Electrovalent bonding

61. Which of the following compounds is ionic

(a) KI (b) CH_4

(c) Diamond (d) H_2

62. Which of the following pairs of species has same electronic configuration

(a) Zn^{2+} and Ni^{2+}

(b) Co^{+3} and Ni^{4+}

(c) Co^{2+} and Ni^{2+}

(d) Ti^{4+} and V^{3+}

63. The energy that opposes dissolution of a solvent is

(a) Hydration energy

(b) Lattice energy

(c) Internal energy

(d) Bond energy

64. Which of the following has highest melting point

(a) $BeCl_2$ (b) $MgCl_2$

(c) $CaCl_2$ (d) $BaCl_2$

65. Which of the following statements is not true for ionic compounds

(a) High melting point

(b) Least lattice energy

(c) Least solubility in organic compounds

(d) Soluble in water

66. Electrolytes are compound containing

(a) Electrovalent bond

(b) Covalent bond

(c) Coordinate bond

(d) Hydrogen bond

67. Which of the following hydrides are ionic

(a) CaH_2 (b) BaH_2

(c) SrH_2 (d) BeH_2

68. Which of the following conduct electricity in the fused state

(a) $BeCl_2$ (b) $MgCl_2$

(c) $SrCl_2$ (d) $BaCl_2$