

a.

Refer back to homework 2, the yield is

$$Y_0 = 0.07 \text{ g cells/g BrO}_3^-$$

The information given from the question

$$\hat{q} = 12 \frac{\text{mgBrO}_3^-}{\text{mgcells}\cdot d} \quad K = 50 \frac{\mu\text{gBrO}_3^-}{L}$$

$$S^o = 200 \text{ mg/L} \quad b = 0.01 \text{ d}^{-1}$$

Then

$$S^{min} = K \times \frac{b}{Y\hat{q} - b} = \frac{50 \mu\text{g}}{L} \times \frac{0.01 \text{ d}^{-1}}{0.07 \frac{\text{mg cells}}{\text{mg BrO}_3^-} \times 12 \frac{\text{mgBrO}_3^-}{\text{mgcells}\cdot d} - 0.01 \text{ d}^{-1}} = 0.60 \mu\text{g/L}$$

$$[\theta^{min}]_{lim} = \frac{1}{Y\hat{q} - b} = \frac{1}{0.07 \frac{\text{mg cells}}{\text{mg BrO}_3^-} \times 12 \frac{\text{mgBrO}_3^-}{\text{mgcells}\cdot d} - 0.01 \text{ d}^{-1}} = 1.20 \text{ d}$$

b.

In my opinion, I think the volume value is missing for this part, because S value depends on the

$$\theta_x = \frac{V}{Q} \text{ directly due to the equation } S = K \frac{1 + b\theta_x}{\theta_x(Y\hat{q} - b) - 1}$$

Or according to the a part, possibly

the $S_{min} = 0.60 \mu\text{g/L} < 10 \mu\text{g/L}$ if the θ_x is much larger

Then this design can achieve the goal concentration of bromide.

c.

If a chemostat can achieve the target treatment goal

$$S = 10 \text{ } \mu\text{g/L}$$

According the relative equation

$$\frac{1}{\theta_x} = Y \frac{\hat{q}S}{K + S} - b = 0.07 \frac{\text{mg cells}}{\text{mg BrO}_3^-} \times \frac{12 \frac{\text{mgBrO}_3^-}{\text{mgcells}\cdot\text{d}} \times 10 \frac{\mu\text{g}}{\text{L}}}{50 \frac{\mu\text{g}}{\text{L}} + 10 \frac{\mu\text{g}}{\text{L}}} - 0.01 \text{d}^{-1} = 0.13 \text{d}^{-1}$$

$$\theta_x = 7.7 \text{ d}$$

If a chemostats cannot achieve the target treatment goal and its safety factor is 10

$$\theta_x^{min} = \frac{K + S^0}{S^0(Y\hat{q} - b) - bK} = 1.20$$

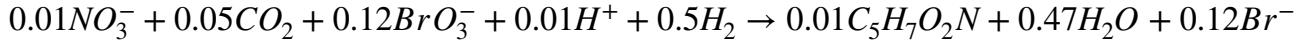
$$SF = \frac{\theta_x}{\theta_x^{min}} = 10 \text{ then } \theta_x = 12 \text{ d}$$

According the relative equation

$$S = K \frac{1 + b\theta_x}{\theta_x(Y\hat{q} - b) - 1} = \frac{50\mu\text{g}}{L} \times \frac{1 + 0.01\text{d}^{-1} \times 12\text{d}}{12\text{d} \times (0.07 \frac{\text{mg cells}}{\text{mg BrO}_3^-} \times 12 \frac{\text{mgBrO}_3^-}{\text{mgcells}\cdot\text{d}} - 0.01 \text{d}^{-1}) - 1} = 6.2\mu\text{g/L}$$

d.

According to the reaction



Then

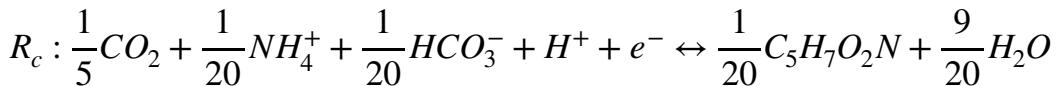
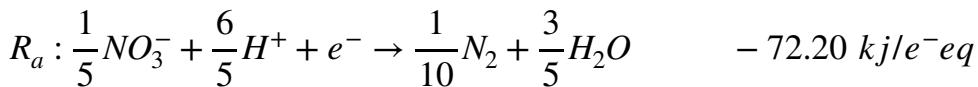
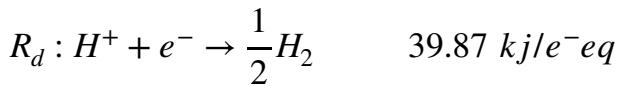
$$Y_{H_2}^0 = 0.5(2)/0.12(129) = 0.0064 \text{ g } H_2/\text{g } BrO_3^-$$

