Matter :- Anything which occupy space and have specific volume is called Matter.

made up of ↓

Particles which are → infinitely small

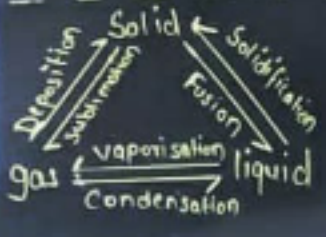
Characteristics

- Space b/w them
- Continuously moving
- Attract each other

Evaporation : Particles are never at rest, some have high K.E. and some have less, and the particles with high KE get convrtd* into vapour This phenomenon is called Evaporation.

Evaporation cause cooling as the particle which have became vapour gain the energy from sorrounding to get move and makes the sorrounding cool.

For eg :- nail polish remover makes the nails feel cool.



* vol^{m} → volume
KE → Kinetic Energy
convrtd → converted

States Of Matter :->

Solid :
- Particles are closely bounded
- They only vibrate
- Have definite shape & volume
- Eg - steel, rubber, iron etc.
- Not get diffused

Liquid :
- Particle are at moderate distance.
- They can move.
- Have definite volume but not definit shape
- Eg - milk, water, ink etc.
- Compare to solid get more diffused.

Gas :
- Particles are at very far distance.
- The can also move
- Neither have definite vol^{m}* nor definite shape.
- Eg. → air, carbon dioxide, Hydrogen etc.
- Diffusion rate is very good.

Change in State of Matter :-
Due To

Effect of temprature : increasing temp. increases K.E* of particles and changes are following
Solid → (Δ) liquid → (Δ) gas

Effect of pressure : incrsng press. decreases KE* of particle and changes are following
gas → (↑ Pres.) liquid → (↑ Pres.) Solid