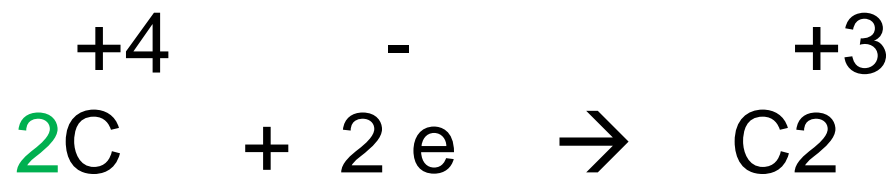


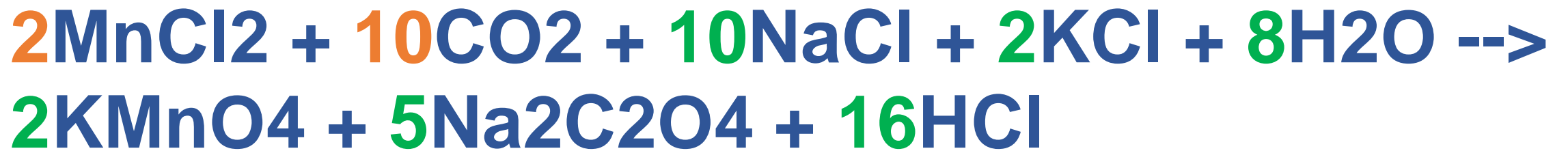
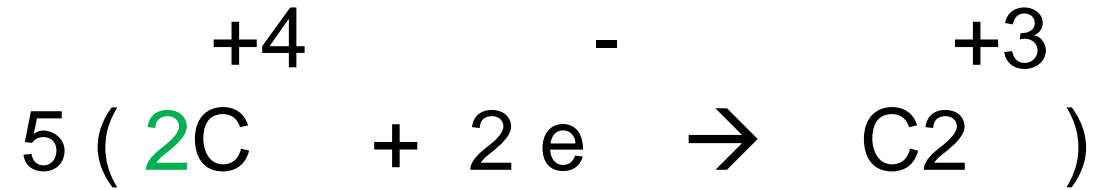
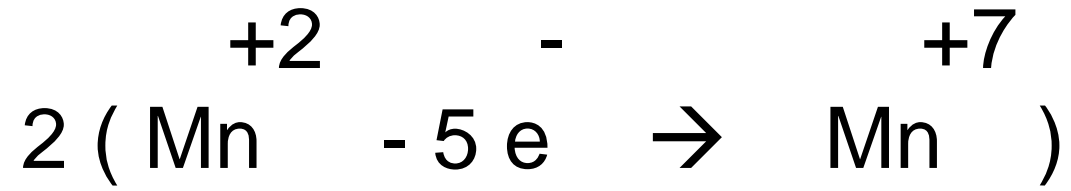
Oxidación

