

Module 3

QUANTUM MECHANICS – NOTES

(As per SVU-R2020 Scheme & Syllabus)

- **Introduction**

Quantum mechanics emerged out of a process to reformulate our concepts and understanding of the nature. Most of our technological applications are still based on the traditional “classical physics” consisting of wave motion, mechanics, electromagnetism, thermodynamics, optics etc. However, the quantum mechanical approach becomes essential as we probe further and further into the nature; particularly for studying the world of “very small”. Masses, dimensions and other physical parameters of objects in this world are orders of magnitudes smaller than our “everyday-life” world. More importantly, these “very small” objects are beyond our senses either directly or even indirectly through instruments. How can we develop the trajectory of a proton if we can't actually see it like a stone going up and coming down? Moreover, we have a tendency to visualize any object by our common preconception. For example, when we talk about an electron, we tend to give it a spherical, ball-bearing type look. But how do we know that it's a particle if we can't see to it?

Facing to very small masses, very small dimensions and a large ensemble of such “identical, indistinguishable” objects (e.g. electrons in a conductor), statistical interpretation becomes inherent in quantum mechanics. Results are given in terms of “probability functions” rather than “deterministic” laws of traditional sciences. Values of physical quantities such as position, energy, momentum etc. are expressed as “most probable distribution” instead of an exact value. This should not be regarded as the shortfall of the theory because the so-called “deterministic” laws of macroscopic world can be shown as an approximation of the quantum mechanical treatment in the limit of large values of physical parameters. Quantum physics is in fact, the “superset” of all of classical sciences.

- **Background**

It would be relevant to discuss some of the important findings that changed the gears of physics at the beginning of the 20th century. Classical physics could not account for many of the experimental findings carried out at that time. Some of them were as follows:

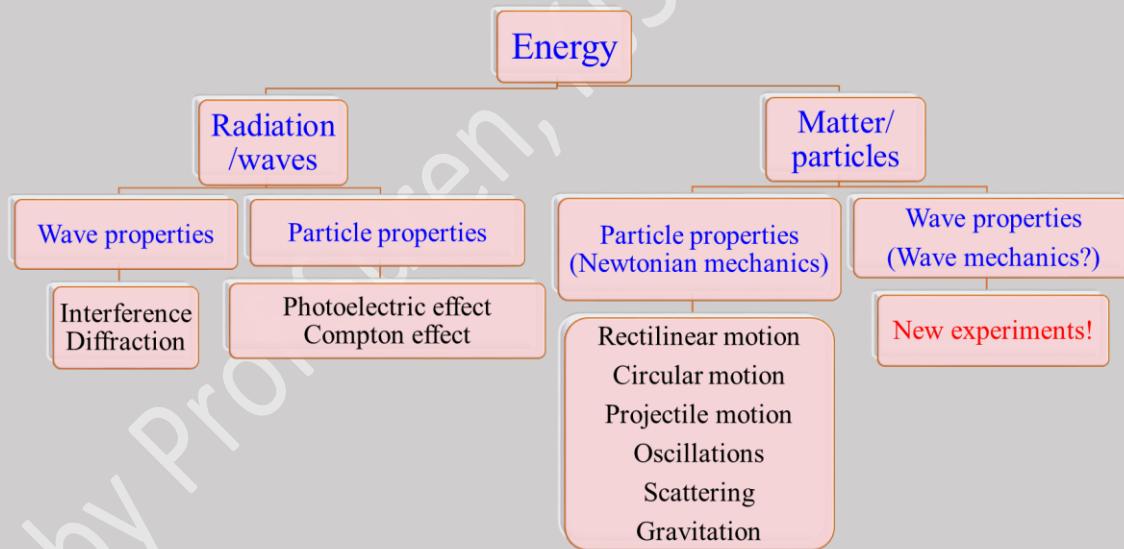
- 1) Blackbody radiation catastrophe: By and large, the Planck's formula for blackbody radiation is considered as the birthplace of quantum physics. It used a novel concept of assuming every matter as made of small oscillating particles (oscillators) and that the energy emitted by their ensemble be in discrete manner rather than at any arbitrary value (continuous). The formula exactly matched with the experimental results at any wavelength and the discreteness of energy was established.
- 2) Specific heats of solids at low temperatures.
- 3) Constituents of matter, stability of atoms.
- 4) Origin of atomic and molecular spectra.
- 5) Fine structure and hyperfine structure splitting of spectral lines in magnetic field.
- 6) Wide range of conductivity of solids (10^{-14} S/m to 10^8 S/m).
- 7) Origin of magnetism, wide range of magnetisation (10^9 A/m to 10^{-6} A/m) and spontaneous magnetisation of ferromagnetic materials.
- 8) Cause of radioactivity.

