



# Cambridge IGCSE™

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\*1182455718\*

**CHEMISTRY****0620/41**

Paper 4 Theory (Extended)

**May/June 2024****1 hour 15 minutes**

You must answer on the question paper.

No additional materials are needed.

**INSTRUCTIONS**

- Answer **all** questions.
- Use a black or dark blue pen. You may use an HB pencil for any diagrams or graphs.
- Write your name, centre number and candidate number in the boxes at the top of the page.
- Write your answer to each question in the space provided.
- Do **not** use an erasable pen or correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.
- You should show all your working and use appropriate units.

**INFORMATION**

- The total mark for this paper is 80.
- The number of marks for each question or part question is shown in brackets [ ].
- The Periodic Table is printed in the question paper.

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This document has **16** pages. Any blank pages are indicated.

1 Name the process that is used to:

- (a) convert sulfur dioxide into sulfur trioxide in the manufacture of sulfuric acid

..... [1]

- (b) obtain water from aqueous sodium chloride

..... [1]

- (c) extract aluminium from purified bauxite

..... [1]

- (d) separate petroleum into useful substances

..... [1]

- (e) produce ethanol from aqueous glucose

..... [1]

- (f) manufacture alkenes and hydrogen from large alkane molecules

..... [1]

- (g) separate a mixture of soluble coloured substances.

..... [1]

[Total: 7]

2 Complete Table 2.1.

**Table 2.1**

atom or ion	number of protons	number of electrons	number of neutrons
$^{37}_{17}\text{Cl}$		17	
$^{63}_{29}\text{Cu}^+$			34
	16	18	17

[5]

3 This question is about the elements sodium and fluorine and the compound sodium fluoride.

(a) Sodium reacts with fluorine to form sodium fluoride.

Write a symbol equation for this reaction.

..... [2]

(b) Some properties of sodium, fluorine and sodium fluoride are shown in Table 3.1.

**Table 3.1**

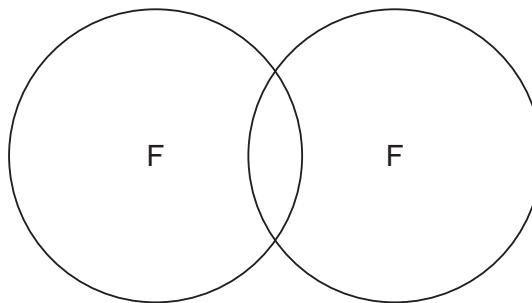
	melting point /°C	boiling point /°C	conduction of electricity when solid	conduction of electricity in aqueous solution
sodium	98	883	good	
fluorine	-220	-188	poor	
sodium fluoride	993	1695	poor	good

(i) Explain why sodium conducts electricity when it is a solid.

..... [1]

(ii) Complete the dot-and-cross diagram in Fig. 3.1 of a molecule of fluorine.

Show outer shell electrons only.



**Fig. 3.1**

[2]

(iii) Deduce the physical state of fluorine at -200 °C. Use the data in Table 3.1 to explain your answer.

physical state .....

explanation .....

..... [2]

- (iv) Explain in terms of structure and bonding why sodium fluoride has a much higher melting point than fluorine.

.....  
.....  
.....  
.....  
..... [3]

- (c) Dilute aqueous sodium fluoride undergoes electrolysis.

Hydrogen is produced at the cathode.

- (i) State what is meant by the term electrolysis.

.....  
.....  
..... [2]

- (ii) Write an ionic half-equation for the production of hydrogen at the cathode.

..... [2]

[Total: 14]

- 4 Hydrogen iodide thermally decomposes into iodine and hydrogen. The reaction is reversible.

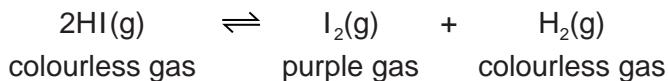
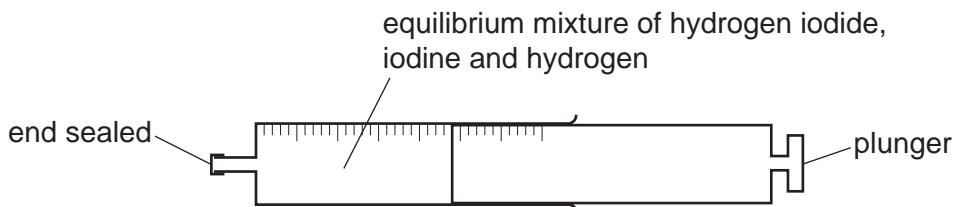


Fig. 4.1 shows a gas syringe containing a mixture of hydrogen iodide, iodine and hydrogen gases. The gas syringe is sealed and the mixture is heated to 300°C. The mixture of gases reaches equilibrium and is purple.



**Fig. 4.1**

- (a) State what is meant by the term equilibrium.

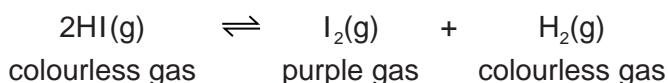
.....  
.....  
.....

[2]

- (b) The pressure of the mixture is increased. All other conditions stay the same. The position of equilibrium does **not** change.

