**ELECTROLYSIS**

FARADAY’S FIRST LAW OF ELECTROLYSIS

It is defined as the amount of substance deposited or liberated at any electrode is directly proportional to the quantity of electricity passed through the electrolytic solution.

m = Z IT

Z is the electrochemical equivalent Z = equivalent wt/ faraday (NOTE –1faraday is charge on 1 mole of electrons ; 96500C)

Z or ECE is defined as the mass of the substance deposited when current of 1 amp is passed for 1 second .

FARADAY’S SECOND LAW OF ELECTROLYSIS

When the same quantity of electricity is passed through different electrolytic solutions connected in series , the weight of the substance produced at the electrode are directly proportional to their chemical equivalent weight .

Wt of Cu deposited /wt of Ag deposited = eq wt of Cu/eq wt of Ag