- 9) Phenomenon of superconductivity.
- 10) Production of X-rays.
- 11) The photoelectric effect.
- 12) Zeeman effect.
- 13) Franck-Hertz Experiment.
- 14) Stern-Gerlach Experiment.

The events indicated above are essentially before de'Broglie put forward his hypothesis. Heisenberg (1925) and Schrödinger (1926) independently developed core theories of quantum mechanics. Later, Dirac (1928) contrived a general formalism and also included relativistic extensions to it. Since then there has been a continuous development in the quantum mechanics at atomic, nuclear and elementary particle scales but discussion on this development is beyond the scope of the syllabus.

- **de'Broglie's hypothesis**

Louis de'Broglie tried to sum up the experimental finding by his time into a set of logical arguments based on symmetry. Energy in the universe is manifested in two forms viz radiation and matter. By then, it was known that radiation shows both, particle properties and wave properties. de'broglie then argued that why can't matter, which is supposed to be made of particles and mostly governed by the laws of Newtonian mechanics, exhibit wave nature? Following chart can be useful to understand de'Broglie's arguments:



At that time, there was no experimental basis to his hypothesis but many took it seriously and soon wave nature of electrons, presumed as perfect particles was demonstrated by two independent experiments; one by Davison and Germer and the other by G P Thomson. Both these experiments were very similar to X-ray diffraction and based on diffraction of electrons by using crystals as three dimensional diffraction gratings.

Prior to this, it was already established that particle aspect of radiation called as "photons" are governed by two important equations:

$$E = hv = \frac{hc}{\lambda} \text{ Also, } E = mc^2 = (mc)c = pc \Rightarrow \frac{hc}{\lambda} = pc \text{ or } \lambda = \frac{h}{p}$$

De Broglie extended this equation to all material particles having mass "m" and moving with speed "v", so the matter has an associated wave whose wavelength can be estimated as

$$\lambda = \frac{h}{mv} \quad (v < c \text{ always}) \quad (1)$$

This relation is called de'Broglie equation.

Note that this equation connects two physical quantities which are usually not used together because the momentum is a typical particle parameter while the wavelength is a typical wave parameter. The notions of a particle and a wave as we learn describe them as opposite to each other. By "particle" we imagine a localized, concise entity and which can be obtained in a discrete fashion while by a "wave" we try to imagine a distributed entity continuous all over. For example, the ocean wave hits the shore simultaneously all over but a marble can hit only at a particular point at a time.

For a particle having kinetic energy K, the de'Broglie wavelength can be written as

$$\lambda = \frac{h}{\sqrt{2mK}} \text{ as } K = \frac{p^2}{2m} = \text{Kinetic energy acquired} \quad (2)$$

For a charged particle of charge "q" accelerated through a potential difference of "V",

$$\lambda = \frac{h}{\sqrt{2mqV}} \text{ as } qV = \text{Potential energy gained} \quad (3)$$

- **Experimental verification of de'Broglie's hypothesis - The Davison-Germer Experiment**

An important demonstration of diffraction of electrons was given by Davison and Germer and independently by G P Thomson in 1927. These experiments confirmed de'Broglie's hypothesis and consequently, the wave nature of material particles. The demonstration is considered to be one of the important experiments in the development of quantum mechanics.

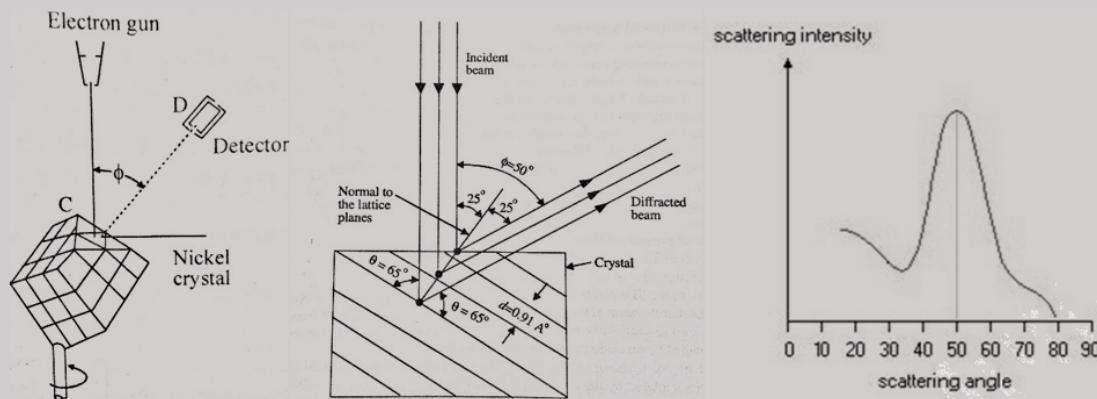
If electrons have wave nature, they should also exhibit diffraction effect. With an electron gun, electrons at energy 50 – 100 eV can easily be obtained. At this energy, according to de'Broglie equation, electrons should have wavelength of a few angstrom. So crystals can act as diffraction grating for electrons also in the same manner they act for X-ray diffraction.

