



## Cambridge IGCSE™

## CHEMISTRY

0620/22

Paper 2 Multiple Choice (Extended)

May/June 2024

45 minutes

You must answer on the multiple choice answer sheet.

You will need: Multiple choice answer sheet  
Soft clean eraser  
Soft pencil (type B or HB is recommended)

## INSTRUCTIONS

- There are **forty** questions on this paper. Answer **all** questions.
- For each question there are four possible answers **A**, **B**, **C** and **D**. Choose the **one** you consider correct and record your choice in soft pencil on the multiple choice answer sheet.
- Follow the instructions on the multiple choice answer sheet.
- Write in soft pencil.
- Write your name, centre number and candidate number on the multiple choice answer sheet in the spaces provided unless this has been done for you.
- Do **not** use correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.

## INFORMATION

- The total mark for this paper is 40.
- Each correct answer will score one mark.
- Any rough working should be done on this question paper.
- The Periodic Table is printed in the question paper.

This document has **16** pages. Any blank pages are indicated.



## 2

1 Which gas has the slowest rate of diffusion?

- A**  $\text{H}_2$                       **B**  $\text{NH}_3$                       **C**  $\text{CH}_4$                       **D**  $\text{CO}_2$

2 Which statements about the position of the elements in the Periodic Table are correct?

- 1 Elements in the same group have similar chemical properties.
- 2 Elements in the same period have similar chemical properties.
- 3 Elements in the same group have the same number of electron shells.
- 4 Elements in the same group have the same number of outer shell electrons.

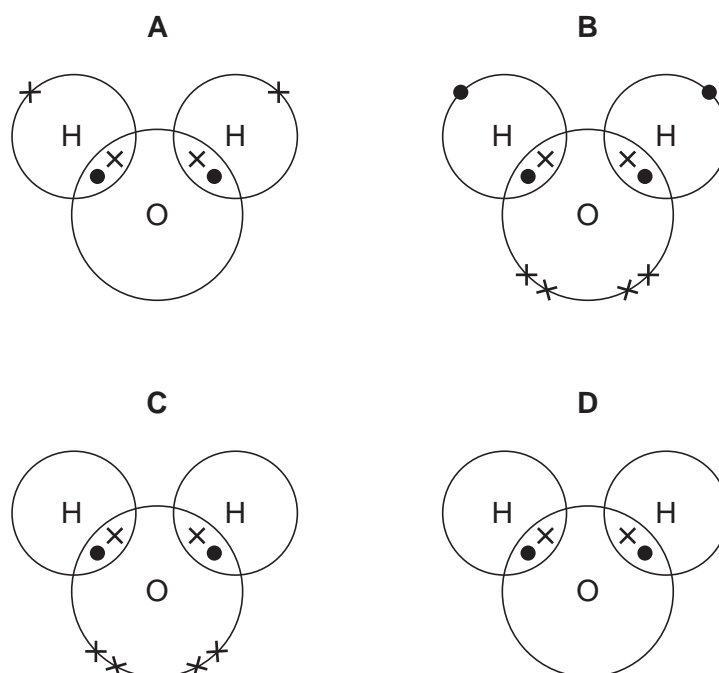
- A** 1 and 3                      **B** 1 and 4                      **C** 2 and 3                      **D** 2 and 4

3 Which statements about isotopes are correct?

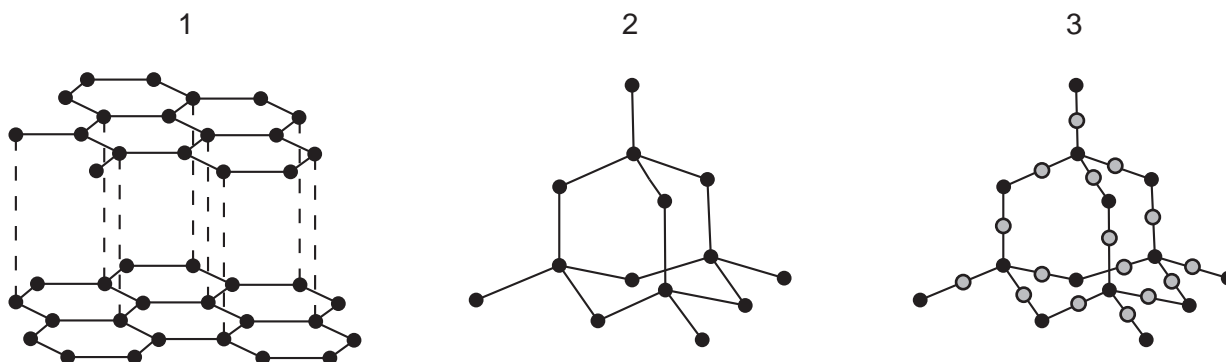
- 1 Isotopes are atoms of different elements with the same number of protons.
- 2 Isotopes of the same element have the same chemical properties.
- 3 Isotopes are atoms with the same relative atomic mass.
- 4 Isotopes of the same element have the same electronic configuration.

