Reviews:

Hoffmann, N. Chem. Rev. 2008, 108, 1052-1103.

Lee-Ruff, E.; Mladenova, G. Chem. Rev. 2003, 103, 1449-1484.

Crimmins, M. T. Chem. Rev. 1988, 88, 1453-1473.

Bach, T. Synthesis 1998, 683-703.

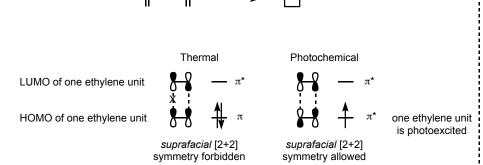
Challenges:

· Cyclobutanes are highly strained:

Wiberg, K. B. Angew. Chem. Int. Ed. 1986, 25, 312-322.

Synthetic Methods For the Construction of Cyclobutanes:

- [2+2] Cycloadditions:
- Suprafacial [2+2] cycloadditions are photochemically allowed but thermally forbidden.
- · Frontier Molecular Orbital Analysis:



- Enones are generally used as they are more easily photoexcited than isolated olefins.
- Photoexcited enones react via the T1 (triplet excited state).
- · Quantum efficiencies are higher in cyclic systems due to rapid intersystem crossing.
- Acyclic and macrocyclic enones are typically not suitable for [2+2] photocycloaddition because
 upon photoexcitation, they undergo cis/trans isomerization.

· Intramolecular [2+2] Cycloadditions:

• Intramolecular [2+2] cycloadducts are readily formed. Tethers are typically 2 to 4 atom:

O CH₃
hv
hexanes
$$98\%$$
, $dr = 92:8$
 H_3C
 OBz
 $Minimize non-bonded interactions$
 $Minimize non-bonded interactions$

$$\begin{array}{c} \text{H}_{3}\text{C} \\ \text{H}_{3}\text{C} \\ \text{E}\text{H}_{3}\text{C} \\ \text{E}\text{E}\text{H}_{3}\text{C} \\ \text{E}\text{H}_{3}\text{C} \\ \text{E}\text{H}_{3}\text{C} \\ \text{E}\text{H}_{$$

Birch, A. M.; Pattenden, G. J. Chem. Soc., Chem. Commun. 1980, 1195–1197. Birch, A. M.; Pattenden, G. J. Chem. Soc., Perkin Trans. 1. 1983, 1913–1917.

· Intermolecular [2+2] Cycloadditions

• Regioselectivity issues: head-to-tail (HT) vs. head-to-head (HH)

· Generally:

HT-isomer favored when R = donor (electron donating)

HH-isomer favored when R = acceptor (electron withdrawing)

An Example of Regioselectivity Governed by Electronics

$$H_3C$$
 H_3C
 H_3C

· Both HT and HH products were obtained as a mixture of two diastereomers.

Suishu, T.; Shimo, T.; Somekawa, K. Tetrahedron 1997, 53, 3545-3556.

- · Stereoselectivity in Intermolecular [2+2] Cycloadditions
- The least hindered transition state usually dominates. There is no "endo" effect, as in the Diels-Alder Reaction:

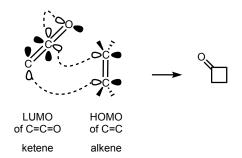
Wilson, S. R.; Phillips, L. R.; Pelister, Y.; Huffman, J. C. J. Am. Chem. Soc. 1979, 101, 7373-7379.

• When the double bond is not constrained within a rigid ring system, the intermediate triplet state diradical can lead to *cis/trans* isomerization of the double bond:

Corey, E. J.; Bass, J. D.; LeMahieu, R.; Mitra, R. B. J. Am. Chem. Soc. 1964, 86, 5570-5583.

· Ketene [2+2] Cycloadditions

- Antarafacial [2+2] cycloadditions are thermally allowed but geometrically disfavored for most substrates.
- Ketenes are linear and sterically less encumbered, making them good substrates for thermal [2+2] cycloadditions.
- Frontier Molecular Orbital Analysis:



· Ketenes can be generated in situ from acid chlorides:

$$CH_3$$
 CH_3 CH_3

Mori, K.; Miyake, M. Tetrahedron 1987, 43, 2229-2239.

Krepski, L. R.; Hassner, A. J. Org. Chem. 1978, 43, 2879-2882.

