

Electrovalent bonding

41. Among the bonds formed by a chlorine atom with atoms of hydrogen, chlorine, sodium and carbon, the strongest bond is formed between
(a) $H - Cl$ (b) $Cl - Cl$
(c) $Na - Cl$ (d) $C - Cl$
42. Which of the following is least soluble
(a) BeF_2 (b) SrF_2
(c) CaF_2 (d) MgF_2
43. Which of the following halides has maximum melting point
(a) $NaCl$ (b) $NaBr$
(c) NaI (d) NaF
44. The high melting point and insolubility in organic solvents of sulphanilic acid are due to its structure.
(a) Simple ionic
(b) Bipolar ionic
(c) Cubic
(d) Hexagonal
45. Out of the following, which compound will have electrovalent bonding
(a) Ammonia
(b) Water
(c) Calcium chloride
(d) Chloromethane
46. The force which holds atoms together in an electrovalent bond is
(a) Vander Waal's force
(b) Dipole attraction force
(c) Electrostatic force of attraction
(d) All the above
47. The main reaction during electrovalent bond formation is
(a) Redox reaction
(b) Substitution reaction
(c) Addition reaction
(d) Elimination reaction
48. Electrovalent compounds are
(a) Good conductor of electricity
(b) Polar in nature
(c) Low M.P. and low B.P.
(d) Easily available
49. Ionic compounds do not have
(a) Hard and brittle nature
(b) High melting and boiling point
(c) Directional properties
(d) Soluble in polar solvents
50. Highest melting point would be of



- (a) He (b) $CsCl$
(c) NH_3 (d) $CHCl_3$
51. What is the effect of more electronegative atom on the strength of ionic bond
(a) Decreases
(b) Increases
(c) Decreases slowly
(d) Remains the same
52. An element X with the electronic configuration $1s^2, 2s^2 2p^6, 3s^2$ would be expected to form the chloride with the formula
(a) XCl_3 (b) XCl_2
(c) XCl (d) X_2Cl
53. Two element have electronegativity of 1.2 and 3.0. Bond formed between them would be
(a) Ionic
(b) Polar covalent
(c) Co-ordinate
(d) Metallic
54. Which of the following is least ionic
(a) C_2H_5Cl
(b) KCl
(c) $BaCl_2$
(d) $C_6H_5N^+H_3Cl^-$
55. Which type of bonding exists in Li_2O and CaF_2 respectively
(a) Ionic, ionic
(b) Ionic, covalent
(c) Covalent, ionic
(d) Coordinate, ionic
56. An atom with atomic number 20 is most likely to combine chemically with the atom whose atomic number is
(a) 11 (b) 14
(c) 16 (d) 10
57. Bond formed in crystal by anion and cation is
(a) Ionic (b) Metallic
(c) Covalent (d) Dipole
58. Atoms or group of atoms which are electrically charged are known
(a) Anions (b) Cations
(c) Ions (d) Atoms
59. Which one is the strongest bond
(a) $Br-F$ (b) $F-F$
(c) $Cl-F$ (d) $Br-Cl$
60. The interionic attraction depends on interaction of





- (a) Solute-Solute
- (b) Solvent-Solvent
- (c) The charges
- (d) Molecular properties

