

## Direct Proportions

Two quantities are **directly proportional** to each other if dividing one by the other gives a constant value. For example, if the masses and volumes of different samples of aluminum are measured, the masses and volumes will be directly proportional to each other. As the masses of the samples increase, their volumes increase by the same factor, as you can see from the data in **Table 7**. Doubling the mass doubles the volume. Halving the mass halves the volume.

When two variables,  $x$  and  $y$ , are directly proportional to each other, the relationship can be expressed as  $y \propto x$ , which is read as “ $y$  is *proportional* to  $x$ .” The general equation for a directly proportional relationship between the two variables can also be written as follows.

$$\frac{y}{x} = k$$

The value of  $k$  is a constant called the proportionality constant. Written in this form, the equation expresses an important fact about direct proportion: the ratio between the variables remains constant. Note that using the mass and volume values in **Table 7** gives a mass-volume ratio that is constant (neglecting measurement error). The equation can be rearranged into the following form.

$$y = kx$$

The equation  $y = kx$  may look familiar to you. It is the equation for a special case of a straight line. If two variables related in this way are graphed versus one another, a straight line, or linear plot that passes through the origin (0,0), results. The data for aluminum from **Table 7** are graphed in **Figure 11**. The mass and volume of a pure substance are directly proportional to each other. Consider mass to be  $y$  and volume to be  $x$ . The constant ratio,  $k$ , for the two variables is density. The slope of the line reflects the constant density, or mass-volume ratio, of aluminum,

**FIGURE 11** The graph of mass versus volume shows a relationship of direct proportion. Notice that the line is extrapolated to pass through the origin.

**TABLE 7** Mass-Volume Data for Aluminum at 20°C

Mass (g)	Volume (cm <sup>3</sup> )	$\frac{m}{V}$ (g/cm <sup>3</sup> )
54.7	20.1	2.72
65.7	24.4	2.69
83.5	30.9	2.70
96.3	35.8	2.69
105.7	39.1	2.70

**Mass Vs. Volume of Aluminum**

