20. An electrochemical cell, reaction shown below, has an E° value of +0.89 V.

$$4 V^{3+}(aq) + 2 H_2O(\ell) + O_2(g) = 4 VO^{2+}(aq) + 4 H^{+}(aq)$$

What is the standard reduction potential for the half-equation below?

$$VO^{2+} + 2 H^+ + e^- \hookrightarrow V^{3+} + H_2O$$

- (a) -0.34 V
- (b) +1.57 V
- (c) +0.34 V
- (d) -1.57 V