

Hydrogen bonding





- (a) NH_3
- (b) P
- (c) As
- (d) Sb

38. The boiling point of a compound is raised by

- (a) Intramolecular hydrogen bonding
- (b) Intermolecular hydrogen bonding
- (c) Covalent bonding
- (d) Ionic covalent

39. The boiling point of water is exceptionally high because

- (a) Water molecule is linear
- (b) Water molecule is not linear
- (c) There is covalent bond between H and O
- (d) Water molecules associate due to hydrogen bonding

40. NH_3 has a much higher boiling point than PH_3 because

- (a) NH_3 has a larger molecular weight
- (b) NH_3 undergoes umbrella inversion
- (c) NH_3 forms hydrogen bond
- (d) NH_3 contains ionic bonds whereas PH_3 contains covalent bonds

