

Boundary Work: 1st TD Law on closed system

Raj Pala,

rpala@iitk.ac.in

Department of Chemical Engineering,
Associate faculty of the Materials Science Programme,
Indian Institute of Technology, Kanpur.

Work, heat & Energy in Thermodynamics

- Generalizing work-energy theorem and conservation of energy beyond mechanics, sign conventions, exact & inexact differentials

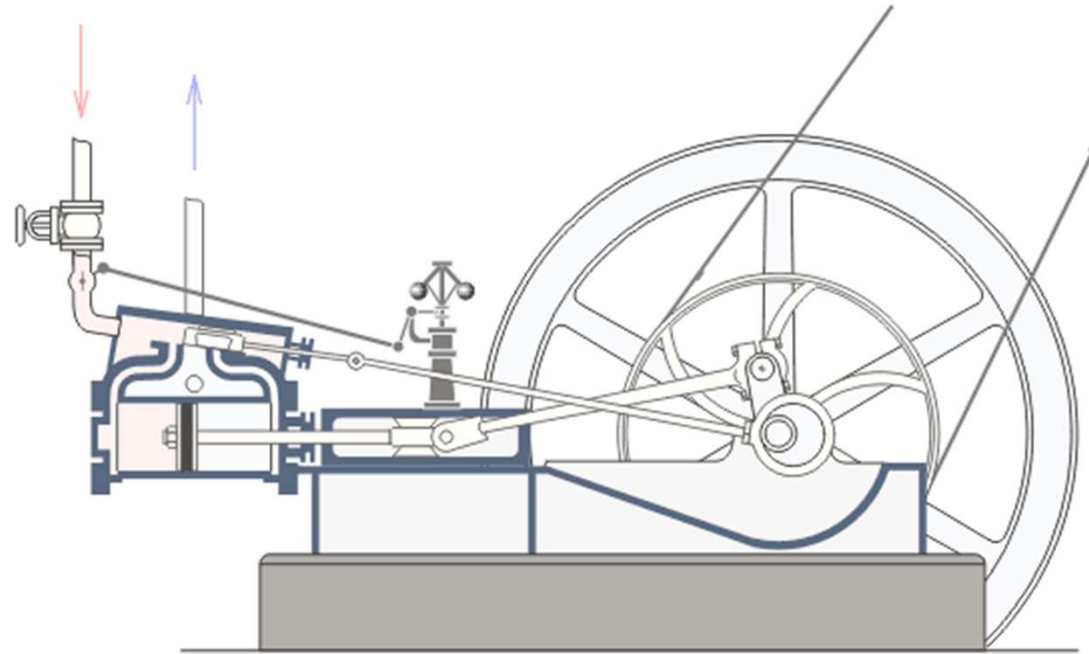
$$\Delta U = \text{Change in Internal Energy } U = \text{Heat \& work exchange} = q - W$$

$$\Delta E = \Delta U + \Delta KE + \Delta PE$$

- Impossibility of perpetual motion machines

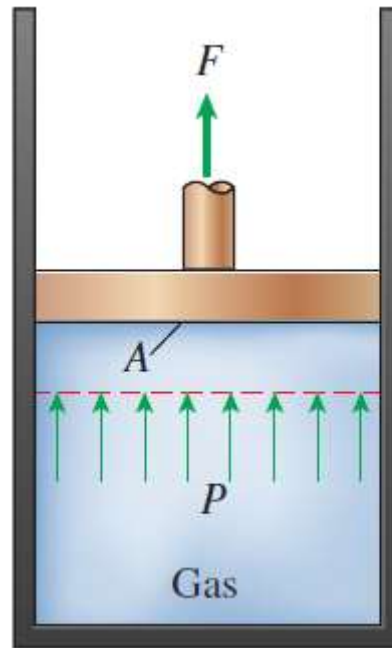
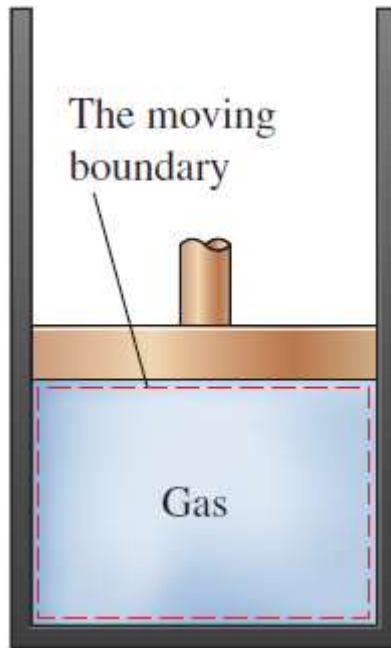


Pistons are coupled to engines and compressors...



