

THE KENYA NATIONAL EXAMINATIONS COUNCIL
Kenya Certificate of Secondary Education

233/2

— CHEMISTRY —

Paper 2



(THEORY)

Nov. 2019 – 2 hours



Name Index Number

Candidate's Signature Date

Instructions to candidates

- (a) Write your name and index number in the spaces provided above.
- (b) Sign and write the date of examination in the spaces provided above.
- (c) Answer all the questions in the spaces provided.
- (d) KNEC mathematical tables and silent non-programmable electronic calculators may be used.
- (e) All working must be clearly shown where necessary.
- (f) This paper consists of 15 printed pages.
- (g) Candidates should check the question paper to ascertain that all the pages are printed as indicated and that no questions are missing.
- (h) Candidates should answer the questions in English.

For Examiner's Use Only

| Question | Maximum Score | Candidate's Score |
|-------------|---------------|-------------------|
| 1 | 12 | |
| 2 | 10 | |
| 3 | 12 | |
| 4 | 12 | |
| 5 | 12 | |
| 6 | 11 | |
| 7 | 11 | |
| Total Score | 80 | |



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1. (a) Alkanes are said to be saturated hydrocarbons.

- (i) What is meant by saturated hydrocarbons.

(1 mark)

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- (ii) Draw the structure of the third member of the alkane homologous series and name it. (2 marks)

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- (b) When the alkane, hexane, is heated to high temperature, one of the products is ethene.

- (i) Write the equation for the reaction.

(1 mark)

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- (ii) Name the process described in (b).

(1 mark)

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- (c) Study the flow chart in **Figure 1** and answer the questions that follow.

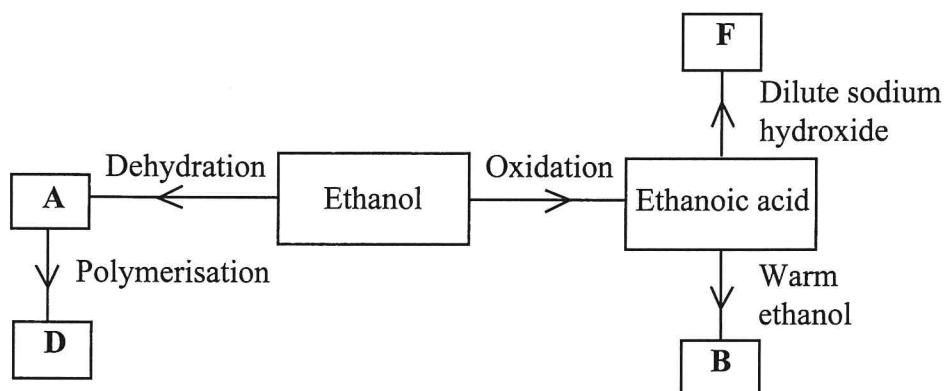


Figure 1

(i) Identify **A**. (1 mark)

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(ii) State **one** physical property of **B**. (1 mark)

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(iii) Draw the structure of **D**. (1 mark)

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(iv) Give a reason why **D** pollutes the environment. (1 mark)

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(v) Write an equation for the formation of **F**. (1 mark)

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(d) Describe an experiment which can be used to distinguish butene from butanol. (2 marks)

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2. (a) Zinc occurs mainly as zinc blende. Name **one** other ore from which zinc can be extracted. (1 mark)
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- (b) The flow chart in **Figure 2** shows the various stages in the extraction of zinc metal. Study it and answer the questions that follow.