The colour of the gaseous mixture turns darker purple.

The temperature remains constant.



- (i) Explain why the position of equilibrium does **not** change.

.....  
.....

[1]

- (ii) Suggest why the colour of the mixture of gases turns darker purple.

.....  
.....

[1]

- (c) The temperature of the mixture of gases is decreased. All other conditions stay the same.

The mixture of gases turns lighter purple.

State what can be deduced about the forward reaction from this information.

..... [1]

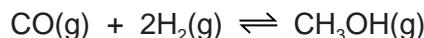
- (d) Deduce the oxidation number of iodine, I, in:

HI .....

I<sub>2</sub> .....

[2]

- (e) Methanol is manufactured by reacting carbon monoxide with hydrogen.



The rate of formation of methanol increases when a catalyst is used.

- (i) Choose from the list the element that is most likely to be used as the catalyst.

Draw a circle around your chosen answer.

calcium      carbon      copper      sodium      sulfur      [1]

- (ii) State the effect on the position of equilibrium when a catalyst is used.

..... [1]

- (iii) State the effect that a catalyst has on the activation energy,  $E_a$ , of a reaction.

..... [1]

[Total: 10]

5 (a) Lead(II) bromide,  $\text{PbBr}_2$ , is an insoluble salt and is made by precipitation.

- (i) Name **two** aqueous solutions that produce a precipitate of lead(II) bromide when they are mixed.

1 .....

2 .....

[2]

- (ii) Describe how to produce a pure sample of lead(II) bromide from the mixture of aqueous solutions in (a)(i).

.....

[2]

- (iii) Write an ionic equation for the precipitation reaction which produces lead(II) bromide. Include state symbols.

..... [3]

(b) When iron(II) sulfate crystals are heated strongly, sulfur dioxide gas is given off.

Describe a test for sulfur dioxide gas.

test .....

observations .....

[2]

(c) Complete the equation for the thermal decomposition of hydrated cobalt(II) nitrate.



[2]

- (d) Hydrated cobalt(II) sulfate,  $\text{CoSO}_4 \cdot \text{xH}_2\text{O}$ , produces water when it is heated.



A student does an experiment to determine the value of  $\text{x}$  in  $\text{CoSO}_4 \cdot \text{xH}_2\text{O}$ .

**step 1** The student weighs a sample of hydrated cobalt(II) sulfate.

**step 2** The student heats the sample of hydrated cobalt(II) sulfate.

**step 3** The student weighs the remaining solid after heating.

- (i)** Describe what else the student should do to ensure that **all** the water has been given off.  
No other substances are required.

.....  
.....  
.....

[2]

- (ii)** In an experiment, 1.405 g of  $\text{CoSO}_4 \cdot \text{xH}_2\text{O}$  is heated until all the water is given off.

The mass of  $\text{CoSO}_4$  that remains is 0.775 g.

$[M_r: \text{CoSO}_4, 155; \text{H}_2\text{O}, 18]$

Determine the value of  $\text{x}$  using the following steps.

- Calculate the number of moles of  $\text{CoSO}_4$  that remains.

..... mol

- Calculate the mass of  $\text{H}_2\text{O}$  given off.

..... g

- Calculate the number of moles of  $\text{H}_2\text{O}$  given off.

..... mol

- Determine the value of  $\text{x}$ .

$\text{x} = \dots$

[4]

[Total: 17]

6 This question is about metals.

(a) Fig. 6.1 shows a blast furnace used to extract iron from its ore.

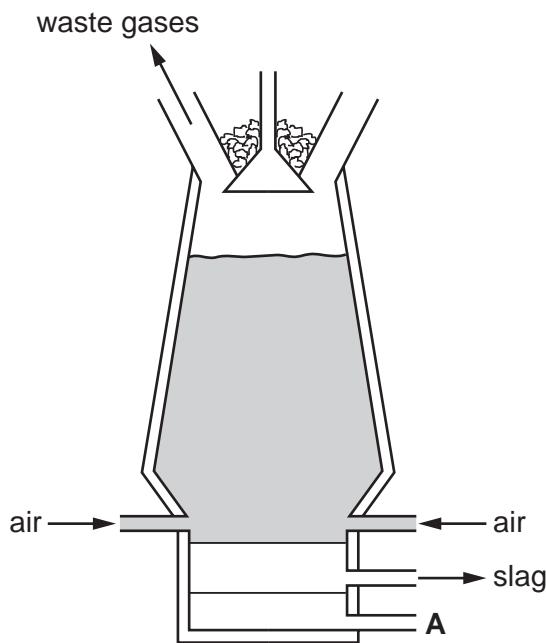


Fig. 6.1

(i) Coke and iron ore are added at the top of the blast furnace.

Name one **other** substance that is added at the top of the blast furnace.

..... [1]

(ii) Name the substance that leaves the blast furnace at A.

..... [1]

(iii) Slag is produced from an impurity in iron ore.

Name the impurity in iron ore that is converted into slag.

..... [1]

(iv) Name **two** substances that react together to produce the high temperature in the blast furnace.