- Cyclic linear motion of the piston → Rotatory motion
- Quasi-static vs. real processes
- Bounds: Work output maximum and minimum work input

Moving boundary: PdV work

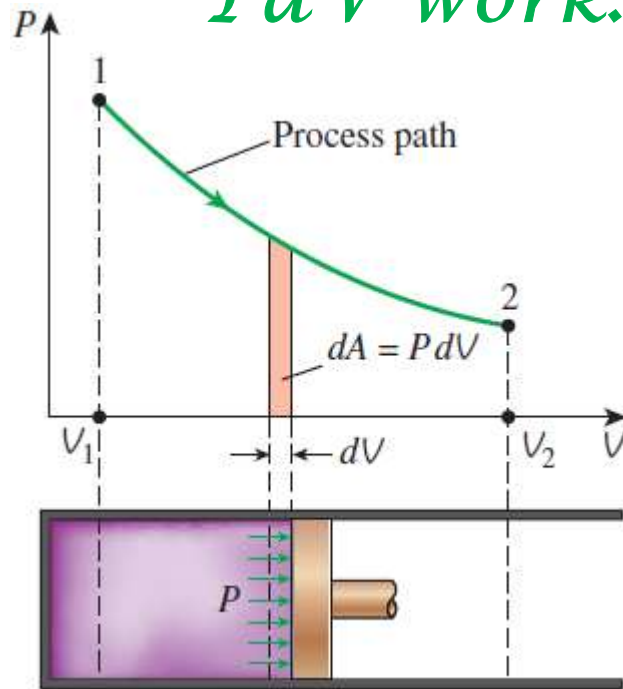


$$\delta W_b = F ds = PA ds = P dV$$

$$W_b = \int_1^2 P dV \quad (\text{kJ})$$

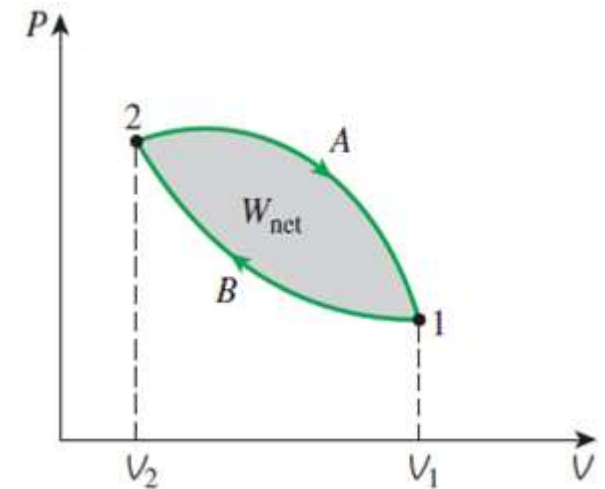
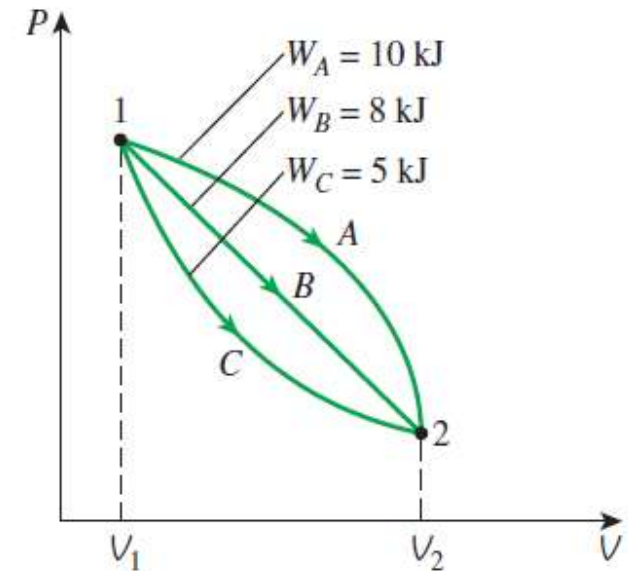
- Quasi equilibrium
- $\Delta U = q - W$
- W_e is positive \rightarrow for expansion
- W_c is negative \rightarrow for compression

