Ionization Energy

An electron can be removed from an atom if enough energy is supplied. Using A as a symbol for an atom of any element, the process can be expressed as follows.

A + energy
$$\rightarrow$$
 A⁺ + e^-

The A^+ represents an ion of element A with a single positive charge, referred to as a 1+ ion. An **ion** is an atom or group of bonded atoms that has a positive or negative charge. Sodium, for example, forms an Na⁺ ion. Any process that results in the formation of an ion is referred to as **ionization.**

To compare the ease with which atoms of different elements give up electrons, chemists compare ionization energies. The energy required to remove one electron from a neutral atom of an element is the ionization energy, IE (or first ionization energy, IE_1). To avoid the influence of nearby atoms, measurements of ionization energies are made on isolated atoms in the gas phase. Figure 15 gives the first ionization energies for the elements in kilojoules per mole (kJ/mol). Figure 16 presents this information graphically.

FIGURE 15 In general, first ionization energies increase across a period and decrease down a group.

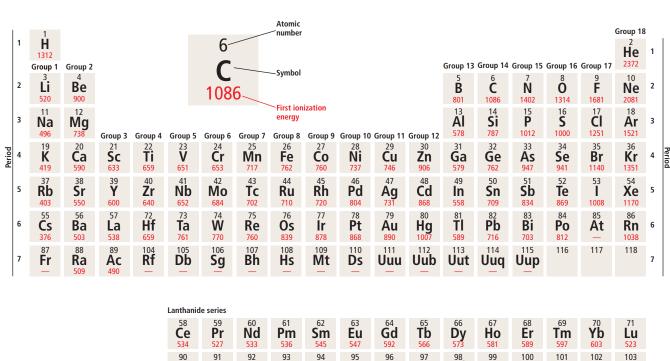
Periodic Table of Ionization Energies (kJ/mol)

Th

587 53 Actinide series

Pa

U



Pu

Am

Cm

Np

Bk

Cf

Es

Lr

Md

No

Fm