- A** 1 and 2                      **B** 1 and 3                      **C** 2 and 4                      **D** 3 and 4

4 Which diagram shows the arrangement of the outer shell electrons in a molecule of water?



5 The structures of three substances are shown.



Which substances are hard and have a high melting point?

- A** 1, 2 and 3      **B** 1 only      **C** 2 only      **D** 2 and 3 only

6 Information about four substances, W, X, Y and Z, is shown.

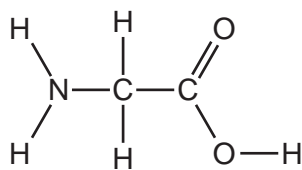
	melting point/ °C	electrical conductivity
W	1710	does not conduct when solid
X	3500	conducts when solid
Y	120	does not conduct
Z	801	conducts when molten

W, X, Y and Z are graphite, poly(ethene), sodium chloride and silicon(IV) oxide but not in that order.

What are W, X, Y and Z?

	W	X	Y	Z
<b>A</b>	graphite	poly(ethene)	silicon(IV) oxide	sodium chloride
<b>B</b>	sodium chloride	graphite	poly(ethene)	silicon(IV) oxide
<b>C</b>	poly(ethene)	sodium chloride	graphite	silicon(IV) oxide
<b>D</b>	silicon(IV) oxide	graphite	poly(ethene)	sodium chloride

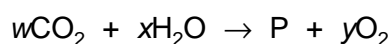
- 7 The structure of glycine is shown.



Which row is correct?

	formula of glycine	number of different elements in glycine
<b>A</b>	$\text{CH}_5\text{O}_2\text{N}$	10
<b>B</b>	$\text{C}_2\text{H}_5\text{O}_2\text{N}$	4
<b>C</b>	$\text{C}_2\text{H}_5\text{O}_2\text{N}$	10
<b>D</b>	$\text{H}_2\text{NCHCOOH}$	4

- 8 The incomplete equation for photosynthesis is shown.



Compound P is a product of the reaction.

Which row describes the values of  $w$ ,  $x$  and  $y$  and gives the empirical formula of compound P?

	values of $w$ , $x$ and $y$	empirical formula of compound P
<b>A</b>	$w$ , $x$ and $y$ are all the same	$\text{CH}_2\text{O}$
<b>B</b>	$w$ , $x$ and $y$ are all the same	$\text{C}_6\text{H}_{12}\text{O}_6$
<b>C</b>	$w$ and $x$ are the same and both are greater than $y$	$\text{CH}_2\text{O}$
<b>D</b>	$w$ and $x$ are the same and both are greater than $y$	$\text{C}_6\text{H}_{12}\text{O}_6$

- 9 The concentration and volume of an aqueous alkali are known.

Which additional information is required to calculate the number of moles of acid needed to neutralise the aqueous alkali?

- A** the concentration of the acid
- B** the equation for the acid–alkali reaction
- C** the formula of the acid
- D** the volume of the acid required for neutralisation

10 Which statement about electrolysis is correct?

- A Electrons move through the electrolyte from the cathode to the anode.
- B Electrons move in the external circuit towards the cathode.
- C Negative ions move in the external circuit towards the anode.
- D Positive ions move through the electrolyte towards the anode.

11 Aqueous copper(II) sulfate is electrolysed using copper electrodes.

What is the half-equation for the reaction at the cathode?

