenylacetylene	-1299.302346	-1298.738111	-1298.834019	-0.048617
G2-pMe-Phenylacetylene	-892.036460	-891.661397	-891.730372	-0.030836
DP_TS-pMe-Phenylacetylene	-1239.388683	-1238.789437	-1238.884475	-0.049468
G1'	-544.473346	-544.241250	-544.290268	-0.017690
G2'	-852.568338	-852.220944	-852.288122	-0.028240

DP_TS'	-1199.950183	-1199.377050	-1199.472487	-0.044945
G3'	-852.135562	-851.798707	-851.865688	-0.024733
O4	-1002.526278	-1002.182337	-1002.255008	-0.028790
O5	-1854.833632	-1854.152278	-1854.269971	-0.065396
O6	-1854.834450	-1854.152981	-1854.271199	-0.061477
O7	-1855.297757	-1854.600007	-1854.717081	-0.062345
O8	-1855.310415	-1854.615085	-1854.730722	-0.069054
O9	-1855.305268	-1854.610517	-1854.725853	-0.072512
O10	-1855.307775	-1854.613188	-1854.730294	-0.067988
O11	-1855.767485	-1855.057259	-1855.175492	-0.069788
O12	-1855.772433	-1855.062694	-1855.176076	-0.075707
O13	-927.898358	-927.545279	-927.613204	-0.028618
O14	-1780.197782	-1779.508445	-1779.622771	-0.071448
O15	-1780.206929	-1779.517973	-1779.632752	-0.074538

**Table S2.** Computed potential energy, enthalpy and free energy profiles (in kcal/mol) for the catalytic cycles, dispersion corrections are included.

Species	Potential	Enthalpy	Free energy
G1	0.0	0.0	0.0
G2	-39.6	-38.0	-24.3
DP_TS	-32.5	-31.8	-6.8
G3	-35.7	-31.7	-19.9
G4	-42.2 (-84.4)	-37.8 (-75.6)	-19.2 (-38.4)
G5 (O16)	-45.5 (-91.0)	-39.7 (-79.4)	-17.2 (-34.4)
RedEl_TS	-41.4 (-82.8)	-33.2 (-72.4)	-13.9 (-27.8)
G1'	0.0	0.0	0.0
G2'	-15.5	-14.4	-2.0
DP_TS'	-26.8	-24.8	-1.4
G3'	-39.1	-34.7	-22.3
G5	-55.9 (-111.8)	-50.4 (-100.8)	-29.4 (-58.8)
RedEl_TS	-51.7 (-103.4)	-47.0 (-94.0)	-26.1 (-52.2)
G1	0.0	0.0	0.0
G2	-39.6	-38.0	-24.3
DP_TS	-32.5	-31.8	-6.8
G3	-35.7	-31.7	-19.9
O4	-38.9	-33.5	-9.5
O5	-46.8 (-93.6)	-41.3 (-82.6)	-16.0 (-32.0)
O6	-45.9 (-91.8)	-40.3 (-80.6)	-15.2 (-30.4)
O7	-42.8 (-85.6)	-37.1 (-74.2)	-11.6 (-23.2)
O8	-48.9 (-97.8)	-44.0 (-88.0)	-18.0 (-36.0)
O9	-48.4 (-96.8)	-43.6 (-87.2)	-17.5 (-35.0)
O10	-47.7 (-95.4)	-43.0 (-86.0)	-17.5 (-35.0)

011	-43.9 (-87.8)	-39.2 (-78.4)	-14.0 (-82.0)
O12	-47.3 (-94.6)	-42.8 (-85.6)	-16.0 (-32.0)
O13	-49.1	-45.7	-26.0
O14	-55.2 (-110.4)	-51.2 (-102.4)	-24.5 (-49.0)
O15	-51.4 (-102.8)	-47.3 (-94.6)	-28.6 (-57.2)
G5	-55.1 (-110.2)	-52.2 (-104.4)	-33.7 (-67.4)
RedEl_TS	-51.0 (-102.0)	-48.8 (-97.6)	-30.4 (-60.8)