Then if a chemostat can achieve the target treatment goal

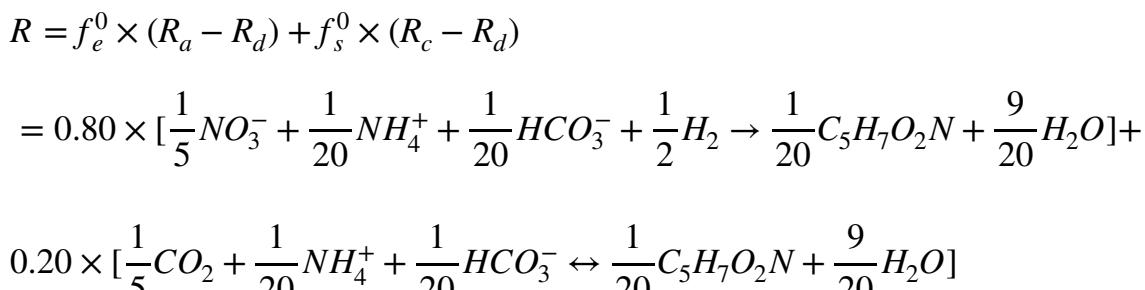
$$C_{min}^{H_2} = (S^0 - S) \times Y_{H_2} = 12.92 \text{ mg/L}$$

e.

When Nitrite as electron acceptor



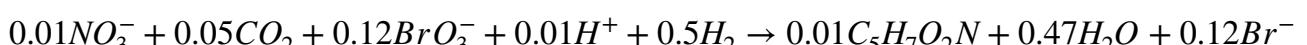
According to the Table 3.1 $f_s^0 = 0.2$



For denitrification process



For Bromide reduction process



$$C_{extra}^{H_2} = C_{min}^{H_2} \times \frac{0.05(113)}{0.4(2)} \times \frac{2(0.5)}{0.01(113)} = 80.75 \text{ mg/L}$$

f.

Refer back to the c part, I choose SF = 10 design, where $\theta_x = 12d$ and $S = 6.2\mu g/L$

$$X_a = Y(S^0 - S) \frac{1}{1 + b\theta_x} = 0.07 \frac{\text{mg cells}}{\text{mg } BrO_3^-} \times (200 - 0.0064) \text{ mg/L} \times \frac{1}{1 + 0.01d^{-1} \times 15.1d} = 12.50 \text{ mg/L}$$

$$X_i = X_i^0 + X_a(1 - f_d)b\theta_x = 0 + X_a \times (1 - 0.8) \times 0.01d^{-1} \times 15.1d = 0.30 \text{ mg/L}$$

$$X_v = X_i + X_a = 12.80 \text{ mg/L}$$

$$r_{ut} = - \frac{S^0 - S}{\theta_x} = - \frac{(200 - 0.0064) \text{ mg/L}}{15.1d} = - 16.66 \text{ mg BOD}_L/\text{L}\cdot\text{d}$$

For SMP calculation,

Where

$$k_1 = 0.12 \text{g COD}_p/\text{g COD}_s \quad \hat{q}_{UAP} = 1.8 \text{g COD}_p/\text{g VSS}_a$$

$$k_2 = 0.09 \text{g COD}_p/\text{g COD}_s \quad K_{UAP} = 100 \text{mg COD}_p/\text{g VSS}_a$$

$$UAP = \frac{-(\hat{q}_{UAP}X_a\theta + K_{UAP} + k_1r_{ut}\theta) + \sqrt{(\hat{q}_{UAP}X_a\theta + K_{UAP} + k_1r_{ut}\theta)^2 - 4K_{UAP}k_1r_{ut}\theta}}{2} = 6.8 \text{ mg/L}$$

$$BAP = \frac{-(K_{BAP} + (\hat{q}_{BAP} - k_2)X_a\theta) + \sqrt{(K_{BAP} + (\hat{q}_{BAP} - k_2)X_a\theta)^2 + 4K_{BAP}k_2X_a\theta}}{2} = 11.69 \text{ mg/L}$$

$$\text{Total COD} = S + SMP + (1.42 \text{ gCOD/g VSS}) \cdot X_v = 79.51 \text{mg/L}$$

$$\text{Total BOD}_L = \text{Total COD} \times 80\% = 63.60 \text{ mg/L}$$

$$BOD_5 = BOD_L \times (1 - e^{-0.23d^{-1} \times 5d}) = 43.46 \text{ mg/L}$$

g.

The N source may be insufficient with time passing by, causing the lower water treatment

efficiency. On the other hand, Br^- is likely to oxide in the real world conditions, thus, oxidation equilibrium should be considered.