

Reversible and Irreversible reaction

1. (b) Reversible reaction always attains equilibrium which proceeds both sides and never go for completion.
2. (c) In a reversible reaction some amount of the reactants remains unconverted into products.
3. (d) In lime kiln CO_2 escaping regularly so reaction proceeds in forward direction.
4. (b)

Explanation: At equilibrium, the rate of forward and reverse reactions becomes equal, so the concentrations of all substances remain constant.

5. (a)

Explanation: The reaction between hydrogen and iodine forming hydrogen iodide is a reversible reaction, while the others go to completion and are irreversible.

6. (d)

Explanation: Decomposition (disintegration) reactions may be reversible or irreversible and can be either endothermic or exothermic depending on the substances involved.

7. (b) The reaction is not reversible.

