

# Cambridge IGCSE™

## CHEMISTRY

0620/22

Paper 2 Multiple Choice (Extended)

October/November 2023

45 minutes

You must answer on the multiple choice answer sheet.



You will need: Multiple choice answer sheet  
Soft clean eraser  
Soft pencil (type B or HB is recommended)

### INSTRUCTIONS

- There are **forty** questions on this paper. Answer **all** questions.
- For each question there are four possible answers **A**, **B**, **C** and **D**. Choose the **one** you consider correct and record your choice in soft pencil on the multiple choice answer sheet.
- Follow the instructions on the multiple choice answer sheet.
- Write in soft pencil.
- Write your name, centre number and candidate number on the multiple choice answer sheet in the spaces provided unless this has been done for you.
- Do **not** use correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.

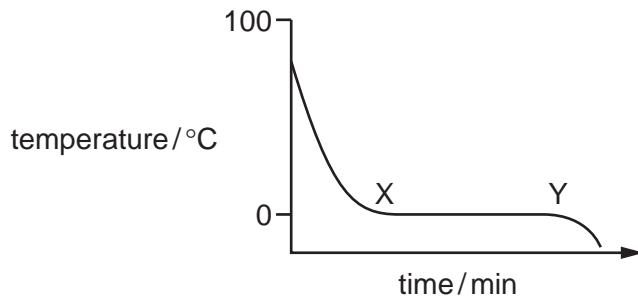
### INFORMATION

- The total mark for this paper is 40.
- Each correct answer will score one mark.
- Any rough working should be done on this question paper.
- The Periodic Table is printed in the question paper.

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This document has **16** pages.

- 1 Part of a cooling curve for water is shown.



What is occurring between points X and Y?

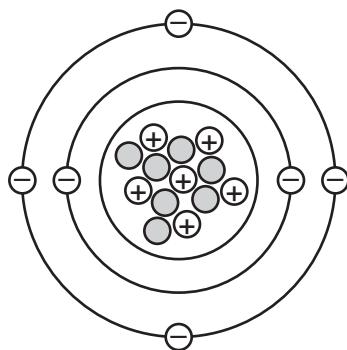
- A Steam is condensing into water.
- B The temperature of the water is decreasing.
- C Ice is melting.
- D Particles are losing heat to the surroundings.

- 2 Which statements about clean, dry air are correct?

- 1 It is a mixture of elements only.
- 2 It is a mixture of elements and compounds.
- 3 It contains only non-metals.

- A 1 and 3
- B 1 only
- C 2 and 3
- D 2 only

- 3 A representation of an atom is shown.



What is the nucleon number of this atom?

- A 6
- B 7
- C 12
- D 13

- 4 The percentage abundances of three isotopes in a sample of neon are shown.

isotope	percentage abundance / %
$^{20}_{10}\text{Ne}$	90.48
$^{21}_{10}\text{Ne}$	0.27
$^{22}_{10}\text{Ne}$	9.25

What is the relative atomic mass,  $A_r$ , of this sample of neon?

- A 10.19      B 20.19      C 21.00      D 30.19

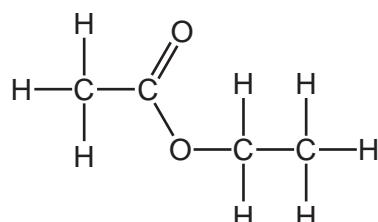
- 5 Potassium reacts with iodine to form potassium iodide.

Which statement about potassium iodide is correct?

- A Each potassium atom shares a pair of electrons with an iodine atom.  
 B In potassium iodide, the particles of potassium have more protons than electrons.  
 C Potassium iodide has a high melting point because it is a covalent compound.  
 D Potassium iodide has a low melting point because it is an ionic compound.
- 6 Which substance has the lowest melting point?

- A graphite  
 B methanol  
 C silicon(IV) oxide  
 D sodium chloride

- 7 The diagram shows the structure of a molecule of ethyl ethanoate.



What is the molecular formula of a molecule of ethyl ethanoate?

- A CHO      B  $\text{C}_4\text{H}_8\text{O}_2$       C  $\text{C}_4(\text{H}_2)_2(\text{O}_2)$       D  $\text{C}_2\text{H}_4\text{O}$

- 8 A hydrocarbon contains 85.7% of carbon by mass.