Experimental set-up:

In the actual experiment, Davison and Germer selected a nickel crystal. A collimated beam of monoenergetic electrons is obtained by an arrangement consisting of electron gun, set of anodes, collimating slits and variable voltage source. The crystal can be rotated on its axis. A detector was placed in the scattering direction, which can also be rotated in accordance to the rotation of the crystal. The detector was attached with a scale to measure angle of scattering (ϕ) and hence to deduce the glancing angle (θ).

Experiment:

Electrons emitted from a hot tungsten filament after getting accelerated and focused fall on the nickel crystal. Electrons scattered from the crystal are collected at the detector, which produces the count in different directions with respect to the incident beam. The intensity of scattering (detector count) is plotted as a function of scattering angle for different values of accelerating voltage (V).



Investigations:

Davison and Germer found that the detector current, which in turn depends upon intensity of scattered electrons increased significantly at certain angle of scattering and at certain values of accelerating voltage. In a particular observation, they found maximum detector current corresponding to scattering angle of 50° for accelerating voltage of 54 volt. With reference to the selected orientation of the nickel crystal (lattice planes), this scattering angle corresponded to glancing angle of 65° . The detector current was very low and random at different angle of scattering. This suggested that there is certain kind of preference to a particular glancing angle when the electrons scatter from the crystal. This could only be explained if it is assumed that electrons have a wave nature and obey the Bragg's law viz $2d \sin\theta = n\lambda$.

Calculations:

The interplanar spacing (d) for the selected orientation of nickel crystal as obtained from X-ray analysis was 0.91 \AA . The glancing angle for which, maximum count obtained was 65° . Using Bragg's law, it gives wavelength of possible waves as $\lambda = \frac{2d \sin \theta}{n}$. The scattering angle corresponded to the first order of diffraction maximum. Hence, $\lambda = 2 \times 0.91 \times \sin 65^\circ = 1.65 \text{ \AA}$. Assuming de'Broglie' hypothesis to be correct, Davison and Germer estimated the wavelength of electron waves using de'Broglie equation $\lambda = \frac{h}{\sqrt{2m\phi}}$. This yielded $\lambda = \frac{6.63 \times 10^{-34}}{\sqrt{2 \times 9.1 \times 10^{-31} \times 1.6 \times 10^{-19} \times 54}} = 1.67 \text{ \AA}$, in excellent closeness to that obtained from the Bragg's law. Thus, de'Broglie equation indeed has significance and wave nature of matter was established experimentally for the first time.

Further confirmation:

The experiment carried independently by G. P. Thomson also confirmed the wave nature of electrons. Later, experiments were performed by using beams of other particles such as protons, neutrons and alpha particles, which also agreed perfectly with the theoretical predictions. Similar experiments were done using heavier ions to prove that wave nature is not restricted only to elementary particles but it applies virtually to all of matter. Thus, wave nature of matter was established. The Davisson-Germer experiment stood as the pioneering experiment in the confirmation of quantum mechanics.

- **Phase velocity (wave velocity), wave packet and group velocity:**

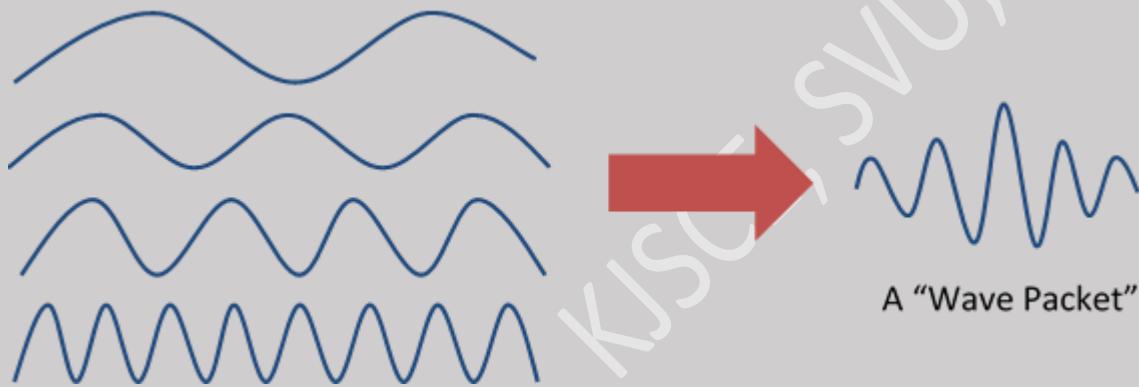
The waves associated with all material particles moving with certain speed are called as "matter waves" but exactly what kind of waves are these? Let's find the velocity with which a wave associated