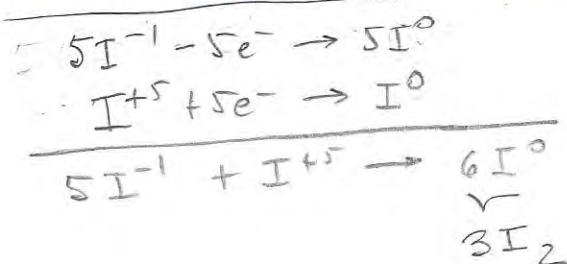
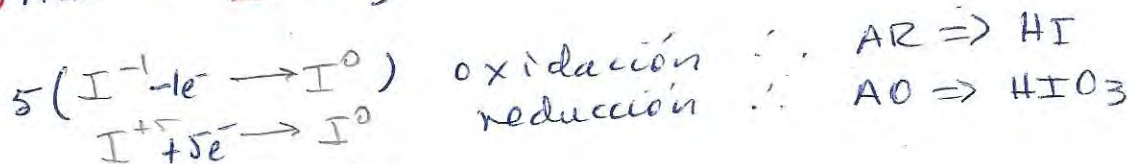
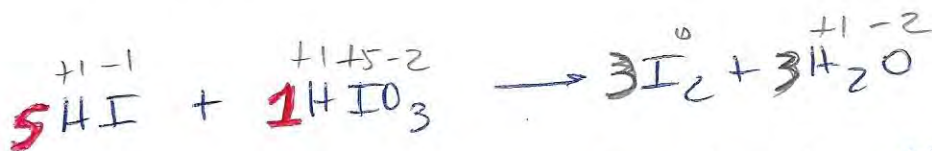
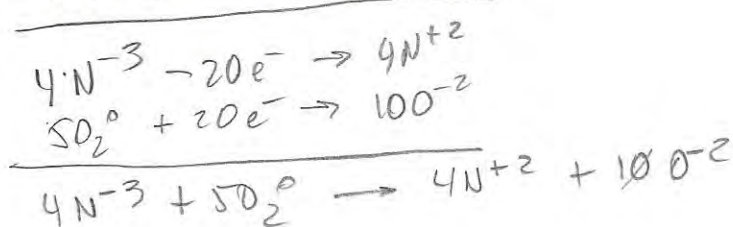
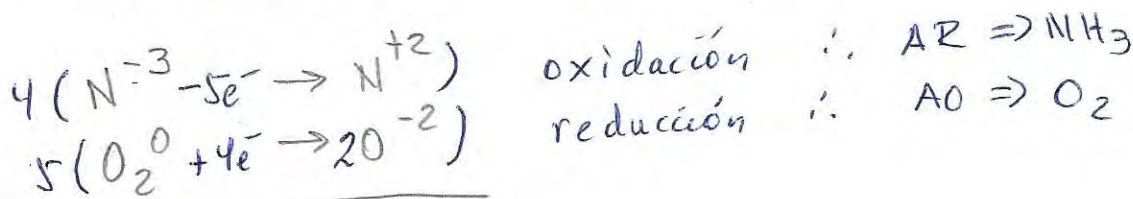
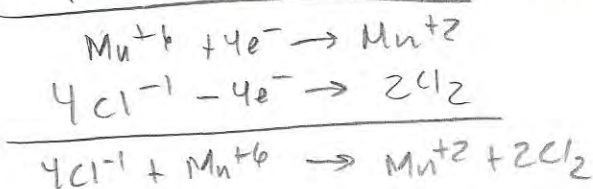
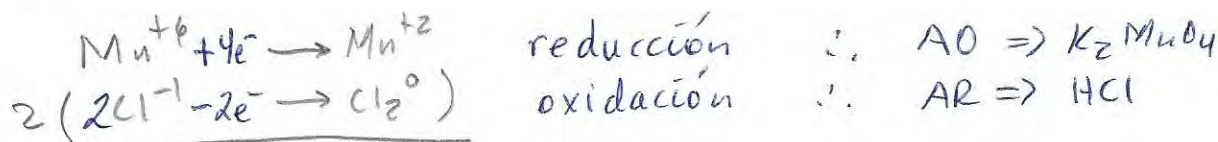
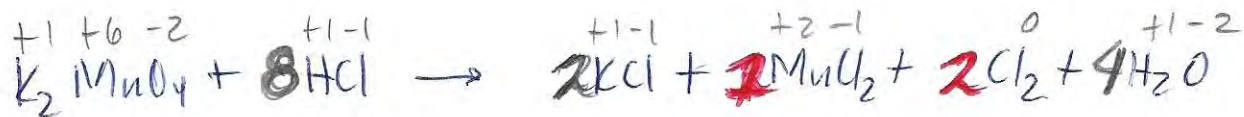
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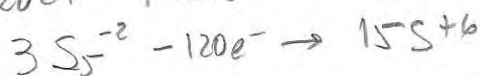
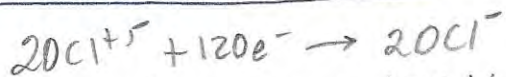
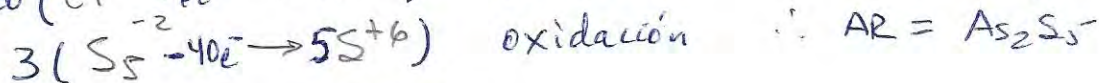
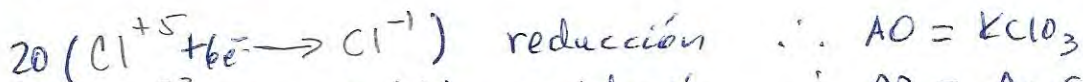
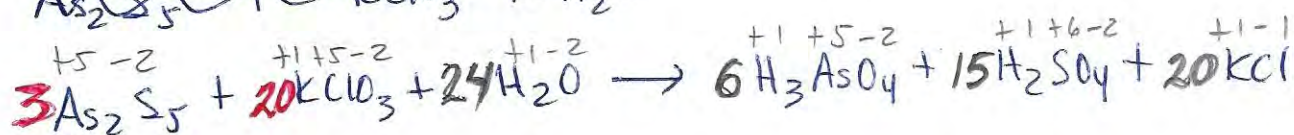
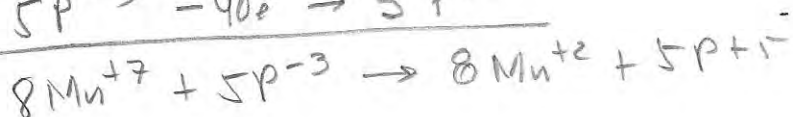
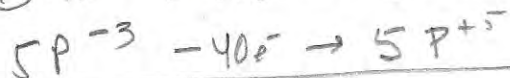
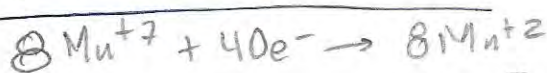
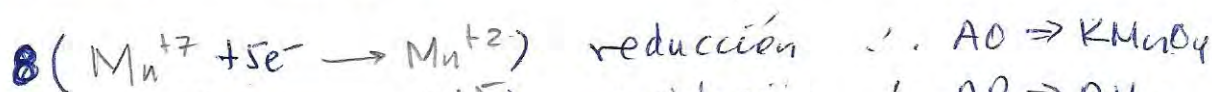
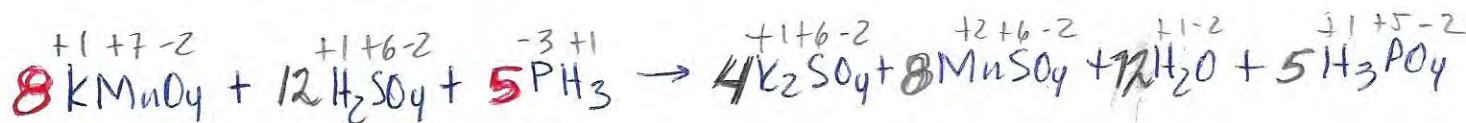
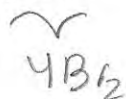
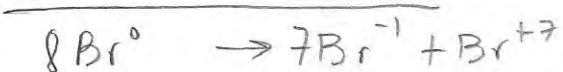
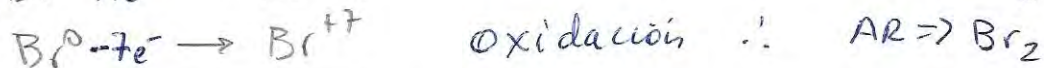


