

## Concept Video 7- Q 1

Give the four quantum numbers that describe the 5<sup>th</sup> electron of a ground state Carbon atom

# Concept Videos 7 – Q 2

Write the ground state electronic configuration (condensed), and the number of valence electrons for the following elements.

- 1) Te
- 2) I
- 3) Rb
- 4) Ga
- 5) As

**Periodic Table of the Elements**

1 1A 1A	2 IIA 2A											13 IIIA 3A	14 IVA 4A	15 VA 5A	16 VIA 6A	17 VIIA 7A	18 VIIIA 8A
1 H Hydrogen 1.008	2 He Helium 4.003											5 B Boron 10.811	6 C Carbon 12.011	7 N Nitrogen 14.007	8 O Oxygen 15.999	9 F Fluorine 18.998	10 Ne Neon 20.180
3 Li Lithium 6.941	4 Be Beryllium 9.012											13 Al Aluminum 26.982	14 Si Silicon 28.086	15 P Phosphorus 30.974	16 S Sulfur 32.066	17 Cl Chlorine 35.453	18 Ar Argon 39.948
11 Na Sodium 22.990	12 Mg Magnesium 24.305	3 III B 3B	4 IV B 4B	5 V B 5B	6 VI B 6B	7 VII B 7B	8 VIII 8	9 VIII 9	10 VIII 10	11 IB 1B	12 IIB 2B	31 Ga Gallium 69.723	32 Ge Germanium 72.631	33 As Arsenic 74.922	34 Se Selenium 78.971	35 Br Bromine 79.904	36 Kr Krypton 83.798
19 K Potassium 39.098	20 Ca Calcium 40.078	21 Sc Scandium 44.956	22 Ti Titanium 47.867	23 V Vanadium 50.942	24 Cr Chromium 51.996	25 Mn Manganese 54.938	26 Fe Iron 55.845	27 Co Cobalt 58.933	28 Ni Nickel 58.693	29 Cu Copper 63.546	30 Zn Zinc 65.38	49 In Indium 114.818	50 Sn Tin 118.711	51 Sb Antimony 121.760	52 Te Tellurium 127.6	53 I Iodine 126.904	54 Xe Xenon 131.294
37 Rb Rubidium 85.468	38 Sr Strontium 87.62	39 Y Yttrium 88.906	40 Zr Zirconium 91.224	41 Nb Niobium 92.906	42 Mo Molybdenum 95.95	43 Tc Technetium 98.907	44 Ru Ruthenium 101.07	45 Rh Rhodium 102.906	46 Pd Palladium 106.42	47 Ag Silver 107.868	48 Cd Cadmium 112.414	81 Tl Thallium 204.383	82 Pb Lead 207.2	83 Bi Bismuth 208.980	84 Po Polonium [209]	85 At Astatine 209.987	86 Rn Radon 222.018
55 Cs Cesium 132.905	56 Ba Barium 137.328	57-71 Lanthanide Series	72 Hf Hafnium 178.49	73 Ta Tantalum 180.948	74 W Tungsten 183.84	75 Re Rhenium 186.207	76 Os Osmium 190.23	77 Ir Iridium 192.217	78 Pt Platinum 195.085	79 Au Gold 196.967	80 Hg Mercury 200.592	113 Nh Nihonium [286]	114 Fl Flerovium [289]	115 Mc Moscovium [289]	116 Lv Livermorium [293]	117 Ts Tennessine [294]	118 Og Oganesson [294]
87 Fr Francium 223.020	88 Ra Radium 226.025	89-103 Actinide Series	104 Rf Rutherfordium [261]	105 Db Dubnium [262]	106 Sg Seaborgium [266]	107 Bh Bohrium [264]	108 Hs Hassium [269]	109 Mt Meitnerium [278]	110 Ds Darmstadtium [281]	111 Rg Roentgenium [280]	112 Cn Copernicium [285]						
			57 La Lanthanum 138.905	58 Ce Cerium 140.116	59 Pr Praseodymium 140.908	60 Nd Neodymium 144.243	61 Pm Promethium 144.913	62 Sm Samarium 150.36	63 Eu Europium 151.964	64 Gd Gadolinium 157.25	65 Tb Terbium 158.925	66 Dy Dysprosium 162.500	67 Ho Holmium 164.930	68 Er Erbium 167.259	69 Tm Thulium 168.934	70 Yb Ytterbium 173.055	71 Lu Lutetium 174.967
			89 Ac Actinium 227.028	90 Th Thorium 232.038	91 Pa Protactinium 231.036	92 U Uranium 238.029	93 Np Neptunium 237.048	94 Pu Plutonium 244.064	95 Am Americium 243.061	96 Cm Curium 247.070	97 Bk Berkelium 247.070	98 Cf Californium 251.080	99 Es Einsteinium [254]	100 Fm Fermium 257.095	101 Md Mendelevium 258.1	102 No Nobelium 259.101	103 Lr Lawrencium [262]

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# Concept Video 9

## Practice Q 1

Rank each of these main-group elements in decreasing atomic size?

1) Br, Cs, Ar

2) Cl, Mg, Rb

# Concept Video 9

## Practice Q4

Among the following which of the following ionizations would require the least amount of energy?

1. From  $\text{Si}_{(\text{g})} ([\text{Ne}]3s^23p^2)$  to  $\text{Si}^+_{(\text{g})} ([\text{Ne}] 3s^13p^2) + e^-$
2. From  $\text{Si}_{(\text{g})} ([\text{Ne}]3s^23p^2)$  to  $\text{Si}^+_{(\text{g})} ([\text{Ne}] 3s^23p^1) + e^-$
3. From  $\text{Si}^+_{(\text{g})} ([\text{Ne}]3s^23p^1)$  to  $\text{Si}^{2+}_{(\text{g})} ([\text{Ne}] 3s^2) + e^-$

# Concept Video 9

## Practice Q5

Using electronic configuration, describe  $IE_1$ ,  $IE_2$ , and  $IE_3$  of the following atoms i.e. which electrons are lost in each IE reaction?

Circle the largest IE (among  $IE_1/IE_2/IE_3$  for that atom) for each atom

C

Na

# Concept Video 9

## Practice Q6

An element belonging to period 3 has the following ionization energies? Predict which element is it?

IE in kJ/mol:

$$IE_1 = 738; IE_2 = 1451; IE_3: 7732; IE_4: 10542$$

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## Practice Question 7

Arrange the following in decreasing ionic size:

$S^{2-}$ ,  $O^{2-}$ ,  $Be^{2+}$ ,  $Li^{+}$

# Concept Video 9

## Practice Question 8

Provide the condensed electronic configuration of the following and determine if they are paramagnetic or diamagnetic

Ti

Zn<sup>2+</sup>

Ti<sup>2+</sup>

Fe<sup>3+</sup>

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