NOMENCLATURE SET 1 A IONIC COMPOUNDS

Provide the matching name or formula for the following ionic compounds.

1. NaI	_____.		2. calcium bromide	_____.
3. AlP	_____.		4. sodium oxide	_____.
5. $RaCl_2$	_____.		6. aluminum iodide	_____.
7. KF	_____.		8. barium chloride	_____.
9. ZnS	_____.		10. beryllium chloride	_____.
11. BF_3	_____.		12. francium sulfide	_____.
13. AlN	_____.		14. hydrogen fluoride	_____.
15. Ag_2S	_____.		16. potassium iodide	_____.
17. Li_3N	_____.		18. magnesium fluoride	_____.
19. Be_3N_2	_____.			

Covalent

Provide the matching name or formula for the following molecules (prefixes!)

1. SO_2	_____.		2. carbon monoxide	_____.
3. PF_3	_____.		4. sulfur hexachloride	_____.
5. CCl_4	_____.		6. diarsenic trisulfide	_____.
7. CO_2	_____.		8. carbon tetrafluoride	_____.
9. Cl_2O	_____.		10. dihydrogen oxide	_____.
11. NO_2	_____.		12. oxygen difluoride	_____.
13. NO	_____.		14. arsenic pentaiodide	_____.
			15. iodine heptafluoride	_____.

NOMENCLATURE SET 1 B IONIC COMPOUNDS

Provide the matching name or formula for the following ionic compounds.

1. sodium iodide	_____.		2. $CaBr_2$	_____.
3. aluminum phosphide	_____.		4. Na_2O	_____.
5. radium chloride	_____.		6. AlI_3	_____.
7. potassium fluoride	_____.		8. $BaCl_2$	_____.
9. zinc sulfide	_____.		10. $BeCl_2$	_____.
11. boron fluoride	_____.		12. Fr_2S	_____.
13. aluminum nitride	_____.		14. HF	_____.
15. silver sulfide	_____.		16. KI	_____.
17. lithium nitride	_____.		18. MgF_2	_____.
19. beryllium nitride	_____.			

Covalent

Provide the matching name or formula for the following molecules (prefixes!)

1. sulfur dioxide	_____.		2. CO	_____.
3. phosphorous trifluoride	_____.		4. SCl_6	_____.
5. carbon tetrachloride	_____.		6. As_2S_3	_____.
7. carbon dioxide	_____.		8. CF_4	_____.
9. dichlorine monoxide	_____.		10. water	_____.
11. nitrogen dioxide	_____.		12. OF_2	_____.
13. nitrogen oxide	_____.		14. AsI_5	_____.
			15. IF_7	_____.

NOMENCLATURE SET 2 A IONIC, MOLECULAR and POLYATOMIC

Ionic and Molecular Compounds:

Determine whether each of the following is an ionic or molecular compound and provide the name or formula using the correct system (IUPAC or prefix)

1. HCl	_____.		2. nitrogen gas	_____.
3. H_2S	_____.		4. gallium carbide	_____.
5. SrO	_____.		6. magnesium oxide	_____.
7. Ga_2O_3	_____.		8. chlorine heptaiodide	_____.
9. $PbCl_4$	_____.		10. calcium sulfide	_____.
11. CO	_____.		12. aluminum phosphide	_____.
13. Ba_3N_2	_____.		14. hydrofluoric acid	_____.
15. Ni_2C	_____.		16. diphosphorus trisulfide	_____.
17. H_2O	_____.		18. strontium nitride	_____.
19. Cs_3N	_____.		20. calcium phosphide	_____.

Polyatomic Compounds:

1. $CaSO_4$	_____.		2. lithium sulfate	_____.
3. $Zn(OH)_2$	_____.		4. potassium chlorite	_____.
5. $AlPO_4$	_____.		6. silver nitrite	_____.
7. $RbNO_3$	_____.		8. sodium carbonate	_____.
9. Na_2SO_3	_____.		10. strontium nitrite	_____.
11. $KClO_3$	_____.		12. magnesium bicarbonate	_____.
13. Ag_3PO_4	_____.		14. calcium phosphate	_____.
15. NaOH	_____.		16. lithium carbonate	_____.
17. $Al(NO_3)_3$	_____.		18. silver hydroxide	_____.

NOMENCLATURE SET 2 B IONIC, MOLECULAR and POLYATOMIC

Ionic and Molecular Compounds:

Determine whether each of the following is an ionic or molecular compound and provide the name or formula using the correct system (IUPAC or prefix)

1. hydrogen chloride	_____.		2. N_2	_____.	
3. hydrogen sulfide	_____.		4. Ga_4C_3	_____.	
5. strontium oxide	_____.		6. MgO	_____.	
7. gallium oxide	_____.		8. ClI_7	_____.	
9. lead (IV) chloride	_____.		10. CaS	_____.	
11. carbon monoxide	_____.		12. AlP	_____.	
13. barium nitride	_____.		14. HF	_____.	
15. nickel (II) carbide	_____.		16. P_2S_3	_____.	
17. water	_____.		18. Sr_3N_2	_____.	
19. cesium nitride	_____.		20. Ca_3P_2	_____.	

Polyatomic Compounds:

1. calcium sulfate	_____.		2. Li_2SO_4	_____.	
3. zinc hydroxide	_____.		4. $KClO_2$	_____.	
5. aluminum phosphate	_____.		6. $AgNO_2$	_____.	
7. rubidium nitrate	_____.		8. Na_2CO_3	_____.	
9. sodium sulfite	_____.		10. $Sr(NO_2)_2$	_____.	
11. potassium chlorate	_____.		12. $Mg(HCO_{3.})_2$	_____.	
13. silver phosphate	_____.		14. $Ca_3(PO_4)_2$	_____.	
15. sodium hydroxide	_____.		16. Li_2CO_3	_____.	
17. aluminum nitrate	_____.		18. AgOH	_____.	