. 1ST LAW: dE = dQ + dW

· ANOTHER WAY TO CHANGE THE ENERGY OF THE SYSTEM IS TO CHANGE THE NO OF PARTICLES

[FOR EXAMPLE, RECALL THE COUPARTHON THEOREM: = LEST PER PARTICLE PER DOF]

ie CHANGE THE No Ni OF VARIOUS CHEMICAE SPECIES.

· IN THIS CASE, CONSERVATION OF ENEMGY TAKES
THE FORM

dE = dQ + dW + & MidNi

CHANGE OF ENCULY DUE
TO CHANGE DIN: OF No°
OF PANTICLES OF
SPECIES i

[FOR A SYSTEM WITH A SINGLE CHEMICAL SPECIES: $dE = dQ + dW + \mu dNI$.]

(2

THE COEFFICIENT Mi IS CALCED THE CHEMICAL POTENTIAL OF SPECIES C'.

CRUDELY: Mi IS THE RATE OF ENERGY

CHANGE OF THE SYSTEM AS WE VARY Ni.

EQUATIONS OF STATE

1st LAW: de = dQ + dW + & MidNi

FOR QUASISTATIC PROCESSES:

tQ = TdS } UNIVERSAL & APPLY
TO ALL SYSTEMS

So, For A QUASISTATIC PROCESS, THE 1ST LAW IS:

dE = TdS - pdV + & MidNi

 $\Rightarrow |dS = \frac{1}{T}dE + \frac{1}{F}dV - \frac{1}{5}\frac{Mi}{T}dNi - (1)$

THIS SHOWS: THE CHANGE IN ENTROPY S IS

COMPLETELY DETERMINED BY THE CHANGES IN

E, V AND No. (THIS IS TRUE IN GENERAL,

ALL WE HAVE USED IS CONSERVATION OF ENERGY)

IN OTHER WORDS: ENTROPY 5 IS "NATURALLY"

A FUNCTION OF E, V, N;

$$S = S(E, V, \vec{N})$$

$$\vec{N} = (N_1, N_2, \dots)$$

· USING THE DEFINITION OF THE DIFFERENTIAL of WE SEE THAT

$$dS(E,V,\vec{N}) = \frac{\partial S}{\partial E} dE + \frac{\partial S}{\partial V} dV + \underbrace{\frac{\partial S}{\partial N_c}}_{c} dN_c$$

-(2,

$$\frac{\partial S(E, V, \vec{N})}{\partial E} = \frac{1}{T} - EQUATIONS OF$$

$$\frac{\partial S(E, V, \vec{N})}{\partial V} = P$$

$$\frac{\partial S(E, V, \vec{N})}{\partial N_i} = -\frac{M_i}{T}$$

$$\frac{\partial S(E, V, \vec{N})}{\partial N_i} = -\frac{M_i}{T}$$

· LATER WE WILL SEE HOW TO COMPUTE S.

THE EQUATIONS OF STATE (3) DETERMINES

T, P, Mi INI TERMS OF $S(E, V, \vec{N})$

• $S(\mathcal{E}, V, \vec{N})$ - DEPENDS ON THE PARTICULAR SYSTEM, & CAN BE DERIVED USING QUANTUM MECHANICS AND $S = k_0 \ln \Omega$ 4