- A  $2\text{H}^+ + 2\text{e}^- \rightarrow \text{H}_2$
- B  $4\text{OH}^- \rightarrow \text{O}_2 + 2\text{H}_2\text{O} + 4\text{e}^-$
- C  $\text{Cu} \rightarrow \text{Cu}^{2+} + 2\text{e}^-$
- D  $\text{Cu}^{2+} + 2\text{e}^- \rightarrow \text{Cu}$

12 Three statements about activation energy,  $E_a$ , are listed.

- 1 Colliding particles must have at least  $E_a$  before they can react.
- 2  $E_a$  for exothermic reactions is always greater than for endothermic reactions.
- 3  $E_a$  is always endothermic.

Which statements are correct?

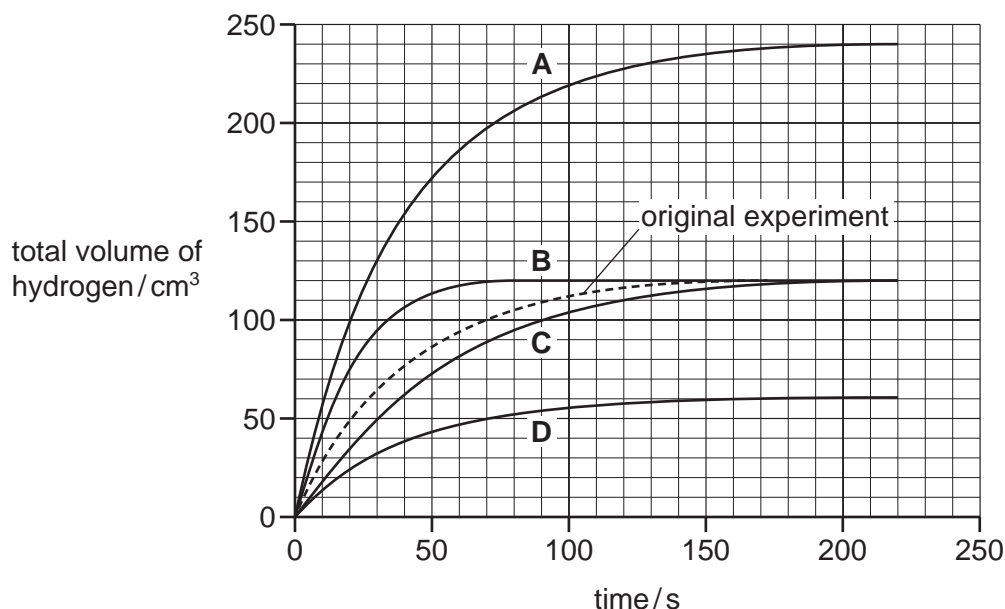
- A 1, 2 and 3      B 1 and 2 only      C 1 and 3 only      D 2 and 3 only

- 13** A student adds excess magnesium ribbon to  $10\text{ cm}^3$  of  $0.5\text{ mol/dm}^3$  sulfuric acid.

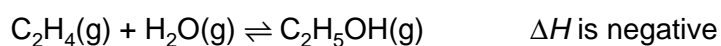
The hydrogen gas produced is collected and its total volume is measured every 10 seconds.

The experiment is repeated with  $5\text{ cm}^3$  of  $0.5\text{ mol/dm}^3$  sulfuric acid added to  $5\text{ cm}^3$  of water using the same mass of magnesium ribbon.

Which line on the graph shows the results of the second experiment?



- 14** The equation represents the reversible reaction between ethene and steam.



Which row describes the conditions that produce the greatest yield of ethanol?

	pressure	temperature
<b>A</b>	low	low
<b>B</b>	low	high
<b>C</b>	high	low
<b>D</b>	high	high

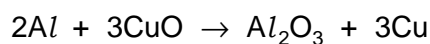
- 15 Which row identifies the pressure and the catalyst used for the conversion of sulfur dioxide to sulfur trioxide in the Contact process?

	pressure/atm	catalyst
<b>A</b>	2	iron
<b>B</b>	2	vanadium(V) oxide
<b>C</b>	200	iron
<b>D</b>	200	vanadium(V) oxide

- 16 Which equation represents a redox reaction?

- A**  $\text{AgNO}_3 + \text{NaCl} \rightarrow \text{AgCl} + \text{NaNO}_3$   
**B**  $\text{PCl}_5 + 4\text{H}_2\text{O} \rightarrow \text{H}_3\text{PO}_4 + 5\text{HCl}$   
**C**  $3\text{Fe} + 4\text{H}_2\text{O} \rightarrow \text{Fe}_3\text{O}_4 + 4\text{H}_2$   
**D**  $\text{MgCO}_3 + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{CO}_2 + \text{H}_2\text{O}$

- 17 The equation for the reaction between aluminium and copper(II) oxide is shown.



