Covalent Bonding

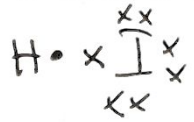
raw the Electron Dot diagrams, then show the resulting covalently bonding compound.

1.Hydrogen and Iodine

State the FORMULA of the compound

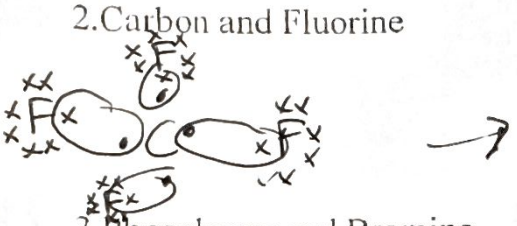
* ONE compound is IONIC ... which one?

NOT →



Hydrogen iodide

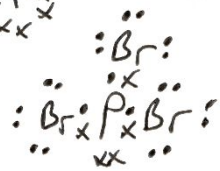
2.Carbon and Fluorine



Both electrons should be in between the symbols!

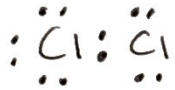
carbon tetrafluoride

3.Phosphorus and Bromine



phosphorous tribromide

4.Chlorine and Chlorine * Final structures only shown for rest.



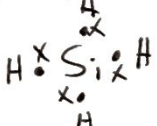
chlorine (Note: it's a HOFBrINCl so Not dichloride)

Hydrogen and Rubidium

→ H:Rb

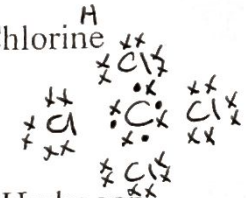
* an exceptional case, not fully covalent OR fully ionic

6.Silicon and Hydrogen



silicon tetrahydride

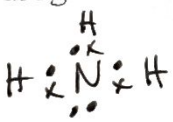
7.Carbon and Chlorine



carbon tetrachloride

* Be sure to memorize prefixes for 1-10!! *

8.Nitrogen and Hydrogen



nitrogen trihydride

9. Oxygen and Chlorine

oxygen dichloride