What is the empirical formula of the hydrocarbon?

- A  $\text{CH}_2$       B  $\text{CH}_4$       C  $\text{C}_2\text{H}_5$       D  $\text{C}_3\text{H}_6$

- 9 The formula of a compound containing element X is  $\text{Na}_2\text{X}_2\text{O}_3$ .

The relative formula mass of the compound is 158.

What is the relative atomic mass of X?

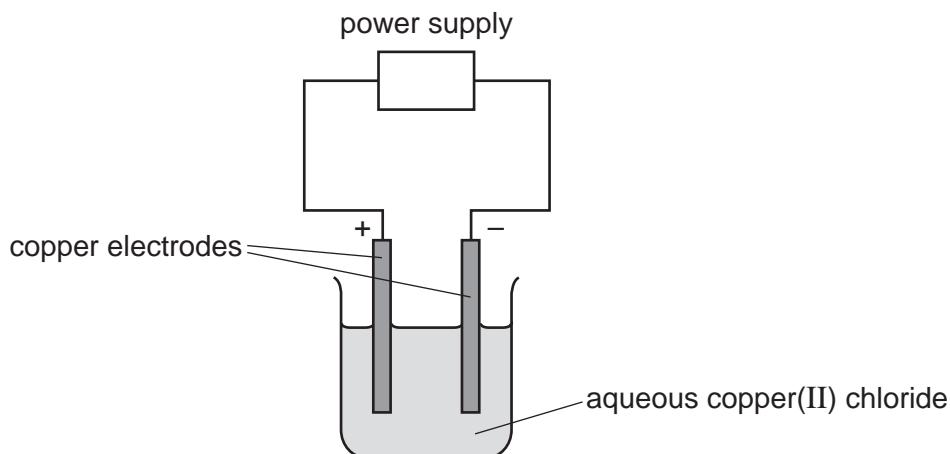
- A 32      B 59.5      C 64      D 119

- 10 Dilute aqueous potassium chloride is electrolysed using platinum electrodes.

Which row identifies the product at each electrode?

	anode	cathode
A	chlorine	hydrogen
B	chlorine	potassium
C	oxygen	hydrogen
D	oxygen	potassium

- 11 Concentrated aqueous copper(II) chloride is electrolysed using copper electrodes, as shown.



What happens to the mass of each electrode during this process?

	positive electrode	negative electrode
<b>A</b>	decreases	decreases
<b>B</b>	decreases	increases
<b>C</b>	increases	decreases
<b>D</b>	increases	increases

- 12 The initial and final temperatures of four different reactions are measured.

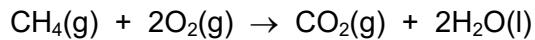
Which reaction is the **least** exothermic?

	initial temperature /°C	final temperature /°C
<b>A</b>	19	25
<b>B</b>	21	18
<b>C</b>	22	17
<b>D</b>	22	26

- 13 Which equation represents an endothermic reaction?

- A**  $\text{Cl}_2(\text{g}) \rightarrow 2\text{Cl}(\text{g})$
- B**  $\text{CH}_4(\text{g}) + 2\text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{l})$
- C**  $\text{H}(\text{g}) + \text{H}(\text{g}) \rightarrow \text{H}_2(\text{g})$
- D**  $2\text{K}(\text{s}) + 2\text{H}_2\text{O}(\text{l}) \rightarrow 2\text{KOH}(\text{aq}) + \text{H}_2(\text{g})$

14 Methane burns in oxygen to form carbon dioxide and water.



The bond energies are shown.

bond	bond energy in kJ/mol
C–H	410
C–O	360
C=O	805
O–H	460
O–O	146
O=O	496

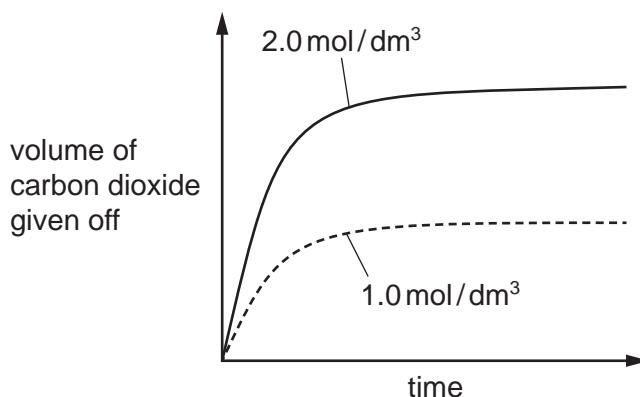
What is the energy change for this reaction?