Which statements about this equation are correct?

- 1 The oxidation number of the aluminium reactant is +2.
- 2 The oxidation number of the aluminium in the product is +6.
- 3 The oxidation number of the copper in the reactant is +2.
- 4 The oxidation number of the copper product is 0.

- A** 1, 2 and 3      **B** 1 and 2 only      **C** 3 and 4      **D** 4 only

- 18 What is the colour of thymolphthalein in dilute hydrochloric acid?

- A** blue  
**B** red  
**C** yellow  
**D** colourless

- 19 Two acids, P and Q, with the same concentration and volume are reacted separately with the same mass of magnesium ribbon.

The reactions produce the same total volume of hydrogen gas but acid Q reacts much more slowly than acid P.

Which explanation for the difference between P and Q is correct?

- A Acid P has a higher pH than acid Q.
  - B Acid P has a lower concentration of hydrogen ions.
  - C Acid Q is partially dissociated and acid P is fully dissociated.
  - D Acid Q is a proton acceptor.
- 20 Which process is **not** used in the preparation of an insoluble salt?
- A filtration
  - B washing
  - C crystallisation
  - D drying

- 21 Lithium and potassium are metals in Group I of the Periodic Table.

Lithium has a melting point of 181 °C and a density of 0.53 g/cm<sup>3</sup>.

Which row describes the melting point and density of potassium?

	melting point in °C	density in g/cm <sup>3</sup>
A	less than 181	less than 0.53
B	less than 181	greater than 0.53
C	greater than 181	less than 0.53
D	greater than 181	greater than 0.53

- 22 Which statements about transition elements are correct?

- 1 They have a low density.
- 2 They form ions with variable oxidation numbers.
- 3 They have a high melting point.
- 4 They form only colourless compounds.

- A 1 and 2
- B 1 and 4
- C 2 and 3
- D 3 and 4



23 Which statements describe the elements in Group VIII of the Periodic Table?

- 1 Their atoms have full outer electron shells.
- 2 They are unreactive metals.
- 3 They are monatomic gases.
- 4 They are diatomic gases.

**A** 1 and 3      **B** 1 and 4      **C** 2 and 3      **D** 2 and 4

24 Four metals, Q, R, S and T, are each added to separate samples of water, steam and dilute hydrochloric acid.

The results are shown.

	observation with water	observation with steam	observation with dilute hydrochloric acid
Q	slow reaction	fast reaction	fast reaction
R	no reaction	no reaction	no reaction
S	no reaction	very slow reaction	slow reaction
T	fast reaction	explodes	explodes

Which statements are correct?

- 1 R is the least reactive metal.
- 2 T could be potassium.
- 3 S is more reactive than Q and R.
- 4 Metals generally react faster with steam than they react with water.

**A** 1, 2 and 4      **B** 1 and 2 only      **C** 2 and 3      **D** 3 and 4

25 Metal X acts as a sacrificial metal to prevent iron from corroding.

Metal X does **not** act as a sacrificial metal to prevent aluminium from corroding.

What is X?

- A** copper
- B** magnesium
- C** silver
- D** zinc

26 Which equation represents the reduction of a compound found in hematite in the blast furnace?

- A  $2\text{Al}_2\text{O}_3 \rightarrow 4\text{Al} + 3\text{O}_2$
- B  $\text{CaO} + \text{SiO}_2 \rightarrow \text{CaSiO}_3$
- C  $2\text{FeO} + \text{CO} \rightarrow \text{Fe}_2\text{O}_3 + \text{C}$
- D  $\text{Fe}_2\text{O}_3 + 3\text{CO} \rightarrow 2\text{Fe} + 3\text{CO}_2$

27 Which statements about the treatment of domestic water supplies are correct?

- 1 Filtration is used to remove insoluble substances from the water.
- 2 Sedimentation is used to remove soluble substances from the water.
- 3 Carbon is used to remove tastes and odours from the water.
- 4 Chlorine is used to lower the pH of the water.

