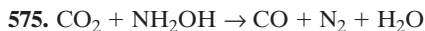


Reactions in Basic Solution

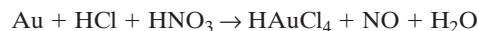
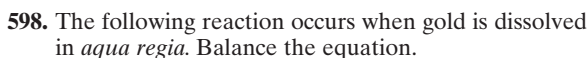
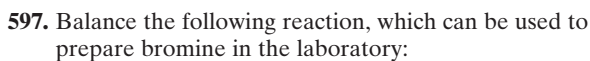
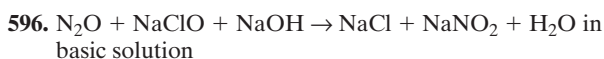
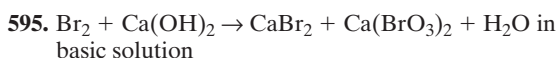
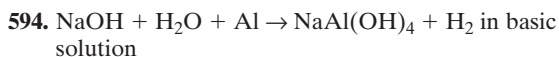
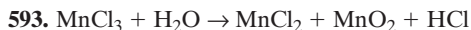
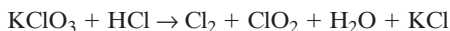
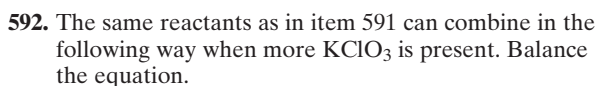
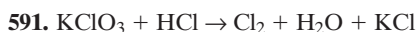
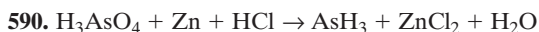
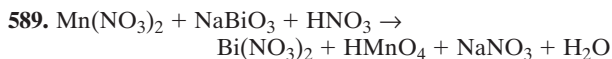
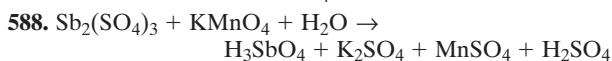
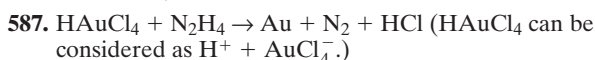
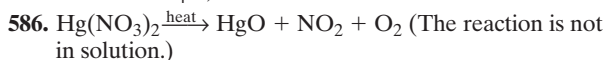
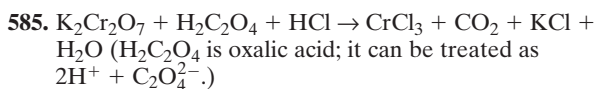
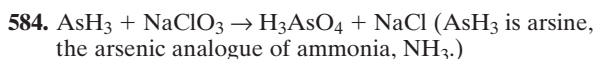
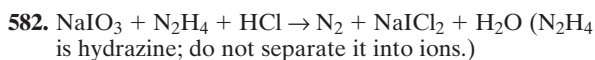
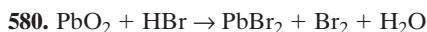
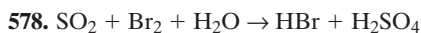
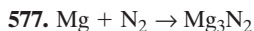
Balance the following redox equations. Assume that all reactions take place in a basic environment where OH^- and H_2O are readily available.



(Both of the potassium-tin-oxygen compounds dissociate into potassium ions and tin-oxygen ions.)

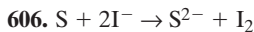
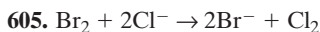
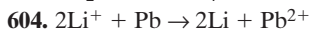
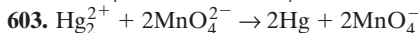
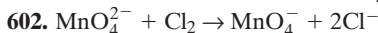
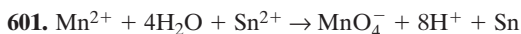
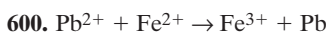
Mixed Review

Balance each of the following redox equations. Unless stated otherwise, assume that the reaction occurs in acidic solution.

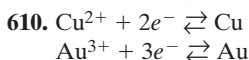
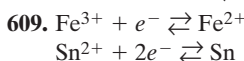
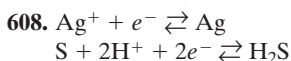
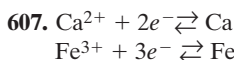


Electrochemistry: Chap. 20, Sec. 2

Use the reduction potentials in the table on page 915 to determine whether the following reactions are spontaneous as written. Report the E_{cell}^0 for the reactions.



If a cell is constructed in which the following pairs of reactions are possible, what would be the cathode reaction, the anode reaction, and the overall cell voltage?



Mixed Review

Use reduction potentials to determine whether the reactions in the following 10 problems are spontaneous.