Danica Rankic, Fan Liu

- Examples of [2+2] Cycloaddition in Synthesis
- · Synthesis of pentacycloanammoxic acid, the principal lipid component of the cell membrane of the anaerobic microbe Candidatus Brocadia anammoxidans:

Mascitti, V.; Corey, E. J. J. Am. Chem. Soc. 2006, 128, 3118-3119.

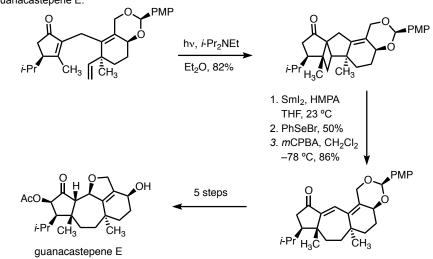
pentacycloanammoxic acid

 A [2+2] cycloaddition/anionic oxy-Cope/electrocyclic ring opening cascade provided a key intermediate to periplanone B. Note that both diastereomers from the [2+2] cycloaddition step can be converted to the final product:

O CH₃ hv, Et₂O O H CH₃
$$\frac{1. \text{ Et}_2\text{O}}{72\%, dr = 2:1}$$
 $\frac{1. \text{ Et}_2\text{O}, \text{MgBr}}{\text{H} \text{ CH}_3}$ $\frac{-78 \text{ °C}, 63\%}{2. \text{ KH}, 18\text{-C-6}}$ $\frac{1. \text{ Et}_2\text{O}, \text{MgBr}}{78\%}$ $\frac{1. \text{ Et}_2\text{O}, \text{MgBr}}{78\%}$

Schreiber, S. L.; Santini, C. J. Am. Chem. Soc. 1984, 106, 4038-4039.

A [2+2] Cycloaddition followed by radical fragmentation provided a key intermediate to guanacastepene E:



Shipe, W. D.; Sorensen, E. J. J. Am. Chem. Soc. 2006, 128, 7025-7035.

· A late-stage coupling reaction provided access to biyouyanagin A:

$$H_3C$$
 CH_3
 H_3C
 CH_3
 H_3C
 CH_3
 H_3C
 CH_3
 H_3C
 CH_3
 H_3C
 CH_3
 H_3C
 H_3C

Nicolaou, K. C.; Sarlah, D.; Shaw, D. M. Angew. Chem. Int. Ed. 2007, 46, 4708-4711.

· A [2+2] photocycloaddition followed by acid-catalyzed rearrangement provided a key intermediate to steviol:

$$H_3C$$
 H_3C
 H_3C

Cherney, E. C.; Green, J. C.; Baran, P. S. Angew. Chem. Int. Ed. 2013, 52, 9019-9022.

Fan Liu

• The synthesis of longifolene was accomplished via a de Mayo reaction — a [2+2] cycloaddition followed by retro-Aldol fragmentation:

OCbz
$$\frac{\text{hv, C}_6\text{H}_{12}}{15-30\text{ °C, 83\%}}$$

$$dr = 2:3$$

$$\frac{\text{H}_3\text{C}}{\text{CH}_3}$$

$$\frac{\text{H}_3\text{C}}{\text{Longifolene}}$$

$$\frac{\text{H}_3\text{C}}{\text{Longifolene}}$$

$$\frac{\text{H}_2\text{, Pd/C}}{\text{AcOH, 25 °C}}$$

$$\frac{\text{H}_2\text{, Pd/C}}{\text{AcOH, 25 °C}}$$

$$\frac{\text{H}_3\text{C} - \text{PPh}_3}{\text{Br}}$$

$$\frac{\text{NaOt-amyl}}{\text{toluene, 25 °C}}$$

$$\frac{\text{NaOt-amyl}}{\text{88\%}}$$

Oppolzer, W.; Godel, T. J. Am. Chem. Soc. 1978, 100, 2583-2584.

• Synthesis of clovene via a ketene [2+2] thermal cycloaddition:

Funk, R. L.; Novak, P. M.; Abelman, M. M. Tetrahedron Lett. 1988, 29, 1493-1496.

Bamford-Stevens-Shapiro Reaction

In the synthesis of ginkgolide B, three contiguous stereogenic centers were established using an intramolecular ketene [2+2] cycloaddition reaction:

Corey, E. J.; Kang, M.-C.; Desai, M. C.; Ghosh, A. K.; Houpis, I. N. *J. Am. Chem. Soc.* **1988**, *110*, 649–651.