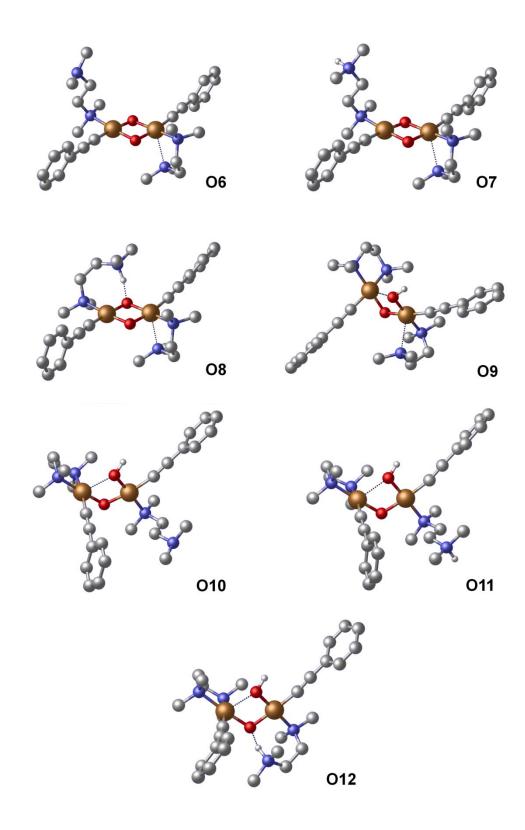


Figure S1. Detailed structures of intermediates O6-O12.

## Outer-sphere catalytic cycle

An alternative reaction pathway, based on outer-sphere mononuclear oxidations, has also been computed since it is known that some Cu/O<sub>2</sub> species act in this fashion. This catalytic cycle consists mainly of a sequence of one-electron transfers and protonation of the derived oxygenated species (Scheme S1).

**Scheme S1.** Proposed outer-sphere mechanism for the Glaser-Hay coupling including the computed free energy values in kcal/mol (Cu(I), Cu(II), Cu(III); for clarity phenylacetylene is depicted as  $\equiv$ ).

The catalytic cycle starts with the 2-electron-transfer from G1 to dioxygen to yield the copper (III)- $\eta^2$ -peroxo intermediate **OS1**. This stage involves also a spin-crossing from triplet to singlet, probably through a low-energy minimum energy crossing point (MECP). In OS1 the O-O distance has been elongated to 1.39 Å, matching the experimentally determined distance for this kind of peroxo complexes. This complex is then capable to act as an outer-sphere oxidant for other copper (I) species such as G3, which can be obtained by acetylene coordination to G1 and subsequent deprotonation with a free TMEDA ligand; it has to be remembered that the barrier for this process is 17.5 kcal/mol. The electron transfer between **OS1** and **G3** affords the copper (II) species G3' and the  $\eta^2$ -peroxocopper (II) OS2, this process requires only 10.1 kcal/mol. The peroxo moiety is then protonated by a TMEDA-H<sup>+</sup>, obtained in the formation of G3, to yield the  $\eta^{1}$ -hydroperoxocopper (II) intermediate **OS3**. This latter compound can, in turn, oxidize another equivalent of G3 to G3' giving rise to the  $\eta^1$ -hydroperoxocopper (I) complex OS4; again, the barrier for this electron-transfer is quite low (11.5 kcal/mol). The protonation of **OS4** yields hydrogen peroxide and the starting catalyst G1; those two reactants can then undergo oxidative addition to form the dihydroxycopper (III) complex **OS5**. As may be seen the free energy of the transition state mediating this stage (OA\_H<sub>2</sub>O<sub>2</sub>) is less than 1 kcal/mol higher than the separated reactants, indicating that this process should be very fast. OS5 is another species capable of oxidizing G3 to G3', in fact, this stage requires only 0.6 kcal/mol and forms the dihydroxycopper (II) intermediate OS6. Adding a proton to OS6 delivers a first water molecule and the hydroxycopper (II) OS7, which can again oxidize another unit of G3 to G3'; this last oxidation requires only 9.8 kcal/mol and affords the hydroxycopper (I) species **OS8**. Finally, the protonation of the hydroxyl forms the second water molecule and regenerates the starting catalyst G1. The formation of diphenyldiacetylene from two G3' units is quite straightforward as described in the main text (Scheme 4).

Once computed the inner- and outer-sphere mechanisms is possible to compare them in free energy terms. For both, the major barrier corresponds to the terminal alkyne deprotonation (17.5) kcal/mol) and thus it is not easy to state which one should be preferred. In terms of oxidation barriers the inner-sphere seems to be somehow favoured since the barrier is just 10.3 kcal/mol (computed between G3 and O4); for the outer-sphere mechanism described here the free energy differences for the electron transfers are: -14.4, 10.1, 11.6, 0.6 and 9.8 kcal/mol. The highest increase in energy is therefore 11.6 kcal/mol, already higher than the 10.3 kcal/mol in the inner sphere mechanism. Moreover, there should be also a small barrier associated to the electron transfer, that could be estimated through Marcus theory, that would further increase the gap in mechanism favor of the inner sphere reported in the main text.

