



36. Methanol is toxic because

- (a) methanol coagulates the protoplasm
- (b) methanol gets oxidised to methanal in liver which coagulates the protoplasm
- (c) methanol gets oxidised to acetic acid in liver which coagulates the protoplasm
- (d) methanol gets oxidised to CO in liver which coagulates the protoplasm



- 6. Which of the following is the general formula for saturated hydrocarbons?
 - (a) $C_n H_{2n+2}$
 - (b) $C_n H_{2n-2}$
 - (c) $C_n H_{2n+1}$
 - (d) $C_n H_{2n-1}$



5. The compound C7H7NO2 has

- (a) 17 atoms in a molecule of the compound
- (b) equal molecules of C and H by mass
- (c) twice the mass of oxygen atoms compared to nitrogen atoms
- (d) twice the mass of nitrogen atoms compared to hydrogen atoms



112. The number of saturated and unsaturated bonds in cyclohexane are:

- (a) 9 and 0 respectively.
- (b) 18 and 3 respectively.
- (e) 18 and 0 respectively.
- (d) 9 and 3 respectively.

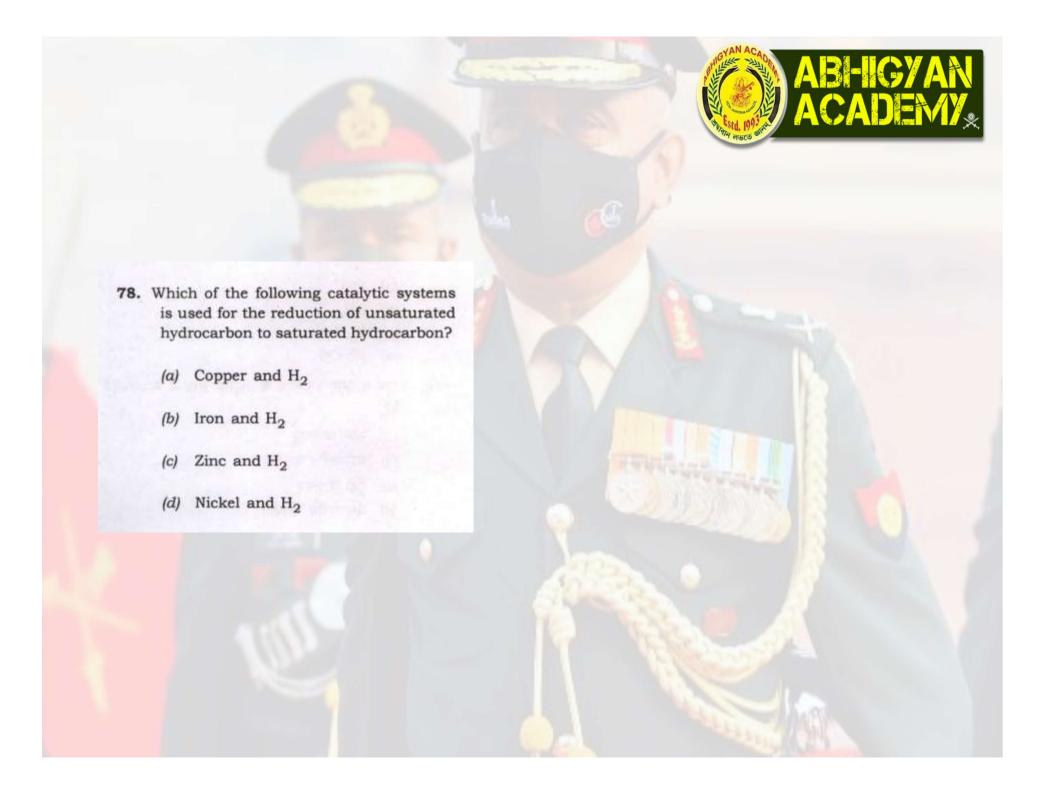


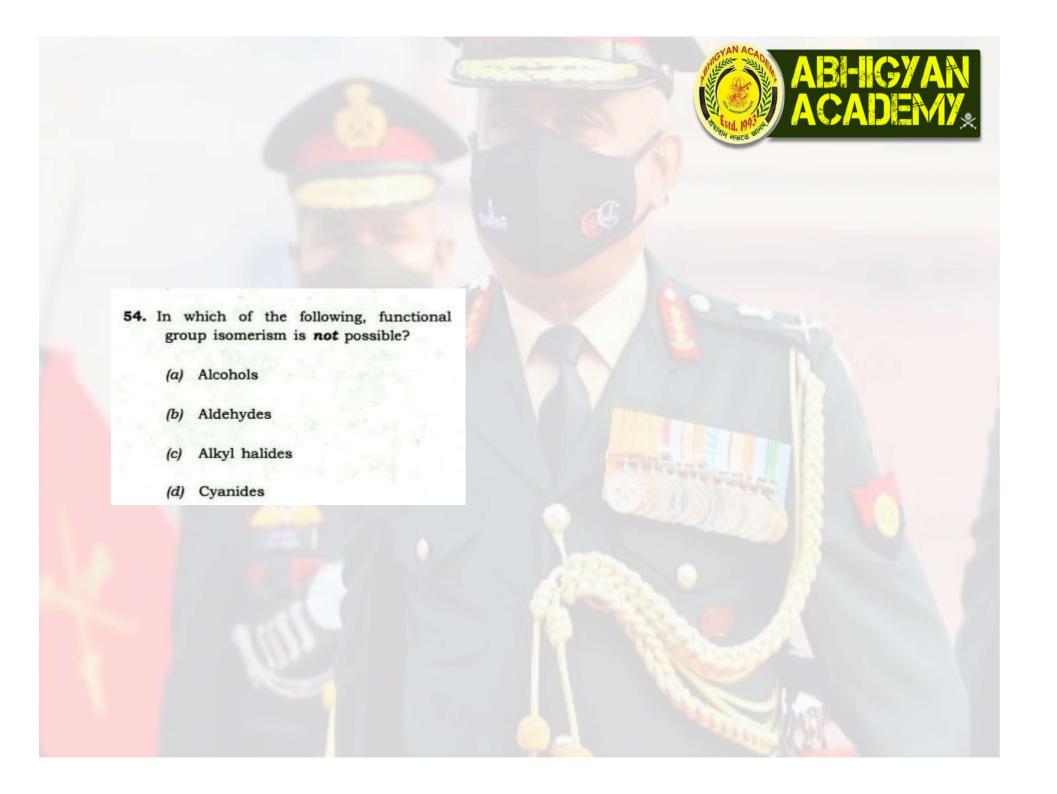
- 34. The correct order of octane number of butane, pentane, hexane and cyclohexane is
 - (a) butane > pentane > hexane > cyclohexane
 - (b) butane > pentane > cyclohexane > hexane
 - (c) butane > cyclohexane > pentane > hexane