 Synthesis of retigeranic acid via a ketene [2+2] cycloaddition followed by cyclobutane ring expansion:

Corey, E. J.; Desaid, M.; Engler, T. A. J. Am. Chem. Soc. 1985, 107, 4339-4341.

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- · Formal [2+2] Cycloadditions: Michael-Aldol Mechanism
- Silyl enol ethers and α,β-unsaturated compounds form cyclobutanes in a stepwise, Lewis acidcatalyzed process.
- In the following example, the bulky tris(trimethylsilyl)silyl (TTMSS) protecting group is required to stabilize the cationic intermediate. This and the use of a bulky aluminum catalyst with the triflimide counterion suppress silyl transfer:

TTMSSO OPh
$$CH_2Cl_2$$
, -40 °C $R2W$, $dr = 10:1$

Boxer, M. B.; Yamamoto, H. Org. Lett. 2008, 7, 3127-3129.

- · Auxiliary-Controlled Stereoselective [2+2] Cycloadditions
- One of the earliest reports involved the use of (–)-phenylmenthol as a chiral auxiliary to achieve a diastereoselective [2+2] cycloaddition:

Greene, A. E.; Charbonnier, F. Tetrahedron Lett. 1985, 26, 5525-5528.

 Upon activation by triflic anhydride, C₂-symmetric chiral pyrrolidine amides form keteniminium salts, which undergo thermal [2+2] cycloadditions with excellent stereoselectivities:

Chen, L.; Ghosez, L. Tetrahedron Lett. 1990, 31, 4467-4470.

• (–)-Phenylmenthol was also found to be an effective chiral auxiliary in formal [2+2] cycloadditions of silyl enol ethers and acrylates:

Takasu, K.; Nagao, S.; Ueno, M.; Ihara, M. *Tetrahedron* **2004**, *60*, 2071–2078.

Takasu, K.; Ueno, M.; Inanaga, K.; Ihara, M. J. Org. Chem. 2004, 69, 517–521.

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93%, 96% ee dr > 20:1

- · Enantioselective [2+2] Cycloadditions with Lewis Acid Catalysts
- In early reports, taddolates were found to catalyze formal [2+2] cycloadditions:

Hayashi, Y.; Narasaka, K. *Chem. Lett.* **1989**, 793–796. Hayashi, Y.; Niihata, S.; Narasaka, K. *Chem. Lett.* **1990**, 2091–2094.

• An aluminum bromide oxazaborolidine complex was found to be a highly efficient catalyst:

Canales, E.; Corey, E. J. J. Am. Chem. Soc. 2007, 129, 12686-12687.

 The reaction is proposed to occur by an asynchronous process. The structures shown below were presented to rationalize the stereochemical outcome:

· Transformations of the cyclobutane products:

TBSO OCH₂CF₃

$$\begin{array}{c}
1. \text{ CH}_3\text{ONHCH}_3 \cdot \text{HCI} \\
\text{CH}_3\text{MgBr} \\
\text{THF, } -30 \, ^{\circ}\text{C}
\end{array}$$

$$\begin{array}{c}
\text{CH}_3\text{MgBr}, \text{ THF} \\
-30 \rightarrow 0 \, ^{\circ}\text{C}, 85\%
\end{array}$$

$$\begin{array}{c}
\text{TBSO} \\
\text{CH}_3
\end{array}$$

$$\begin{array}{c}
\text{CH}_3
\end{array}$$

$$\begin{array}{c}
\text{TBAF, THF} \\
0 \rightarrow 23 \, ^{\circ}\text{C}
\end{array}$$

$$\begin{array}{c}
\text{TBAF, THF} \\
0 \rightarrow 23 \, ^{\circ}\text{C}
\end{array}$$

$$\begin{array}{c}
\text{CH}_3
\end{array}$$

Canales, E.; Corey, E. J. J. Am. Chem. Soc. 2007, 129, 12686-12687.

