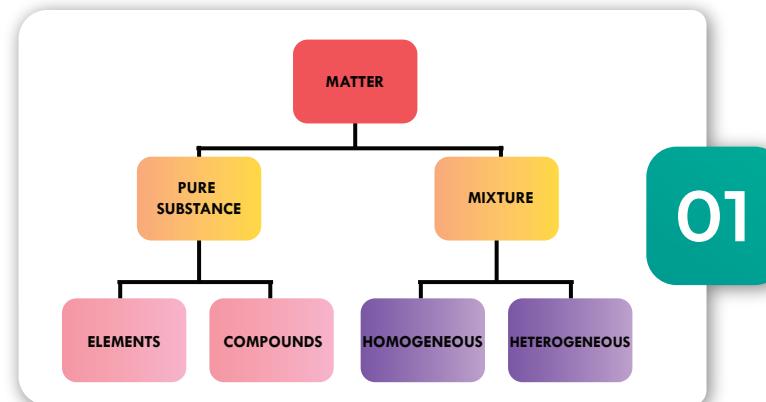
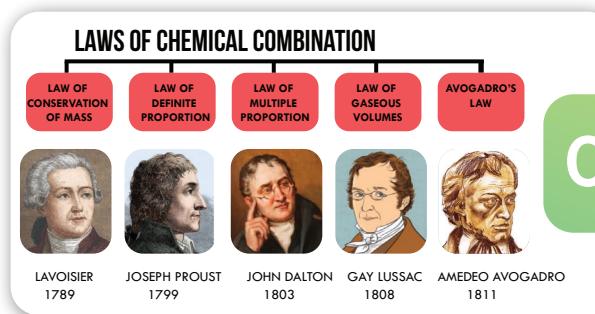


CHEMISTRY



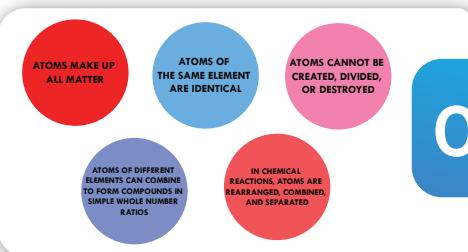
01

- Q) Which one of the following is not a mixture ?
 (A) Tap water (B) Distilled water
 (C) Salt in water (D) Oil in water



02

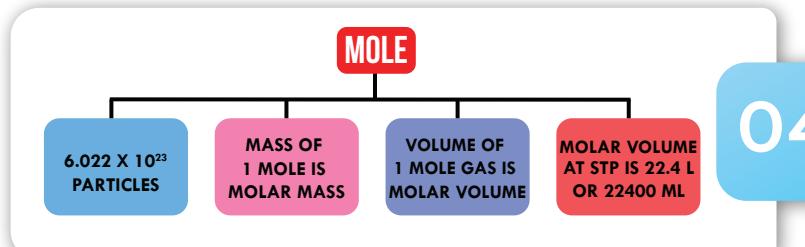
- Q) Which one of the following pairs of compound illustrate the law of multiple proportions ?
 (A) H_2O , Na_2O (B) MgO , Na_2O
 (C) Na_2O , BaO (D) SnCl_2 , SnCl_4



03

DALTON'S ATOMIC THEORY

MOLE CONCEPT



04

- Q) Which one of the followings has maximum number of atoms ?
 (A) 1 g of Mg(s) [Atomic mass of Mg = 24]
 (B) 1 g of O_2 [Atomic mass of O=16]
 (C) 1 g of Li(s) [Atomic mass of Li = 7]
 (D) 1 g of Ag(s) [Atomic mass of Ag = 108]

04

PERCENTAGE COMPOSITION

05

05

Percentage composition of H_2O		Percentage composition of CO_2	
MASS % OF H $\frac{2}{18} \times 100$	MASS % OF O $\frac{16}{18} \times 100$	MASS % OF C $\frac{12}{44} \times 100$	MASS % OF O $\frac{32}{44} \times 100$

- Q) Mass % of carbon in ethanol is :
 (A) 52 (B) 13 (C) 34 (D) 90

01

NATURE OF MATTER

08

STOICHIOMETRIC CALCULATIONS

07

LIMITING REACTANT

07



- Q) The number of moles of hydrogen molecules required to produce 20 moles of ammonia through Haber's process is :
 (A) 40 (B) 10 (C) 20 (D) 30

06

EF & MF

06

ACTUAL FORMULA	SIMPLEST FORMULA
Molecular Formula	Empirical Formula
$\text{C}_3\text{H}_6\text{O}_3$	CH_2O
$\text{C}_{10}\text{H}_{14}\text{N}_2$	$\text{C}_5\text{H}_7\text{N}$
$\text{C}_{12}\text{H}_{22}\text{O}_{11}$	$\text{C}_{12}\text{H}_{22}\text{O}_{11}$

- Q) An organic compound contains 80% (by wt.) C & the remaining percentage of H . The empirical formula of this compound is :

- (A) CH_3 (B) CH_4 (C) CH (D) CH_2