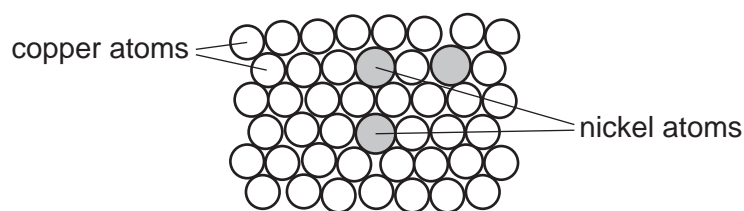
- A 1 and 3      B 1 and 4      C 2 and 3      D 2 and 4

28 Which statement about water is correct?

- A It turns anhydrous copper(II) sulfate from pink to blue.
- B It turns anhydrous copper(II) sulfate from white to blue.
- C It turns cobalt(II) chloride paper from blue to white.
- D It turns cobalt(II) chloride paper from pink to blue.

29 Cupronickel is used to make coins.

The arrangement of atoms in cupronickel is shown.



Which kind of substance is cupronickel?

- A an alloy
- B an isotope
- C a compound
- D a transition element

**30** Which physical properties are typical of **all** metals?

- 1 good heat conductivity
- 2 low density
- 3 malleability

**A** 1 and 2      **B** 1 and 3      **C** 2 and 3      **D** 3 only

**31** Which oxide is used to neutralise acidic gases in flue gas desulfurisation?

- A** calcium oxide  
**B** carbon dioxide  
**C** nitrogen oxide  
**D** sulfur dioxide

**32** Some information about three gases, P, Q and R, is listed.

- Gas P forms when magnesium reacts with dilute hydrochloric acid.
- Gas Q makes up 78% of the air.
- Gas R forms when gas P reacts with gas Q.

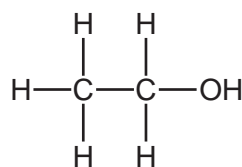
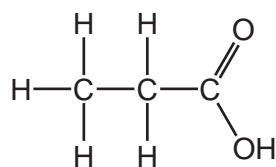
What is gas R?

- A** ammonia  
**B** methane  
**C** nitrogen dioxide  
**D** water vapour

**33** Which formula represents an alkene?

**A**  $\text{CH}_4$       **B**  $\text{C}_2\text{H}_4$       **C**  $\text{C}_2\text{H}_6$       **D**  $\text{C}_2\text{H}_5\text{OH}$

34 The structures of two compounds used to make an ester are shown.



What is the name of the ester?

- A ethyl propanoate
  - B propyl ethanoate
  - C ethyl ethanoate
  - D propyl propanoate
- 35 Which statement about a homologous series is correct?
- A All members have the same general formula.
  - B All members have the same molecular formula.
  - C All members have similar physical properties.
  - D Members show a trend in their chemical properties.
- 36 Which statements about aqueous ethanoic acid are correct?
- 1 It contains the functional group  $\text{-COOH}$ .
  - 2 It reacts with carbonates to produce hydrogen.
  - 3 It turns universal indicator paper blue.
  - 4 It has a pH lower than pH 7.
- A 1 and 2      B 1 and 3      C 1 and 4      D 2 and 4

**37** Alkenes react with steam in an addition reaction.

Some alkenes produce only one alcohol product.

Some alkenes produce two different alcohols which are structural isomers.

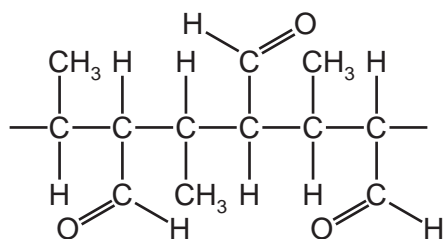
Which row gives the number of alcohol isomers formed when ethene reacts with steam and when propene reacts with steam?

	number of alcohol isomers formed	
	ethene + steam	propene + steam
<b>A</b>	1	1
<b>B</b>	1	2
<b>C</b>	2	1
<b>D</b>	2	2

**38** What is an advantage of manufacturing ethanol by fermentation rather than by the addition of steam to ethene?

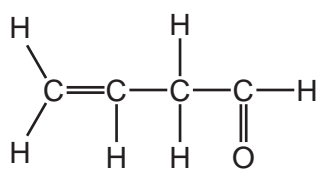
- A** No gases that cause global warming are produced.
- B** The ethanol that is produced is pure.
- C** Fermentation is a fast process.
- D** Fermentation uses renewable raw materials.