· Photoredox Catalysis

 Photoredox catalysts absorb light in the visible region and provide a pathway for cyclization mediated by electron transfer rather than direct photoexcitation of the substrate:

 MV^{2+} = methyl viologen (N,N'-dimethyl-4,4'-bypyridinium)

· Mechanism:

visible light strong reductant
$$H_3C \stackrel{+}{-}N \longrightarrow N \stackrel{+}{-}CH_3 (MV^{2+})$$
 $Ru(bpy)_3^{2+}$ strong $(MV-radical)$
 $Ru(bpy)_3^{2+}$ product strong oxidant MV^{2+}
 $P-CH_3OC_6H_4$
 $P-CH_3OC_6H_4$

- Upon photoexcitation, the photocatalyst $(Ru(bpy)_3(PF_6)_2)$ acts as a strong reductant, reducing MV by a single electron.
- The resulting Ru(III) species acts as a strong oxidant, which oxidizes the electron-rich styrene to produce a radical cation and regenerates photocatalyst.
- The resulting radical cation undergoes cyclization.

Ischay, M. A.; Lu, Z.; Yoon, T. P. J. Am. Chem. Soc. 2010, 132, 8572-8574.

Cis-olefins are isomerized during the course of the reaction:

• Limitation of the method: at least one of the styrenes must bear an electron-donating substituent at the *para* or *ortho* position. Aliphatic olefins are not suitable reaction partners.

• Photoredox catalysis can also be used for [2+2] cycloadditions of enones:

Ph Visible light
$$Ru(bpy)_3Cl_2$$
 (5 mol%)

LiBF₄, *i*-Pr₂NEt
MeCN, 89%, $dr > 10:1$

Ph HHH

Ph H

Visible light
 $Ru(bpy)_3Cl_2$ (5 mol%)

LiBF₄, *i*-Pr₂NEt
MeCN, 84%, $dr > 10:1$

Ischay, M. A.; Anzovino, M.E.; Du, J.; Yoon, T.P. *J. Am. Chem. Soc.* **2008**, *130*, 12886–12887. Du, J.; Yoon, T.P. *J. Am. Chem. Soc.* **2009**, *131*, 14604–14605.

· Mechanism:

- Lewis acid coordination decreases the reduction potential of the enone, facilitating single electron transfer.
- Limitation of the method: one coupling partner must be an aromatic ketone. Aliphatic ketones and esters do not under go cycloaddition because of their higher reduction potentials.

Ischay, M. A.; Anzovino, M.E.; Du, J.; Yoon, T.P. *J. Am. Chem. Soc.* **2008**, *130*, 12886–12887. Du, J.; Yoon, T.P. *J. Am. Chem. Soc.* **2009**, *131*, 14604–14605.

α,β-Unsaturated 2-imidazolyl ketones also undergo photochemically induced cycloadditions:

Visible light
$$Ru(bpy)_3Cl_2 (2.5 mol\%)$$

$$LiBF_4, i-Pr_2NEt$$

$$MeCN, 73\%, dr > 10:1$$

$$I-Pr$$

$$i-Pr$$

$$i-Pr$$

$$i-Pr$$

- · An intramolecular variant was also developed.
- The product can be converted to acids, esters, thioesters and amides:

visible light
$$Ru(bpy)_3Cl_2$$
 (2.5 mol%)

LiBF₄, *i*-Pr₂NEt MeCN, 87%, *dr* >10:1

1. MeOTf, CH₂Cl₂
67%
2. BnNH₂, DBU
CH₂Cl₂, 98%

Tyson, E. L.; Farney, E. P.; Yoon, T. P. Org. Lett. 2012, 14, 1110-1113.

• Styrenes and alkenes undergo [2+2] cycloaddition in the presence of an Ir catalyst (III) and light. The catalyst functions as a photosensitizer.

(±)-cannabiorcicyclolic acid

EtO₂C
$$H_3$$
 visible light H_3 C CH_3 H_3 C CH_3

Lu, Z.; Yoon, T. P. Angew. Chem. Int. Ed. 2012, 51, 10329-10332

- · Other Methods for Cyclobutane Synthesis
- · Brook Rearrangement of 1,4-Dicarbonyls
- Treatment of keto acylsilanes with organolithium reagents produces highly functionalized cyclobutanes favoring cis-stereochemistry between newly formed alcohols

TBS
$$\longrightarrow$$
 i-Pr \longrightarrow PhLi, THF \longrightarrow TBSO, OH \longrightarrow HO, OTBS \longrightarrow 67% \longrightarrow 10%

Mechanism:

Takeda, K.; Haraguchi, H.; Okamoto, Y. Org. Lett. 2003, 5, 3705-3707.

