

## Worksheet: Measure heat capacity with a calorimeter

Name(s) \_\_\_\_\_

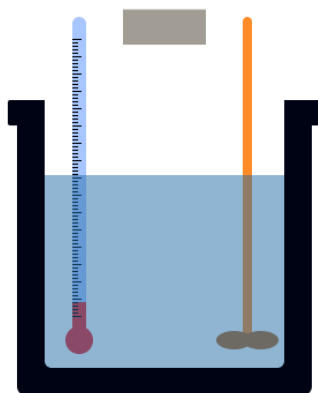
In this experiment, a hot metal cube is dropped into an insulated container of water and the increase in the water temperature is used to determine the heat capacity of the metal.

### Student Learning Objectives- Measure heat capacity with a calorimeter

1. Be able to explain the concept of heat capacity.
2. Be able to apply the concept of an adiabatic process.
3. Be able to apply the first law for a closed system.
4. Be aware of assumptions made and sources of error in using a static calorimeter.
5. Learn proper use and calibration of calorimeters.
6. Apply statistics to determine heat capacity from repeat experiments and determine standard deviation.

### Equipment

A well-insulated calorimeter that contains 1000 mL of water and a stirrer. A thermometer to record the water temperature.



### Questions to answer before starting experiment

In this experiment, 1000 mL of water is in an insulated container.

What is the advantage of using more water?

What is the disadvantage of using more water?

What is the advantage and disadvantage of using more metal?

What is the advantage and disadvantage of heating the metal to a higher temperature?

### Assumptions

- The calorimeter is well insulated.
- The heat capacity of the metal sample is independent of temperature.

### Calibrate the calorimeter.

1. Turn on the stirrer.
2. Measure the initial temperature of water.
3. Record mass of water in calorimeter
4. Take a sample of a known metal (iron), record its mass and its temperature, and drop it into the calorimeter.
5. Record final equilibrium temperature.
6. Calculate heat absorbed by calorimeter to determine its heat capacity

$$m_{HW} C_{PW} (T_{HW} - T_{final}) = m_{CW} C_{PW} (T_{final} - T_{CW}) + m_{cal} C_{Pcal} (T_{final} - T_{CW})$$

7. where  $m_{HW}$  = mass of hot water added to calorimeter
8.  $C_{PW}$  = water heat capacity
9.  $T_{final}$  = final equilibrium temperature
10.  $m_{CW}$  = mass of cold water in calorimeter initially
11.  $T_{CW}$  = initial temperature water in calorimeter
12.  $m_{cal} C_{Pcal}$  = mass of calorimeter (not including water) x heat capacity of calorimeter. This is unknown and is calculated from the above equation.

Value of  $m_{cal} C_{Pcal}$  = \_\_\_\_\_

### Measure heat capacity of unknown sample

1. Select a sample (A, B, C, D, E) from the drop-down menu.
2. Turn on stirrer.
3. Record mass of water in calorimeter  $m_{CW}$ .
4. Record starting temperature of water  $T_{CW}$ .
5. Record the starting temperature of the solid sample  $T_{sample}$ .
6. Record mass of solid sample  $m_{sample}$ .
7. Drop solid into water.
8. Record temperature as function of time.
9. Determine when temperature stabilizes and record final temperature  $T_{final}$ .

### Calculate the heat capacity of sample.

Use this energy balance to calculate the heat capacity  $C_{Psample}$  of the sample:

$$m_{sample} C_{Psample} (T_{sample} - T_{final}) = m_{CW} C_{PW} (T_{final} - T_{CW}) + m_{cal} C_{Pcal} (T_{final} - T_{CW})$$

Repeat the experiment at least 3 times and calculate the average value of heat capacity (report units) and standard deviation.

$$C_{Psample} = \underline{\hspace{2cm}} \pm \underline{\hspace{2cm}}$$

### What can cause errors in the value of heat capacity you measured?

1. Does the stirrer add energy to the water?
2. How might you determine how good the assumption is that the system is adiabatic?