



Name _____

Year 10 Science

Chemistry 2 Test: Chemical reactions

Multiple choice answer section.

Circle your choice:

1. A B C D
2. A B C D
3. A B C D
4. A B C D
5. A B C D

6. A B C D
7. A B C D
8. A B C D
9. A B C D
10. A B C D

SECTION 2: WRITTEN

Write your answers in the spaces provided.

1. When David added some hydrochloric acid solution to some zinc metal, he noticed that the zinc bubbled and produced the flammable gas, hydrogen. Zinc chloride was also made.

For this reaction

- a) Name the reactants (2)

hydrochloric acid and zinc(metal)

- b) Name the products (2)

hydrogen (gas) and zinc chloride

- c) Write a word equation for this reaction (1)

hydrochloric acid + zinc \rightarrow hydrogen + zinc chloride

- d) Write a balanced chemical equation for this reaction (2)



- e) List three ways that David could speed up the rate of this reaction? (3)

\rightarrow increase temp/concentration HCl.
 \rightarrow powdered / cut up zinc.
 \rightarrow add catalyst
 \rightarrow stirring / agitation.

2. When fuel, such as ethanol, is burned in oxygen, it releases a lot of heat and light.

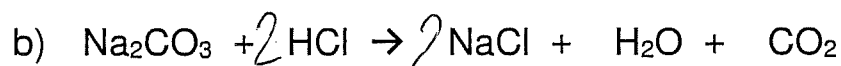
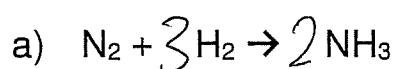
Tick any that apply to this reaction (You can tick more than one) (6)

- ☒ This is classified as combustion.
- ☒ This is a type of exothermic reaction.
- ☒ Ethanol is a reactant.
- ☐ This is a type of decomposition reaction.
- ☒ The reactants contain more chemical potential energy than the products.
- ☐ This reaction would go faster if there was less oxygen in the air.

1.5 each

-1.5 for wrong answer.

3. Balance the following equations (2)



4. What is a catalyst? (2)

→ speeds up reaction
→ not used up / not reactant or product.

End of Test (Out of 30)