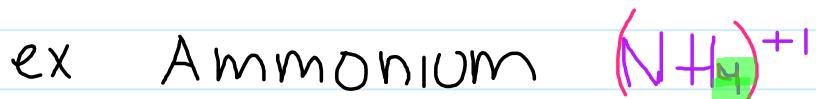


Polyatomic Compounds

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Polyatomic ions form compounds like other ions



Naming

- 1 positive polyatomic ions are written first, like metals
- 2 Negative polyatomic ions are written second, and the name is **NOT** changed

Ex. #1

- (a) Calcium + nitrate \rightarrow calcium nitrate
- (b) Hydrogen + dichromate \rightarrow hydrogen dichromate
- (c) $\text{K}_2\text{SO}_4 \rightarrow$ Potassium sulfate
- (d) $\text{KMnO}_4 \rightarrow$ Potassium permanganate

Writing formulas

- 1 Write the metal and non-metal elements in their ion form (polyatomic is always written in brackets)
- 2 Re-write the elements without ion charges, criss-cross the numbers
 - ↳ common factor \rightarrow REDUCE!!
 - ↳ if there is only 1 of the polyatomic group, omit the brackets

Ex. #2

