

Covalent bonding

61. The acid having $O - O$ bond is
 (a) $H_2S_2O_3$ (b) $H_2S_2O_6$
 (c) $H_2S_2O_8$ (d) $H_2S_4O_6$
62. The following salt shows maximum covalent character
 (a) $AlCl_3$ (b) $MgCl_2$
 (c) $CsCl$ (d) $LaCl_3$
63. Which type of bond is present in H_2S molecule
 (a) Ionic bond
 (b) Covalent bond
 (c) Co-ordinate
 (d) All of three
64. H_2S is more acidic than H_2O , due to]
 (a) O is more electronegative than S
 (b) $O - H$ bond is stronger than $S - H$ bond
 (c) $O - H$ bond is weaker than $S - H$ bond
 (d) None of these
65. Which of the following has covalent bond
 (a) Na_2S (b) $AlCl_3$
 (c) NaH (d) $MgCl_2$
66. The following element forms a molecule with eight its own weight atoms
 (a) Si (b) S
 (c) Cl (d) P
67. In H_2O_2 , the two oxygen atoms have
 (a) Electrovalent bond
 (b) Covalent bond
 (c) Coordinate bond
 (d) No bond
68. Carbon has a valency of 2 in CO and 4 in CO_2 and CH_4 . Its valency in acetylene (C_2H_2) is
 (a) 1 (b) 2
 (c) 3 (d) 4
69. Number of electrons in the valence orbit of nitrogen in an ammonia molecule are
 (a) 8 (b) 5
 (c) 6 (d) 7
70. Hydrogen atoms are held together to form hydrogen molecules by
 (a) Hydrogen bond (b) Ionic bond
 (c) Covalent bond (d) Dative bond
71. Strongest bond is
 (a) $C - C$ (b) $C - H$



- (c) $C - N$ (d) $C - O$ (b) NaH and CaH_2
(c) NaH and NH_3

72. The major binding force of diamond, silicon and quartz is
(a) Electrostatic force
(b) Electrical attraction
(c) Co-valent bond force
(d) Non-covalent bond force

73. Multiple covalent bonds exist in a molecule of

- (a) H_2 (b) F_2
(c) C_2H_4 (d) N_2

74. Which of the following does not obey the octet rule

- (a) CO (b) NH_3
(c) H_2O (d) PCl_5

75. Which of the following statements is correct for covalent bond

- (a) Electrons are shared between two atoms
(b) It may be polar or non-polar
(c) Direction is non-polar
(d) Valency electrons are attracted

76. Among CaH_2 , NH_3 , NaH and B_2H_6 , which are covalent hydride

- (a) NH_3 and B_2H_6