**Table S3.** Computed potential energies, enthalpies, free energies and dispersion correction terms (in Hartrees) for the outer-sphere catalytic cycle.

Species	Potential	Enthalpy	Free energy	Dispersion
OS1	-694.912111	-694.671500	-694.724661	-0.020024
OS2	-695.046531	-694.808456	-694.866349	-0.019163
OS3	-695.520425	-695.268675	-695.325293	-0.019569
OS4	-695.654253	-695.404699	-695.463318	-0.020052
$H_2O_2$	-151.460359	-151.430832	-151.456559	-0.000170
$OA_H_2O_2$	-696.105737	-695.845574	-695.903083	-0.021476
OS5	-696.194779	-695.932025	-695.986514	-0.022896
OS6	-696.345198	-696.085114	-696.143161	-0.022135
OS7	-620.429391	-620.182199	-620.234971	-0.019362
OS8	-620.566560	-620.321972	-620.377707	-0.017857

### Alternative mechanisms for diphenyldiacetylene formation

A couple of alternative pathways have been found and analyzed, both of them involving a bimetallic reductive elimination from Cu(III) intermediates. These mechanisms are described in Scheme S2 and S3.

**Scheme S2.** Proposed alternative mechanisms for the Glaser Hay coupling including the computed free energy values in kcal/mol (Cu(II), Cu(III); for clarity phenylacetylene is depicted as  $\equiv$ ).

In Scheme S2, after the proton is transferred from the nitrogen to the bridging oxygen in O12 the dinuclear complex is not stable anymore and rearranges into O13', where both alkynyl groups act as bridging ligands, or into O13'' where a hydrogen bond interaction is established between both hydroxyl groups. Bimetallic reductive elimination from either compound is quite easy, forming a first diphenyldiacetylene molecule and the copper (II) intermediate O14'. After that, the reaction follows the pathway shown in Scheme S3 where a second molecule of diphenyldiacetylene is formed and the catalyst is recovered.

**Scheme S3.** Proposed alternative mechanism for the product formation in the Glaser-Hay coupling including the computed free energy values in kcal/mol (Cu(I), Cu(II); for clarity phenylacetylene is depicted as  $\equiv$ ).

Complex **O14'** can easily coordinate a free alkyne to form **O15'**, then the intramolecular deprotonation happens with an energy requirement of less than 10 kcal/mol to yield the copper (II)— $\sigma$ —acetylide—H<sub>2</sub>O **O16'**. This latter intermediate is able to dimerize to form **O17'** (the same complex as **G5**) by liberating two molecules of water. In fact, the optimization of the dinuclear species where the water molecules remained on the copper atoms always ended up with their dissociation along the formation of the bimetallic complex. Once **O17'** is formed, a reductive elimination that requires less than 7 kcal/mol, the same as **RedEl\_TS** in the outer sphere mechanism, produces the second diphenyldiacetylene molecule and regenerates the initial copper (I) catalyst.

**Table S4.** Computed potential energies, enthalpies, free energies and dispersion correction terms (in Hartrees) for the alternative pathways.

Species	Potential	Enthalpy	Free energy	Dispersion
O13'	-1855.799260	-1855.091562	-1855.205391	-0.070853
RedEl_TS'	-1855.774432	-1855.067731	-1855.178994	-0.076709
O14'	-620.429360	-620.182152	-620.234562	-0.019343
O13"	-1855.749385	-1855.044092	-1855.161144	-0.079984
RedEl_TS''	-1855.742761	-1855.037839	-1855.151319	-0.080964
O15'	-928.508727	-928.146521	-928.217891	-0.031001
DP_TS-OH'	-928.492964	-928.135776	-928.206332	-0.029629
O16'	-928.517040	-928.153459	-928.225472	-0.029053

### Alternative intramolecular deprotonation

This alternative pathway, depicted in Scheme S4, consists in the coordination of TMEDA to G2 to form intermediate G2', which is substantially higher in free energy than the separated counterparts. Nevertheless, the intramolecular deprotonation affords a barrier that is very close to the one proposed in the main text, suggesting that both pathways could be operating.

