complicated. First, the ending of the element's name is dropped. Then the ending *-ide* is added to the root name, as illustrated by the examples at right.

The names and symbols of the common monatomic cations and anions are organized according to their charges in **Table 1.** The names of many of the ions in the table include Roman numerals. These numerals are part of the *Stock system* of naming chemical ions and elements. You will read more about the Stock system and other systems of naming chemicals later in this chapter.

Examples of Anions

Element	Anion
F	F^-
Fluor*ine*	Fluor*ide* anion
N	N^{3-}
Nitr*ogen*	Nitr*ide* anion

TABLE 1 *Some Common Monatomic Ions*

Main-group elements

1+		2+		3+		
lithium	Li^+	beryllium	Be^{2+}	aluminum	Al^{3+}	
sodium	Na^+	magnesium	Mg^{2+}			
potassium	K^+	calcium	Ca^{2+}			
rubidium	Rb^+	strontium	Sr^{2+}			
cesium	Cs^+	barium	Ba^{2+}			
1–		**2–**		**3–**		
fluoride	F^-	oxide	O^{2-}	nitride	N^{3-}	
chloride	Cl^-	sulfide	S^{2-}	phosphide	P^{3-}	
bromide	Br^-					
iodide	I^-					

***d*-Block elements and others with multiple ions**

1+		2+		3+		4+	
copper(I)	Cu^+	vanadium(II)	V^{2+}	vanadium(III)	V^{3+}	vanadium(IV)	V^{4+}
silver	Ag^+	chromium(II)	Cr^{2+}	chromium(III)	Cr^{3+}	tin(IV)	Sn^{4+}
		manganese(II)	Mn^{2+}	iron(III)	Fe^{3+}	lead(IV)	Pb^{4+}
		iron(II)	Fe^{2+}	cobalt(III)	Co^{3+}		
		cobalt(II)	Co^{2+}				
		nickel(II)	Ni^{2+}				
		copper(II)	Cu^{2+}				
		zinc	Zn^{2+}				
		cadmium	Cd^{2+}				
		tin(II)	Sn^{2+}				
		mercury(II)	Hg^{2+}				
		lead(II)	Pb^{2+}				