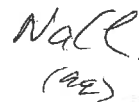
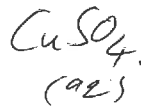


# Suggested.

## SET 1



Id'd.  
CuSO<sub>4</sub>

1. CuSO<sub>4</sub>(aq) is blue put aside.

2. Add some drops of the CuSO<sub>4</sub>(aq) to each of the other 3 sol's.

The BaCl<sub>2</sub>(aq) will produce a ppt.

BaCl<sub>2</sub>

3. Now Add some drops of the BaCl<sub>2</sub>(aq) to each of the remaining...

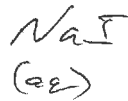
a) When added to H<sub>2</sub>SO<sub>4</sub>(aq) ~~precipitate~~ ∴ H<sub>2</sub>SO<sub>4</sub>

BaSO<sub>4</sub> (s) ppt forms

b) No ppt forms in NaCl(aq)

∴ NaCl.

## SET 2



Insol.  
AgI - pale yellow  
AgCl - white  
BaSO<sub>4</sub> - white  
AgSO<sub>4</sub> - white (slightly sol.)

1. a) ~~Take~~ Take solution G add 2-3 drops to H, I, J. Note no. of ppts.

b) Take H .. add to G I J  
" I " " G H J.  
" J " " G H I

ONE of these will produce 2 OR 3 ppts

This will be AgNO<sub>3</sub> (aq)

2. Observe ppts from 1.

a) The soln which produced a creamy - pale yellow ppt is NaI(aq)

3. The remaining two solutions

are BaCl<sub>2</sub>(aq) — produces the most (whitest) ppt. AgCl(s)

and ZnSO<sub>4</sub>(aq) — produces the slightly soluble (less ppt) AgSO<sub>4</sub>(s).