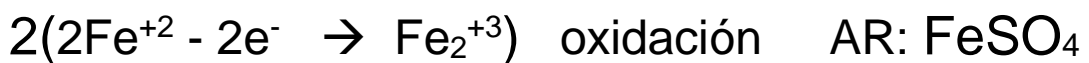
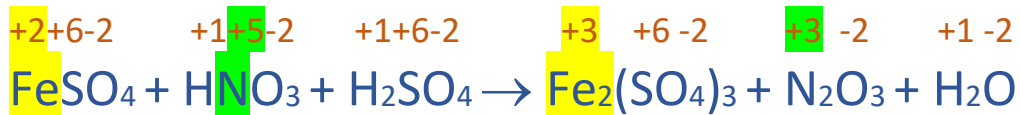
Reducción

AO: CO_2





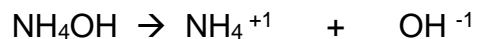
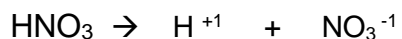




MÉTODO REDOX; ION-ELECTRÓN

a) Medio ácido.

1.- En la ecuación molecular global, disociar los compuestos posibles (ÁCIDOS, HIDRÓXIDOS Y SALES), recuerda que los óxidos, peróxidos, hidruros y excepciones, **NO SE DISOCIAN**.



