

**TABLE A-7 COMMON IONS**

Cation	Symbol	Anion	Symbol
<b>aluminum</b>	$\text{Al}^{3+}$	<b>acetate</b>	$\text{CH}_3\text{COO}^-$
<b>ammonium</b>	$\text{NH}_4^+$	<b>bromide</b>	$\text{Br}^-$
<b>arsenic(III)</b>	$\text{As}^{3+}$	<b>carbonate</b>	$\text{CO}_3^{2-}$
<b>barium</b>	$\text{Ba}^{2+}$	<b>chlorate</b>	$\text{ClO}_3^-$
<b>calcium</b>	$\text{Ca}^{2+}$	<b>chloride</b>	$\text{Cl}^-$
<b>chromium(II)</b>	$\text{Cr}^{2+}$	<b>chlorite</b>	$\text{ClO}_2^-$
<b>chromium(III)</b>	$\text{Cr}^{3+}$	<b>chromate</b>	$\text{CrO}_4^{2-}$
<b>cobalt(II)</b>	$\text{Co}^{2+}$	<b>cyanide</b>	$\text{CN}^-$
<b>cobalt(III)</b>	$\text{Co}^{3+}$	<b>dichromate</b>	$\text{Cr}_2\text{O}_7^{2-}$
<b>copper(I)</b>	$\text{Cu}^+$	<b>fluoride</b>	$\text{F}^-$
<b>copper(II)</b>	$\text{Cu}^{2+}$	<b>hexacyanoferrate(II)</b>	$\text{Fe}(\text{CN})_6^{4-}$
<b>hydronium</b>	$\text{H}_3\text{O}^+$	<b>hexacyanoferrate(III)</b>	$\text{Fe}(\text{CN})_6^{3-}$
<b>iron(II)</b>	$\text{Fe}^{2+}$	<b>hydride</b>	$\text{H}^-$
<b>iron(III)</b>	$\text{Fe}^{3+}$	<b>hydrogen carbonate</b>	$\text{HCO}_3^-$
<b>lead(II)</b>	$\text{Pb}^{2+}$	<b>hydrogen sulfate</b>	$\text{HSO}_4^-$
<b>magnesium</b>	$\text{Mg}^{2+}$	<b>hydroxide</b>	$\text{OH}^-$
<b>mercury(I)</b>	$\text{Hg}_2^{2+}$	<b>hypochlorite</b>	$\text{ClO}^-$
<b>mercury(II)</b>	$\text{Hg}^{2+}$	<b>iodide</b>	$\text{I}^-$
<b>nickel(II)</b>	$\text{Ni}^{2+}$	<b>nitrate</b>	$\text{NO}_3^-$
<b>potassium</b>	$\text{K}^+$	<b>nitrite</b>	$\text{NO}_2^-$
<b>silver</b>	$\text{Ag}^+$	<b>oxide</b>	$\text{O}^{2-}$
<b>sodium</b>	$\text{Na}^+$	<b>perchlorate</b>	$\text{ClO}_4^-$
<b>strontium</b>	$\text{Sr}^{2+}$	<b>permanganate</b>	$\text{MnO}_4^-$
<b>tin(II)</b>	$\text{Sn}^{2+}$	<b>peroxide</b>	$\text{O}_2^{2-}$
<b>tin(IV)</b>	$\text{Sn}^{4+}$	<b>phosphate</b>	$\text{PO}_4^{3-}$
<b>titanium(III)</b>	$\text{Ti}^{3+}$	<b>sulfate</b>	$\text{SO}_4^{2-}$
<b>titanium(IV)</b>	$\text{Ti}^{4+}$	<b>sulfide</b>	$\text{S}^{2-}$
<b>zinc</b>	$\text{Zn}^{2+}$	<b>sulfite</b>	$\text{SO}_3^{2-}$