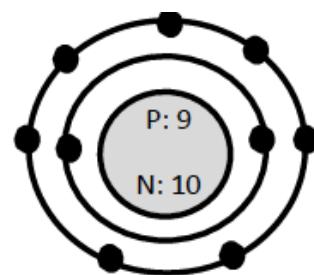
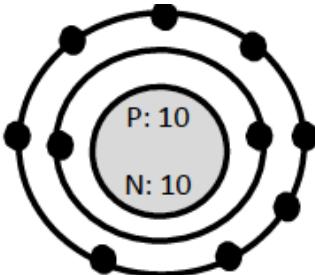
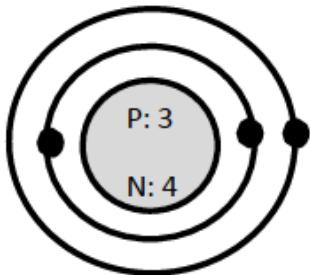


Name: _____ Date: _____ Period: _____

Periodic Table Review

Directions: Complete the chart using a periodic table.

M- metals NM- Nonmetals MT- metalloids



Element		
M, NM, MT		
Reactivity		

Directions: fill in the blanks to the statements below.

Vertical columns on the periodic table are called _____.

Horizontal rows on the periodic table are called _____.

The number of protons in an atom is that element's _____ number.

The total of the protons and neutrons is that atom's _____ number.

The elements in group _____ are the most reactive metals.

The elements in group _____ are the most reactive nonmetals.

The elements in group _____ are very unreactive.

Directions: Complete the chart using a periodic table.

	Element Symbol	M, NM, or MT	Valence Electrons	Reactivity (high, medium, non)
group 14, period 2				
group 2, period 2				
group 1, period 1				
group 18, period 3				
group 13, period 5				
group 17, period 4				
group 1, period 3				

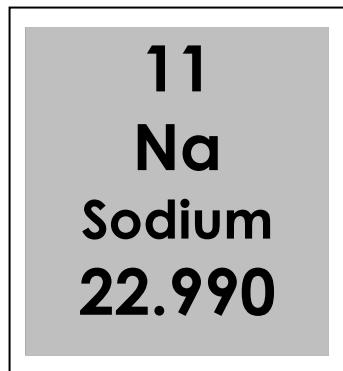
Directions: For each of the following, label as a metal (M), nonmetal (NM), and/or metalloid (MT)

Poor conductor of electricity		Silicon	
Usually a solid at room temp		Most are a gas at room temp.	
Ductile		Cobalt	
Chlorine		Good Conductor	
Semi-conductor		Brittle	
Malleable		Oxygen	

List the 3 subatomic particles in an atom.

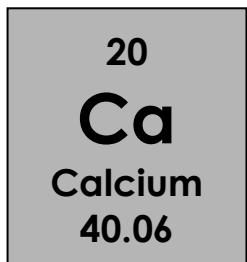
Atomic Number is equal to...	Atomic Mass is equal to...

Label the parts in the periodic square of sodium.

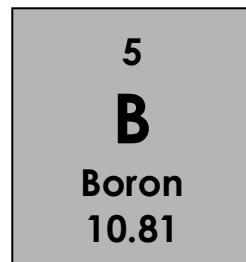


A		M	
P		A	
E		N	

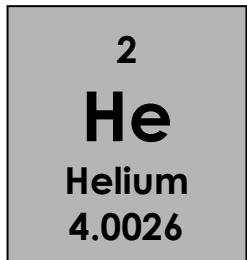
Directions: Complete APEMAN for the elements below



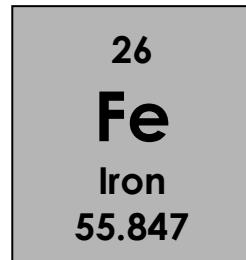
$$\begin{array}{ll} A= & M= \\ P= & A= \\ E= & N= \end{array}$$



$$\begin{array}{ll} A= & M= \\ P= & A= \\ E= & N= \end{array}$$



$$\begin{array}{ll} A= & M= \\ P= & A= \\ E= & N= \end{array}$$



$$\begin{array}{ll} A= & M= \\ P= & A= \\ E= & N= \end{array}$$

Nucleus of Atom	Element	Electrons	Atomic Mass

Directions: label the trends of the periodic table by adding arrows and descriptions.