39 The diagram shows the structure of a polymer.

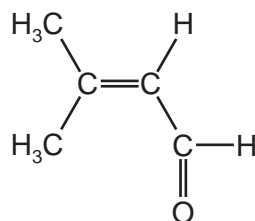


Which structure represents the monomer for this polymer?

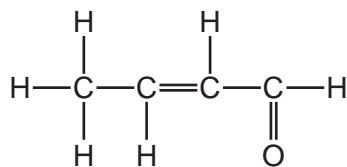
**A**



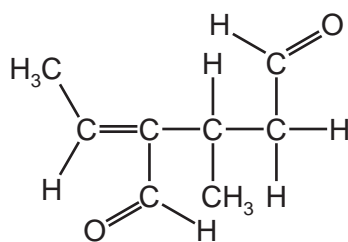
**B**



**C**



**D**



40 Which piece of apparatus is used to measure  $24.5 \text{ cm}^3$  of gas produced during a reaction?

- A** beaker
- B** conical flask
- C** measuring cylinder
- D** volumetric pipette

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The Periodic Table of Elements

Group																		
I	II											III	IV	V	VI	VII	VIII	
<div>1Hhydrogen1</div> <div>Key<div>atomic number atomic symbol name relative atomic mass</div></div>																		
3Li lithium 7	4Be beryllium 9											5B boron 11	6C carbon 12	7N nitrogen 14	8O oxygen 16	9F fluorine 19	10Ne neon 20	
11Na sodium 23	12Mg magnesium 24											13Al aluminium 27	14Si silicon 28	15P phosphorus 31	16S sulfur 32	17Cl chlorine 35.5	18Ar argon 40	
19K potassium 39	20Ca calcium 40	21Sc scandium 45	22Ti titanium 48	23V vanadium 51	24Cr chromium 52	25Mn manganese 55	26Fe iron 56	27Co cobalt 59	28Ni nickel 59	29Cu copper 64	30Zn zinc 65	31Ga gallium 70	32Ge germanium 73	33As arsenic 75	34Se selenium 79	35Br bromine 80	36Kr krypton 84	
37Rb rubidium 85	38Sr strontium 88	39Y yttrium 89	40Zr zirconium 91	41Nb niobium 93	42Mo molybdenum 96	43Tc technetium —	44Ru ruthenium 101	45Rh rhodium 103	46Pd palladium 106	47Ag silver 108	48Cd cadmium 112	49In indium 115	50Sn tin 119	51Sb antimony 122	52Te tellurium 128	53I iodine 127	54Xe xenon 131	
55Cs caesium 133	56Ba barium 137	57–71 lanthanoids		72Hf hafnium 178	73Ta tantalum 181	74W tungsten 184	75Re rhenium 186	76Os osmium 190	77Ir iridium 192	78Pt platinum 195	79Au gold 197	80Hg mercury 201	81Tl thallium 204	82Pb lead 207	83Bi bismuth 209	84Po polonium —	85At astatine —	86Rn radon —
87Fr francium —	88Ra radium —	89–103 actinoids		104Rf rutherfordium —	105Db dubnium —	106Sg seaborgium —	107Bh bohrium —	108Hs hassium —	109Mt meitnerium —	110Ds darmstadtium —	111Rg roentgenium —	112Cn copernicium —	113Nh nihonium —	114Fl flerovium —	115Mc moscovium —	116Lv livermorium —	117Ts tennessine —	118Og oganeson —

lanthanoids	57 La lanthanum 139	58 Ce cerium 140	59 Pr praseodymium 141	60 Nd neodymium 144	61 Pm promethium —	62 Sm samarium 150	63 Eu europium 152	64 Gd gadolinium 157	65 Tb terbium 159	66 Dy dysprosium 163	67 Ho holmium 165	68 Er erbium 167	69 Tm thulium 169	70 Yb ytterbium 173	71 Lu lutetium 175
	89 Ac actinium —	90 Th thorium 232	91 Pa protactinium 231	92 U uranium 238	93 Np neptunium —	94 Pu plutonium —	95 Am americium —	96 Cm curium —	97 Bk berkelium —	98 Cf californium —	99 Es einsteinium —	100 Fm fermium —	101 Md mendelevium —	102 No nobelium —	103 Lr lawrencium —

The volume of one mole of any gas is 24dm<sup>3</sup> at room temperature and pressure (r.t.p.).