CHEMISTRY SINGLE REPLACEMENT REACTION WORKSHEET

Using the Activity Series Table, complete the following reactions by writing the products that are formed. Be sure to Balance each equation. If No single replacement reaction occurs, write NR to the right of the arrow.

1. $Ag + KNO_3 \rightarrow$
2. $Zn + AgNO_3 \rightarrow$
3. $Al + H_2SO_4 \rightarrow$
4. $Cl_2 + KI \rightarrow$
5. $Li + H_2O \rightarrow$
6. $Cu + FeSO_4 \rightarrow$
7. $Na + H_2O \rightarrow$
8. $Fe + Pb(NO_3)_2 \rightarrow$
9. $Cu + H_2O \rightarrow$
10. $Cu + Al_2(SO_4) \rightarrow$
11. $Al + Pb(NO_3)_2 \rightarrow$
12. $Cl_2 + NaI \rightarrow$
13. $Fe + AgCH_3COO \rightarrow$
14. $Al + CuCl_2 \rightarrow$
15. $Br_2 + CaI_2 \rightarrow$
16. $Al + HCl \rightarrow$
17. $Mg + HCl \rightarrow$
18. $Zn + H_2SO_4 \rightarrow$
19. $Fe + CuSO_4 \rightarrow$
20. $Cl_2 + MgI_2 \rightarrow$

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ACTIVITY SERIES REFERENCE CHART

	METALS	HALOGENS
MOST REACTIVE	Lithium	Fluorine
	Rubidium	Chlorine
	Potassium	Bromine
	Caesium	Iodine
	Barium	
	Strontium	
	Calcium	
	Sodium	
	Magnesium	
	Aluminum	
	Titanium	
	Manganese	
	Zinc	
	Chromium	
	Iron	
	Cobalt	
	Nickel	
	Tin	
	Lead	
	* Hydrogen	
	Copper	
	Mercury	
	Silver	
	Platinum	
	Gold	
LEAST REACTIVE		

In general, a metal can displace any of the metals which are lower in the reactivity series.
Same rule applies for Halogens.

*Metals from Li to Na will replace H from acid and water
Metals from Mg to Pb will replace H from acids only.