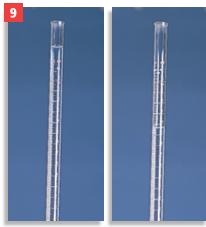
Fill the other buret with the standard base solution to a point above the calibration mark. The concentration of the standard base is known to a certain degree of precision because the base was previously titrated with an exact mass of solid acid, which is the primary standard.





Release some base from the buret to remove any air bubbles and to lower the volume to the calibrated portion of the buret.



Record the volume of the base to the nearest 0.01 mL as your initial volume. Remember to read the volume at the bottom of the meniscus.



Place the Erlenmeyer flask under the base buret as shown. Notice that the tip of the buret extends into the mouth of the flask.



Slowly release base from the buret into the flask while constantly swirling the contents of the flask. The pink color of the indicator should fade with swirling.



The titration is nearing the end point when the pink color stays for longer periods of time. At this point, add base drop by drop.



The equivalence point is reached when a very light pink color remains after 30 seconds of swirling.



Subtract the initial volume reading on the buret from the final reading. This is the exact volume of the base released into the flask. Record it to the nearest 0.01 mL.