PdV work: Path/Inexact differential

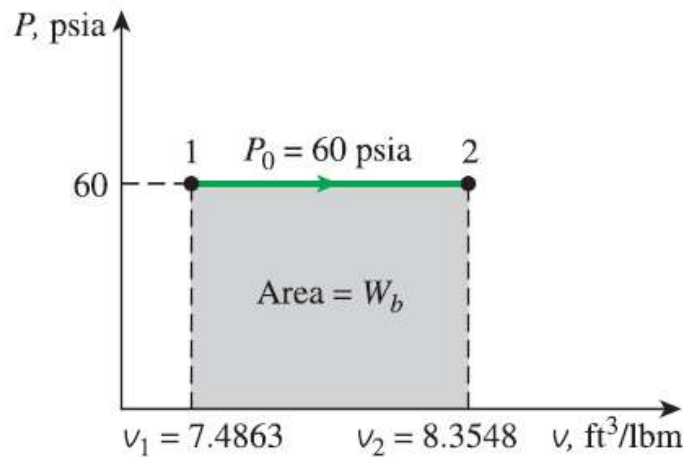
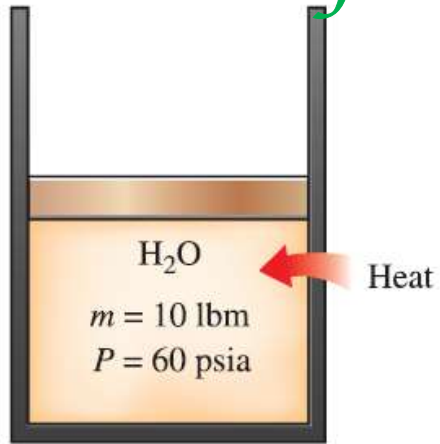


$$\text{Area} = A = \int_1^2 dA = \int_1^2 P dV$$

- Net work in a cycle=Work done (by system-on system)
- If work were not a path fxn, No net work is possible from cyclic processes!

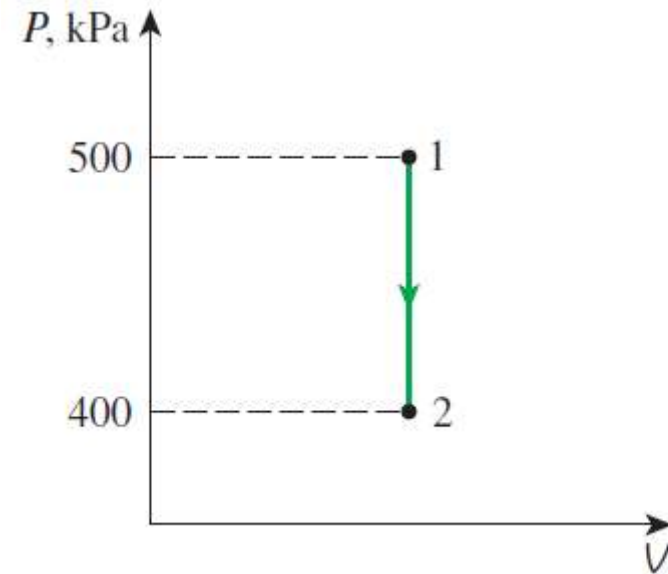
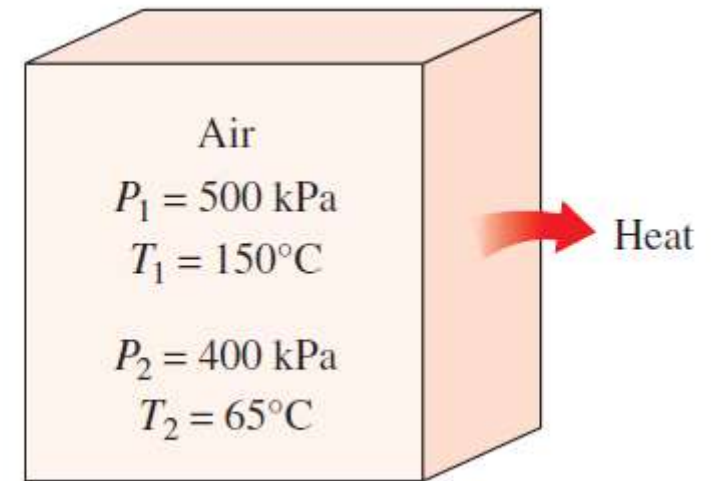