- A –818 kJ/mol    B –102 kJ/mol    C +102 kJ/mol    D +818 kJ/mol

- 15 Hydrochloric acid is added to excess calcium carbonate in two separate experiments.

Two different concentrations of hydrochloric acid are used but the temperature is the same in both experiments.

The graph of the results shows the volume of carbon dioxide gas given off over time.

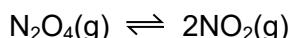


Which row is correct?

	particles in 2.0 mol / dm <sup>3</sup> compared to 1.0 mol / dm <sup>3</sup>	
	collision rate	collision energy
A	higher	no change
B	higher	higher
C	lower	no change
D	lower	higher

- 16 The decomposition of dinitrogen tetroxide, N<sub>2</sub>O<sub>4</sub>, into nitrogen dioxide, NO<sub>2</sub>, is a reversible reaction.

The equation for the reaction is shown.

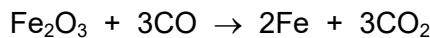


The forward reaction is endothermic.

Which row shows the effect on the position of equilibrium and the rate of the reverse reaction when the temperature is increased?

	position of equilibrium	rate of the reverse reaction
A	shifts to the left	decreases
B	shifts to the left	increases
C	shifts to the right	decreases
D	shifts to the right	increases

- 17** In a blast furnace, iron(III) oxide is converted to iron and carbon monoxide is converted to carbon dioxide.



What happens to each of these reactants?

- A** Both iron(III) oxide and carbon monoxide are oxidised.
  - B** Both iron(III) oxide and carbon monoxide are reduced.
  - C** Iron(III) oxide is oxidised and carbon monoxide is reduced.
  - D** Iron(III) oxide is reduced and carbon monoxide is oxidised.
- 18** Which row describes what happens to  $\text{Fe}^{2+}$  ions when they are oxidised?

	electron movement	oxidation number of iron
<b>A</b>	they gain electrons	decreases
<b>B</b>	they gain electrons	increases
<b>C</b>	they lose electrons	decreases
<b>D</b>	they lose electrons	increases

- 19** In which reaction does an acid react with a base?

- A** Dilute sulfuric acid is added to a piece of magnesium ribbon producing hydrogen.
- B** Dilute sulfuric acid is added to aqueous barium chloride producing a white precipitate of barium sulfate.
- C** Aqueous sodium hydroxide is added to aqueous copper(II) sulfate producing a blue precipitate of copper(II) hydroxide.
- D** Aqueous sodium hydroxide is added to solid ammonium sulfate producing gaseous ammonia.

- 20** Which element forms an oxide that reacts with an aqueous solution of a base?

- A** argon
- B** sulfur
- C** magnesium
- D** copper

**21** Which method is used to produce insoluble salts?

- A** addition of excess insoluble base to an acid
- B** addition of excess metal to an acid
- C** precipitation using two aqueous solutions
- D** titration using an acid and an alkali

**22** The noble gases are in Group VIII of the Periodic Table.

Some properties of the first four noble gases are shown.

noble gas	boiling point in °C	density in g/dm <sup>3</sup>
helium	-267	0.179
neon	-246	0.900
argon	-186	1.782
krypton	-152	3.708

Which row identifies the trends in boiling point and in density as Group VIII is descended?

	boiling point	density
<b>A</b>	decreasing	increasing
<b>B</b>	increasing	increasing
<b>C</b>	decreasing	decreasing
<b>D</b>	increasing	decreasing

**23** Some properties of element R are shown.

melting point in °C	98
boiling point in °C	883
reaction with cold water	gives off H <sub>2</sub> gas
reaction when heated with oxygen	burns to give a white solid

In which part of the Periodic Table is R found?

