9.   The net ionic equation for the predominant hydrolysis reaction occurring in a 1.00 mol L⁻¹ potassium hydrogensulfate solution is:

(a) $KHSO_4(aq) + H_2O(\ell) \rightleftharpoons K^+(aq) + H_2SO_4(aq) + OH^-(aq)$
(b) $K^+(aq) + HSO_4^-(aq) \rightleftharpoons K^+(aq) + H^+(aq) + SO_4^{2-}(aq)$
(c) $HSO_4^-(aq) + H_2O(\ell) \rightleftharpoons H_2SO_4(aq) + OH^-(aq)$
(d) $HSO_4^-(aq) + H_2O(\ell) \rightleftharpoons SO_4^{2-}(aq) + H_3O^+(aq)$