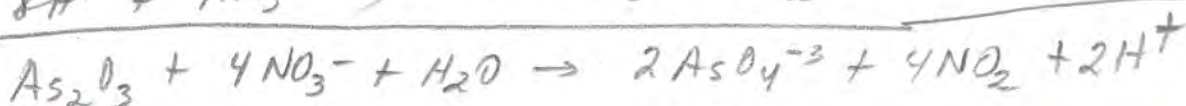
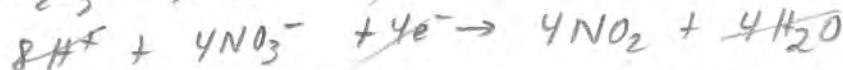
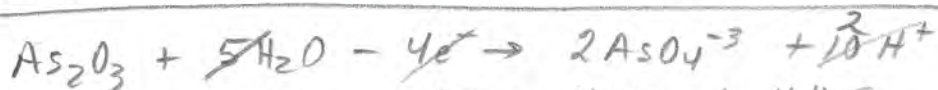
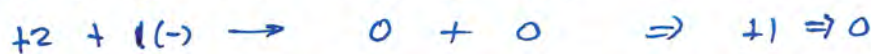
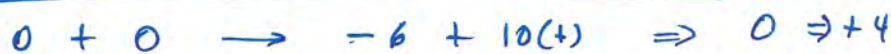
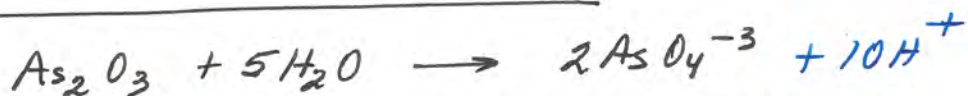
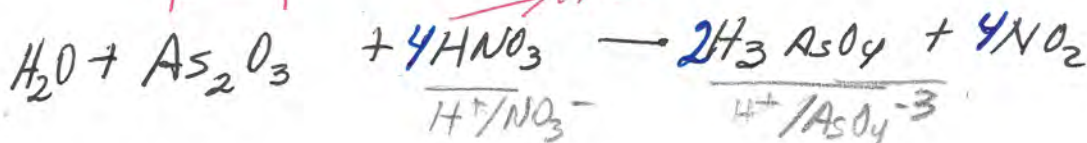
2. Identificar las especies que están sufriendo un cambio en estructura.
3. Ajustar (balancear) los átomos de elementos participantes (H/O; no).
4. Ajustar átomos de oxígeno agregando moléculas de agua en el lado que sea necesario.
5. Ajustar átomos de hidrógeno agregando iones hidrógeno (H^{+1}), en el lado que sea necesario.
6. Realizar sumatoria de cargas; agregar o eliminar electrones para igualar ambos lados.
7. Identificar proceso de oxidación y proceso de reducción, así como el agente oxidante y agente reductor.
8. Proceder a realizar la(s) operación(es), para eliminar los electrones ganados y perdidos.
9. Utilizar los coeficientes en la reacción iónica total, en la ecuación molecular y ajustar finalmente por observación (tanteo).

b) Medio básico.

- i. Se identifica por la presencia de hidróxidos, derivados de NH_3 , PH_3 , mientras no exista presencia de ácidos.
- ii. Se precede igual que en medio ácido del numeral **1 al 5**.
- iii. Considerando la cantidad de iones hidrógeno (H^{+1}), se agregan tantos OH^- , como sean necesarios para formar agua ($\text{H}^{+1}/\text{OH}^{-1}$), en **AMBOS** lados de la semirreacción.
- iv. Una vez que se complementaron las moléculas de agua, se debe realizar reducción de especies iguales en cada semirreacción, para proceder ajustar cargas agregando o eliminando electrones.
- v. Se procede igual que en medio ácido del numeral **6 a 9**.

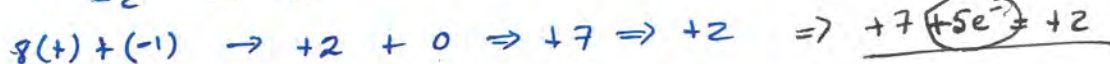
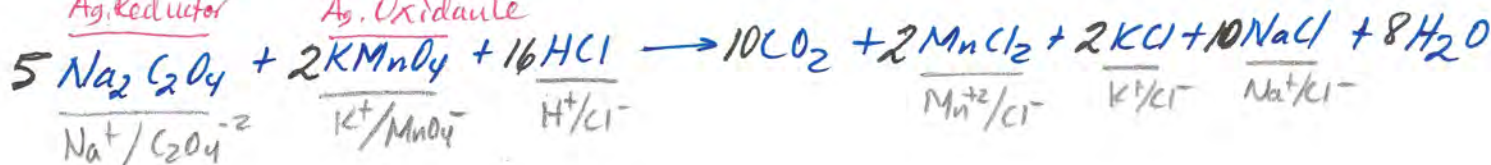
Agente Reductor