- A** Group I
- B** Group VII
- C** Group VIII
- D** transition elements

**24** Which pair of compounds shows that transition elements have variable oxidation states?

- A** Cr<sub>2</sub>O<sub>3</sub> and CrBr<sub>3</sub>
- B** CuSO<sub>4</sub> and CuCl<sub>2</sub>
- C** Fe<sub>2</sub>O<sub>3</sub> and FeCl<sub>2</sub>
- D** NiO and NiCl<sub>2</sub>

**25** The list gives the order of some metals and hydrogen in the reactivity series.

Metal X is also included.

most reactive	K
Mg	
Zn	
H	
X	
least reactive	Cu

Which row shows the properties of metal X?

	reacts with dilute acids	oxide reduced by carbon
<b>A</b>	no	no
<b>B</b>	no	yes
<b>C</b>	yes	no
<b>D</b>	yes	yes

**26** When zinc is added to an aqueous solution containing magnesium ions, there is no reaction.

Which species has the greatest tendency to lose electrons?

- A** Mg
- B** Mg<sup>2+</sup>
- C** Zn
- D** Zn<sup>2+</sup>

**27** Which gas in the air is needed for iron to rust?

- A** argon
- B** carbon dioxide
- C** nitrogen
- D** oxygen

**28** Which coating prevents iron from rusting even when the coating is damaged?

- A** grease
- B** paint
- C** plastic
- D** zinc

**29** Why is limestone added to the blast furnace?

- A** It neutralises the molten slag produced.
- B** It reacts with impurities to form slag.
- C** It releases carbon dioxide which reduces the iron(III) oxide.
- D** It removes acidic gases such as carbon dioxide.

**30** The flow chart shows stages in the treatment of river water to produce drinking water.



What occurs at stages J and K?

	J	K
<b>A</b>	distillation	chlorination
<b>B</b>	distillation	filtration
<b>C</b>	filtration	chlorination
<b>D</b>	filtration	distillation

- 31** Carbon dioxide acts as a greenhouse gas by interacting with a particular type of energy that radiates from the Earth's surface into the atmosphere.

Which type of energy is involved and what happens when this energy interacts with carbon dioxide molecules?

	type of energy involved	what happens
<b>A</b>	thermal	carbon dioxide molecules increase the Earth's energy loss to space
<b>B</b>	thermal	carbon dioxide molecules absorb the energy
<b>C</b>	light	carbon dioxide molecules increase the Earth's energy loss to space
<b>D</b>	light	carbon dioxide molecules absorb the energy

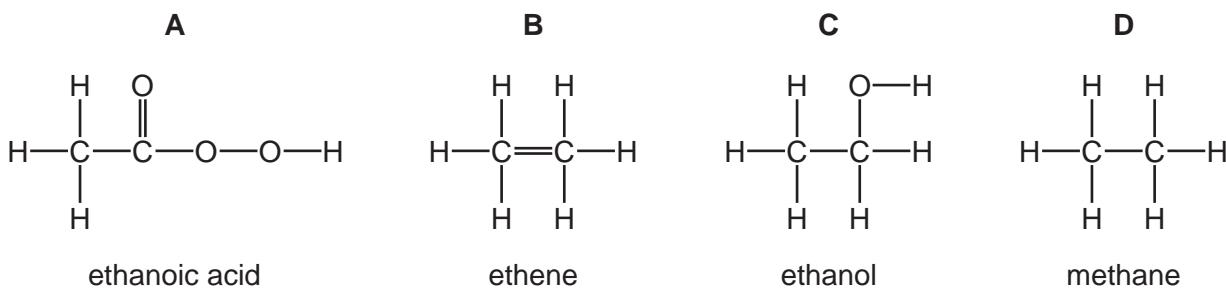
- 32** Oxides of nitrogen, such as NO and  $\text{NO}_2$ , are formed in the petrol engines of cars.

They are removed from the exhaust gases by reactions in the car's catalytic converter.

Which row describes how oxides of nitrogen are formed in a petrol engine and a reaction that happens in the catalytic converter?

	how oxides of nitrogen are formed	a reaction that happens in the catalytic converter
<b>A</b>	by the reaction between nitrogen and oxygen from the air	$2\text{NO} + 2\text{CO} \rightarrow \text{N}_2 + 2\text{CO}_2$
<b>B</b>	by the reaction between nitrogen and oxygen from the air	$2\text{NO} + 2\text{H}_2 \rightarrow \text{N}_2 + 2\text{H}_2\text{O}$
<b>C</b>	by the reaction between nitrogen compounds in petrol and oxygen from the air	$2\text{NO} + 2\text{CO} \rightarrow \text{N}_2 + 2\text{CO}_2$
<b>D</b>	by the reaction between nitrogen compounds in petrol and oxygen from the air	$2\text{NO} + 2\text{H}_2 \rightarrow \text{N}_2 + 2\text{H}_2\text{O}$

- 33** Which diagram shows the displayed formula for the named organic compound?



