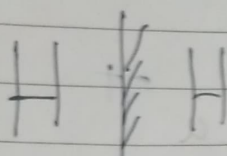


2 Optical Isomerism:-

For a compound to be optically active:-

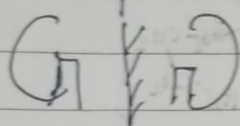
- (i) It should not be superimposable on its mirror image
- (ii) It should not contain any element of symmetry

Case 1:-



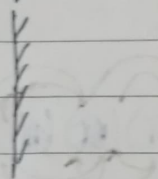
Superimposable mirror image  
Optically inactive


Case 2:-

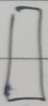


Non-superimposable mirror image  
Optically active

Case 3:-

Polarized Light:-

  
Unpolarized light

  
Nicol prism  
(calcite  $\text{CaCO}_3$ )

To check if compound is optically active/not

→ Polarimeter →

Plane  $P_0$

Polarized Light (PPL)

Rotates in Right direction  
Dextro (d) (clockwise)

Rotates in Left direction  
Levo (l) (anticlockwise)

#

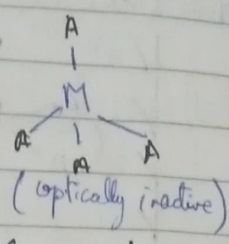
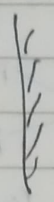
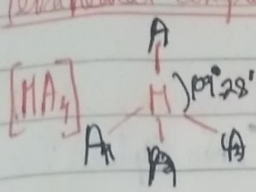
Plane Polarized Light (PPL):-

- 1) If the compound rotates PPL in clockwise/Right direction  
Then it is dextro [denoted by d/+]
- 2) If the compound rotates PPL in anti-clockwise/Left direction  
Then it is levo [denoted by l/-]

CN=4

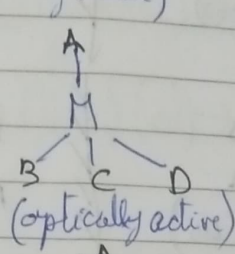
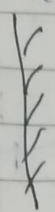
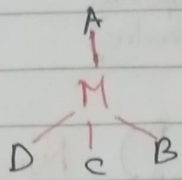
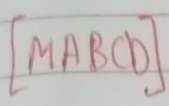
# (A) Tetrahedral complexes

Case 1:-



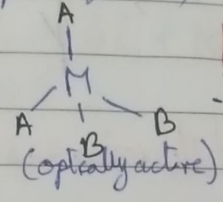
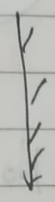
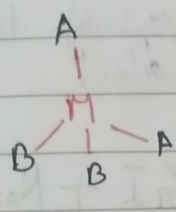
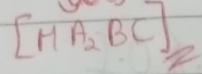
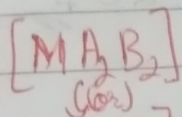
Superimposable Mirror Image  
∴ No Optical Isomerism

Case 2:-



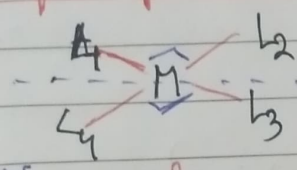
Non-Superimposable Mirror Image  
∴ Has 2 Optical Isomers  
(OI)

Case 3:-



Non-Superimposable Mirror Image  
∴ Has 2 O.I

# (B) Square planar complexes



Plane of symmetry

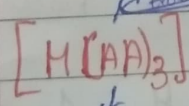
∴  $\angle$  bet<sup>n</sup>  $\begin{matrix} L_1, L_2 \\ L_2, L_3 \\ L_3, L_4 \\ L_4, L_1 \end{matrix}$   $\begin{cases} 90^\circ \\ 180^\circ \end{cases}$   $\begin{matrix} L_1, L_3 \\ L_2, L_4 \end{matrix}$

Hence, square planar complexes never show Optical Isomerism  
(due to presence of plane of symmetry  $\rightarrow$  P.O.S)

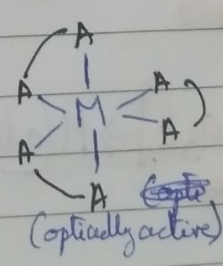
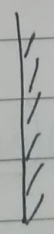
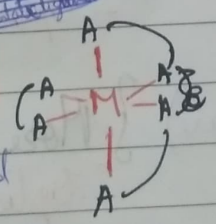
CN=6

# Octahedral complexes

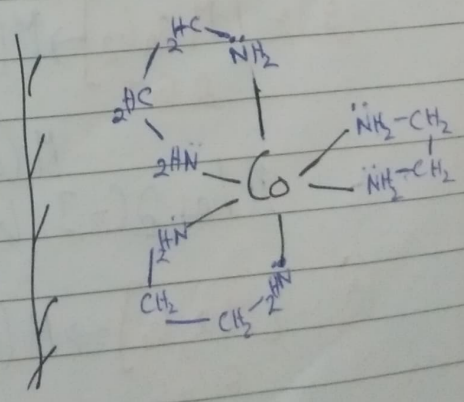
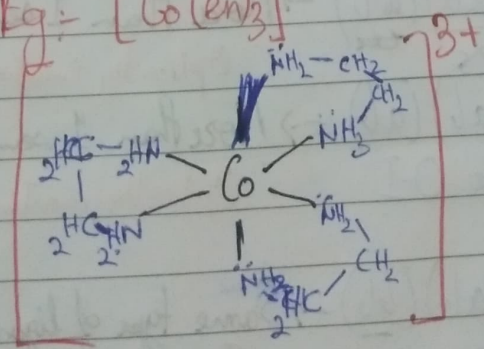
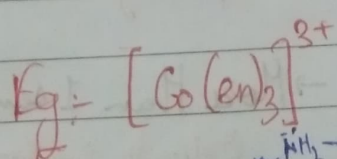
Case 1:-



