

### **Nomenclature Worksheet # 3**

**Name the following**

**Write the formula**

- |  |                                  |
|--|----------------------------------|
| 1. $\text{Al}_2(\text{SO}_4)_3$ _____  | 17. barium bromate _____         |
| 2. $\text{Mg}(\text{NO}_3)_2$ _____    | 18. potassium permanganate _____ |
| 3. $\text{NH}_4\text{ClO}_3$ _____     | 19. copper(I) nitrite _____      |
| 4. $\text{NaCH}_3\text{COO}$ _____     | 20. sodium oxide _____           |
| 5. $\text{NaIO}_3$ _____               | 21. barium phosphate _____       |
| 6. $\text{HgClO}_2$ _____              | 22. tin(IV) chlorate _____       |
| 7. $\text{Fe}_3(\text{PO}_4)_2$ _____  | 23. lead(II) nitrate _____       |
| 8. $\text{Mg}(\text{NO}_2)_2$ _____    | 24. sodium perchlorate _____     |
| 9. $\text{CaSO}_3$ _____               | 25. aluminum hydroxide _____     |
| 10. $\text{K}_2\text{O}_2$ _____       | 26. mercury(II) nitrite _____    |
| 11. $\text{Fe}_2(\text{CO}_3)_3$ _____ | 27. stannous perchlorate _____   |
| 12. $\text{Al}(\text{ClO}_4)_3$ _____  | 28. plumbic dichromate _____     |
| 13. $\text{Al}(\text{OH})_3$ _____     | 29. copper(I) sulfite _____      |
| 14. $(\text{NH}_4)_2\text{S}$ _____    | 30. aluminum cyanide _____       |
| 15. $\text{Ba}(\text{OH})_2$ _____     | 31. tin(IV) hypobromite _____    |
| 16. $\text{Ge}(\text{SO}_2)_2$ _____   | 32. auric chlorite _____         |