**34** What is the total number of covalent bonds in a molecule of butane,  $C_4H_{10}$ ?

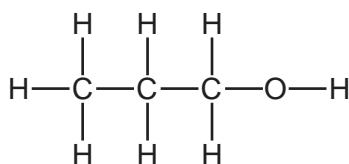
- A** 3      **B** 10      **C** 13      **D** 14

**35** Propane reacts with chlorine in a substitution reaction.

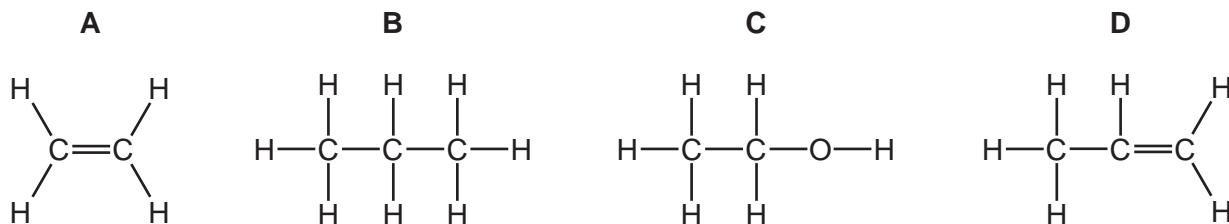
Which reaction condition is required for the reaction to occur?

- A** acid catalyst
- B** iron catalyst
- C** temperature of  $400\text{ }^{\circ}\text{C}$
- D** ultraviolet light

**36** The structure of an organic compound is shown.



Which structure represents a molecule that reacts with steam to produce this product?

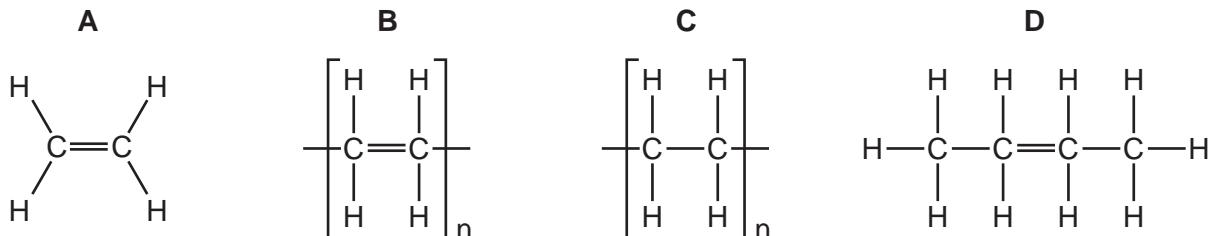


**37** Which term describes nylon?

- A** addition polymer
- B** natural polymer
- C** polyamide
- D** polyester

**38** Ethene can be polymerised.

Which diagram represents the structure of the product formed?



**39** An acid–base titration is described.

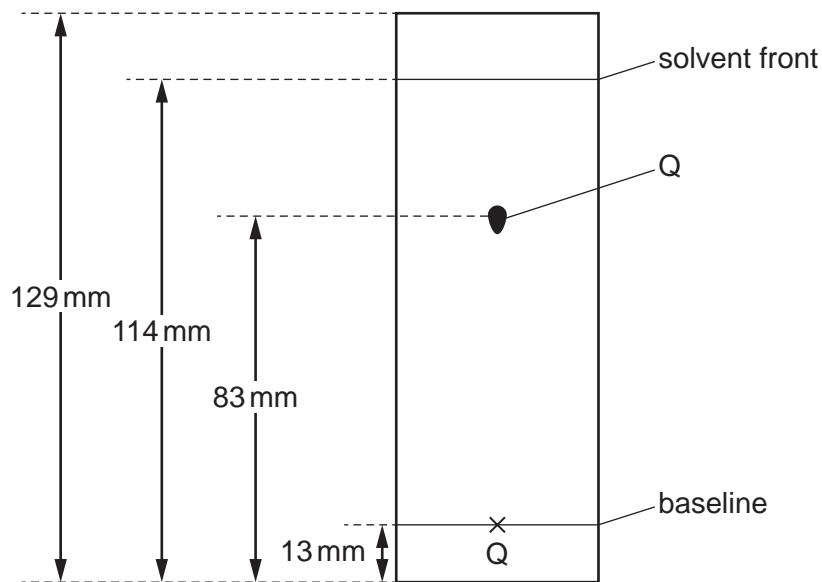
- 25.0 cm<sup>3</sup> of dilute aqueous alkali is put into a conical flask.
- Indicator is added to the flask.
- Dilute acid is added to the aqueous alkali until the indicator changes colour.
- The volume of acid used is then recorded.