bidentate ligand

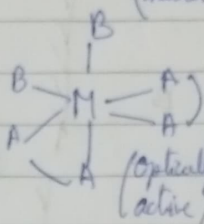
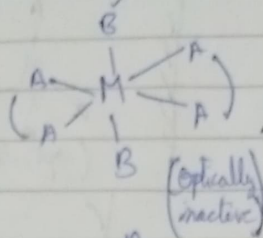
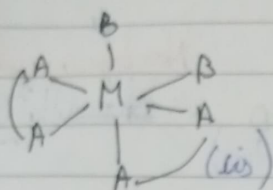
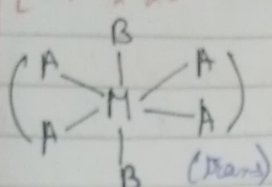


Non-Superimposable M.I  
∴ Has 2 O.I





Case 2:  $[M(AA)_2B_2] \rightarrow 2 \text{ G.I. (cis \& trans)}$

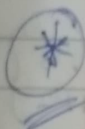


Superimposable M.I.  
∴ No Optical Isomer

Non-Superimposable M.I.  
Has 2 O.I.



Total no. of stereoisomers = No. of G.I. + No. of O.I.  
= 2 + 2 = 4



Trick to check Optical Isomerism:-

If the pairs have same type of ligand [eg. (aa) type]

more than one same pair [eg. (ab)(ab) type]

Then, the compound is **Optically inactive**

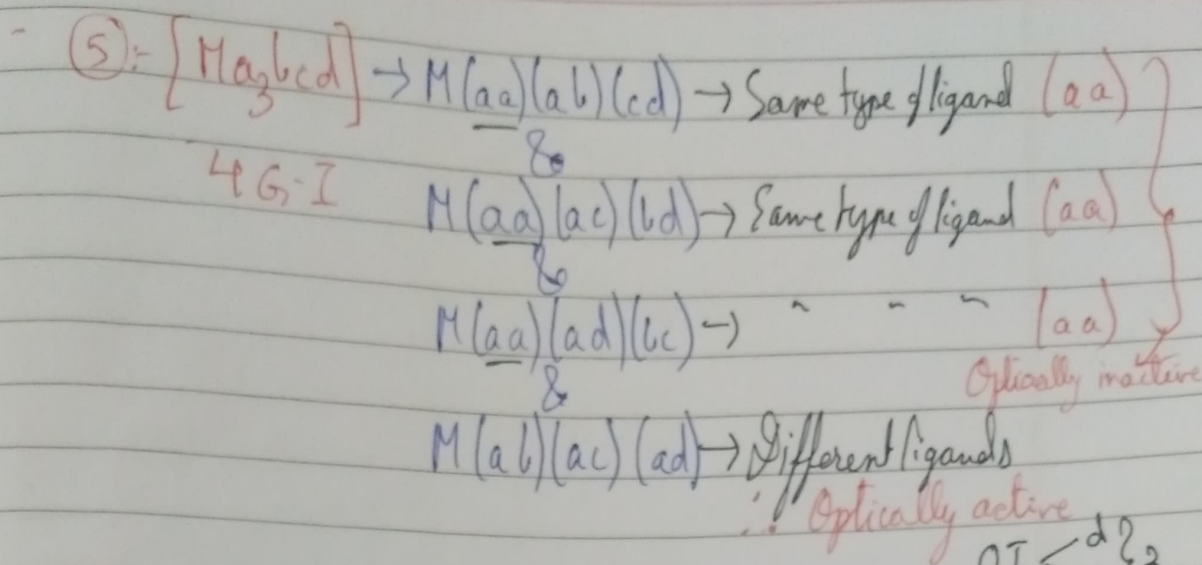
Cases:-

Ex ①:  $[Ma_3] \rightarrow M(\underline{aa})(\underline{aa})(\underline{aa})$  } More than 1 same pair  
Optically inactive

②:  $[Ma_2b] \rightarrow M(\underline{ab})(\underline{aa})(\underline{ab})$  } More than 1 same pair  
Optically inactive

③:  $[Ma_2b_2] \rightarrow M(\underline{aa})(\underline{aa})(\underline{bb})$  } More than 1 same pair  
Optically inactive  
2 G.I.  
8  
 $M(\underline{aa})(\underline{ab})(\underline{ab}) \rightarrow$  More than 1 same pair  
It has 2 G.I. but no O.I. Optically inactive

④:  $[Ma_2b_2] \rightarrow M(\underline{aa})(\underline{ab})(\underline{bb})$  } Same type of ligand [aa], [bb]  
Optically inactive  
2 G.I.  
8  
 $M(\underline{ab})(\underline{ab})(\underline{ab}) \rightarrow$  More than 1 same pair  
Optically inactive



$$\begin{aligned}
 \therefore \text{Total no. of stereoisomers} &= \text{No. of G.I} + \text{No. of O.I} \\
 &= 4 + 2 \\
 &= 6
 \end{aligned}$$

<u>Complex</u>	<u>No. of Optical Isomers</u>
$[MA_4B_2]$	0
$[MA_4BC]$	0
$[MA_3B_3]$	0
$[MA_3B_2C]$	0
$[MA_3BCD]$	2
$[MA_2B_2C_2]$	2
$[MA_2B_2CD]$	4
$[MA_2BCDE]$	12
⊕ $[MAB CDEF]$	30
$[M(AA)_2B_2]$	2
$[M(AA)_2BC]$	2
$[M(AB)_3]$	4
$[M(AA)_3]$	2