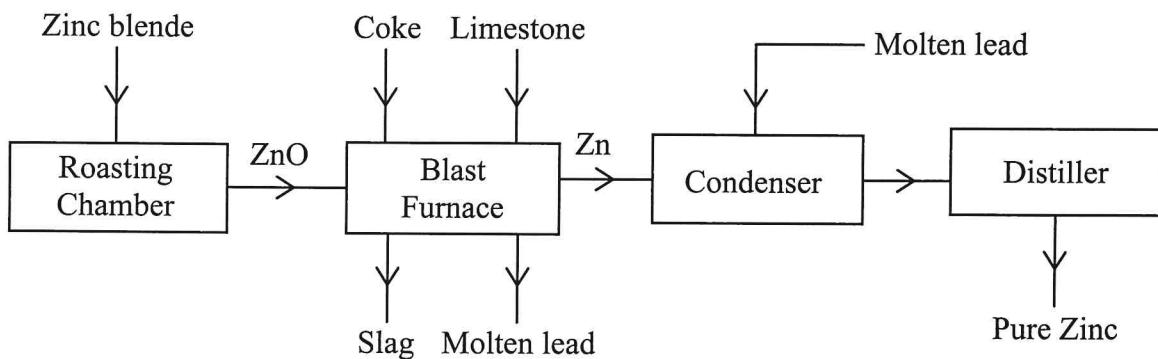


Figure 2

- (i) Write an equation for the reaction which occurs in the roasting chamber. (1 mark)
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- (ii) Describe the process that takes place in the blast furnace. (3 marks)
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- (iii) Explain why molten lead is added to the condenser. (1 mark)
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(iv) State **two** uses of zinc.

(1 mark)

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(v) Give **one** reason why the extraction of zinc causes pollution to the environment.
(1 mark)

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(c) Explain the observations made when zinc metal is added to hot sodium hydroxide.

(2 marks)

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3. Figure 3 is a flow chart that shows the process that occurs in the manufacture of nitric(V) acid.

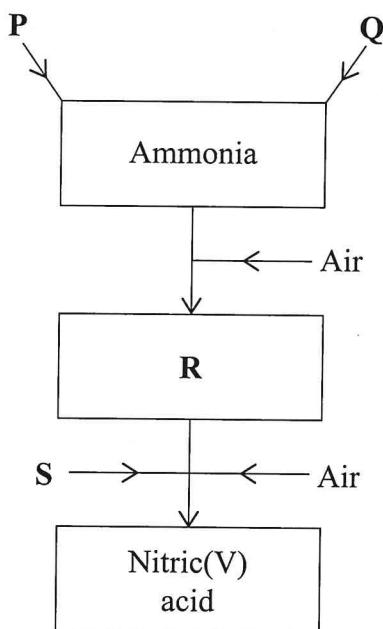


Figure 3

- (a) Name substance **P**, **Q**, **R** and **S**.

P (1 mark)

Q (1 mark)

R (1 mark)

S (1 mark)

- (b) To obtain substance **R**, ammonia is heated at 900 °C in the presence of air and a catalyst. The product is then cooled in air.

- (i) Name the catalyst for the reaction. (1 mark)

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- (ii) Write the equations for the two reactions described in (b). (2 marks)

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(iii) Other than nitric(V) acid, name another product that is formed. (1 mark)

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(c) When ammonia is reacted with nitric(V) acid, it produces a nitrogenous fertiliser.

(i) Explain why fertilisers play a major role in food production. (2 marks)

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(ii) State **two** problems associated with the use of nitrogenous fertilisers. (2 marks)

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4. (a) Explain the following observations:

- (i) The colour of aqueous copper(II) sulphate fades when a piece of magnesium metal is dropped into the solution. (2 marks)

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- (ii) A piece of iron bar is coated with a brown substance when left in the open on a rainy day. (2 marks)

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- (b) A sample of water is suspected to contain aluminium ions (Al^{3+}) . Describe a laboratory experiment that can be carried out to show that Al^{3+} ions are present in the water sample. (3 marks)

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(c) In an experiment to determine the number of moles of water of crystallisation of a hydrated compound, $\text{Na}_2\text{SO}_4 \cdot X \text{H}_2\text{O}$, 5 g of the compound were heated strongly to a constant mass.