**Scheme S4.** Alternative intramolecular deprotonation mechanism including the computed free energy values in kcal/mol (for clarity phenylacetylene is depicted as ≡).

# Optimized geometries

Phenylacetylene	C -1.681060 1.371765 -0.534711	H 0.556323 1.414203 -1.441388
C 3.250962 0.000004 0.000002	H -0.708197 1.851763 -0.731609	H 1.224212 2.342947 -0.073716
H 4.328734 0.000022 0.000005	H -2.308505 1.499350 -1.434023	H -0.556319 1.414198 1.441393
C 2.027329 -0.000013 -0.000001	H -2.177130 1.915331 0.306836	H -1.224206 2.342948 0.073725
C 0.597380 -0.000005 0.000000	C -2.789710 -0.727068 -0.173328	N -1.513935 0.214422 0.001520
C -0.119663 -1.221882 0.000000	H -3.421362 -0.349689 0.668456	N 1.513938 0.214420 -0.001520
C -0.119665 -1.221882 0.000000 C -0.119652 1.221877 0.000000	H -3.357985 -0.588650 -1.110073	C 2.579622 0.023609 -1.011121
C -1.519061 -1.215391 0.000000		H 3.256652 0.902662 -1.050772
	H -2.636847 -1.809931 -0.018652	
H 0.431149 -2.167473 0.000000	TMEDA H	H 3.171721 -0.867235 -0.748572
C -1.519051 1.215397 0.000000	TMEDA-H <sup>+</sup> C 0.677296 0.271605 1.139708	H 2.126675 -0.125884 -2.003859
H 0.431165 2.167465 0.000000		C 2.130197 0.361208 1.337995
C -2.222634 0.000007 0.000000	C -0.730657 -0.342975 1.080664	H 2.696502 -0.552235 1.577550
H -2.063564 -2.165311 0.000000	H 0.635760 1.364029 1.261343	H 2.820379 1.230713 1.360446
H -2.063544 2.165323 0.000000	H 1.291074 -0.154928 1.946451	H 1.353554 0.503045 2.104952
H -3.317602 0.000010 0.000000	H -0.662272 -1.441071 1.155704	C -2.130199 0.361220 -1.337992
	H -1.317321 0.013465 1.952354	H -1.353558 0.503061 -2.104951
Diphenyldiacetylene	N -1.352811 -0.007367 -0.212223	H -2.696506 -0.552221 -1.577551
C 0.677808 -0.000121 -0.000102	Н 0.505546 0.007793 -0.835399	H -2.820379 1.230726 -1.360434
C 1.915102 -0.000097 -0.000063	N 1.347543 0.002384 -0.187740	C -2.579616 0.023605 1.011123
C -0.677808 -0.000015 -0.000025	C 1.982978 -1.350081 -0.255746	H -3.256645 0.902658 1.050782
C -1.915102 0.000034 0.000026	H 1.247249 -2.113020 0.032703	H -3.171716 -0.867237 0.748572
C -3.335504 0.000023 0.000019	H 2.315822 -1.525147 -1.288374	H -2.126666 -0.125894 2.003860
C -4.054538 0.865129 0.866121	H 2.840571 -1.369158 0.432481	
C -4.054518 -0.865092 -0.866088	C 2.299665 1.077173 -0.595640	G2
C -5.452744 0.859406 0.860389	H 2.709837 0.823901 -1.583821	Cu -0.757578 -0.500483 0.007956
Н -3.503525 1.533036 1.534914	H 1.755687 2.030693 -0.644887	C -2.254116 1.887187 0.292165
C -5.452725 -0.859387 -0.860370	H 3.108700 1.136069 0.147221	C -3.186046 0.885691 -0.385656
Н -3.503491 -1.532993 -1.534875	C -1.936781 1.340462 -0.212864	Н -2.384629 1.852582 1.386007
C -6.155933 0.000005 0.000006	H -1.181189 2.089165 0.076544	H -2.498274 2.915286 -0.039857
H -5.997540 1.530548 1.532351	H -2.287824 1.585465 -1.229093	H -3.098220 0.959361 -1.482151
H -5.997505 -1.530536 -1.532338	H -2.797152 1.425224 0.487123	H -4.238211 1.108419 -0.119354
H -7.250775 -0.000002 0.000001	C -2.345158 -1.002875 -0.631107	N -2.833698 -0.513431 -0.004249