with a particle moves. Traditionally, the rate at which the plane containing waves having the same phase (viz. the wave front) advances in a certain direction is called the phase velocity. It is the same as the wave velocity in ordinary sense in a wave motion. Thus,

$$v_{\text{phase}} = \frac{\omega}{k}; \text{ Where } \omega = 2\pi v \text{ and } k = \frac{2\pi}{\lambda}$$

$\therefore v_{\text{phase}} = v \lambda = \frac{E}{h} \times \frac{h}{p} = \frac{E}{p} = \frac{mc^2}{mv} = \frac{c^2}{v} > c!$ as $v < c$ always. Here, "v" is particle velocity in the traditional sense.

Estimation above gives an unusual and absurd result that the wave associated with particle moves with an extraordinary velocity leaving the particle far behind. Since this is impossible, it is concluded that particle cannot be associated with a single wave in ordinary sense. Next, it is proposed that a number of waves of different frequencies are associated with the particle.



The reason for this proposal is that a large number of waves can be added and it results in a repetitive pattern called as the "wave packet". This wave packet essentially carries the energy associated with the system (a particle; in this case). The collection of these waves is called as the wave group and the speed at which this entire group moves is different than velocities of individual waves. It is not just the average velocity but it gives the rate at which frequency varies as a function of the wave vector "k" within the group. In mathematical terms,

Group velocity $V_g = \frac{d\omega}{dk}$. Let's estimate what is this group velocity.

Let $V_g = \frac{d\omega}{dk} = \frac{d\omega/dv}{dk/dv}$; Where, "v" is the particle velocity in traditional sense.

We know that $\omega = 2\pi v \therefore \omega = \frac{2\pi E}{h} = \frac{2\pi mc^2}{h} = \frac{2\pi m_0 c^2}{h\sqrt{1-v^2/c^2}}$; By using Planck's formula, Mass-energy equivalence and relativistic expression for mass.

$$\therefore \frac{d\omega}{dv} = \frac{2\pi m_0 v}{h(1-v^2/c^2)^{3/2}}$$

Also, $k = \frac{2\pi}{\lambda} \therefore k = \frac{2\pi p}{h} = \frac{2\pi mv}{h} = \frac{2\pi m_0 v}{h\sqrt{1-v^2/c^2}}$; By using de'Broglie equation, definition of momentum and relativistic expression for mass.

$$\therefore \frac{dk}{dv} = \frac{2\pi m_0}{h(1-v^2/c^2)^{3/2}}.$$

$$\therefore \frac{d\omega}{dk} = \frac{d\omega/dv}{dk/dv} = v \text{ i. e. } v_{\text{particle}}.$$

Thus the velocity with which the group moves in the form a wave packet is the same as particle velocity. So it is inferred that a particle can be represented by a wave packet.

- Matter waves and their properties**

Since all the material particles show wavelike properties such as interference or diffraction, we try to associate wave nature to matter in general. But it should be understood that it is an entirely different concept that our usual understanding of the waves and wave motion. In introductory section it was stated that quantum mechanical treatment of solving any problem is rather statistical in nature and it involves calculation of probability factors. For example, classically, the radius of first Bohr orbit in hydrogen atom is 0.53 Å exactly but according to quantum mechanics, it is just the most probable value of the radius. When we attempt to measure the radius a number of times, many times we may obtain this value but occasionally we may get some other values also. Thus, any value is expressed with some distribution about the most probable value instead of an exact value. The exact value of a physical quantity is somehow a representation of pure particle picture used in classical mechanics while the probabilistic value with some spread (often taken as the uncertainty) is a kind of wave picture used in quantum mechanics. The wave nature of matter and the concept of wave packet are linked with this probabilistic interpretation. This connection was first proposed by Max Born. In quantum mechanics, this probabilistic nature is carried by an abstract mathematical quantity called as the "wave function".

- Differences between electromagnetic waves and matter waves**

Electromagnetic waves	Matter waves
1) Electromagnetic waves are associated with photon, which has zero rest mass.	Matter waves are associated with all moving material particles having non-zero rest mass.
2) A single de'Broglie wave can be associated with the particle (photon).	A single de'Broglie wave cannot be associated with the material particle.
3) The quantities those vary periodically with space and time are the electric and magnetic fields.	The quantity that varies periodically with space and time is called the wave function.
4) Electric and magnetic fields are real physical quantities and can be measured experimentally.	The wave function is an abstract mathematical quantity and has no direct physical interpretation.
5) Square of field amplitude gives intensity of electromagnetic waves.	Square of wave function gives probability of locating the particle in a given interval.

