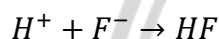


161. (b) Gaseous HCl does not give H^+ but liquid HCl gives H^+ in aqueous solution there for gaseous HCl is not a Arrhenius acid due to covalent bonding in gaseous condition.

162. (d) $H_2O \rightarrow H^+ + OH^-$ [Acid due to donation of proton]
Acid

$H_2O + H^+ \rightarrow H_3O^+$ [Basic due to gaining of proton]
Base

163. (c) F^- strongest conjugate base due to it smallest size in a group and gain proton due to most electronegative capacity.



164. (d) $H_2PO_4^- \rightarrow H^+ + HPO_4^{2-}$
Conjugate acid Conjugate base

165. (a) $HSO_4^- \rightarrow H^+ + SO_4^{2-}$
Conjugate acid Conjugate base

