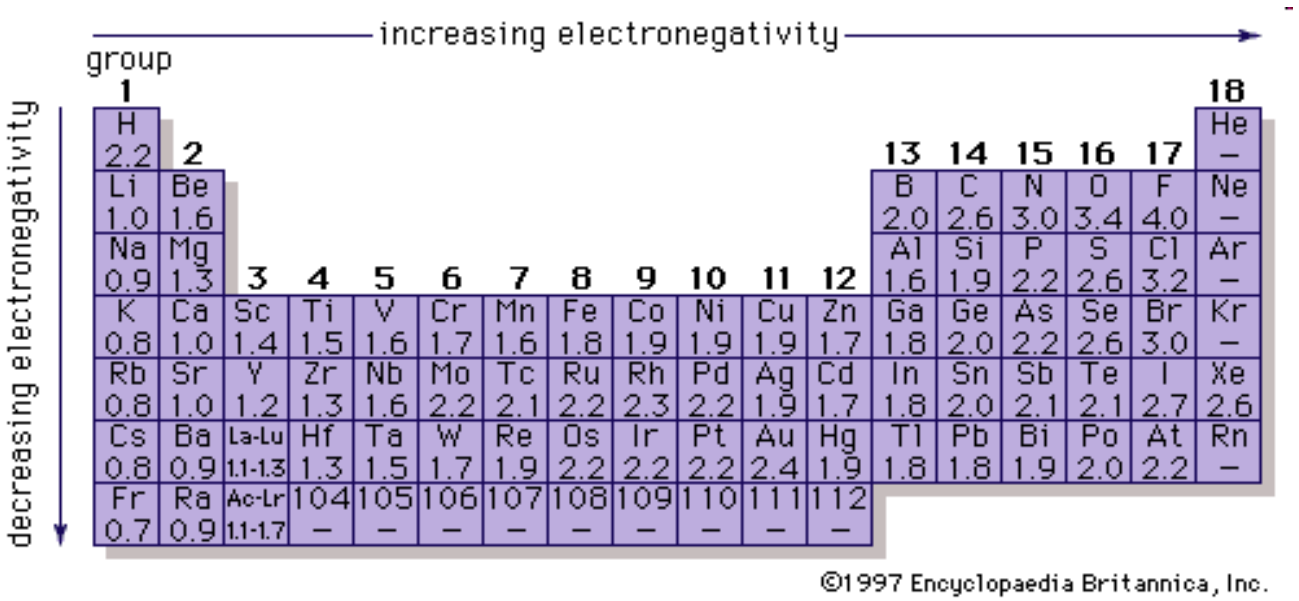
Bond Type Practice



Directions: Determine the type of bond that will form between each pair of atoms in the table below. Use the Electronegativity Chart above. Then provide the name of the compound or molecule.

	Chemical formula	**Compound Name**	**Atom 1**	**Atom 2**	**Electronegativity Difference** **(show work!)**	**Bond Type** (Nonpolar Covalent (NPC), Polar Covalent (PC), or Ionic (I))
1.	$CuBr_2$		Copper	Bromine		
2.	SeI_2		Selenium	Iodine		
3.	SiF_2		Silicon	Fluorine		
4.	K_3N		Potassium	Nitrogen		
5.	Li_2O		Lithium	Oxygen		
6.	H_2O		Hydrogen	Oxygen		
7.	CaS		Calcium	Sulfur		
8.	NBr_3		Nitrogen	Bromine		
9.	CO_2		Carbon	Oxygen		

10.	PF_3		Phosphorus	Fluorine		
11.	NaCl		Sodium	Chloride		
12.	Mg_3N_2		Magnesium	Nitrogen		
13.	Al_2S_3		Aluminum	Sulfur		

14. For each of the following molecules:
    a. Draw the Lewis dot structure
    b. Predict the VSEPR shape
    c. Build the molecule to confirm the shape.

A. NCl_3 Shape: _____	E. CH_3Cl Shape: _____
B. SCl_2 Shape: _____	F. HCl Shape: _____
C. CH_4 Shape: _____	G. CO_2 Shape: _____
D. H_2O Shape: _____	H. CH_2O Shape: _____