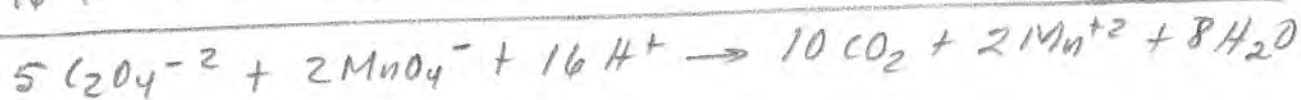
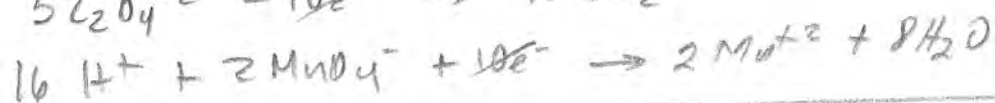
Agente Oxidante



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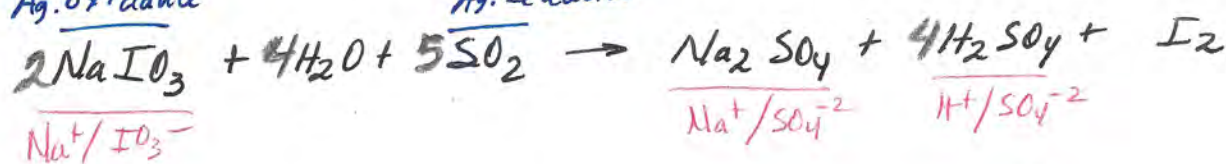
Ag. Oxidante



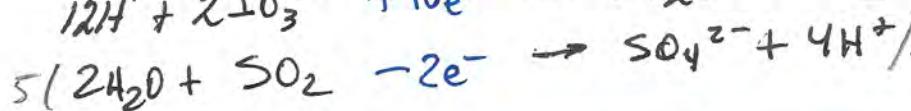
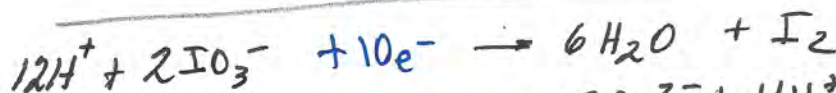


Ag. oxidante

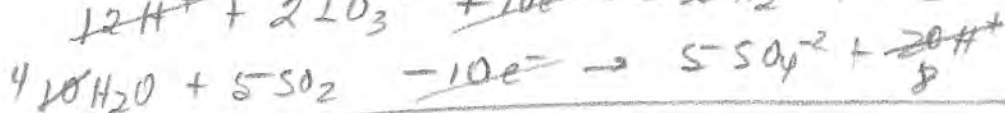
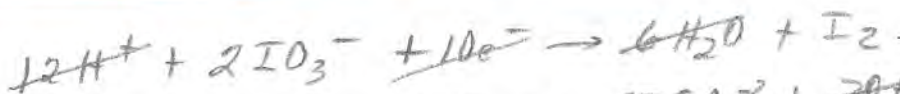
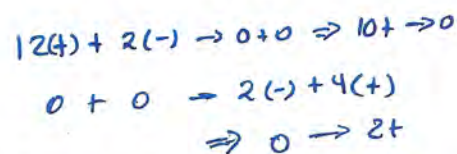
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Reducción