- **Properties of matter waves:**

- 1) Matter waves are associated with all moving particles.
- 2) Their wavelength is inversely proportional to momentum of the particle.
- 3) They are independent of charge of the particle.
- 4) They are neither mechanical nor electromagnetic kind of waves.
- 5) A particle cannot be represented by a single wave. Instead, it is associated with a large number of waves forming a wave packet.
- 6) The quantity that varies periodically with space and time is taken as the wave function.
- 7) Square of the wave function gives probability of locating the particle.

- **Wave packet, wave-particle duality and the uncertainty**

The idea of a wave packet is central and very important in quantum mechanics. The wave packet has finite extent and it is neither “localised” like an ideal particle nor spread all over from $-\infty$ to $+\infty$ like an ideal wave. This leads to an idea of wave-particle dualism. By wave-particle duality it means that matter as well as radiation cannot be assigned only wave nature or only particle nature but in principle, matter as well as radiation possesses dual nature. However, at a time only one nature is exhibited and never both simultaneously. This wave-particle duality can be explained by using the concept of a wave packet as follows:

- 1) In quantum mechanics, all material particles are represented by a wave packet of finite extent. Mathematically, the wave packet is formed by coupling a large number of waves by means of Fourier analysis.
- 2) Spread of the wave packet is linked with uncertainty in the exact location of the particle.
- 3) Classically, a particle has an exact location at a given time. Now if we try to have exact location of the particle, the wave packet should be restricted in its extent. But, in doing so, the wave nature becomes in-deterministic. When the wave packet shrinks to a point giving exact particle picture, the wave nature collapses completely.
- 4) On the other hand, if we try to have a perfect wave picture, the wave packet should be extended further. But in doing so, the position of the particle and hence the particle nature itself becomes highly uncertain. When the wave packet extends to infinity giving exact wave picture, the particle nature collapses completely.

Since all the material particles show wavelike properties such as interference or diffraction effects, we try to associate wave nature to particle or matter in general.

- **The uncertainty principle**

In 1927, with his matrix formulation of quantum mechanics, Heisenberg proposed the uncertainty principle. It puts forward the fundamental theoretical limit to the exact determination of physical quantities. It is particularly applicable to simultaneous measurement of certain pairs of dynamical variables. In classical physics also it was long known that certain pairs of physical quantities such as position-momentum are “non-commutative” (i.e. $AB \neq BA$). Heisenberg showed that the product of

uncertainties of such pairs is greater than or equal to the factor $\frac{h}{2\pi}$ (in 2-Dimensions). In mathematical form,

$$\Delta x \Delta p_x \geq \frac{h}{2\pi}$$

It states that "It is impossible to exactly determine the position and momentum of an object simultaneously such that the product of their uncertainties is never better than the factor $\frac{h}{2\pi}$ ".

(In 3-Dimensions, the factor is $\frac{h}{4\pi}$).

- Other forms of Uncertainty Principle:

1. Energy-time: $\Delta E \Delta t \geq \frac{h}{2\pi}$

2. Angular position-angular momentum: $\Delta \theta \Delta L \geq \frac{h}{2\pi}$

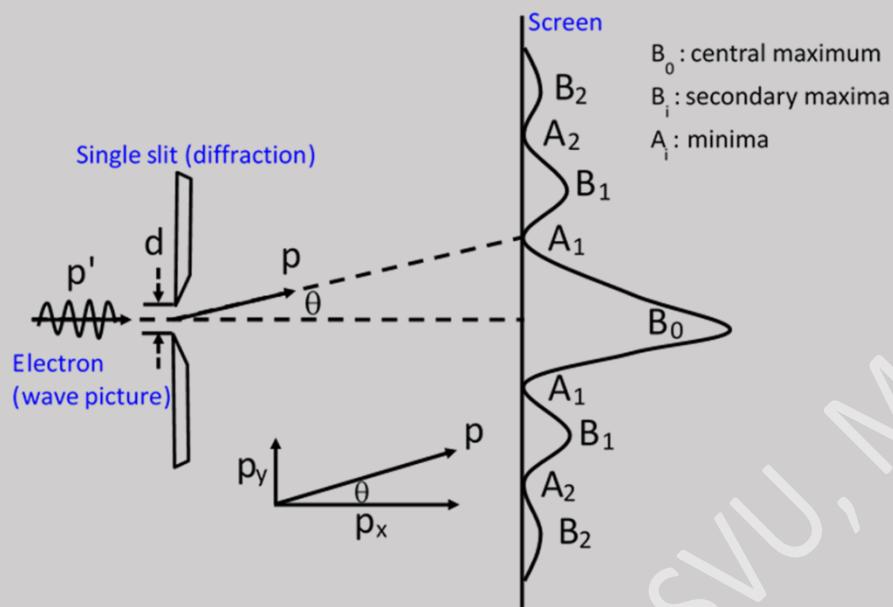
At first glance, the uncertainty principle may appear to be a negation since it puts forward the fundamental limit on exact measurements. It also leads to a misunderstanding that quantum mechanical approach does not yield exact value of any dynamical variable and hence it is incomplete. But it should be understood that the limitation comes due to the process of measurement itself. The fact is that the uncertainty principle is not just applicable at microscopic levels but virtually to *all* kinds of measurements; macroscopic as well as microscopic. At macroscopic levels, the uncertainties in the measurements are so small that we are usually unable to detect them and the results appear to be exact. But at the microscopic levels, the uncertainties become significant as compared to the actual values and the uncertainty principle becomes necessary. Uncertainty principle is not restricted just to put fundamental limit on measurements but it proved to be far more applicable to clarify a number of theoretical gaps as well as experimental findings.

- Proof-of-concept derivation for uncertainty principle

Uncertainty principle first appeared through the matrix algebra developed by Heisenberg for quantum mechanics and its actual derivation is beyond the scope of this syllabus. Hence, one of the qualitative ways to arrive at the uncertainty principle is given here.

Single slit diffraction of electrons – the wave picture:

Here, we consider that electron is wave of wavelength λ be incident on a narrow slit of width d . Let p' be the correspondong initial momentum of electron as per de'Broglie equation. The single narrow slit has the same effect as it would have on a monochromatic beam of light. As an electron passes through the slit, its position becomes uncertain by an amount say Δy . Since it can pass anywhere from the slit, we can take $\Delta y \approx d$. Further, the electron gains y-component of momentum as it can reach anywhere on the screen on either sides of central maximum B_0 . Let p be the momentum of electron after passing through the slit. The y-component of momentum is given by $p_y = p \sin \theta$. For convenience, let the electron reach at point A_1 on the screen. As this point corresponds to first minimum in the diffraction pattern, it is given by an equation $\sin \theta = \frac{\lambda}{d}$, where θ is angle of diffraction.



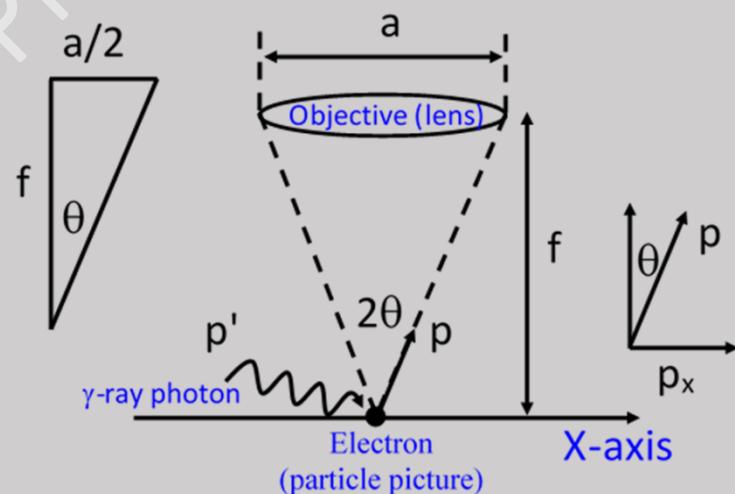
Now, letting the uncertainty in the y -component of momentum to be at the most equal to the y -component of momentum itself,

$\Delta p_y \approx p_y \approx p \sin \theta \approx \frac{h \lambda}{\lambda d} \approx \frac{h}{\Delta y}$; By using de'Broglie's equation, expression for the first minimum in the diffraction pattern and the substitution $\Delta y \approx d$.

Thus, $\Delta y \Delta p_y \approx h$, as required.

Seeing an electron with a gamma-ray microscope – the particle picture:

Here, we consider an electron as a perfect particle. For convenience, we assume that it is almost at rest and we try to locate its position by bouncing off a γ -ray photons from it. This γ -ray is a suitable choice, as one needs a beam having wavelength smaller than the object in order to resolve it. Only γ -rays have wavelength lesser than the de'Broglie wavelength of an electron. We imagine at least a single photon having momentum p' strikes the electron and it is scattered as per the Compton scattering mechanism.