..... and ..... [1]

(v) Name **two** waste gases that leave the blast furnace.

1 .....

2 .....

[2]

(b) Zinc is produced from zinc oxide in a furnace.

The zinc is produced as a gas. It then forms molten zinc.

(i) Suggest why the zinc produced inside the furnace is a gas.

..... [1]

(ii) State the name of the physical change that occurs when gaseous zinc is converted into molten zinc.

..... [1]

(c) Zinc is used to coat iron to prevent rusting.

(i) Name the process used to coat iron with zinc as a method of rust prevention.

..... [1]

(ii) When the zinc coating is scratched, the iron underneath does not rust.

Explain why the iron underneath the zinc does **not** rust.

.....  
..... [2]

(d) Zinc oxide neutralises both acids and bases.

(i) State the general name given to oxides that neutralise both acids and bases.

..... [1]

(ii) When zinc oxide neutralises aqueous sodium hydroxide, sodium zincate is formed.

The formula of the zincate ion is  $\text{ZnO}_2^{2-}$ .

Deduce the formula of sodium zincate.

..... [1]

(iii) Name the zinc compound that forms when zinc oxide neutralises dilute sulfuric acid.

..... [1]

[Total: 14]

- 7** Many organic compounds contain carbon and hydrogen only.

(a) (i) Organic compound A has the following composition by mass.

C, 82.76%; H, 17.24%

Calculate the empirical formula of compound A.

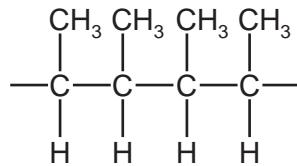
empirical formula = ..... [3]

(ii) Compound B has the empirical formula  $\text{CH}_2$  and a relative molecular mass of 70.

Determine the molecular formula of compound B.

molecular formula = ..... [1]

**(b)** Fig. 7.1 shows a section of polymer Q.



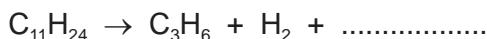
**Fig. 7.1**

- Draw the displayed formula of the monomer that forms polymer Q.
  - Name the monomer used to form polymer Q.

[3]

- (c) Propene,  $C_3H_6$ , can be produced by heating  $C_{11}H_{24}$ . The products of the reaction are propene, hydrogen and one other product in a 1:1:1 mole ratio.

Complete the symbol equation for this reaction.



[1]

- (d) Carboxylic acids and esters contain carbon, hydrogen and oxygen only.

An ester **X** and a carboxylic acid **Y** both contain 3 carbon atoms.

**X** and **Y** have the same molecular formula.

- (i) State the name given to compounds with the same molecular formula but different structural formulae.

..... [1]

- (ii) Esters are made by the reaction between carboxylic acids and alcohols.

Ester **X** is methyl ethanoate.

Name the carboxylic acid and the alcohol used to make methyl ethanoate.

carboxylic acid .....

alcohol .....

[2]

- (iii) Draw the displayed formula of carboxylic acid **Y**. Name the carboxylic acid.

name .....

[2]

[Total: 13]

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## The Periodic Table of Elements

I		II		Group																												
				I						II																						
				Key																												
				atomic number name relative atomic mass		atomic symbol																										
3	Li	4	Be	beryllium 9	9	1	H	hydrogen 1																								
11	Na	12	Mg	magnesium 24	23	20	Sc	scandium 45	21	22	23	24	Cr	Mn	Fe	Co	Cu	Zn	Ga	Ge	As	Se	S	Cl	F	O	Ne	He				
19	K	20	Ca	calcium 40	40	48	Ti	titanium 48	51	51	52	52	Cr	Mn	Fe	Co	copper 64	65	Ni	nickel 59	59	59	73	31	32	33	34	Br	selenium 80	Fr	helium 4	
37	Rb	38	Sr	strontium 88	88	89	Y	yttrium 89	40	40	41	41	Mo	Tc	Ru	Rh	ruthenium 101	103	Pd	palladium 106	46	47	48	In	In	In	In	I	Kr	Xe	Neon 84	
55	Cs	56	Ba	barium 137	137	57-71	Hf	hafnium 178	72	72	73	74	W	Ta	Ta	Ir	osmium 190	192	Pt	platinum 195	78	79	80	Hg	mercury 201	Pb	Bi	Te	Te	At	Rn	radon –
87	Fr	88	Ra	radium –	–	89-103	Rf	rutherfordium –	104	105	106	107	Sg	Bh	Bh	Hs	meitnerium –	–	Ds	roentgenium –	110	109	111	Mc	Fl	Lv	Ts	Yb	Og	ogenesson –		
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57	La	58	Ce	59	Pr	60	Nd	61	Pm	62	Sm	63	Eu	64	Gd	65	Tb	66	Dy	67	Ho	68	Er	69	Tm	70	Yb	71	Lu
89	Ac	90	Th	91	Pa	92	U	93	Np	94	Pu	95	Am	96	Cm	97	Bk	98	Cf	99	Fm	100	Md	101	No	102	Lu	103	lawrencium –
actinoids																													

The volume of one mole of any gas is  $24\text{dm}^3$  at room temperature and pressure (r.t.p.).