

**TABLE A-13 SOLUBILITY OF COMPOUNDS**

Solubilities are given in grams of solute that can be dissolved in 100 g of water at the temperature (°C) indicated.

Compound	Formula	0°C	20°C	60°C	100°C
<b>aluminum sulfate</b>	Al <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub>	31.2	36.4	59.2	89.0
<b>ammonium chloride</b>	NH <sub>4</sub> Cl	29.4	37.2	55.3	77.3
<b>ammonium nitrate</b>	NH <sub>4</sub> NO <sub>3</sub>	118	192	421	871
<b>ammonium sulfate</b>	(NH <sub>4</sub> ) <sub>2</sub> SO <sub>4</sub>	70.6	75.4	88	103
<b>barium carbonate</b>	BaCO <sub>3</sub>	—*	0.0022 <sup>18°</sup>	—*	0.0065
<b>barium chloride dihydrate</b>	BaCl <sub>2</sub> • 2H <sub>2</sub> O	31.2	35.8	46.2	59.4
<b>barium hydroxide</b>	Ba(OH) <sub>2</sub>	1.67	3.89	20.94	101.40 <sup>80°</sup>
<b>barium nitrate</b>	Ba(NO <sub>3</sub> ) <sub>2</sub>	4.95	9.02	20.4	34.4
<b>barium sulfate</b>	BaSO <sub>4</sub>	—*	0.000 246 <sup>25°</sup>	—*	0.000 413
<b>calcium carbonate</b>	CaCO <sub>3</sub>	—*	0.0014 <sup>25°</sup>	—*	0.0018 <sup>75°</sup>
<b>calcium fluoride</b>	CaF <sub>2</sub>	0.0016 <sup>18°</sup>	0.0017 <sup>26°</sup>	—*	—*
<b>calcium hydrogen carbonate</b>	Ca(HCO <sub>3</sub> ) <sub>2</sub>	16.15	16.60	17.50	18.40
<b>calcium hydroxide</b>	Ca(OH) <sub>2</sub>	0.189	0.173	0.121	0.076
<b>calcium sulfate</b>	CaSO <sub>4</sub>	—*	0.209 <sup>30°</sup>	—*	0.1619
<b>copper(II) chloride</b>	CuCl <sub>2</sub>	68.6	73.0	96.5	120
<b>copper(II) sulfate pentahydrate</b>	CuSO <sub>4</sub> • 5H <sub>2</sub> O	23.1	32.0	61.8	114
<b>lead(II) chloride</b>	PbCl <sub>2</sub>	0.67	1.00	1.94	3.20
<b>lead(II) nitrate</b>	Pb(NO <sub>3</sub> ) <sub>2</sub>	37.5	54.3	91.6	133
<b>lithium chloride</b>	LiCl	69.2	83.5	98.4	128
<b>lithium sulfate</b>	Li <sub>2</sub> SO <sub>4</sub>	36.1	34.8	32.6	30.9 <sup>90°</sup>
<b>magnesium hydroxide</b>	Mg(OH) <sub>2</sub>	—*	0.0009 <sup>18°</sup>	—*	0.004
<b>magnesium sulfate</b>	MgSO <sub>4</sub>	22.0	33.7	54.6	68.3
<b>mercury(I) chloride</b>	Hg <sub>2</sub> Cl <sub>2</sub>	—*	0.000 20 <sup>25°</sup>	0.001 <sup>43°</sup>	—*
<b>mercury(II) chloride</b>	HgCl <sub>2</sub>	3.63	6.57	16.3	61.3
<b>potassium bromide</b>	KBr	53.6	65.3	85.5	104
<b>potassium chlorate</b>	KClO <sub>3</sub>	3.3	7.3	23.8	56.3
<b>potassium chloride</b>	KCl	28.0	34.2	45.8	56.3
<b>potassium chromate</b>	K <sub>2</sub> CrO <sub>4</sub>	56.3	63.7	70.1	74.5 <sup>90°</sup>
<b>potassium iodide</b>	KI	128	144	176	206
<b>potassium nitrate</b>	KNO <sub>3</sub>	13.9	31.6	106	245
<b>potassium permanganate</b>	KMnO <sub>4</sub>	2.83	6.34	22.1	—*
<b>potassium sulfate</b>	K <sub>2</sub> SO <sub>4</sub>	7.4	11.1	18.2	24.1
<b>silver acetate</b>	AgC <sub>2</sub> H <sub>3</sub> O <sub>2</sub>	0.73	1.05	1.93	2.59 <sup>80°</sup>
<b>silver chloride</b>	AgCl	0.000 089 <sup>10°</sup>	—*	—*	0.0021
<b>silver nitrate</b>	AgNO <sub>3</sub>	122	216	440	733
<b>sodium acetate</b>	NaC <sub>2</sub> H <sub>3</sub> O <sub>2</sub>	36.2	46.4	139	170
<b>sodium chlorate</b>	NaClO <sub>3</sub>	79.6	95.9	137	204
<b>sodium chloride</b>	NaCl	35.7	35.9	37.1	39.2
<b>sodium nitrate</b>	NaNO <sub>3</sub>	73.0	87.6	122	180
<b>sucrose</b>	C <sub>12</sub> H <sub>22</sub> O <sub>11</sub>	179.2	203.9	287.3	487.2

\*Dashes indicate that values are not available.