 - (d) cyclohexane > butane > pentane > hexane



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113. In graphite, each carbon atom is bonded to three other carbon atoms

- (a) forming a three-dimensional structure
- (b) in the same plane giving a hexagonal array
- (c) in the same plane giving a square array
- in the same plane giving a pentagonal array



9. During white-washing of walls, slaked lime reacts slowly with carbon dioxide in air to form a thin layer of calcium carbonate on the walls. Which of the following reactions represents this correctly?

(a)
$$CaO(s) + CO_2(g) \rightarrow CaCO_3(s)$$

(b)
$$\operatorname{CaO}(l) + \operatorname{CO}_2(g) \rightarrow \operatorname{CaCO}_3(s)$$

(c)
$$\operatorname{Ca(OH)}_{2}(l) + \operatorname{CO}_{2}(l) \rightarrow \operatorname{CaCO}_{3}(s) + \operatorname{H}_{2}O(l)$$

(d)
$$Ca(OH)_2(aq) + CO_2(g) \rightarrow CaCO_3(s) + H_2O(l)$$



27. In which one of the following reactions, the maximum quantity of H₂ gas is produced by the decomposition of 1 g of compound by H₂O/O₂?

(a)
$$CH_4 + H_2O \rightarrow CO + 3H_2$$

(b)
$$CO + H_2O \rightarrow CO_2 + H_2$$

(c)
$$CH_4 + \frac{1}{2}O_2 \rightarrow CO + 2H_2$$

(d)
$$C_{12}H_{24} + 6O_2 \rightarrow 12CO + 12H_2$$

