



Science Department Multiple Choice Answer Sheet

Test Topic: Chem 2 Name: ANSWERS Year:

Multiple Choice – 20 questions.

Circle your choice. If you change your mind, scrub your choice out and circle the one you want. If it is messy, clearly write your choice next to question.

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|-----|--------------|--------------|--------------|--------------|-----|--------------|--------------|--------------|--------------|
| 1. | A | B | C | D | 11. | A | B | C | D |
| 2. | A | B | C | D | 12. | A | B | C | D |
| 3. | A | B | C | D | 13. | A | B | C | D |
| 4. | A | B | C | D | 14. | A | B | C | D |
| 5. | A | B | C | D | 15. | A | B | C | D |
| 6. | A | B | C | D | 16. | A | B | C | D |
| 7. | A | B | C | D | 17. | A | B | C | D |
| 8. | A | B | C | D | 18. | A | B | C | D |
| 9. | A | B | C | D | 19. | A | B | C | D |
| 10. | A | B | C | D | 20. | A | B | C | D |

Written Section:

Write your answers for the written section below. Ask your teacher if you need more paper.

SECTION 2: WRITTEN ANSWER (1 mark each)

Write your answer in the spaces provided on this sheet.

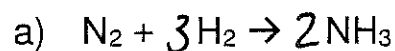
1. Match the chemical equation to the reaction type by writing the reaction type's number next to each equation. (5)

5	a) $\text{Ag}^+_{(\text{aq})} + \text{Cl}^-_{(\text{aq})} \rightarrow \text{AgCl}_{(\text{s})}$	1 Combination
2	b) $\text{CuCO}_{3(\text{s})} \rightarrow \text{CuO}_{(\text{s})} + \text{CO}_{2(\text{g})}$	2 Decomposition
4	c) $2\text{C}_8\text{H}_{18(\text{l})} + 25\text{O}_{2(\text{g})} \rightarrow 16\text{CO}_{2(\text{g})} + 18\text{H}_2\text{O}_{(\text{g})}$	3 Displacement
1	d) $\text{Zn}_{(\text{s})} + \text{I}_{2(\text{s})} \rightarrow \text{ZnI}_{2(\text{s})}$	4 Combustion
3	e) $\text{Mg}_{(\text{s})} + 2\text{HCl}_{(\text{aq})} \rightarrow \text{MgCl}_{2(\text{aq})} + \text{H}_{2(\text{g})}$	5 Precipitation

2. Write the formula of each of the following substances, then classify each of these as acid, base, salt, carbonate or gas (10)

	Formula	Classification
a) sodium hydroxide	NaOH	base
b) potassium chloride	KCl	salt
c) sulphuric acid	H ₂ SO ₄	acid.
d) copper (II) carbonate	CuCO ₃	carbonate
e) carbon dioxide	CO ₂	gas.

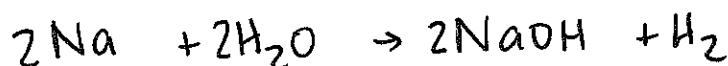
3. Balance the following equations (2)



4. Write a balanced chemical equation from the following word equations

(2)

a) sodium + water \rightarrow sodium hydroxide + hydrogen



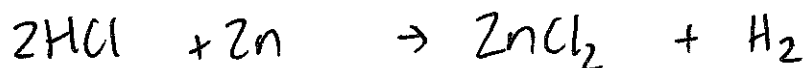
b) calcium carbonate + hydrochloric acid \rightarrow calcium chloride + carbon dioxide + water



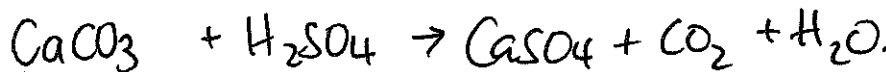
5. Complete the word equation by predicting the products and then write a balanced chemical equation for the full reaction

(4)

a) hydrochloric acid + zinc \rightarrow zinc chloride + hydrogen



b) ^{sodium* change for next year} calcium carbonate + sulphuric acid \rightarrow calcium sulfate + carbon dioxide + water.



6. Define nanochemicals and give one example of a current application.

(2)

particles 1-100 nm ($\frac{1}{100}$ billion to $\frac{100}{100}$ billion m).

eg dirt proof clothing

sunscreen

medication that targets cells

carbon capture filters

etc.

End of Test (Out of 40)



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