The incident photon imparts some momentum on the electron. The microscope is held in a direction so as to collect the scattered photon which carries the “information” of position of the electron. The formation of “image” by the photon (i.e. position of the electron) is limited by the inherent diffraction effect. The position of the electron (i.e. the “image”) has a spread equal to the width of the central maximum, which is given by

$$\Delta x \approx \frac{\lambda f}{a} \quad \text{---- (1)}$$

Where f is the focal length of the lens and a is aperture size. Both are microscope parameters. This value represents uncertainty in the position of the electron.

Next, let the momentum imparted on the electron along the X-direction be approximately equal to the X-component of the momentum of scattered photon, which is scattered within an angle 2θ towards the microscope. From the diagram, this is $p_x = p \sin \theta$. Taking uncertainty in the momentum of the electron Δp to be at the most equal to p_x , we can write, $\Delta p \approx p \sin \theta \approx \frac{h}{\lambda} \sin \theta$. From the diagram, for very small apertures, $a \ll f$ and $\sin \theta = \frac{a/2}{\sqrt{(a/2)^2 + f^2}} \approx \frac{a}{2f}$. Therefore,

$$\Delta p \approx \frac{ha}{2\lambda f} \quad \text{---- (2)}$$

$$\text{From (1) and (2), } \Delta x \Delta p \approx \frac{\lambda f}{a} \times \frac{ha}{2\lambda f}$$

Thus, $\Delta x \Delta p \approx h/2$, as required.

- Implications of the uncertainty principle**

The uncertainty principle could answer many unsolved questions in different branches of physics such as nuclear, atomic and molecular, solid state, spectroscopy, astrophysics and cosmology. Creation of very short-lived elementary particles, wave-particle duality, non-existence of electrons in the nucleus, zero-point energy possessed by atoms in solids at absolute zero, finite width of spectral lines, limits of measurements of frequencies of radiations, beta-decay (emission of an electron by the nucleus), sharing of electrons in covalent bonds, exact energy value of the ground state and many other effects can be satisfactorily explained with the help of the uncertainty principle. We take a look at some common implications of the uncertainty principle.

- (a) non-existence of electrons in the nucleus:**

The maximum kinetic energies possessed by electrons during beta decay are about 100 keV. Let an electron be a part of the nucleus. The average nuclear dimensions are about 10^{-15} m. Taking $\Delta x \approx 10^{-15}$ m, $\Delta p \approx 1.055 \times 10^{-19}$ kg-m/s. Assuming the momentum possessed by such electron to be at least equal to uncertainty in it, $p \approx 1.055 \times 10^{-19}$ kg-m/s. This electron is indeed a relativistic electron as $v \rightarrow c$. Using the expression of kinetic energy, $K = \sqrt{m_0^2 c^4 + p^2 c^2} - m_0 c^2$. Here, $(m_0 c^2)^2 \ll (pc)^2$. Therefore, $K \approx pc = 3.165 \times 10^{-11}$ J = 197.8 MeV! Electrons never have so large energy and it contradicts the experimental results. Hence the initial argument is wrong and it can be concluded that electrons cannot be constituents of the nucleus.

(b) Minimum energy possessed by an electron in the atom:

Here, the average atomic dimensions are about 10^{-10} m and an electron is a part of the atom. Taking $\Delta x \approx 10^{-10}$ m, $\Delta p \approx 1.055 \times 10^{-24}$ kg-m/s. Assuming the momentum possessed by such electron to be at least equal to uncertainty in it, $p \approx 1.055 \times 10^{-24}$ kg-m/s. This electron is not a relativistic electron as $v \ll c$ for this electron. Using the normal expression for kinetic energy, $K = \frac{p^2}{2m}$, $K = 6.1 \times 10^{-19}$ J = 3.8 eV. It suggests that the minimum energy possessed by any electron in any atom cannot be less than 3.8 eV. Indeed, for the ground state electron in a hydrogen atom, the electron energy is 13.6 eV, which is well-above this limit.

(c) Finite width of spectral lines

Wavelengths (or frequencies) of radiation emitted from any monochromatic source are never sharp at a particular value but always have some “spread”, which is usually referred as “line width”. This is true even for a laser source. In spectroscopy, for many years, it was a wonder but finally, the explanation was provided by the uncertainty principle. The average lifetime of an excited state is typically 10^{-8} sec. When an electron makes a downward jump from such state, it introduces an uncertainty in the exact time interval for de-excitation. Taking maximum uncertainty in the measurement of time to be the lifetime of the excited state, $\Delta E \approx \frac{h/2\pi}{\Delta t}$.

$$\therefore \Delta v = \frac{\Delta E}{h} \approx \frac{1}{2\pi\Delta t} \approx 16 \times 10^6 \text{ Hz} = 16 \text{ MHz.}$$

This is the theoretical line width of any spectral line emitted from a perfect monochromatic source. The important point is that it is *non-zero*.