- (i) Explain how a constant mass was obtained. (2 marks)

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- (ii) During the experiment, the mass of the residue was found to be 2.205 g.
Determine the number of moles of water of crystallisation in the compound.
(Na = 23.0 ; O = 16.0 ; S = 32.0 ; H = 1.0) (3 marks)

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5. (a) What is meant by molar heat of neutralisation? (1 mark)

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- (b) In an experiment to determine the molar heat of neutralisation, 50 cm^3 of 1M hydrochloric acid was neutralised by adding 10 cm^3 portions of dilute sodium hydroxide. During the experiment, the data in **Table 1** was obtained.

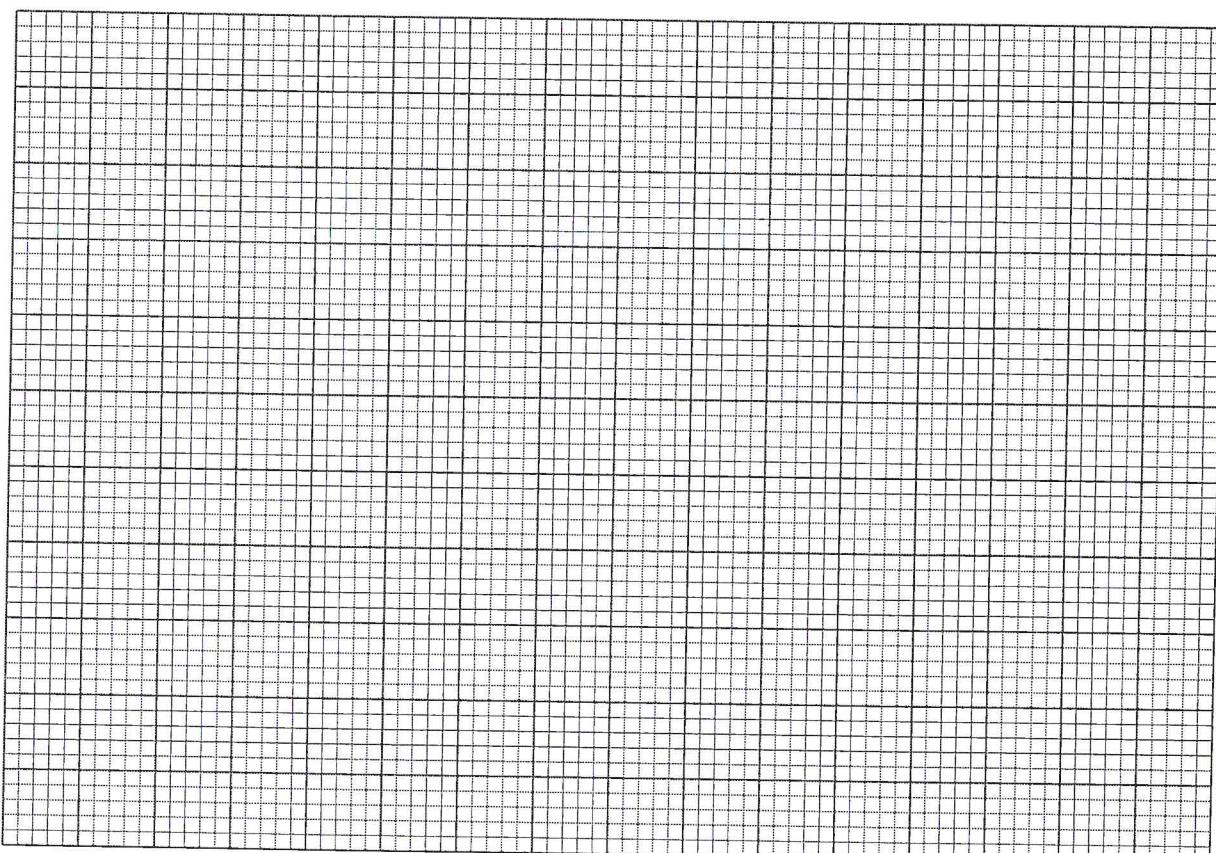
Table 1

| | | | | | | | |
|--|------|------|------|------|------|------|------|
| Volume of Sodium hydroxide (cm^3) | 0 | 10 | 20 | 30 | 40 | 50 | 60 |
| Temperature of mixture ($^\circ\text{C}$) | 25.0 | 27.0 | 29.0 | 31.0 | 31.0 | 30.0 | 29.0 |