Which use of apparatus is correct?

- A** The 25.0 cm<sup>3</sup> of aqueous alkali is measured using a volumetric pipette.
- B** The 25.0 cm<sup>3</sup> of aqueous alkali is measured using the lines on the conical flask.
- C** The volume of acid is measured using a measuring cylinder.
- D** The volume of acid is measured using a volumetric pipette.

- 40 Substance Q is investigated using chromatography.

The chromatogram is shown. The diagram is not drawn to scale.



What is the  $R_f$  value of Q?

- A** 0.60      **B** 0.64      **C** 0.69      **D** 0.72

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16

The Periodic Table of Elements

		Group											
I	II	1				2				VII	VIII		
		Key		atomic number		atomic symbol		III		IV	V	VI	VII
				name	relative atomic mass								
3	Li	4	Be	beryllium	9		H						
7	lithium						hydrogen	1					
11	Na	12	Mg	magnesium	24								
19	K	20	Ca	calcium	40	21	Sc	scandium	45	Ti	titanium	48	
39	potassium					22				V	vanadium	51	
37	38	38	Sr	strontium	88	39	Zr	zirconium	91	Cr	chromium	52	
85	Rb	rubidium				40	Y	yttrium	89	Fe	iron	56	
55	56	56	Ba	barium	137	57-71	Zr	niobium	93	Mn	manganese	55	
133	Cs	caesium				lanthanoids				Co	cobalt	59	
87	Fr	88	Ra	radium	—	actinoids				Ni	nickel	59	
—	francium									Cu	copper	64	
										Zn	zinc	65	
										Ge	germanium	73	
										As	arsenic	75	
										Se	selenium	79	
										Br	bromine	80	
										Te	tellurium	128	
										Sb	antimony	122	
										In	tin	119	
										Sn	tin	115	
										Pd	cadmium	112	
										Ag	silver	108	
										Co	palladium	106	
										Rh	rhodium	103	
										Ir	rhenium	101	
										Os	osmium	190	
										Re	rhenium	186	
										W	tungsten	184	
										Ta	tantalum	181	
										Hf	hafnium	178	
										Tl	thallium	204	
										Pt	platinum	195	
										Hg	mercury	201	
										Au	gold	197	
										Ag	silver	108	
										Pb	lead	207	
										Bi	bismuth	209	
										Po	polonium	—	
										At	astatine	—	
										Fr	francium	—	
										Mc	moscovium	—	
										Nh	nihonium	—	
										Lv	livemoreum	—	
										Ts	tennessine	—	
										Og	oganesson	—	
										He	helium	4	
										F	fluorine	19	
										Ne	neon	20	
										Kr	krypton	84	
										Xe	xenon	131	
										Rn	radon	—	

57	La	58	Ce	59	Pr	60	Nd	61	Pm	62	Sm	63	Eu	64	Gd	65	Tb	66	Dy	67	Ho	68	Er	69	Tm	70	Yb	71	Lu
	lanthanum		cerium		praseodymium		neodymium		promethium		samarium		euroopium		gadolinium		terbium		dysprosium		holmium		erbium		ytterbium		lutetium		
139		140		141		144		—		150		152		157		159		163		165		167		173		175			
89	Ac	90	Th	91	Pa	92	U	93	Pu	94	Am	95	Cm	96	Bk	97	Cf	98	Fm	99	Md	100	No	101	102	103	Lr		
	actinium		thorium		protactinium		uranium		plutonium		americium		curium		berkelium		—		californium		einsteinium		terbium		mendelevium		nobelium		
—		232		231		238		—		—		—		—		—		—		—		—		—		—	lawrencium	—	

The volume of one mole of any gas is  $24\text{ dm}^3$  at room temperature and pressure (r.t.p.).