

**Extended or long form of periodic table**

- Which of the following statement is not correct for the element having electronic configuration $1s^2, 2s^2p^6, 3s^1$
 - It is a monovalent electropositive
 - It forms basic oxide
 - It is a non-metal
 - It has low electron affinity
- Which of these does not reflect the periodicity of the elements
 - Bonding behaviour
 - Electronegativity
 - Ionization energy
 - Neutron/proton ratio
- If an atom has electronic configuration $1s^2 2s^2 2p^6 3s^2 3p^6 3d^3 4s^2$, it will be placed in

(a) Second group	(b) Third group
(c) Fifth group	(d) Sixth group
- All the s-block elements of the periodic table are placed in the groups

(a) IA and IIA	(b) IIIA and IVA
(c) B sub groups	(d) VA to VIIA
- The electronic configuration of halogen is

(a) $ns^2 np^6$	(b) $ns^2 np^3$
(c) $ns^2 np^5$	(d) ns^2
- Hydrogen by donating one electron forms H^+ . In this property, it resembles with
 - Transitional metals
 - Alkaline earth metals
 - Alkali metals
 - Halogens
- The tenth elements in the periodic table resembles with the
 - First period
 - Second period
 - Fourth group
 - Ninth group
- The element with quantum numbers $n = 2, l = 1, m = 1, s = -1/2$ has the following position in the periodic table
 - Group VII-A, period II
 - Group 0, period II
 - Group VII-A, period III
 - Group 0, period III
- Who developed the long form of periodic table

(a) Lothar Meyer	(b) Niels Bohr
(c) Mendeleeff	(d) Moseley
- The electronic configuration of an element is $1s^2, 2s^2 2p^6, 3s^2 3p^3$. What is the atomic number of the element which is just below the above element in the periodic table

(a) 33	(b) 34
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- (c) 31 (d) 49
- (b) Seven horizontal rows and eighteen vertical columns
- (c) Seven horizontal rows and seven vertical columns
- (d) Eight horizontal rows and eight vertical columns
11. In the periodic table, the element with atomic number 16 will be placed in the group
- (a) Third (b) Fourth
- (c) Fifth (d) Sixth
12. The first element of rare-earth metals is
- (a) Cerium (b) Actinium
- (c) Uranium (d) Lanthanum
13. The *d*-block elements consists mostly of
- (a) Monovalent metals
- (b) All non-metals
- (c) Elements which generally form stoichiometric metal oxide
- (d) Many metals with catalytic properties
14. "The 6 properties of the elements are periodic function of their atomic numbers." The statement was given by
- (a) N. Bohr
- (b) J.W. Dobereiner
- (c) D.I. Mendeleef
- (d) H.G.J. Moseley
15. The long form of periodic table has
- (a) Eight horizontal rows and seven vertical columns
16. The telluric helix was given by
- (a) De Chan Courtois (b) Newlands
- (c) L. Meyer (d) Mendeleef
17. Which one of the following belongs to representative group of elements in the periodic table
- (a) Lanthanum (b) Argon
- (c) Chromium (d) Aluminium
18. An element of atomic number 29 belongs to
- (a) *s*-block (b) *p*-block
- (c) *d*-block (d) *f*-block
19. The element californium belongs to the family
- (a) Actinide series
- (b) Alkali metal family
- (c) Alkaline earth family
- (d) Lanthanide series
20. On moving from left to right across a period in the table the metallic character
- (a) Increases
- (b) Decreases





- (c) Remains constant
(d) First increases and then decreases
- (b) Atomic number
(c) Number of neutrons
(d) None of the above
21. An element with atomic number 20 will be placed in which period of the periodic table
(a) 4 (b) 3
(c) 2 (d) 1
22. The electronic structure $(n-1)d^{1-10}ns^{0-2}$ is characteristic of
(a) Transition elements
(b) Lanthanides
(c) Actinides
(d) Rare gases
23. The elements with atomic number 10, 18, 36, 54 and 86 are all
(a) Light metals (b) Inert gases
(c) Halogens (d) Rare-earths
24. Elements of atomic number 6 is placed in
(a) IV group (b) IV period
(c) VI group (d) III group
25. Which of the following elements is a lanthanide (Rare-earth element)
(a) Cadmium (b) Californium
(c) Cerium (d) Cesium
26. Mendeleef's periodic law is based on
(a) Atomic weight
27. The heaviest atom amongst the following is
(a) *U* (b) *Ra*
(c) *Pb* (d) *Hg*
28. Which of the following pairs has both members from the same group of the periodic table
(a) *Mg* – *Ba* (b) *Mg* – *Na*
(c) *Mg* – *Cu* (d) *Mg* – *K*
29. Which of the following pairs has both members from the same period of the periodic table
(a) *Na* – *Ca* (b) *Na* – *Cl*
(c) *Ca* – *Cl* (d) *Cl* – *Br*
30. Diagonal relationship is shown by
(a) Elements of first period
(b) Elements of second period
(c) Elements of third period
(d) (b) and (c) both

