

Question 31**(6 marks)**

- (a) Select **one** basic, **one** acidic and **one** neutral salt from the list below to complete the table. (3 marks)

KCN, NH_4Cl , $\text{Mg}_3(\text{PO}_4)_2$, NaNO_3 , KHCO_3 , NaCH_3COO , KCl

Acidic salt	Neutral salt	Basic salt

- (b) Use the Brønsted-Lowry model to explain why the pH of ammonia solution is greater than 7.0 at 25 °C. Incorporate at least **one** appropriate equation in your answer. (3 marks)