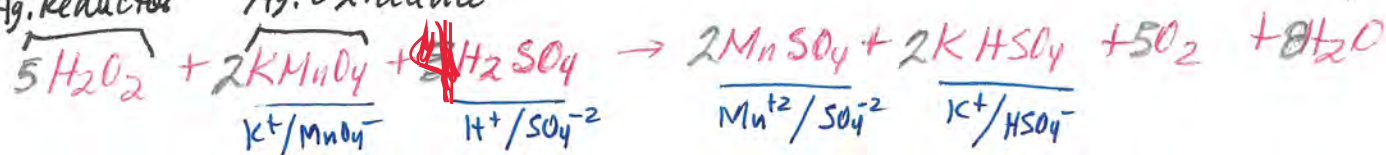


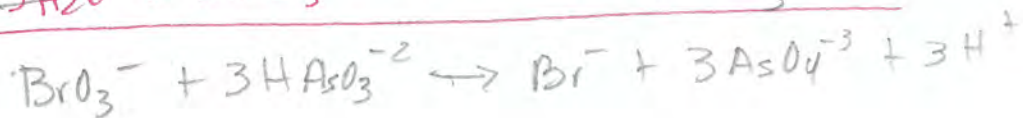
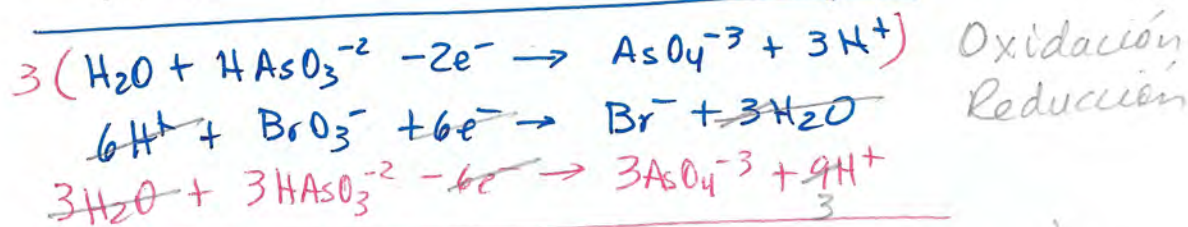
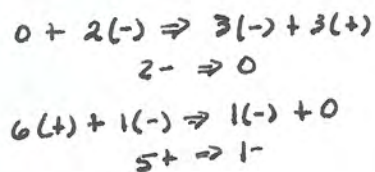
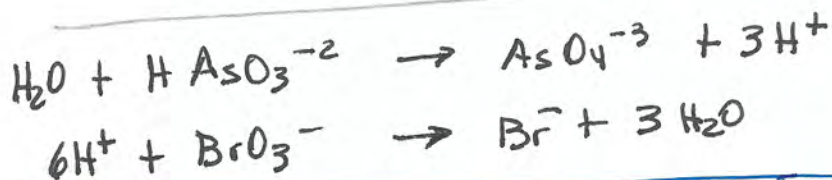
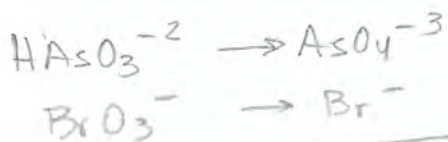
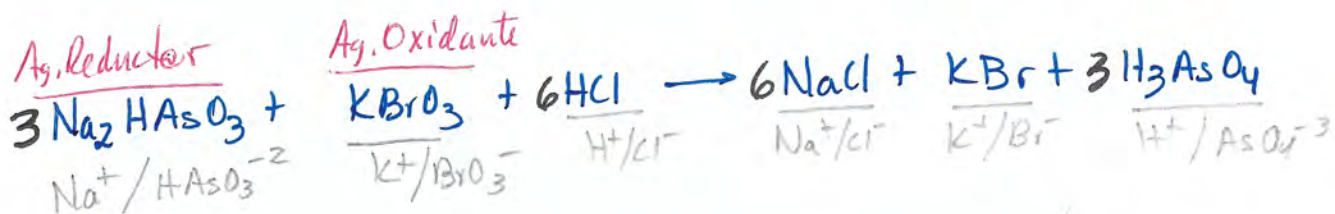
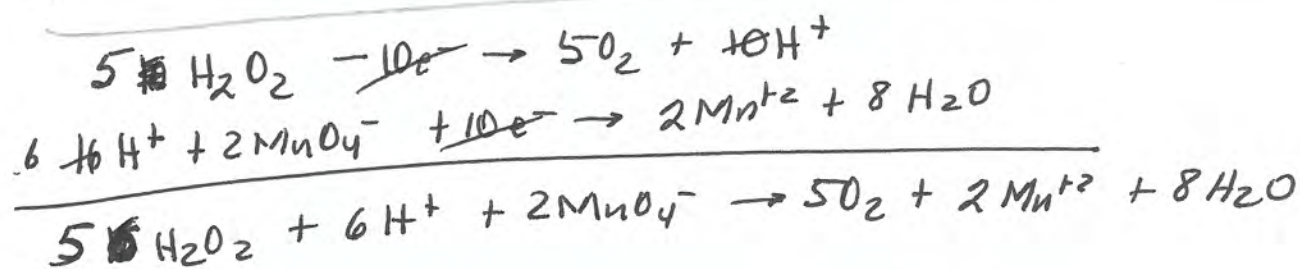
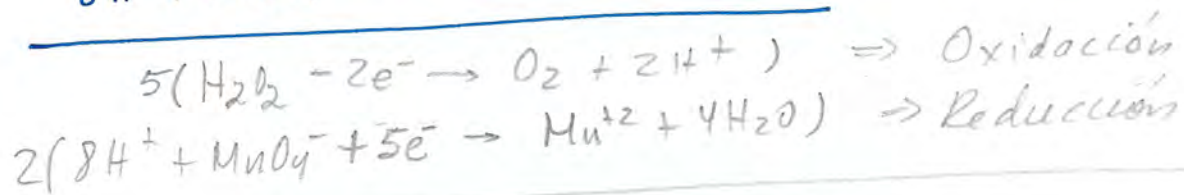
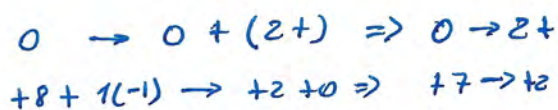
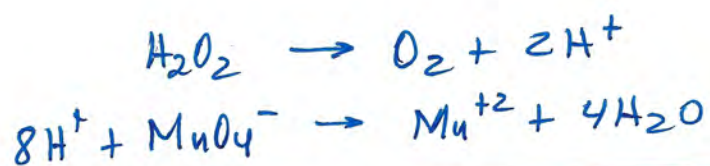
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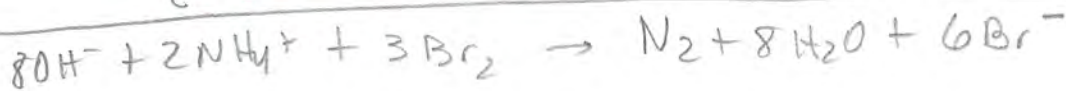
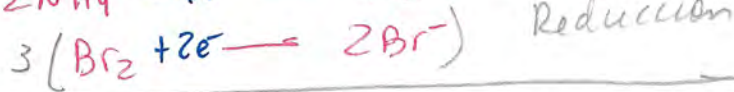
