

17. Three metals, A, B and C, were tested to compare their reactivity. Samples of each metal were placed separately into test tubes each containing a nitrate solution of the other metal ions. The following results were obtained.

	<b>A(s)</b>	<b>B(s)</b>	<b>C(s)</b>
$A^{2+}(aq)$		No visible reaction	Solid A forms
$B^{2+}(aq)$	Solid B forms		Solid B forms
$C^{2+}(aq)$	No visible reaction	No visible reaction	

From these results, the metals arranged in order of decreasing strength as reducing agents can be concluded to be

- (a)  $C > A > B$ .
- (b)  $B > C > A$ .
- (c)  $B > A > C$ .
- (d)  $A > C > B$ .