

CHEMICAL REACTIONS > 2 mg o (s) - Combination $1.2mg(s) + O_2(g)$ (Yellow) +2KNO3(aq) 2. 2. Pb(NO2) 2(aq) + ZKI (aq) Zn (2(aq) + H2(1) - Displacement 3. Zn(s) +2H(l(ag)) (a(OH)2(aq) + Heat -> (ombination 4. CaO(s) + H2O(2) (Slaked Lime)

(Slaked Lime) (Quick lime) 5.2 fe SOy(s) -3 (Gureen (olows) 6. Ca(O3(s) Heat (aO(s) + (O2(1) - Decomposition (Limestone) (Quick lime) 7. 2Pb(NO3)2(5) Heat > 2PbO(s) + 4NO2(g) + O2(g) (Yellow) (Byoun) Electrolytic secomposition 8. 2 H2O(1) Electricity 2 H2 (g) + O2(g) -9. 2Rg Cl(s) Sunlight, 2Rg(s) + Cl2(g)--> Photolytic decomposition (White) (Gorey) 10. 2AgBH(s) Sunlight > 2Ag(s) + BH2(g) -> Black-white photographs 11. fe(s) + (usoy(ag) --) fesoy(ag) + (u(s) -) Displacement (Blue) (Colowdess) 12. Nozsoy(ag) + Bacl2(ag) - Basoy(s) + 2Nacl(ag) - Cwhite)

Double Double 13. $(u(s) + O_2(g) \longrightarrow (uO(s) - DCombination)$ Displacem (Black) 14. (00(s) + Hz(g) Heat Cu + HzO - Displacement 15. Zn(s) + Cusoy (aq) -> ZnSoylag) + Culs)>