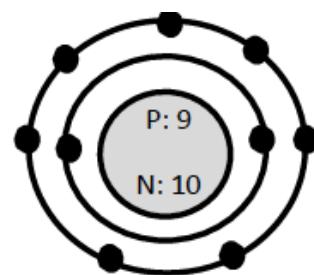
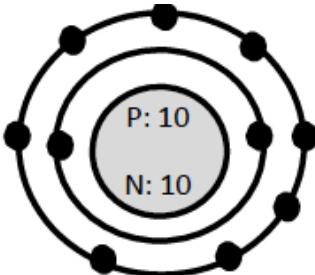
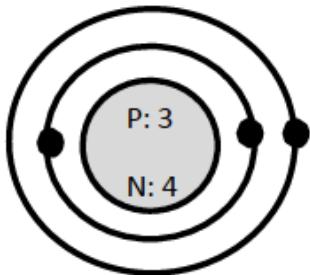
1 1A	1 H 1.008 Hydrogen	2 2A	14 Si 28.086 Silicon	13 B 10.812 Boron	14 C 12.011 Carbon	15 N 14.007 Nitrogen	16 O 15.999 Oxygen	17 F 18.998 Fluorine	18 Ne 20.193 Helium
3 Li 6.941 Lithium	4 Be 9.012 Beryllium	5 Mg 24.306 Magnesium	6 6B	7 7B	8 8B	9 9B	10 10B	11 11B	12 12B
11 Na 22.990 Sodium	12 Mg 24.306 Magnesium	3 3B	4 4B	5 5B	6 6B	7 7B	8 8B	9 9B	10 10B
19 K 39.098 Potassium	20 Ca 40.079 Calcium	21 Sc 44.966 Scandium	22 Ti 47.937 Titanium	23 V 50.942 Vanadium	24 Cr 51.996 Chromium	25 Mn 54.939 Manganese	26 Fe 55.845 Iron	27 Co 58.933 Cobalt	28 Ni 58.933 Nickel
37 Rb 85.469 Rubidium	38 Sr 87.62 Strontium	39 Y 88.906 Yttrium	40 Zr 91.224 Zirconium	41 Nb 92.908 Niobium	42 Mo 95.96 Molybdenum	43 Tc 95.96 Technetium	44 Ru 101.07 Ruthenium	45 Rh 102.908 Rhodium	46 Pd 104.62 Palladium
55 Cs 132.905 Cesium	56 Ba 137.329 Barium	57 Lu 174.96 Lutetium	58 Hf 178.49 Hafnium	59 Ta 180.94 Tantalum	60 W 183.904 Tungsten	61 Re 186.207 Rhenium	62 Os 190.23 Osmium	63 Ir 192.217 Iridium	64 Pt 195.095 Platinum
87 Fr (223) Francium	88 Ra (226) Radium	71 Lr (228) Lawrencium	72 Rf (227) Rutherfordium	73 Hf (228) Hafnium	74 Ta (229) Tantalum	75 W (230) Tungsten	76 Re (231) Rhenium	77 Os (232) Osmium	78 Ir (233) Iridium
103 Fr (229) Francium	104 Ra (228) Radium	105 Lr (228) Lawrencium	106 Rf (227) Rutherfordium	107 Db (229) Darmstadtium	108 Sg (229) Seaborgium	109 Bh (229) Bohrium	110 Hs (229) Hassium	111 Mt (229) Meitnerium	112 Ds (229) Dimitrievium
Mass numbers in parentheses are those of the most stable or most common isotope.									
Updated Spring 2011									
Lanthanide Series									
57 La 126.906 Lanthanum	58 Ce 140.116 Cerium	59 Pr 140.908 Praseodymium	60 Nd 144.24 Neodymium	61 Pm 145.0 Promethium	62 Sm 150.904 Samarium	63 Eu 151.904 Europium	64 Gd 157.25 Gadolinium	65 Tb 160.905 Terbium	66 Dy 162.500 Dysprosium
69 Ac (227) Actinium	70 Th 226.026 Thorium	91 Pa 221.026 Protactinium	92 U 226.031 Uranium	93 Np (237) Neptunium	94 Pu (244) Plutonium	95 Am (243) Americium	96 Cm (247) Curium	97 Bk (247) Berkelium	98 Cf (251) Californium
99 Es (250) Einsteinium	100 Fm (257) Fermium	101 Md (258) Mendelevium	102 No (259) Nobelium						

Name: _____ Date: _____ Period: _____

Periodic Table Review Answer Key

Directions: Complete the chart using a periodic table.

