

Acids and Bases



- 131.** An aqueous solution of aluminium sulphate shows
(a) A basic nature
(b) An acidic nature
(c) A neutral nature
(d) Both acidic and basic nature

132. Neutralization of an acid with a base invariably results in the production of
(a) H_3O^+
(b) OH^-
(c) H_2O
(d) H^+ and OH^-

133. The conjugate acid of HPO_4^{2-} is
(a) $H_2PO_4^-$
(b) PO_4^{3-}
(c) H_3PO_4
(d) H_3PO_3

134. Which of the following is not used as a Lewis acid
(a) $SnCl_4$
(b) $FeCl_3$
(c) KCl
(d) BF_3

135. Orthoboric acid in aqueous medium is
(a) Monobasic
(b) Dibasic
(c) Tribasic
(d) All are correct

136. According to Lewis concept which one of the following is not a base
(a) OH^-
(b) H_2O
(c) Ag^+
(d) NH_3

137. The aqueous solution of ammonium chloride is
(a) Neutral
(b) Basic
(c) Acidic
(d) Amphoteric

138. In the process $BCl_3 + PH_3 \rightarrow BCl_3 \cdot PH_3$ The Lewis acid is
(a) PH_3
(b) BCl_3
(c) Both
(d) None

139. The conjugate acid of NH_3 is
(a) NH_3
(b) NH_4^+
(c) N_2H_4
(d) NH_2OH

140. Which halide of nitrogen is least basic
(a) NBr_3
(b) NI_3
(c) NCl_3
(d) NF_3