C 3.335504 -0.000041 -0.000024	H -3.216820 -1.056550 0.057512	N -0.822236 1.569993 0.001224
C 4.054516 0.866140 -0.865068	H -2.717323 -0.740996 -1.635934	C 1.035733 -1.461156 0.010470
C 4.054540 -0.866167 0.865055	Н -1.877539 -1.999903 -0.682269	C 0.046056 -2.253549 0.006417
C 5.452722 0.860468 -0.859321	572 (G2 7) 3 <sup>2</sup>	H -0.377329 -3.255086 0.002642
Н 3.503486 1.534947 -1.532946	$[Fe(CN)_6]^{3-}$	C -3.400265 -1.471532 -0.980889
C 5.452747 -0.860388 0.859375	Fe 0.000000 0.000000 0.000000	H -3.117143 -2.495664 -0.691768
Н 3.503530 -1.535015 1.532909	C 0.000000 0.000000 1.925760	H -4.506182 -1.400340 -1.009130
C 6.155933 0.000068 0.000045	C 0.000000 1.934524 0.000000	H -2.998672 -1.257236 -1.983693
Н 5.997501 1.532492 -1.530416	C -1.934524 0.000000 0.000000	C -3.329858 -0.844847 1.353491
Н 5.997544 -1.532369 1.530496	C 1.934524 0.000000 0.000000	H -4.436627 -0.789487 1.389197
H 7.250775 0.000111 0.000072	C 0.000000 -1.934524 0.000000	H -3.012237 -1.866565 1.614101
	C 0.000000 0.000000 -1.925760	H -2.912096 -0.147520 2.095512
$O_2$	N 0.000000 0.000000 3.113291	C -0.432986 2.061454 -1.344256
O 0.000000 0.000000 0.614021	N 3.120881 0.000000 0.000000	H -0.514690 3.165897 -1.394993
O 0.000000 0.000000 -0.614021	N 0.000000 -3.120881 0.000000	Н 0.608619 1.767976 -1.548683
	N 0.000000 3.120881 0.000000	H -1.082084 1.619115 -2.115326
$H_2O$	N -3.120881 0.000000 0.000000	C 0.055647 2.191476 1.020972
O 0.000000 0.000000 0.123172	N 0.000000 0.000000 -3.113291	H 1.105500 1.948352 0.795370
Н 0.000000 0.761276 -0.492687		H -0.063848 3.293488 1.026206
Н 0.000000 -0.761276 -0.492687	$[Fe(CN)_6]^{4-}$	H -0.198751 1.795840 2.016757
	Fe 0.000000 0.000000 0.000000	C 2.332497 -0.837525 0.006369
TMEDA	C 0.000000 0.000000 1.914115	C 2.975171 -0.533256 -1.220201
C 0.703041 0.307856 0.931681	C 0.000000 1.914115 0.000000	C 2.976891 -0.514424 1.226926
C -0.703043 -0.307939 0.931651	C -1.914115 0.000000 0.000000	C 4.237274 0.071714 -1.218108
H 0.630991 1.404563 1.050976	C 1.914115 0.000000 0.000000	H 2.477845 -0.780185 -2.163429
H 1.216661 -0.071119 1.851687	C 0.000000 -1.914115 0.000000	C 4.239228 0.090109 1.213523
H -0.630992 -1.404657 1.050846	C 0.000000 0.000000 -1.914115	H 2.480628 -0.745008 2.174816
H -1.216665 0.070952 1.851691	N 0.000000 0.000000 1.514115 N 0.000000 0.000000 3.108955	C 4.872804 0.385206 -0.005120
N -1.497764 -0.053784 -0.275012	N 3.108955 0.000000 0.000000	H 4.728473 0.298289 -2.170048
	N 0.000000 -3.108955 0.000000 N 0.000000 -3.108955 0.000000	H 4.731785 0.331499 2.161112
N 1.497765 0.053811 -0.275003		
C 1.681057 -1.371716 -0.534831	N 0.000000 3.108955 0.000000	Н 5.859262 0.859767 -0.009460
H 0.708194 -1.851693 -0.731769	N -3.108955 0.000000 0.000000	DD TO
H 2.308500 -1.499221 -1.434156	N 0.000000 0.000000 -3.108955	DP_TS
H 2.177129 -1.915357 0.306666	61	Cu -1.229171 -0.653345 -0.180883
C 2.789713 0.727080 -0.173257	G1	C -3.909499 -1.563831 0.225515
H 3.357986 0.588750 -1.110016	Cu -0.000003 -1.181400 -0.000004	C -3.189541 -1.886919 1.536325
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