- (i) Write the equation for the reaction in this experiment. (1 mark)

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- (ii) On the grid provided, plot a graph of temperature (Y-axis) against volume of sodium hydroxide (X-axis) added. (3 marks)



(iii) Determine from the graph the:

- I. volume of sodium hydroxide which completely neutralises 50 cm^3 of 1M hydrochloric acid. (1 mark)

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- II. change in temperature, ΔT , when complete neutralisation occurred.

(1 mark)

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(iv) Calculate:

- I. the heat change, ΔH when complete neutralisation occurred.

(Specific heat capacity = $4.2 \text{ J g}^{-1} \text{ K}^{-1}$, density of solution 1.0 g cm^{-3})

(2 marks)

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- II. molar heat of neutralisation of hydrochloric acid with sodium hydroxide.

(1 mark)

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- (v) How would the value of molar heat differ if 50 cm^3 of 1M ethanoic acid was used instead of 1M hydrochloric acid? Give a reason. (2 marks)

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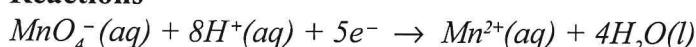
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6. (a) What is meant by standard electrode potential of an element? (1 mark)

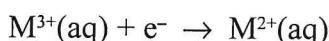
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- (b) Use the standard electrode potentials given below to answer the questions that follow.

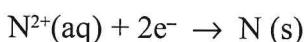
Reactions



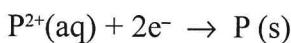
E^θ (V)
+ 1.49



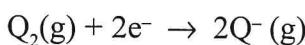
+ 0.77



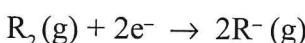
+ 0.34



- 0.23



+ 2.87



+ 1.36

- (i) State whether acidified MnO_4^- can oxidise M^{2+} . Give a reason. (2 marks)

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- (ii) Select two half-cells which when combined will give the highest e.m.f. (1 mark)

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- (iii) Write the cell representation for the cell formed in b (ii). (1 mark)

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- (iv) Calculate the E^θ value for the cell formed in b (iii). (2 marks)

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- (c) A mass of 1.24 g of a divalent metal was deposited when a current of 6A was passed through a solution of the metal sulphate for 12 minutes. Determine the relative atomic mass of the metal. (1 Faraday = 96,500 C mol⁻¹) (3 marks)

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- (d) State **two** applications of electrolysis. (1 mark)

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7. (a) What is meant by rate of reaction. (1 mark)

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- (b) In the space provided, sketch the diagram of a set-up that can be used to determine the rate of reaction between manganese(IV) oxide and hydrogen peroxide. (3 marks)

- (c) A student placed a small amount of liquid bromine at the bottom of a sealed gas jar of air as shown in **Figure 4**.

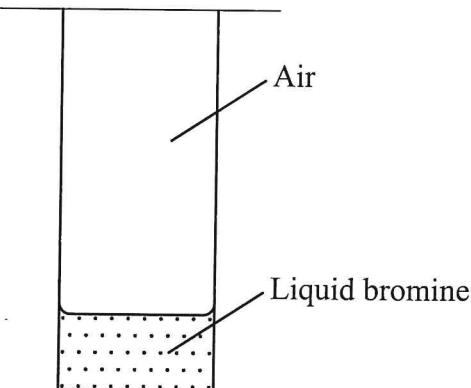


Figure 4

- (i) Describe what will be observed: (1 mark)

I. after two minutes

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II. after 30 minutes

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(ii) Use the Kinetic theory to explain the observations: (2 marks)

I. after 2 minutes

II. after 30 minutes

(d) Some plants have seeds that contain vegetable oil.

(i) Describe how the oil can be obtained from the seeds. (3 marks)

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(ii) Explain how it could be confirmed that the liquid obtained from the seeds is oil. (1 mark)

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