

***Ionic and Molecular:***

NOTE: Watch out for multivalent metals in the ionic compounds!! Use common names!

1. Na <sub>3</sub> N	_____.	2. nitrogen trihydride	_____.
3. HgCl <sub>2</sub>	_____.	4. water	_____.
5. N <sub>2</sub>	_____.	6. lithium chloride	_____.
7. ZrO <sub>2</sub>	_____.	8. cobalt(III) iodide	_____.
9. Ca <sub>2</sub> C	_____.	10. arsenic trioxide	_____.
11. NH <sub>3</sub>	_____.	12. gold(I) phosphide	_____.
13. KCl	_____.	14. selenium hexabromide	_____.
15. HF	_____.	16. gallium fluoride	_____.
17. N <sub>2</sub> O <sub>3</sub>	_____.	18. bromine	_____.

***Polyatomic (Set 3):***

1. HNO <sub>2</sub>	_____.	2. magnesium sulfate	_____.
3. Ca <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub>	_____.	4. ammonium chlorate	_____.
5. Ba(ClO <sub>3</sub> ) <sub>2</sub>	_____.	6. magnesium sulfite	_____.
7. Cu <sub>2</sub> SO <sub>3</sub>	_____.	8. copper(II) nitrate	_____.
9. F <sub>2</sub>	_____.	10. aluminum sulfate	_____.
11. NH <sub>4</sub> OH	_____.	12. iron(III) phosphate	_____.

**IONIC SPOT THE ERROR**

Name > Formula	Error	Corrected!
manganese (II) oxide > MnO <sub>2</sub>		
Fe(III)Cl > FeCl <sub>3</sub>		
potassium (I) nitrate > KNO <sub>3</sub>		
Potassium chloride > PCl		
Iron (II) fluoride > Fe(II)F <sub>2</sub>		