Extended or long form of periodic table

61. The electronic configuration of an atom A is $1s^2$, $2s^2p^6$, $3s^2p^6d^{10}, 4s^2p^3$. The chemistry of A is therefore likely to be similar to that of
(a) Chlorine (b) Nitrogen
(c) Oxygen (d) Boron

62. The element having the electronic configuration $1s^2, 2s^2 2p^6, 3s^2 3p^1$ is
(a) A transition element
(b) A representative element
(c) An inert gas
(d) An inner–transition element

63. The element with configuration $1s^2, 2s^2p^6, 3s^2$ would be
(a) A metal (b) A non-metal
(c) An inert gas (d) A metalloid

64. The long form of periodic table is based on
(a) Shape of the atom
(b) Mass of the atom
(c) Atomic number of the atom
(d) Electronegativity

65. Chloride of an element A gives neutral solution in water. In the periodic table, the element A belongs to
(a) First group
b) Third group
(c) Fifth group
d) First transition series

66. The fundamental basis of the present-day Periodic Table is that elements are
(a) Arranged in the order of increasing atomic weights
(b) Grouped according to chemical properties
(c) Arranged in the order of increasing number of neutrons in the atomic nucleus
(d) Arranged in the order of increasing number of protons in the nucleus

67. All the elements in a group in the periodic table have the same
(a) Atomic number
(b) Electronic configuration
(c) Atomic weight
(d) Number of electrons in the outermost shell or number of electrons for bonding

68. The most predominantly ionic compounds will be obtained from the combination of elements belonging to
(a) 1 and 7 groups
(b) 2 and 6 groups
(c) 3 and 5 groups
(d) 0 and 7 groups

69. An atom with atomic number 21 belongs to the category of
(a) s -block elements

(b) p -block elements
(c) d -block elements
(d) f -block elements

70. Which metal has 2 electrons in the outermost orbit
(a) *Na* (b) *Cu*
(c) *Au* (d) *Be*

71. In the modern periodic table, elements are arranged in
(a) Increasing mass
(b) Increasing volume
(c) Increasing atomic number
(d) Alphabetically

72. Alkali metals in each period have
(a) Smallest size
(b) Lowest ionization potential
(c) Highest ionization potential
(d) Highest electronegativity

73. The elements on the right side of the periodic table are
(a) Metals
(b) Metalloids
(c) Non-metals
(d) Transition elements

74. The screening effect of *d*-electons is
(a) Equal to that of *p*-electrons
(b) More than that of *p*-electrons
(c) Same as *f*-electrons
(d) Less than *p*-electrons

75. Chemical behaviour of an atom is determined by
(a) Atomic number
(b) Mass number
(c) Binding energy
(d) Number of isotopes

76. Which of the following is a inert element
(a) *Na* (b) *Fe*
(c) *Li* (d) *He*

77. The lightest metal is
(a) Li (b) Mg
(c) Ca (d) Na

78. Choose the typical element
(a) *K* (b) *Na*
(c) *Sc* (d) *He*

79. Of the following pairs, the one containing example of metalloid elements in the periodic table is
(a) Sodium and potassium
(b) Fluorine and chlorine
(c) Calcium and magnesium
(d) Boron and silicon

80. The number of elements in each of the long periods in the periodic table is
(a) 2 (b) 8
(c) 18 (d) 32

81. In the long form of the periodic table, all the non- metals are placed under
(a) *s*-block (b) *p*-block
(c) *d*-block (d) *f*-block

82. Elements with outer electronic configuration ns^2np^6 are
(a) Alkaline earth metals
(b) Transition elements
(c) Chalcogenes
(d) Noble gases

83. Highest density is of
(a) Ir (b) Os
(c) Pb (d) Hg

84. Lithium shows diagonal relationship with
(a) *Al* (b) *Mg*
(c) *Be* (d) *B*

85. $1s^2 2s^2 2p^6 3s^2$ is the electronic configuration of the metal
(a) Na (b) Mg
(c) Fe (d) Al

86. Element having atomic number 17 is placed in
(a) *I*-group (b) *V*-group
(c) *VIII*-group (d) *VII*-group

87. The most importasnt active step in the development of periodic table was taken by
(a) Mendeleef (b) Dalton
(c) Avogadro (d) Cavendish

88. Who is called the father of chemistry
(a) Faraday (b) Priestley
(c) Rutherford (d) Lavosier

89. The total number of rare–earth elements is
(a) 8 (b) 32
(c) 14 (d) 10

90. Which is metalloid
(a) Pb (b) Sb
(c) Bi (d) Zn
(e) Mg