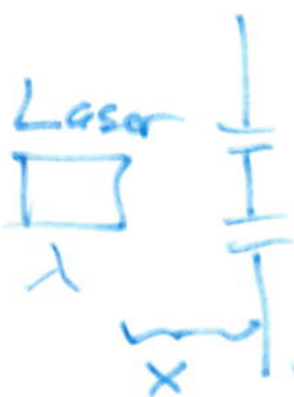


Exam 3 Review

See updated equation sheet.

2x Slit



$x = 51$
cm

$\lambda = 650 \text{ nm}$

Find dist (on screen) from center to 3rd dark spot on one side.

$$d \sin \theta = m \lambda$$

$$a \sin \theta = p \lambda$$

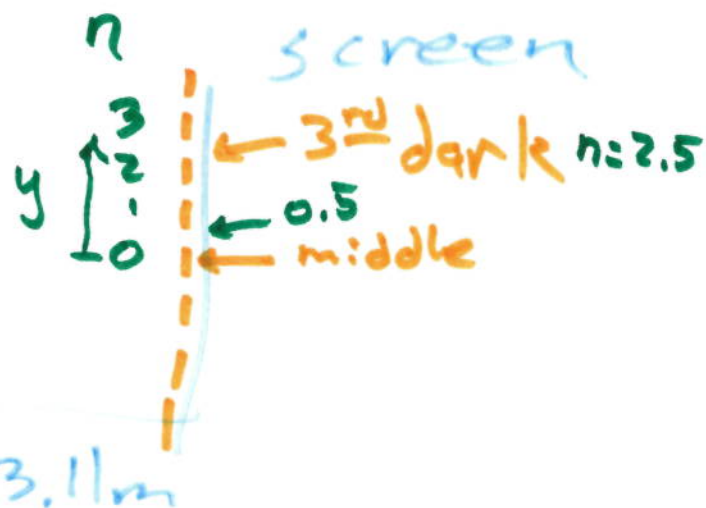
$m = 0, \pm 1, \pm 2, \dots$ Bright

$p = \pm 1, \pm 2, \pm 3, \dots$ Dark

$d \sin \theta = n \lambda$

double slit bright spots

$0.617 \text{ mm} \frac{y}{L} = 2.5 (650 \text{ nm})$



$$y = \frac{2.5 (650 \text{ nm}) \times 3.11 \text{ m}}{0.617 \text{ mm}} = 8.19 \text{ mm}$$

$10^{-9} \quad 10^{-3}$

2 states:

Ⓐ	E $3.2 \frac{\text{Kcal}}{\text{mol}}$	S
Ⓑ	0	0
		$-0.015 \frac{\text{Kcal}}{\text{mol} \cdot \text{K}}$

at $T=300\text{K}$, find P_A .

2 state systems

$$P_A = \frac{1}{1 + e^{-\Delta G / RT}}$$

$$\Delta G = G_B - G_A$$

$$= E_B - TS_B - E_A + TS_A$$

$$= 0 - 300\text{K}(-0.015 \frac{\text{Kcal}}{\text{mol} \cdot \text{K}}) - 3.2 \frac{\text{Kcal}}{\text{mol}} + T \cdot 0$$

$$= 4.5 \frac{\text{Kcal}}{\text{mol}} - 3.2 \frac{\text{Kcal}}{\text{mol}} = 1.3 \frac{\text{Kcal}}{\text{mol}}$$

$$P_A = \frac{1}{1 + e^{-1.3 \frac{\text{Kcal}}{\text{mol}} / (0.002 \frac{\text{Kcal}}{\text{mol} \cdot \text{K}} \cdot 300\text{K})}} = 0.897$$

$$P_B = 1 - P_A = 0.102$$

Check: $E(x,t) = E_0 (Ax^3 - Bt^3)$

is a sol'n of $\boxed{c^2 \frac{d^2 E}{dx^2} = \frac{d^2 E}{dt^2}}$

If yes, find c
If no, why no?

$$\frac{\partial E}{\partial x} = E_0 A 3x^2$$

$$\frac{d^2 E}{dx^2} = E_0 A 6x$$

$$\frac{\partial E}{\partial t} = -E_0 B 3t^2$$

$$\frac{d^2 E}{dt^2} = -E_0 B 6t$$

$$c^2 E_0 A 6x \stackrel{?}{=} -E_0 B 6t$$

$$c^2 A x \stackrel{?}{=} -Bt$$

$$c^2 = \frac{-Bt}{Ax} \quad \times \quad \underline{\text{No}} \quad c^2 = \text{const} \quad \text{not a} \quad \text{fun}(x,t)$$

$Q = cm\Delta T$ yes. Know this

$Q = Lm$ No. Not on eqv. sheet.

You have 300g unknown liquid.
at 17°C . Add 45g hot (100°C)
copper $c = \frac{J}{0.385 \text{ g}^\circ\text{C}}$. Find c of unknown liquid.
If final T of mixture is 20.9°C , find c of unknown liquid.

$$Q = cm\Delta T$$

Assume no heat loss (meaning all heat from copper goes into liquid).

$$0 = C_1 m_1 \Delta T_1 + C_c m_c \Delta T_c$$

$$\Delta T = T_f - T_{\text{init}}$$

$$0 = C_1^* 300\text{g} (20.9^\circ\text{C} - 17^\circ\text{C}) + 0.385 \frac{\text{J}}{\text{g}^\circ\text{C}} 45\text{g} (20.9^\circ\text{C} - 100^\circ\text{C})$$

$$0 = C_1 (1170 \text{ g}^\circ\text{C}) + -1370.4 \text{ J}$$

$$\frac{1170}{1370.4}$$

$$= C_1 = \frac{1370 \text{ J}}{1170 \text{ g}^\circ\text{C}} =$$
$$C_1 = 1.2 \text{ J/g}^\circ\text{C}$$

AC: $V = 9.17V$ $f = 546Hz$

$R = 123\Omega$ $\underbrace{\hspace{1cm}}_{\text{series}} C = 8.76 \times 10^{-7}F$

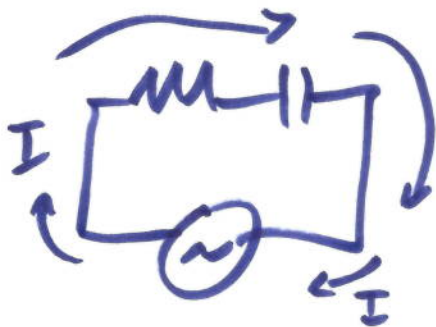
find I , V across each element.

$X_R = R$ $X_C = \frac{1}{\omega C}$

$\omega = 2\pi f = 3430 \frac{1}{s}$

$X_C = 333.0\Omega$

$X_L = L\omega$



$V = XI$

$V = ZI$

$Z = \sqrt{R^2 + (X_L - X_C)^2} = \sqrt{(123\Omega)^2 + (333\Omega)^2}$

$Z = 355\Omega$

$I = \frac{V}{Z} = \frac{9.17V}{355\Omega} = 2.58 \times 10^{-2}A$
 $= 25.8mA$

$V = XI$

$V_R = RI = 123\Omega (2.58 \times 10^{-2}A) = \cancel{0.318V}$
 $V_C = X_C I = 333\Omega (2.58 \times 10^{-2}A) = \cancel{3.17V}$
 $= 8.59V$