- · Ring Expansion of Cyclopropanes via Pinacol-Type Rearrangements
- · Hydroxycyclopropyl carbinols can be ring-expanded by treatment with protic or Lewis acids.
- Either cis- or trans-substituted cyclobutanones could be produced from a single diastereomer of a an α-hydroxycyclopropyl carbinol:

Hussain, M. M.; Li, H.; Hussain, N.; Urena, M.; Carroll, P. J.; Walsh, P. J. *J. Am. Chem. Soc.* **2009**, *131*, 6516–6524.

- · Gold(I)-Catalyzed Ring Expansion of Cyclopropanes
- Alkynyl cyclopropanols can undergo ring expansion upon treatment with catalytic Au(I):

• The mechanism is proposed to involve a stereospecific 1,2-alkyl shift:

Markham, J. P.; Staben, S. T.; Toste, D. F. J. Am. Chem. Soc. 2005, 127, 9708-9709.

- · Cu-catalyzed 1,4-Ring closure
- Homoallylic sulfonates can undergo borylation/cyclization using a Cu^I catalyst:

$$\begin{array}{c} \text{CuCl (5 mol\%)} \\ \text{dppp (5 mol\%)} \\ \text{B}_{2}\text{pin}_{2}, \text{ KO} \text{t-Bu} \\ \text{THF, 23 °C} \end{array} \begin{array}{c} \text{H}_{3}\text{C} \\ \text{Ph} \\ \text{O} \\ \text{CH}_{3} \\ \text{CH}_{3} \\ \text{CH}_{3} \\ \text{CH}_{3} \\ \text{CH}_{3} \\ \text{B}_{2}\text{pin}_{2} \\ \text{B}_{2}\text{pin}_{2} \\ \end{array} \begin{array}{c} \text{CH}_{3} \\ \text{H}_{3}\text{C} \\ \text{CH}_{3} \\ \text{H}_{3}\text{C} \\ \text{CH}_{3} \\ \text{B}_{2}\text{pin}_{2} \\ \end{array}$$

$$(H_3C)_2 PhSi$$
OMs
$$\frac{CuCl (5 \text{ mol}\%)}{dppp (5 \text{ mol}\%)}$$

$$\frac{(H_3C)_2 PhSi}{B_2 pin_2, KOt\text{-Bu}}$$

$$\frac{(H_3C)_2 PhSi}{CH_3}$$

$$\frac{(H_3C)_2 PhSi}{CH_3}$$

$$\frac{(H_3C)_2 PhSi}{CH_3}$$

· Mechanism:

$$\begin{array}{c} R_{2} \\ \text{pinB-O}t\text{-Bu} \\ \text{pinB-Bpin} \\ \text{CuCl} \\ \text{OMs} \\ \text{CuCl} \\ \text{OMs} \\ \text{CuCl} \\ \text{Important pinB-Bpin} \\ \text{OMs} \\ \text{OMs} \\ \text{CuCl} \\ \text{Important pinB-D}t\text{-Bu} \\ \text{OMs} \\ \text{OMs} \\ \text{CuCl} \\ \text{Important pinB-D}t\text{-Bu} \\ \text{OMs} \\ \text{Important pinB-D}t\text{-Bu} \\ \text{I$$

• Limitation of the method: only aryl- or silyl-substituted olefins react; alkyl olefins do not undergo borocupration.

Ito, H.; Toyoda, T.; Sawamura, M. J. Am. Chem. Soc. 2010, 132, 5990-5992.