M- metals NM- Nonmetals MT- metalloids



Element	lithium	neon	Fluorine
M, NM, MT	M	NM	NM
Reactivity	Highly Reactive	Non-Reactive	Highly Reactive

Directions: fill in the blanks to the statements below.

Vertical columns on the periodic table are called groups.

Horizontal rows on the periodic table are called periods.

The number of protons in an atom is that element's atomic number.

The total of the protons and neutrons is that atom's mass number.

The elements in group 1 are the most reactive metals.

The elements in group 17 are the most reactive nonmetals.

The elements in group 18 are very unreactive.

Directions: Complete the chart using a periodic table.

	Element Symbol	M, NM, or MT	Valence Electrons	Reactivity (high, medium, non)
group 14, period 2	Si	MT	4	medium
group 2, period 2	Be	M	2	medium
group 1, period 1	H	NM	1	high
group 18, period 3	Ar	NM	8	non
group 13, period 5	In	M	3	medium
group 17, period 4	Br	NM	7	high
group 1, period 3	Na	M	1	high

Directions: For each of the following, label as a metal (M), nonmetal (NM), and/or metalloid (MT)

Poor conductor of electricity	NM	Silicon	MT
Usually a solid at room temp	M	Most are a gas at room temp.	NM
Ductile	M	Cobalt	M
Chlorine	NM	Good Conductor	M
Semi-conductor	MT	Brittle	NM, MT
Malleable	M	Oxygen	NM

List the 3 subatomic particles in an atom.

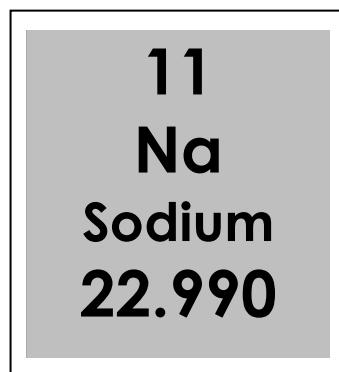
Protons

Neutrons

Electrons

Atomic Number is equal to...	Atomic Mass is equal to...
Protons	(nucleus of the atom) Protons + Neutrons

Label the parts in the periodic square of sodium.



Atomic Number

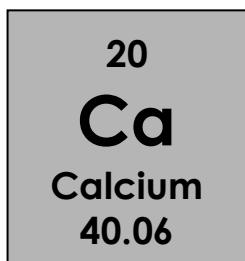
Symbol

Element Name

Mass Number

A	Atomic Number	M	Mass Number
P	Protons	A	Atomic Number (Subtract)
E	Electrons	N	Neutrons

Directions: Complete APEMAN for the elements below



$$A = 20$$

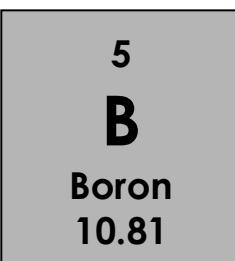
$$P = 20$$

$$E = 20$$

$$M = 40$$

$$A = 20$$

$$N = 20$$



$$A = 5$$

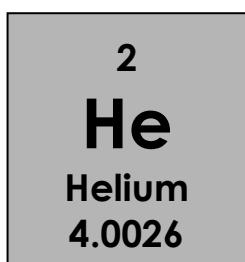
$$P = 5$$

$$E = 5$$

$$M = 11$$

$$A = 5$$

$$N = 6$$



$$A = 2$$

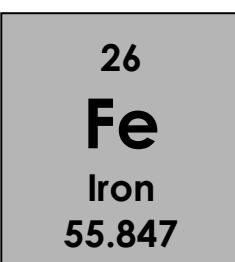
$$P = 2$$

$$E = 2$$

$$M = 4$$

$$A = 2$$

$$N = 2$$



$$A = 26$$

$$P = 26$$

$$E = 26$$

$$M = 56$$

$$A = 26$$

$$N = 30$$

Nucleus of Atom	Element	Electrons	Atomic Mass
	Li	Total= 3 Valence= 1	P= 3 N= 4 Mass= 7
	C	Total= 6 Valence= 4	P= 6 N= 6 Mass= 12
	H	Total= 1 Valence= 1	P= 1 N= 0 Mass= 1
	Be	Total= 4 Valence= 2	P= 4 N= 5 Mass= 9

Directions: label the trends of the periodic table by adding arrows and descriptions.

Metallic properties DECREASE