- **Wave Mechanics**

Material particles follow laws of Newtonian mechanics. These laws and equations are based on classical concept of “perfect particles”. But at the beginning of 20th century, many experimentalists showed that material particles had effects which were typically regarded for many years as “wave properties” (for example, diffraction of electrons). There was no explanation to these newly demonstrated wave properties of matter in the framework of classical Newtonian mechanics. Hence, there was a need of mechanics, which could incorporate the wave nature of matter. This was developed independently by Werner Heisenberg and Erwin Schrödinger. Heisenberg’s formulation contains matrix algebra while that of Schrödinger’s is based on wave equations. Schrödinger’s approach is called as the wave mechanics.

For this work, quantum mechanics introduced new physical properties and concepts. For example, the wave packet was introduced to explain wave nature shown by material particles and later it was supported by a foundation theory. A general practice in quantum mechanics in the early stage was to assign certain properties to microscopic objects, to form some models, to predict some effects and to link them with some observable/measureable macroscopic parameter. Further, this model should be consistent for any other kind of such experimentations. For example, all elementary particles like electrons and protons have been assigned an additional property called the spin and it is modelled that atoms possess magnetic moments due to intrinsic spins as well as orbital motion of electrons. Now, although we can’t see how an electron is, whether it actually “spins” and “revolves” and in what sense it interacts with a magnetic field, we can verify effect of the spin and orbital magnetic moments due to a collection of large number atoms in terms of magnetisation possessed by certain ferromagnetic materials. The amount of magnetisation can be measured experimentally and it

becomes the “observed/measured macroscopic parameter” in this case. Later of course, the concept of spin was supported by a foundation theory developed by Dirac based on relativistic quantum mechanics.

- **The Wave function $\psi(x, t)$**

A wave is said to be doubly periodic phenomenon meaning, in any kind of a wave, we can associate at least one physical quantity which is a periodic function of space (position) and time. For example, in electromagnetic waves, they are the electric and magnetic fields while in sound waves, it is the acoustic pressure that varies and gives rise to a compression or a rarefaction. In matter waves, this job has been assigned to the wave function. The only difference is that unlike electric/magnetic fields or pressure, which are real, physical quantities, the wave function is an abstract mathematical term and it can be a complex quantity to have a more general approach. However, the square of the wave function i.e. $\psi^*\psi$ or $|\psi|^2$ is called as the “probability” of finding the particle and this is a real number between 0 and 1 (or 0% and 100%). Thus, the matter waves or wave associated with material particles are somewhat related to the probabilistic outcome for any kind of measurement performed on the particle. Therefore, sometimes, matter waves are called as “waves of probability”.

- **Conditions on the wave function**

The wave function should satisfy following requirements:

- 1) It must be finite, single valued and continuous (and hence differentiable) everywhere.
- 2) Its derivatives ($\frac{\partial\psi}{\partial x}, \frac{\partial\psi}{\partial t}$ etc.) must also be finite, single valued and continuous.
- 3) It must be “normalised” meaning as $x \rightarrow \pm\infty$, $\psi(x, t) \rightarrow 0$.
- 4) It must have at least one physically acceptable solution.
- 5) It must obey the principle of superposition of waves i.e. if ψ_1 and ψ_2 are solutions of the wave equation then $A\psi_1 + B\psi_2$ is also a solution.

- **Schrödinger's time dependent wave equation (STDE)**

With the wave function being complex and representing periodicity in space and time, we can think of the most general expression for the wave function as $\psi(x, t) = Ae^{i(kx-\omega t)}$; Where, A is amplitude.

Let us develop Schrödinger's equation from this expression of wave function. Taking partial derivatives with respect of space (x) and time (t),

$$\begin{aligned} \frac{\partial\psi}{\partial t} &= Ae^{i(kx-\omega t)} \times (-i\omega) = -i\omega\psi = -i(2\pi\nu)\psi = -i\left(\frac{2\pi E}{h}\right)\psi = -\frac{iE}{\hbar}\psi \\ \therefore i\hbar\frac{\partial\psi}{\partial t} &= E\psi. \end{aligned} \quad (1)$$

Similarly,

$$\frac{\partial\psi}{\partial x} = Ae^{i(kx-\omega t)} \times (ik) = ik\psi = i\left(\frac{2\pi p}{\lambda}\right)\psi = i\left(\frac{2\pi p}{h}\right)\psi = \frac{ip}{\hbar}\psi$$