Boundary work: Constant P & V processes

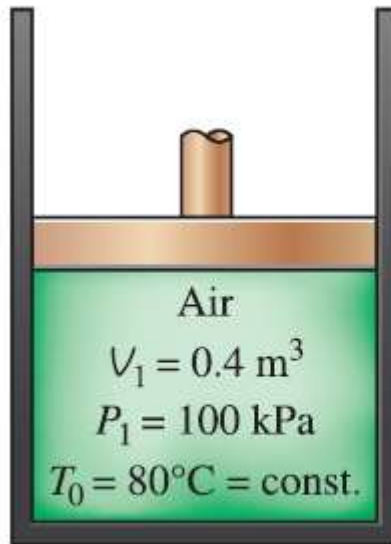


$$W_b = \int_1^2 P dV = P_0 \int_1^2 dV = P_0(V_2 - V_1)$$

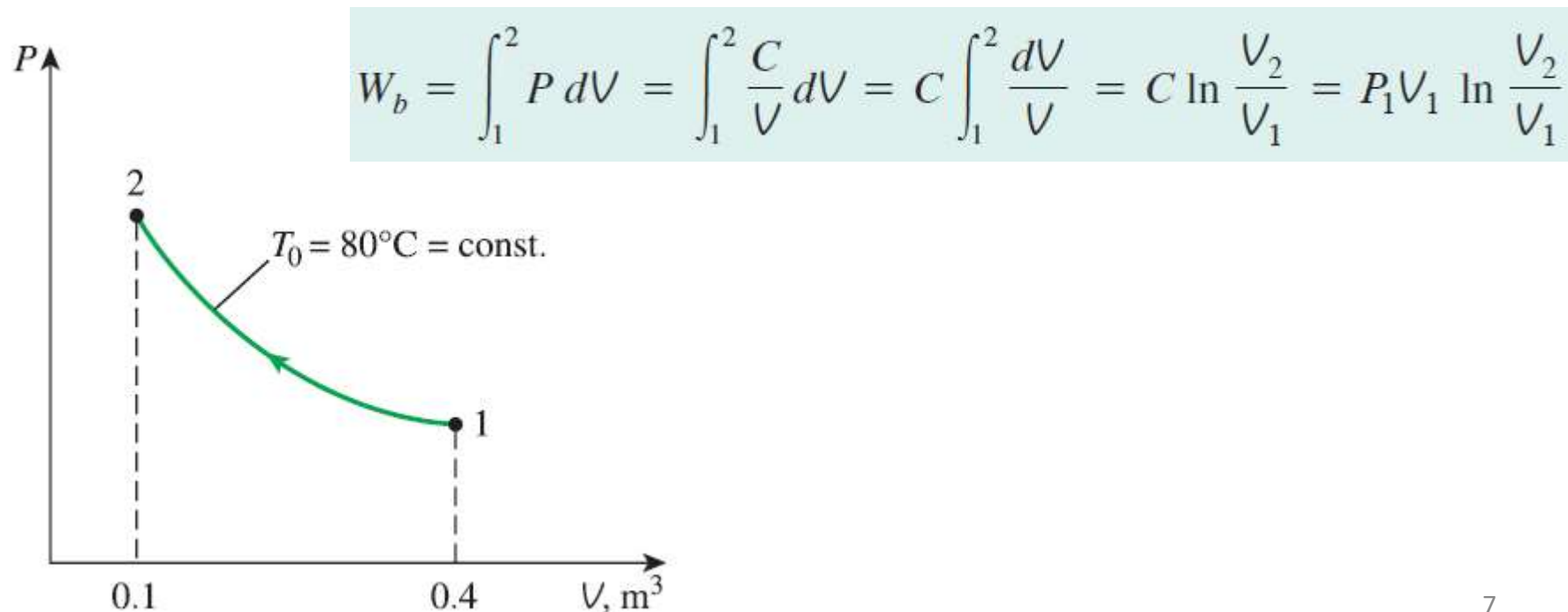
$$W_b = mP_0(v_2 - v_1)$$



Boundary work: Isothermal compression process



$$PV = mRT_0 = C \quad \text{or} \quad P = \frac{C}{V}$$



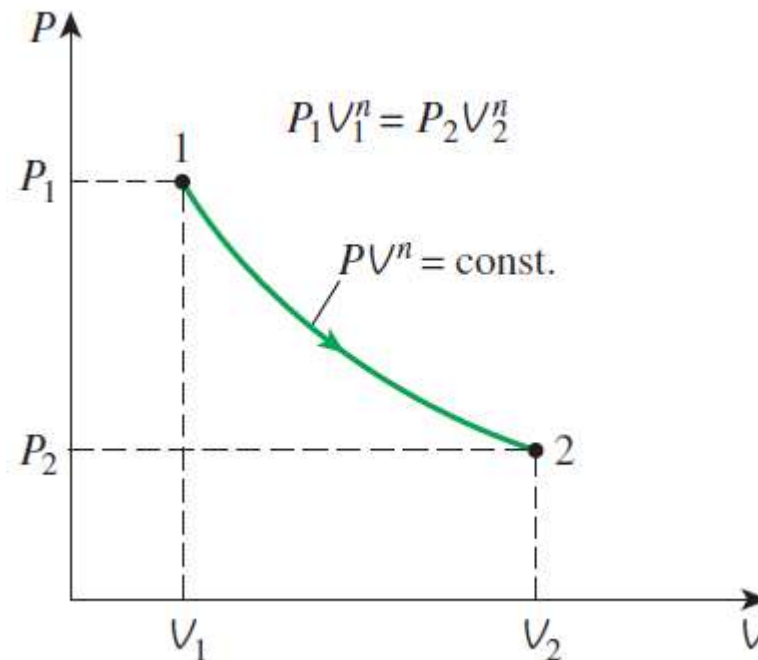
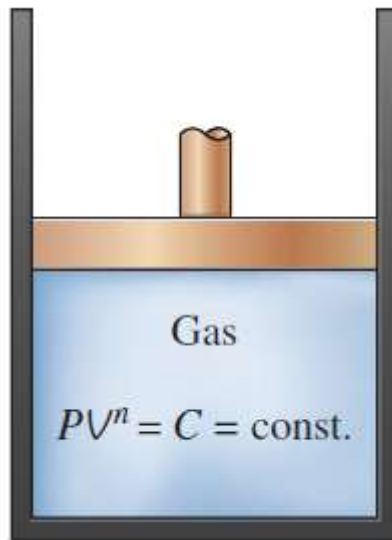
$$W_b = \int_1^2 P dV = \int_1^2 \frac{C}{V} dV = C \int_1^2 \frac{dV}{V} = C \ln \frac{V_2}{V_1} = P_1 V_1 \ln \frac{V_2}{V_1}$$

Boundary work: Polytropic process

$$PV^n = C, \quad P = CV^{-n}$$

$$W_b = \int_1^2 P dV = \int_1^2 CV^{-n} dV = C \frac{V_2^{-n+1} - V_1^{-n+1}}{-n+1} = \frac{P_2 V_2 - P_1 V_1}{1-n}$$

$$W_b = \frac{mR(T_2 - T_1)}{1-n} \quad \text{For ideal gas}$$



Expansion against spring