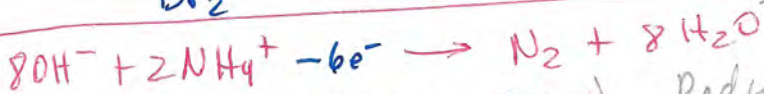
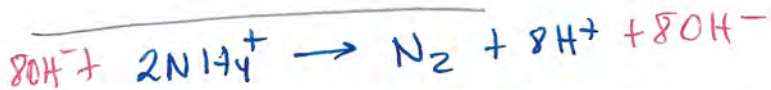
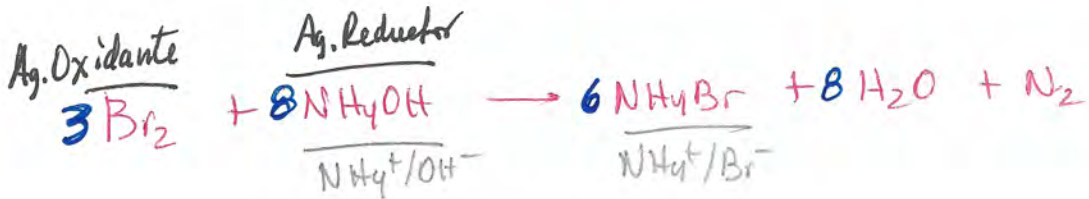
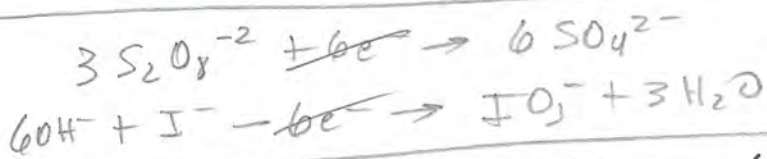
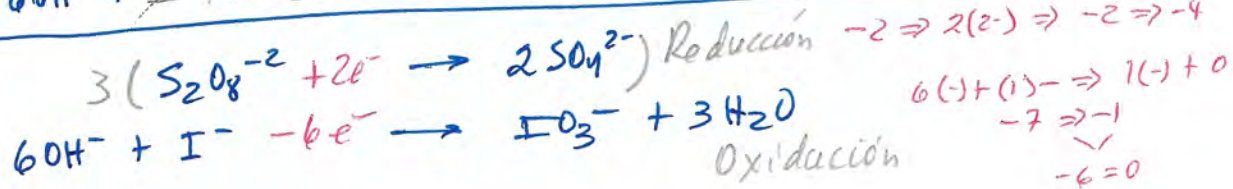
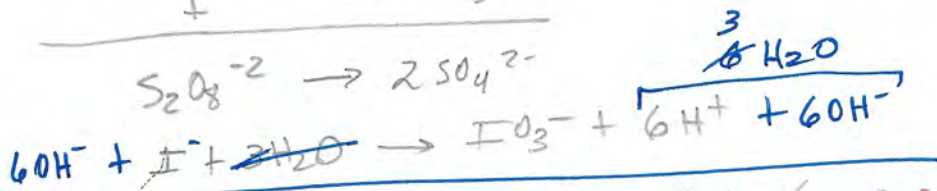
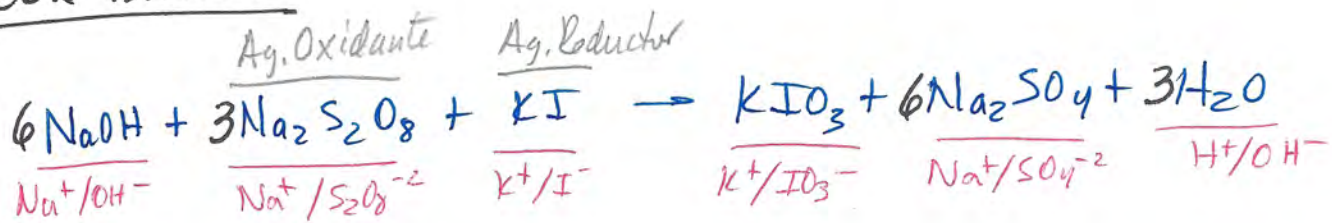
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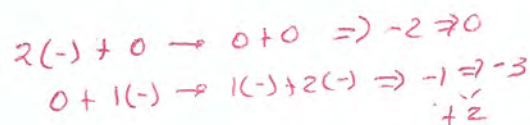
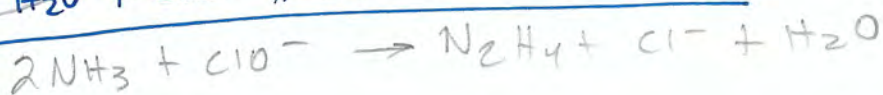
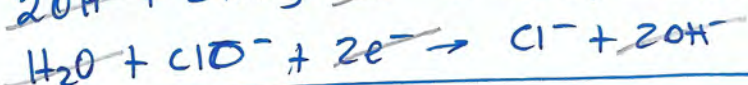
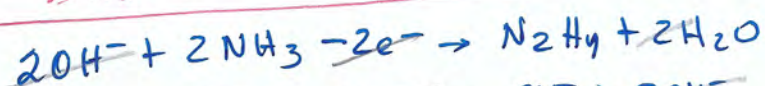
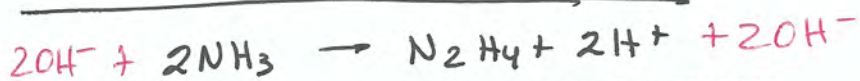
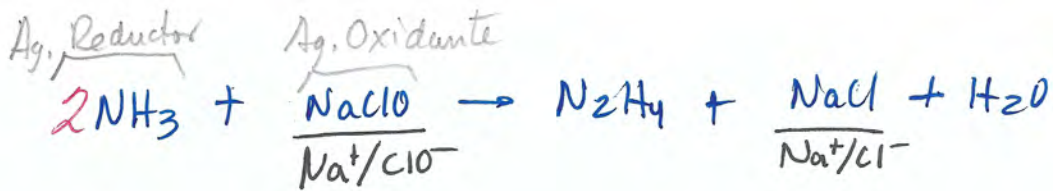
Ag. Oxidante



