Presentation Template

G. V. V. Sharma Dept. of Electrical Engg., IIT Hyderabad.

January 21, 2020

Problem

- Solution
 - Linear Equation
 - Matrix Equation
 - Row Reduction
 - Balanced Equation

Problem Statement

Balance the following chemical equation

$$Fe + H_2O \rightarrow Fe_3O_4 + H_2$$
 (2.1)

using a matrix equation.

Linear Equation

Let the balanced version of (2.1) be

$$x_1Fe + x_2H_2O \rightarrow x_3Fe_3O_4 + x_4H_2$$
 (3.1)

which results in the following equations

$$(x_1 - 3x_3) Fe = 0$$

 $(2x_2 - 2x_4) H = 0$
 $(x_2 - 4x_3) H = 0$ (3.2)

Matrix Equation

The linear equations can be expxressed as

$$x_1 + 0.x_2 - 3x_3 + 0.x_4 = 0$$

$$0.x_1 + 2x_2 + 0.x_3 - 2x_4 = 0$$

$$0.x_1 + x_2 - 4x_3 + 0.x_4 = 0$$
(3.3)

resulting in the matrix equation

$$\begin{pmatrix} 1 & 0 & -3 & 0 \\ 0 & 2 & 0 & -2 \\ 0 & 1 & -4 & 0 \end{pmatrix} \mathbf{x} = \mathbf{0}$$
 (3.4)

where

$$\mathbf{x} = \begin{pmatrix} x_1 \\ x_2 \\ x_3 \\ x_4 \end{pmatrix} \tag{3.5}$$

Row Reduction

(3.4) can be row reduced as follows

$$\begin{pmatrix} 1 & 0 & -3 & 0 \\ 0 & 2 & 0 & -2 \\ 0 & 1 & -4 & 0 \end{pmatrix} \xrightarrow{R_2 \leftarrow \frac{R_2}{2}} \begin{pmatrix} 1 & 0 & -3 & 0 \\ 0 & 1 & 0 & -1 \\ 0 & 1 & -4 & 0 \end{pmatrix}$$
(3.6)

$$\stackrel{R_3 \leftarrow R_3 - R_2}{\longleftrightarrow} \begin{pmatrix} 1 & 0 & -3 & 0 \\ 0 & 1 & 0 & -1 \\ 0 & 0 & -4 & 1 \end{pmatrix}$$

$$\stackrel{R_1 \leftarrow 4R_1 - 3R_3}{\longleftrightarrow} \begin{pmatrix} 4 & 0 & 0 & -3 \\ 0 & 1 & 0 & -1 \\ 0 & 0 & 4 & 1 \end{pmatrix}$$

$$\begin{array}{c}
R_1 \leftarrow \frac{1}{4} \\
 \leftarrow R_3 \leftarrow -\frac{1}{4}R_3
\end{array}
\begin{pmatrix}
1 & 0 & 0 & -\frac{3}{4} \\
0 & 1 & 0 & -1 \\
0 & 0 & 1 & -\frac{1}{4}
\end{pmatrix}$$

(3.7)

(3.8)

Balanced Equation

Thus,

$$x_1 = \frac{3}{4}x_4, x_2 = x_4, x_3 = \frac{1}{4}x_4$$
 (3.10)

$$\implies \mathbf{x} = x_4 \begin{pmatrix} \frac{3}{4} \\ 1 \\ \frac{1}{4} \\ 1 \end{pmatrix} = \begin{pmatrix} 3 \\ 4 \\ 1 \\ 4 \end{pmatrix} \tag{3.11}$$

upon substituting $x_4 = 4$. (3.1) then becomes

$$3Fe + 4H_2O \rightarrow Fe_3O_4 + 4H_2$$
 (3.12)

The code in

https://github.com/gadepall/school/raw/master/training/chemistry/codes/chembal.py

verifies (3.10).