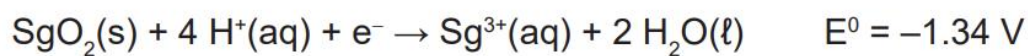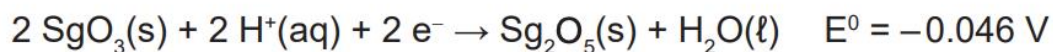19. The following half-equations show some predicted standard reduction potentials for seaborgium (Sg) oxides:

$$2\ SgO_3(s) + 2\ H^+(aq) + 2\ e^- \rightarrow Sg_2O_5(s) + H_2O(\ell) \quad E^0 = -0.046\ V$$

$$Sg_2O_5(s) + 2\ H^+(aq) + 2\ e^- \rightarrow 2\ SgO_2(s) + H_2O(\ell) \quad E^0 = +0.11\ V$$

$$SgO_2(s) + 4\ H^+(aq) + e^- \rightarrow Sg^{3+}(aq) + 2\ H_2O(\ell) \quad E^0 = -1.34\ V$$

The strongest reducing agent is

(a) SgO_3
(b) Sg_2O_5
(c) SgO_2
(d) Sg^{3+}