The products can be derivatized:

Bn(H₃C)₂Si BPin
$$= \frac{1. \text{ H}_2\text{O}_2, \text{NaOH}}{1. \text{ H}_5\text{O} \circ \text{C}}$$
Bn(H₃C)₂Si OBz $= \frac{1. \text{ H}_2\text{O}_2, \text{NaOH}}{1. \text{ H}_5\text{O} \circ \text{C}}$
Bn(H₃C)₂Si OBz $= \frac{1. \text{ H}_2\text{O}_2, \text{NaOH}}{1. \text{ H}_5\text{O} \circ \text{C}}$
C2. PhCOCI, C₅H₅N CH₂Cl₂, 0 °C $= \frac{1. \text{ TBAF, THF}}{23 \circ \text{C}}$
C3. PhCOCI, C₅H₅N CH₂Cl₂, 0 °C $= \frac{1. \text{ TBAF, THF}}{23 \circ \text{C}}$
C4. H₂O₂, KHCO₃ CH₃OH, 23 °C $= \frac{1. \text{ TBAF, THF}}{23 \circ \text{C}}$
C5. H₂O₂, KHCO₃ CH₃OH, 23 °C $= \frac{1. \text{ TBAF, THF}}{23 \circ \text{C}}$
C1. TBAF, THF C3 °C $= \frac{1. \text{ TBAF, THF}}{23 \circ \text{C}}$
C1. TBAF, THF C3 °C $= \frac{1. \text{ TBAF, THF}}{23 \circ \text{C}}$
C2. H₂O₂, KHCO₃ CH₃OH, 23 °C $= \frac{1. \text{ TBAF, THF}}{23 \circ \text{C}}$
C3. H₂O₂, KHCO₃ CH₃OH, 23 °C $= \frac{1. \text{ TBAF, THF}}{23 \circ \text{C}}$
C4. H₂O₂, KHCO₃ CH₃OH, 23 °C $= \frac{1. \text{ TBAF, THF}}{23 \circ \text{C}}$
C5. H₂O₂, KHCO₃ CH₃OH, 23 °C $= \frac{1. \text{ TBAF, THF}}{23 \circ \text{C}}$
C5. H₂O₂, KHCO₃ CH₃OH, 23 °C $= \frac{1. \text{ TBAF, THF}}{23 \circ \text{C}}$
C6. H₂O₂, KHCO₃ CH₃OH, 23 °C $= \frac{1. \text{ TBAF, THF}}{23 \circ \text{C}}$
C7. H₂O₂, KHCO₃ CH₃OH, 23 °C $= \frac{1. \text{ TBAF, THF}}{23 \circ \text{C}}$
C8. H₂O₂, KHCO₃ CH₃OH, 23 °C $= \frac{1. \text{ TBAF, THF}}{23 \circ \text{C}}$
C9. H₂O₂, KHCO₃ CH₃OH, 23 °C $= \frac{1. \text{ TBAF, THF}}{23 \circ \text{C}}$
C9. H₂O₂, KHCO₃ CH₃OH, 23 °C $= \frac{1. \text{ TBAF, THF}}{23 \circ \text{C}}$
C9. H₂O₂, KHCO₃ CH₃OH, 23 °C $= \frac{1. \text{ TBAF, THF}}{23 \circ \text{C}}$
C9. H₂O₂, KHCO₃ CH₃OH, 23 °C $= \frac{1. \text{ TBAF, THF}}{23 \circ \text{C}}$
C9. H₂O₂, KHCO₃ CH₃OH, 23 °C $= \frac{1. \text{ TBAF, THF}}{23 \circ \text{C}}$
C9. H₂O₂, KHCO₃ CH₃OH, 23 °C $= \frac{1. \text{ TBAF, THF}}{23 \circ \text{C}}$
C9. H₂O₂, KHCO₃ CH₃OH, 23 °C $= \frac{1. \text{ TBAF, THF}}{23 \circ \text{C}}$
C9. H₂O₂, KHCO₃ CH₃OH, 23 °C $= \frac{1. \text{ TBAF, THF}}{23 \circ \text{C}}$
C9. H₂O₂ CH₃OH, 23 °C $= \frac{1. \text{ TBAF, THF}}{23 \circ \text{C}}$
C9. H₂O₂ CH₃OH, 23 °C $= \frac{1. \text{ TBAF, THF}}{23 \circ \text{C}}$
C9. H₂O₂ CH₃OH, 23 °C $= \frac{1. \text{ TBAF, THF}}{23 \circ \text{C}}$
C9. H₂O₂ CH₃OH, 23 °C $= \frac{1. \text{ TBAF, THF}}{23 \circ \text{C}}$
C9. H₃OH, 2

• Note that in each case the stereochemistry of the starting material is preserved in the product.

Markham, J. P.; Staben, S. T.; Toste, D. F. J. Am. Chem. Soc. 2005, 127, 9708–9709.