

Solutions to Pathria's Statistical Mechanics

Chapter 6

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Problem 6.1

Problem 6.13

For classical ideal gases, the velocity distribution of the atoms are the Maxwell distribution, which can be obtained by canonical distribution:

$$f(\mathbf{u}) = \left(\frac{m}{2\pi kT}\right)^{3/2} e^{-\frac{m\mathbf{u}^2}{2kT}} \quad (1)$$

a) So the number of impacts made by gas molecules on a unit area of the wall in a unit time for which the angle of incidence lies between θ and $\theta + d\theta$ will be:

$$\frac{dN}{dAdt} = \frac{N}{V} \int_0^\infty du \int_0^{2\pi} d\phi \left(\frac{m}{2\pi kT}\right)^{3/2} e^{-\frac{m\mathbf{u}^2}{2kT}} u^3 \sin \theta \cos \theta d\theta = \frac{N}{V} \sqrt{\frac{2kT}{\pi m}} \sin \theta \cos \theta d\theta \quad (2)$$

b) The number of impact for which the speed lies between u and $u + du$ is

$$\frac{dN}{dAdt} = \frac{N}{V} \int_0^{2\pi} d\phi \int_0^{\pi/2} d\theta \left(\frac{m}{2\pi kT}\right)^{3/2} e^{-\frac{m\mathbf{u}^2}{2kT}} u^3 \sin \theta \cos \theta du = \frac{\pi N}{V} \left(\frac{m}{2\pi kT}\right)^{3/2} e^{-\frac{m\mathbf{u}^2}{2kT}} u^3 du \quad (3)$$

c) The rate of dissociate is given by the following impact number with following condition $u > \sqrt{2\epsilon_0/m} = u_0$ in which $\epsilon_0 = 10^{-19}\text{J}$, so the impacting number will be:

$$\begin{aligned} \frac{dN}{dAdt} &= \frac{N}{V} \int_0^{2\pi} d\phi \int_0^{\pi/2} d\theta \int_{u_0}^\infty \left(\frac{m}{2\pi kT}\right)^{3/2} e^{-\frac{m\mathbf{u}^2}{2kT}} u^3 \sin \theta \cos \theta \\ &= \frac{\pi N}{V} \int_{u_0}^\infty du u^3 \left(\frac{\epsilon_0}{u_0^2 kT}\right)^{3/2} \exp\left(-\frac{\epsilon_0}{u_0^2 kT} u^2\right) \end{aligned} \quad (4)$$

Then define a function $f(u, T)$ as shown:

$$f(u, T) = \left(\frac{\epsilon_0}{u_0^2 kT}\right)^{3/2} \exp\left(-\frac{\epsilon_0}{u_0^2 kT} u^2\right)$$

so for two different temperature, we can calculate the ratio of the function $f(u, T)$:

$$\begin{aligned} \frac{f(u, T_2 = 310\text{K})}{f(u, T_1 = 300\text{K})} &= \left(\frac{T_1}{T_2}\right)^{3/2} \exp\left[\frac{\epsilon_0 u^2}{u_0^2 k} \left(\frac{1}{T_1} - \frac{1}{T_2}\right)\right] \\ &\geq \left(\frac{T_1}{T_2}\right)^{3/2} \exp\left[\frac{\epsilon_0}{k} \left(\frac{1}{T_1} - \frac{1}{T_2}\right)\right] = 2.075 > 2 \end{aligned} \quad (5)$$

so the integral of $f(u, T_2)$ will be twice larger than that of $f(u, T_1)$, which means the dissociation probability will be more than twice.

Problem 6.23

The first step is to find the equilibrium point:

$$\frac{\partial V(r)}{\partial r} = 0 \quad \Rightarrow \quad r_{\text{eq}} = r_0$$

Then assume that the oscillating mode will not be excited at the lowest energy scale of rotation mode, and calculate the moment of inertia:

$$I = \frac{mr_0^2}{2}$$

in which m is the mass of a hydrogen atom and r_0 is the equilibrium length. So the energy scale of the rotation mode will be

$$\epsilon_{\text{rot}} = \frac{\hbar^2}{mr_0^2} l(l+1) \quad (6)$$

and the characteristic temperature of rotation will be

$$T_{\text{rot}} = \frac{\hbar^2}{kmr_0^2} \sim 78\text{K}$$

Then talk about the oscillation mode. Expand the potential energy around r_0 and we can find the potential will be:

$$V(r) \sim -V_0 + \frac{1}{2} \frac{2V_0}{a^2} (\delta r)^2 \quad (7)$$

since the reduced mass of the hydrogen molecule is $\mu = m/2$, the oscillating frequency is

$$\omega = \sqrt{\frac{2V_0}{a^2\mu}} = \sqrt{\frac{4V_0}{a^2m}} \quad (8)$$

and its characteristic temperature is

$$T_{\text{oc}} = \frac{\hbar}{k} \sqrt{\frac{4V_0}{a^2m}} \sim 6394\text{K}$$

Since $T_{\text{oc}} \gg T_{\text{rot}}$, the approximation that the molecule is rigid when we calculate the moment of inertia is acceptable. And our conclusion is that the oscillating mode will begin to contribute to the heat capacity at temperature $\sim 6000\text{K}$ and vibration mode will begin to contribute to heat capacity at temperature $\sim 70\text{K}$.

Problem 6.33

Assumes that the initial density of O_2 and CH_4 is

$$\begin{aligned} n(\text{O}_2) &= n_{\text{O}_2} n_0 \\ n(\text{CH}_4) &= n_{\text{CH}_4} n_0 \end{aligned}$$

then the equilibrium equation tells us that the remained density of methane will be:

$$\delta = n_0 \frac{4n_{\text{CH}_4}^3}{(n_{\text{O}_2} - 2n_{\text{CH}_4})K(T)} \quad (9)$$

And from the given condition, we can calculate the equilibrium constant will be:

$$K(T) = \exp\left(-\beta\delta\mu^{(0)}\right) = 6.75 \times 10^{27}. \quad (10)$$