

Corrosion

- Corrosion is basically a
 - Altered reaction in presence of H_2O
 - Electrochemical phenomenon
 - Interaction
 - Union between light metal and heavy metal
- Rusting of iron is catalysed by which of the following
 - Fe
 - O_2
 - Zn
 - H^+
- Which of the following is a highly corrosive salt
 - $FeCl_2$
 - $PbCl_2$
 - Hg_2Cl_2
 - $HgCl_2$
- Corrosion of iron is essentially an electrochemical phenomenon where the cell reactions are
 - Fe is oxidised to Fe^{2+} and dissolved oxygen in water is reduced to OH^-
 - Fe is oxidised to Fe^{3+} and H_2O is reduced to O_2^{2-}
 - Fe is oxidised to Fe^{2+} and H_2O is reduced to O_2^-
 - Fe is oxidised to Fe^{2+} and H_2O is reduced to O_2