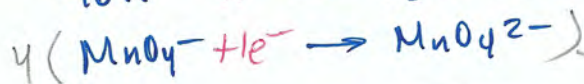
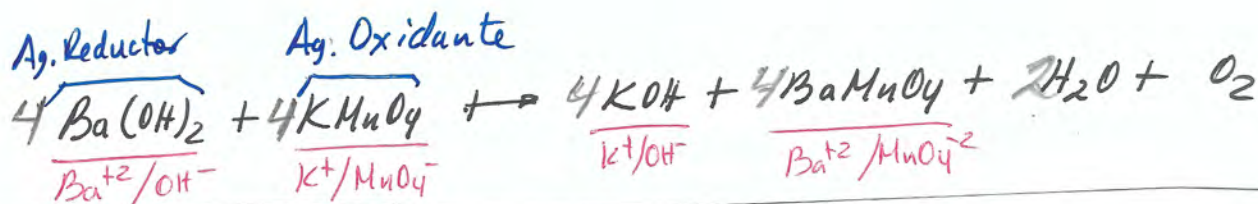


MEDIO BÁSICO

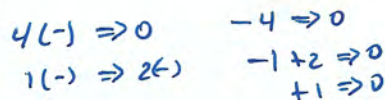




Oxidación
Reducción



Oxidación



Reducción

