TAREA 1

$m = 10\,g$

$Z = 47$

$M = 107.87\,g/mol.$

$$N = \left(\frac{10\,g}{107.87\,g/mol}\right)\left(N_A\left[\frac{atoms}{mol}\right]\right)\left(47\,\frac{elect.}{atom}\right)$$

Hatomos

$$N = 2.62 \times 10^{24}\ electrons.$$

$q = -1\,mC$

$$\frac{q}{e} = 6.24 \times 10^{15}\ electrons.$$

Antes 2.62×10^{24} el.

Agregamos. 6.24×10^{15} el.

paquetes 10^9 elec./paquete.

$$\#\text{paquetes\_Antes} = \frac{2.62 \times 10^{24}\ \text{el.}}{10^9\ \text{el./paq.}} = 2.62 \times 10^{15}\ \text{paq.}$$

$$\#\text{el/paq.} = \frac{6.24 \times 10^{15}\ \text{el}}{2.62 \times 10^{15}\ \text{paq}} = 2.38\ \text{el./paq.}$$